

Lithium-Oxygen Batteries and Related Systems: Potential, Status and Future

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Abstract

The goal of limiting Global Warming to 1.5 C requires a drastic reduction in CO₂ emissions across many sectors of the world economy. Batteries are vital to this endeavour, whether used in electric vehicles, to store renewable electricity or in aviation. Present lithium-ion technologies are preparing the public for this inevitable change, but their maximum theoretical specific capacity presents a limitation. Their high cost is another concern for commercial viability. Metal-air batteries have the highest theoretical energy density of all possible secondary battery technologies and could yield step changes in energy storage - if their practical difficulties could be overcome. The scope of this review is to provide an objective, comprehensive and authoritative assessment of the intensive work invested in non-aqueous rechargeable metal-air batteries over the last few years, which identified the key problems and guides directions to solve them. We focus primarily on the challenges and outlook for Li-O₂ cells, but include Na-O₂, K-O₂ and Mg-O₂ cells for comparison. Our review highlights the interdisciplinary nature of this field that involves a combination of materials chemistry, electrochemistry, computation, microscopy, spectroscopy, and surface science. The mechanisms of O₂ reduction and evolution are considered in the light of recent findings, along with developments in positive and negative electrodes, electrolytes, electro-catalysis on surfaces and in solution, and the degradative effect of singlet oxygen, which is typically formed in Li-O₂ cells.

Section 1: Introduction

1.1. Contemporary Energy Needs

Over the last decades, the intensive use of fossil fuels and electricity has dramatically increased the standard of living. Energy, particularly electrical energy, is core to a modern society. Presently, over 70% of the world's energy requirement is met by fossil fuels including coal, oil and natural gas, but this is damaging our environment. Serious issues like global warming and air pollution make it extremely clear that it is imperative to reduce our dependence on fossil fuels to meet our energy requirements. In order to implement the necessary changes scientific advances and technological solutions are urgently required.

Electrochemical power sources and energy storage systems will play a vital role in shifting the paradigm of the future energy network towards clean, renewable sources. This is because such systems form a vital bridge between dispatchable energy generation, and intermittent supply from renewable sources such as wind and solar power.^{1,2} The most impressive scientific and applicative successes of modern electrochemistry are lithium-ion batteries (LIBs), which today power most of our mobile electronic devices. Building on this progress, the scientific community has set ambitious targets in the hope of substantially reducing fossil fuel use in the near future and promote the propulsion of ground transportation by electric power. It is well known that in powering the 3Cs - computers, communication devices and consumer electronics - LIBs have deeply penetrated every corner of our daily lives.^{1,3} Despite LIBs successful history so far, there has been a continuous call to improve their capabilities and performance, primarily driven by their inability to meet society's emerging needs.² For example, the present LIBs, with a maximum specific energy of ca. 250 Wh kg⁻¹ (at the cell level), do not possess sufficient energy, rate, durability or affordability to match the performance of traditional automotive gasoline/internal combustion engines (Figure 1).² In addition, detrimental climatic change has hastened the need to electrify transport and develop reliable energy storage systems. These urgent requirements drive the scientific

community to further increase the energy and power density of electrochemical power sources and not least the ecological footprint beyond the limits of LIBs.

Relying on recent research efforts from both academia and industry, it is believed that further advancements in LIB technology and composition will eventually push their specific energy to $\sim 300 \text{ Wh kg}^{-1} \text{ cell}^{-1}$.¹⁻⁴ With such an upper limit, advanced LIBs can meet the requirement of mid-sized passenger or light goods electric vehicles. Nevertheless, for the realization of full-electric vehicles (EVs) with long driving range (500 miles), and especially for long range small vehicles and HGVs, new electrochemical couples and chemistries, with a targeted specific energy of $500 \text{ Wh kg}^{-1} \text{ cell}^{-1}$ and beyond, need to be developed. In addition, the low reserves (0.0017 wt.%) of Li in the earth's crust has spurred extensive research in innovative battery platforms exploiting other active metals such as Na, Mg or K.

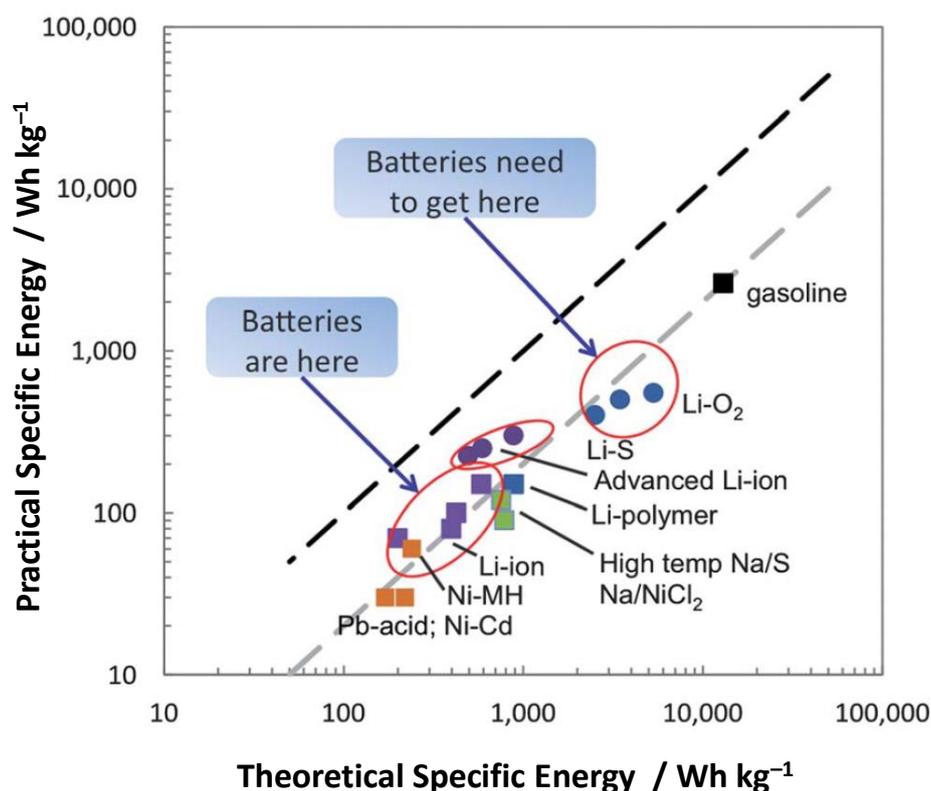


Figure 1. The graph shows the present position of various batteries on a plot of practical specific energy versus theoretical energy density. This illustrates the enormous challenge of increasing the practical energy density of advanced batteries beyond that of present state-of-the-art LIBs if they have to compete with the energy stored and delivered by fossil fuels like gasoline. Reproduced with permission from ref². Copyright 2012 Royal Society of Chemistry.

1.2. Brief overview of electrochemical energy storage

The electrification of many technologies has motivated researchers to develop advanced batteries replete with innovation. Currently, the development of Li-ion batteries is a topic of prime importance. Significant progress has been achieved with the discovery and development of some high energy layered cathode materials of the general formula $\text{Li}_{1+x}\text{Ni}_z\text{Co}_w\text{Mn}_y\text{O}_2$, denoted as Li and Mn-rich NCM compounds ($x > 0$, $y > 0.5$, $x+y+z+w = 1$). These cathodes can deliver a specific capacity $> 260 \text{ mAh g}^{-1}$, almost twice that of conventional layered, spinel, or olivine type cathode materials.³⁻⁸ Unfortunately, these high specific capacity cathode materials require high charging potentials $> 4.6 \text{ V}$, which challenge the stability of most relevant electrolyte solutions. Also, they suffer from stability problems, plagued by both capacity and average voltage fading during their cycling due to their complex structure and operation mechanism.^{6,9,10} Another approach to boost the energy densities of the LIBs is to use layered transition metal oxide cathodes with a high nickel content, denoted as Ni-rich NCM cathode materials ($\text{LiNi}_x\text{Co}_y\text{Mn}_z\text{O}_2$, $x > 0.8$)^{8,11-13} As the Ni content approaches 100%, their specific capacity can reach 240 mAh g^{-1} , with a maximal charging potential $< 4.3 \text{ V}$. These materials also suffer from capacity fading problems upon cycling; however, those issues can be mitigated by using judicious doping and coating.^{7,10,14-18} Using high voltage cathode materials ($> 4.7 \text{ V}$) like LiCoPO_4 (olivine) and $\text{LiNi}_{0.5}\text{Mn}_{1.5}\text{O}_4$ may also be attractive for some applications; however, the relatively low specific capacity of these cathodes (practically $< 140 \text{ mAh g}^{-1}$) weakens their importance.¹⁹⁻²¹ Since LIBs use graphite anodes (with a theoretical capacity of 372 mAh g^{-1}) it is possible to increase their energy density by moving to high capacity Si based anodes (theoretically $> 3000 \text{ mAh g}^{-1}$).²²⁻²⁵ However, Si based anodes suffer from intrinsic instability issues related to SEI formation.^{23,24} Nonetheless, since the limiting factor is the cathode side, changing the anodes in LIBs cannot drastically increase their energy density. The durability and cycle life of LIBs depends on complex

interfacial interactions between the electrodes and the electrolyte solution and development of passivation phenomena on the electrode surfaces.^{26,27} Thereby, much attention has been devoted to developing suitable electrolyte solutions for LIBs.^{28,29} The success in developing commercial LIBs which currently conquer the markets and the fear about a possible shortage in lithium if the electro-mobility revolution succeeds, promotes development of more battery technologies based on other active metals. An excellent example are the intensive efforts to develop Na-ion batteries in recent years.³⁰⁻³³ Hence, despite the constant progress in LIBs performance, and encouraged by their success, worldwide efforts are underway to find new types of even more promising redox couples.^{22,34,35} Relying on the benefits of metallic anodes, the choice of sulfur and oxygen as the cathode is based on their high specific capacity (1168 mAh g⁻¹ for Li₂O₂). Yet, the practical performance of batteries based on these redox couples lags significantly behind their theoretical promise. Nevertheless, metal-sulfur and metal-oxygen batteries may outperform all other possible choices from a cost perspective and have attracted global attention from both academia and industries. This is because of their 3-5 fold higher specific energy, as seen in **Figure 2**.^{22,36-38} The key determinant in the advancement of these batteries relies on combating parasitic reactions that take a toll in terms of efficiency, cyclability and safety.² There is much research activity ongoing in the field of Li-S batteries,^{39,40} but this topic is beyond the scope of this review, which focuses on the alluring lithium-oxygen (Li-O₂) batteries, based on their formal specific energy of 3500 Wh kg⁻¹. First, however, we introduce the reader to metal-air batteries in the next section.

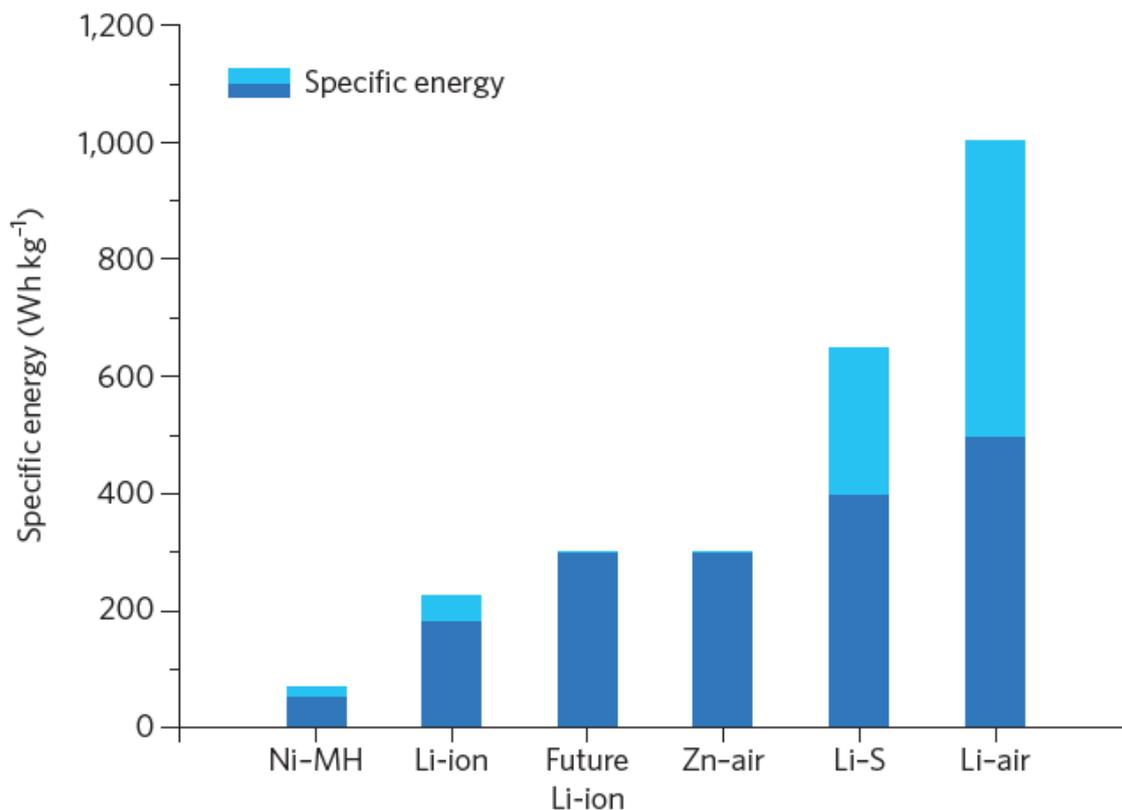


Figure 2. Bar graph demonstrating the practical specific energies in the battery world. These are reduced from the theoretical values by current collectors, cell packaging and system overheads. Light blue boxes indicate the range of specific energies. Reproduced with permission from ref³⁷. Copyright 2016 Springer Nature.

1.3. Metal-air batteries

Metal-air batteries^{36,41} promise a higher specific energy in comparison to the state of art LIBs and thus have arguably gained the most attention under the topic of “*post Li-ion batteries*”. Based on the pairing of a metal anode and an oxygen/air electrode (called as anode and cathode respectively in this paper), metal-air batteries present a hybrid architecture combining the features of both batteries and fuel cells. To date, several metal-air batteries using alkali metals (Li, Na and K), alkaline earth metal (Mg), and first row transition metals (Fe, Zn) or Al as the anode have been reported and their formal specific energies are presented in **Figure 3**. At this point it is important to note that, despite being the motivation for the research, there is lots of ambiguity with what energy metal-O₂ batteries could potentially store. The problem arises from confusing *formal capacity* (1168 mAh·g⁻¹, 2500 mAh·cm⁻³ Li₂O₂) with *theoretical*

capacity (Li_2O_2 including the minimum electron and ion conductor to allow the storage process $\text{O}_2 + 2 \text{e}^- + 2 \text{Li}^+ \leftrightarrow \text{Li}_2\text{O}_2$ to take place) [Ref 136, 185]. As broadly discussed in Section 3, Li_2O_2 can by far not be cycled in pure substance which cuts the theoretical well below formal capacity. Yet there is no theoretical barrier for the Li- O_2 cathode to achieve higher true capacity than LIB cathodes (see Section 3.4(c)).

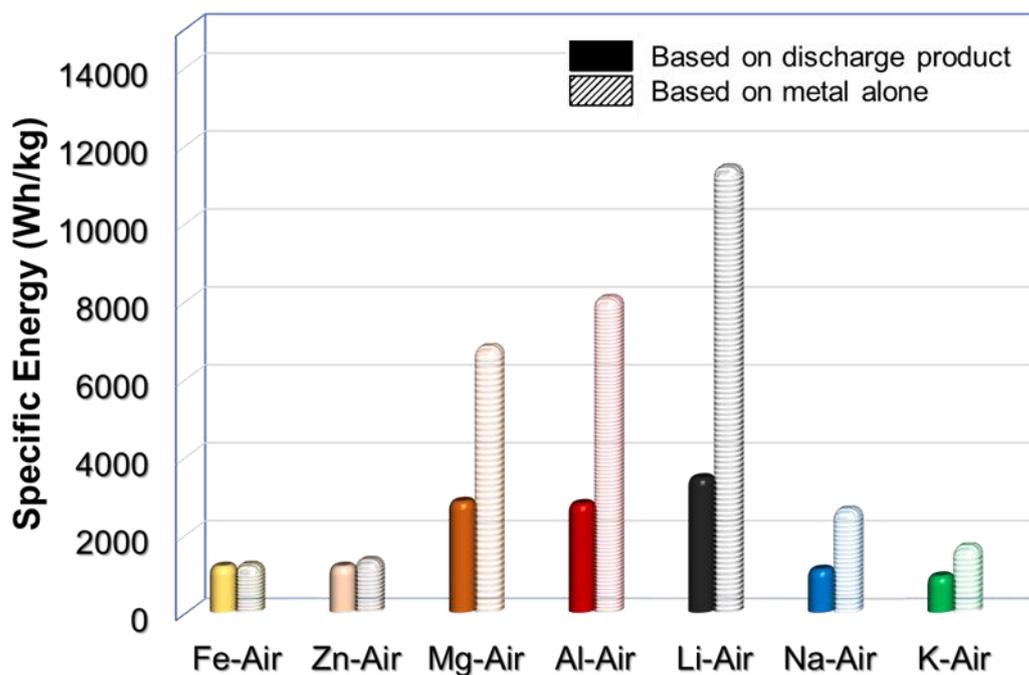


Figure 3. Specific energies of metal-air batteries. Solid colored bars represent specific energies densities based on the discharge product and one equivalent of the metal anode whereas patterned bars are based on the formal specific energies taking only metal into account. Reproduced with permission from ref⁴². Copyright 2017 American Chemical Society.

The non-aqueous systems have gained recent attention specifically due to the high formal specific energy of Li- O_2 batteries. However, aqueous electrolyte solution-based configurations have been researched since the early 1900's.⁴³ The commercialization of metal-air batteries commenced with the primary Zn-air battery in 1932.⁴⁴ Later, other aqueous systems like Al-air^{45,46} and Mg-air⁴⁷ came into existence in 1960s. Amongst the aqueous counterparts, the aluminium-air battery was considered to be an attractive candidate as a power source for electric vehicles (EVs) because of its high formal energy density (2800 Wh kg^{-1}).

Notably, in 2016, an Al-air battery weighing 100 kg was fabricated by Alcoa and Phinergy Co. Ltd. and shown to be capable of running a car on air for more than 1000 miles.⁴⁵ Unfortunately, most of the aqueous metal-air cells, including the Al-air battery, function as primary cells and can only be mechanically charged by replacing the used metal anodes and the electrolyte solution. Additionally, corrosion of metallic anodes in presence of water, high self-discharge rate, sluggish discharge kinetics, and short shelf life render these systems inappropriate for demanding applications like full EV vehicles.

On the other hand, non-aqueous metal air systems⁴⁸ are comparatively new but offer rechargeability with extremely high theoretical capacity. In 1996, Abraham et al.⁴⁹ reported the first rechargeable non-aqueous lithium-O₂ battery, which opened up a completely new field for researchers that led to development of other analogous systems such as Na-O₂ and K-O₂ batteries. Amongst these, the Li-O₂ battery is the most attractive since it offers the highest operational voltage (~2.96 V) with a formal specific energy of ~3500 Wh kg⁻¹ (based on formation of Li₂O₂ on discharge). The societal need for energy storage that exceeds Li-ion and the possibility that Li-air could deliver the step-change required has convinced the research community of its importance. Significant effort was devoted a decade ago to commercialize Li-air, which failed, but prompted research aimed at a deeper understanding of the poorly understood underpinning chemistry and electrochemistry, including the processes of degradation.

1.4. Scope of this review article

In the past, diverse aspects of metal-air batteries, in general and the most promising Li-O₂ batteries in particular, have been reviewed by various leading research groups.^{37,48,50-56} Detailed reviews with emphasis on the strategies towards a stable metal anode,⁵⁷⁻⁵⁹ methodologies for mechanistic studies,⁶⁰⁻⁶² effect of contaminants,⁶³⁻⁶⁵ stability of the electrolyte solutions⁶⁶ and the role of redox mediators on the performance of Li-O₂ batteries

^{54,67} have appeared. Even overviews questioning the practical reality of metal air batteries have been published recently.^{42,68,69} The intensive work in this area over the last decade has identified the key problems and guided directions to solve them. Thus, the preparation of a comprehensive report covering the challenges, possible solutions and recent developments associated with the progress of these batteries is called for. It is hoped that such a review will give the interested and educated reader a clear picture of the potential prospects of Li-O₂ and related batteries.

This review has four major goals:

(1) To clearly describe the current status of Li-O₂ batteries with respect to: new mechanistic insights; the stability and degradation of the electrolyte solutions and the electrodes in the presence of reactive oxygen species (ROS), singlet oxygen and other contaminants; and the catalytic ability of anchored and mobile catalysts (redox mediators) to address the problematic oxygen reduction and evolution reactions.

(2) To extend the knowledge gained by in-depth research in the field of Li-O₂ batteries to analogous systems like Na-O₂, K-O₂ and Mg-O₂ batteries which are still in their infancy.

(3) To introduce novel electrolyte solutions and electrode architectures which stabilize extremely reactive systems without compromising their promising energy density.

(4) To present a critical, yet realistic overview of the accumulated knowledge in this rapidly expanding field in order to overcome the gap between breakthroughs and commercialization. We have, therefore, tried avoiding redundant areas while carefully maintaining the comprehensive nature of the review to keep it relevant for readers of broad interest.

Section 2: Li-O₂ batteries: Diversity and Evolution

2.1. Configuration of Li-O₂ cells

The Li-O₂ battery - composed of a metallic lithium as the negative electrode, namely the anode, air/O₂ as the cathode active mass and a Li⁺ containing electrolyte solution - is one of the most studied class of metal-air batteries. The cathode in these systems is a composite electronically conducting porous matrix that enables the electrochemical contact between oxygen gas and Li ions in the electrolyte solution phase. It primarily functions with the dissolution/deposition of lithium metal at the anode and an oxygen reduction (ORR)/ oxygen evolution reaction (OER) at the cathode,⁵⁵ as schematically outlined in **Figure 4**.

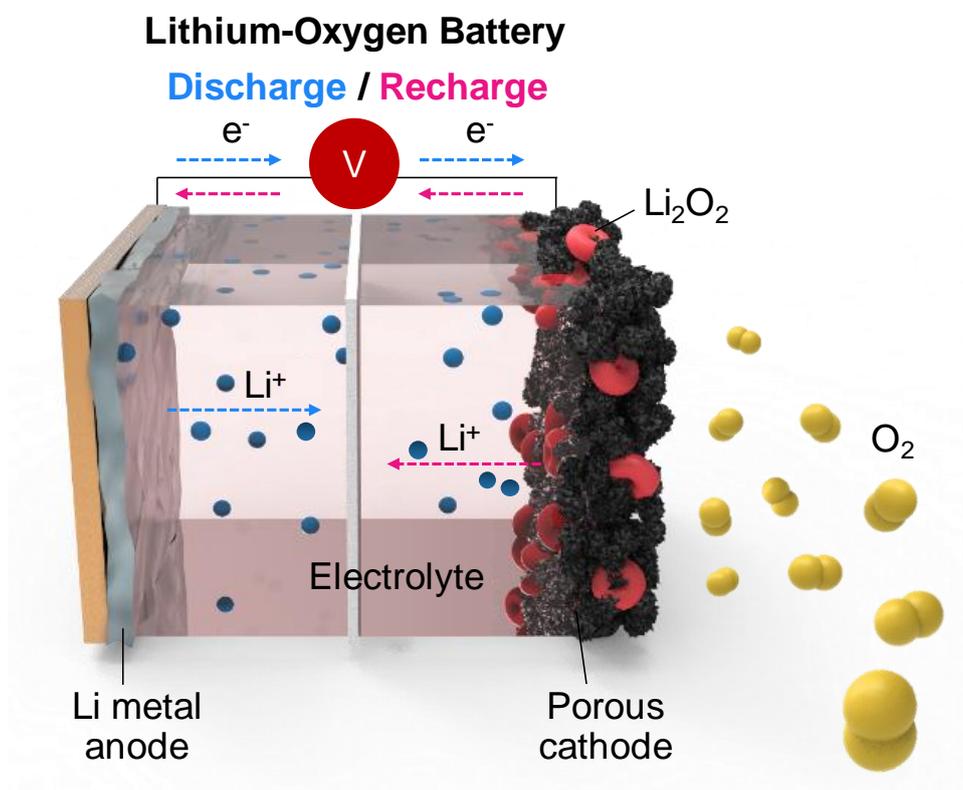
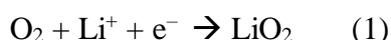


Figure 4. Schematic presentation of Li-O₂ battery.

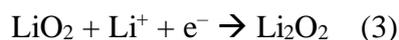
There are currently four categories of Li-O₂ systems under investigation, referred to as aprotic, aqueous, hybrid and solid-state batteries. The four types differ from each other with regard to the electrolyte species involved. The latter, in turn, determines the specific

electrochemical reactions during energy storage and release. A schematic illustration of these four types of batteries is provided in **Figure 5**.⁵⁶

The “aprotic LOB”^{49,70} consists of a lithium anode, a porous air cathode, and an electrolyte composed of a lithium salt dissolved in non-aqueous, organic solvent. Electrochemical reaction pathways for this type of LOB were proposed by Abraham et al.,⁷⁰⁻⁷² and not surprisingly, involve the reduction of molecular oxygen to lithium superoxide (LiO_2) via one electron transfer, as depicted in eq. 1:



This reduction is followed by either disproportionation (eq. 2) or receipt of a second electron (eq. 3) – both pathways forming lithium peroxide (Li_2O_2) as the main discharge product.



The resulting lithium peroxide is insoluble in the aprotic solvents⁷³ and thus deposits on the cathode either as an insulating thin film, or crystallizes in the shape of large toroidal precipitates. Consequently, the capacity of this battery is limited by the thickness of the Li_2O_2 film (~ 7 nm for this insulating solid) in case of an electrolyte solution that forms a Li_2O_2 film, and by the volume of the porosity of the cathode if the product is toroidal particles. During charge, the deposited Li_2O_2 undergoes oxidation at the positive electrode surface, regenerating molecular oxygen.

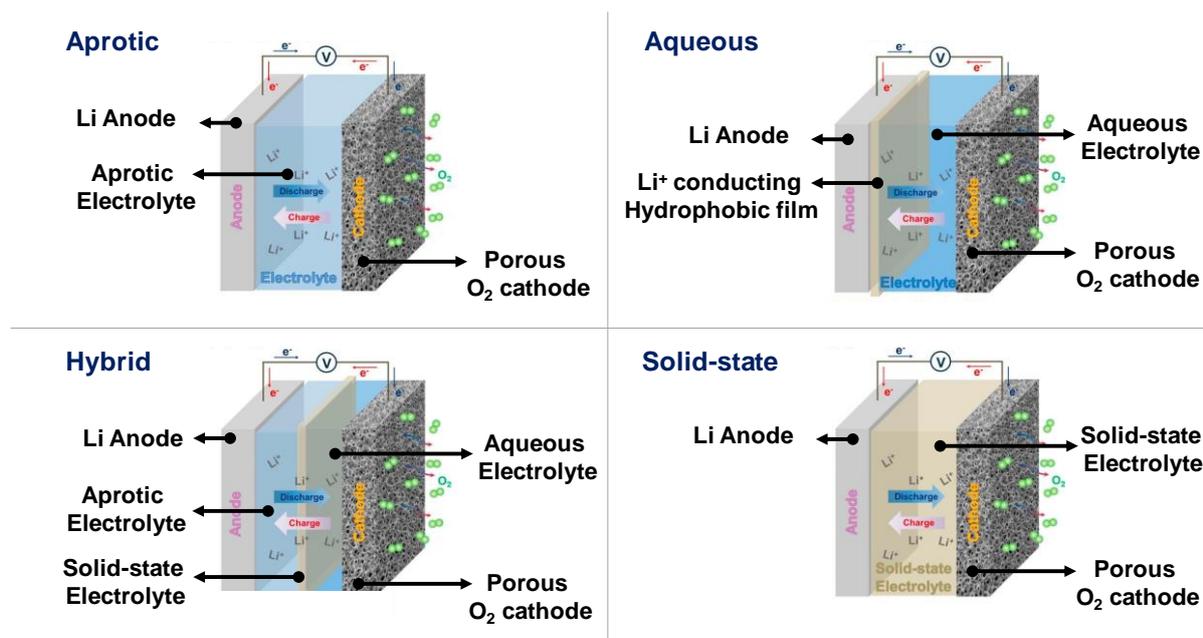


Figure 5. Schematic configuration of all types of Li-O₂ cells. Reproduced with permission from ref⁵⁶. Copyright 2017 Elsevier.

Researchers at the PolyPlus Battery Company eliminated the solid, insoluble discharge product of Aprotic LOB by replacing the non-aqueous solvent with an aqueous solution, and employing an Ohara glass Li-ion conducting membrane to protect the Li anode from corrosion.⁷⁴ In 2004, such a cell was reported by Visco et al (see **Figure 5**).^{50,75} The use of an aqueous electrolyte resulted in the formation of soluble lithium hydroxide (LiOH) as the discharge product, in contrast to solid Li₂O₂. This system has not yet become a commercial reality.

Another attempt to implement an aqueous system brought the “Hybrid LOB” into existence.^{73,76–78} In the proposed design, the anode side of the battery was filled with an aprotic solvent, whereas the air electrode was in contact with an aqueous electrolyte. The two electrolyte solutions were de-coupled by a solid electrolyte membrane. The hybrid LOB avoided direct contact between the lithium and the solid electrolyte membrane, and indeed resulted in improved lithium ion conductivity. However, the design leads to added complications and different diffusion kinetics in the two solutions.

“Solid-state LOBs” were also developed which do not use any liquid electrolytes, as shown in **Figure 5**. The development and application of solid-state lithium ion conducting materials for use in LOBs gained pace around 2010.⁷⁹ However, a truly all-solid-state LOB is challenged by how much Li_2O_2 can be stored in the cathode.

Since only the aprotic configuration of a LOB has shown a considerable rechargeability, it has attracted the largest effort worldwide. Consequently, in the following section, we will focus our discussion on this configuration.

2.2. The History and Early-Stage Research on Aprotic Li–O₂ Batteries

A Li and O₂ electrochemical couple for batteries was proposed in early 1970s; however, the first experimental observation was made in 1996 by Abraham et al.⁴⁹ – and that in a serendipitous manner.⁷⁰ During the in-situ investigation of the electrochemical intercalation of graphite in a Li/graphite cell with a polymer electrolyte, O₂ was accidentally injected while collecting the evolved gases for infrared spectroscopy. On resuming the experiment, a higher OCV and increased capacity was observed which was attributed to the formation of Li_2O_2 and other products as a result of reaction between Li and O₂. These unexpected conclusions led them to build the first Li-air cell and study this hybrid fuel cell. The field took off a decade later to unravel many electrochemical and mechanistic details. In 2002, Read et al.^{80–82} illustrated that the discharge capacity, rate performance and cyclability of Li–O₂ cells depends largely on the electrolyte solution and the air cathode formulation. In 2006, Ogasawara et al.,⁸³ demonstrated for the first time that Li_2O_2 could be oxidized electrochemically, using mass spectrometry to prove O₂ evolution. Many of these early pioneering studies used carbonate electrolytes and manganese dioxide as the electrocatalyst, which have since been shown to be unstable.^{84,85} Overall, it can be said unequivocally that by 2008, the pioneering work of the research groups of Abraham, Read and Bruce laid the foundation which attracted immense attention to Li–O₂ systems.

Section 3: Current Status

3.1. A short introduction on important aspects of non-aqueous electrochemistry

As already noted, the use of Li metal anodes in Li-O₂ batteries is mandatory to achieve high energy density. This dictates the use of cell components that can enable safe and controllable operation of anodes comprising this reactive alkali metal. It is very likely that a Li anode protected by a solid electrolyte membrane will be required as the criteria for a liquid electrolyte in direct contact with Li metal is a major challenge. We can summarize a few of the criteria that are relevant to any chosen system:

- 1) The polarity of the solvent should be sufficient to enable facile dissolution of salts.
- 2) The overall physical properties (polarity, viscosity) should enable high ionic conductivity.
- 3) The electrochemical window available must be wide enough to allow the full operation of the battery.
- 4) The systems must be stable against side reactions (e.g. via development of effective passivation phenomena).
- 5) The components must be of high purity, due to possible undesirable side reactions with contaminants such as atmospheric species, acidic moieties etc.
- 6) The electrolyte should exhibit appropriate safety features (e.g. low volatility, flammability, and opportunity for thermal run-away routes).

While many possible combinations of aprotic solvents and lithium salts answer the first and second criteria, only a few candidates at best can fit criteria 3 - 6. In fact, all electrolyte solutions are only metastable toward the electroactive species and working potentials of standard lithium-oxygen batteries. The empirical physical and chemical properties of various aprotic solvents and lithium salts that affect (or even determine) the mechanisms of oxygen electrochemistry in them, became more and more understood, thanks to systematic studies

during the last decade. We note, however, that non-aqueous electrolyte solutions that are truly suitable alone for controllable Li-O₂ electrochemistry have yet to be found. Nevertheless, the insights and fundamental knowledge gleaned over the last decade attests to the great progress made in understanding the limitations of many types of electrolytes examined to date in Li-O₂ cells and hence, inform the properties required from suitable systems.

3.2. Mechanistic aspects of non-aqueous oxygen electrochemistry:

3.2(a) Li₂O₂ formation on discharge

Studies of oxygen electrochemistry in non-aqueous electrolyte solutions took place much before the appearance of the Li-O₂ battery concept.^{86,87} These studies were conducted with the aim of understanding the behavior of reactive oxygen species (ROS) in the absence of proton donors such as water. It was found (even by simple voltammetric measurements) that oxygen electrochemistry is very dependent on the nature of polar aprotic solvent and the electrolyte salt (especially the identity of the cation). In some cases, oxygen exhibited reversible redox behavior, while in others the response was more complex (irreversible in many cases). Therefore, to properly describe the oxygen reduction reaction (ORR) and oxygen evolution reaction (OER), one must consider the media in which the reaction is taking place.⁸⁸

The important parameters that govern the mechanism of O₂ reduction to Li₂O₂ are:

- 1) The donor number (DN) of the solvent and salt. These are critically important as they determine the solubility and dissociation of LiO₂ and hence whether Li₂O₂ grows on the electrode (low DN) or in solution (high DN).
- 2) The Lewis acidity of the solvated lithium cation.
- 3) Contaminants and additives in the electrolyte solution.
- 4) The stability of the solvent toward the reduced oxygen species generated and toward possible formation of very a reactive molecular oxygen state (as discussed further below).

The above parameters can have a synergetic effect; therefore, the second and third two parameters must be described together. The stability of various aprotic solvents is described in the next section of this review.

The first step in O_2 reduction is the one electron reduction to form the superoxide radical (O_2^-) as confirmed by early experiments in nonaqueous solutions with solvents such as DMSO.^{86,87} The subsequent fate of the superoxide is dependent on the nature of the counter cation. With large cations such as K^+ and tetrabutylammonium (TBA^+) the metal peroxides are less stable compared to Na_2O_2 and Li_2O_2 (the latter being the most stabilized peroxide). The level of stability of the metal peroxide affects the lifetime of superoxide intermediates in these systems. Hence, with cations like TBA^+ or Na^+ it is possible to observe reversible reduction of oxygen to superoxide moieties, while with the lithium cation such a reversible behaviour is not seen. Using methods such as Raman and UV spectroscopy coupled with electrochemical measurements, it was demonstrated that the superoxide does form in lithium containing solutions (Eq. 1 above).^{89,90} However, this unstable radical quickly disproportionates to form lithium peroxide, Li_2O_2 (Eq. 2 above). While the superoxide is somewhat soluble and can diffuse in the electrolyte solution, the peroxide moieties are insoluble in aprotic solvents (if their DN are not high enough).^{72,71}

The mechanisms of O_2 reduction in different aprotic solvents that vary in their donor numbers were extensively studied in recent years. The mechanism of O_2 reduction to form Li_2O_2 is summarised in Figure 6. In all solvents of high donor number (DN), following reduction of O_2 to O_2^- at the electrode surface, the O_2^- is present in solution due to the strong solvation of the Li^+ ion by coordination by the solvent molecules. In other words, in a high DN solvent, LiO_2 does not exist and is instead present as solvated ions. However, as Li_2O_2 is more stable than LiO_2 , disproportionation occurs for LiO_2 to grow relatively large toroidal Li_2O_2 particles directly from solution. In contrast, when the DN is sufficiently low, Li^+ is not

strongly solvated and LiO_2 (ion coupled electron transfer) forms on the electrode surface. Either rapid disproportionation or a second electron transfer forms Li_2O_2 that grows as a film on the electrode. This understanding of the role of the DN in determining whether LiO_2 will dissolve in solution or will deposit on the electrode surface following the first electron transfer, was a milestone in understanding the Li- O_2 cell. It showed that high DN electrolyte solutions will promote growth of Li_2O_2 from solution, thus forming large toroidal particles in the pores of the electrode that do not passivate the electrode. This results in relatively high rates and capacities of ORR in high DN solvents, whereas, low DN electrolyte solutions, like ethers, lead to Li_2O_2 film growth and consequently low rates and capacities.

Given the important role of LiO_2 solvation/stabilization and the importance of the interactions of the solvent molecules with the Li ions, the effective DN and AN of the solutions determines whether Li_2O_2 forms via a solution or surface mediated pathways. High DN and AN additives were shown to favour solution pathways. Considering first the DN, the counter-anion can strongly tune the DN of the electrolyte solution. The degree in which the anion can interact with the solvated Li cation can be quantified by the association strength of the salt. Strongly dissociated salts such as LiTFSI form solvated complexes where the Li cation hardly interacts with the TFSI anion. In weakly dissociated or associated salts such as LiTf and LiNO_3 , the negatively charged functional groups of the anions interact with the Li cation, forming complexes such as $\text{NO}_3^- \cdots \text{Li}^+ \cdots \text{Solvent}$, are strongly solvated. The stabilization of the Li cations by the anion will promote LiO_2 in solution. We note that the effect of the anion was shown to be relevant only for lower DN solvents, such as polyethers.

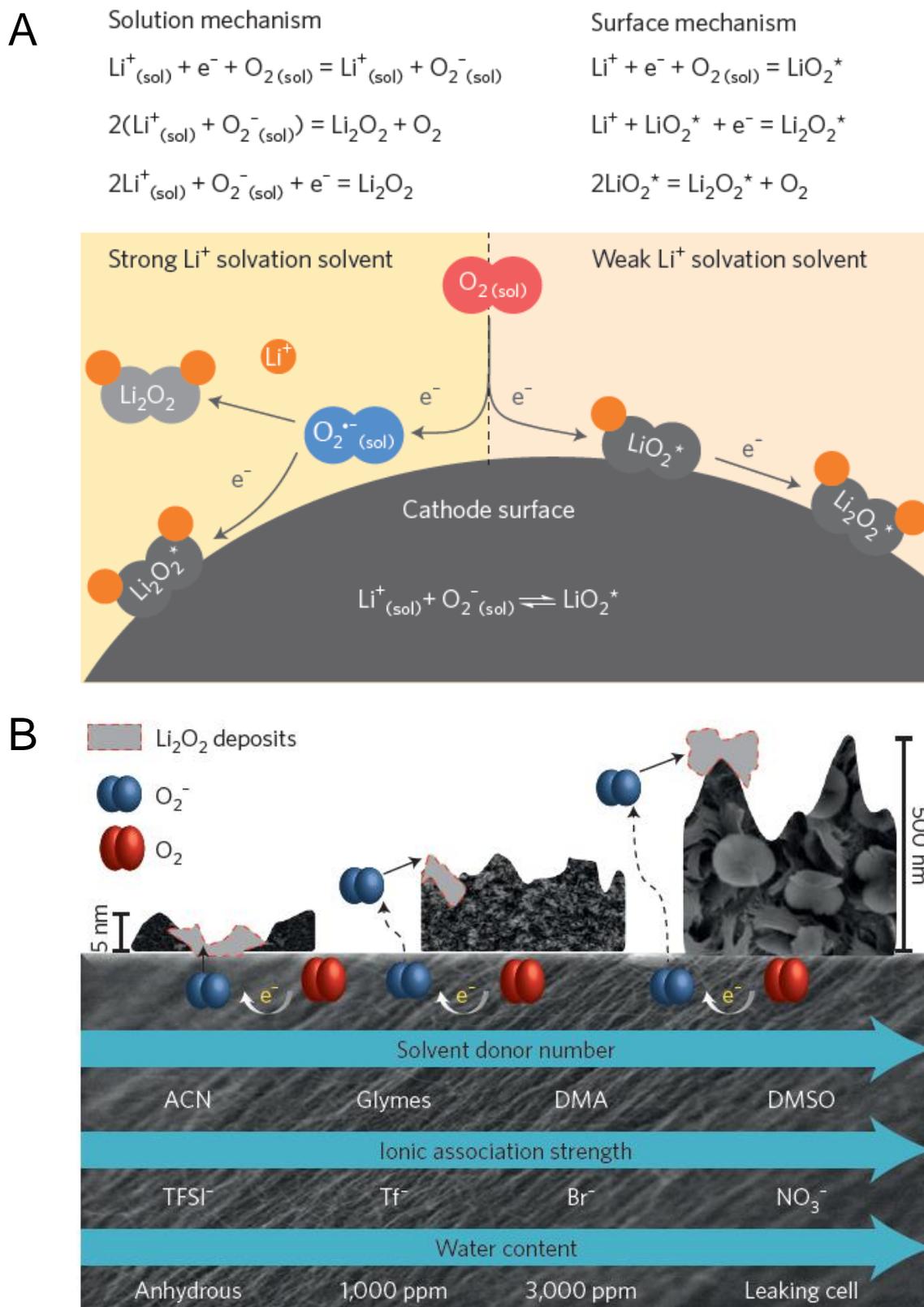


Figure 6. (A) Oxygen reduction mechanistic pathways in non-aqueous Li-O₂ batteries. Reproduced with permission from ref⁹⁰. Copyright 2014 Springer Nature. (B) The influence of solvents, counter anions and water content on the Li₂O₂ growth mechanism during ORR. Reproduced with permission from ref⁹⁷. Copyright 2016 Springer Nature.

In the solution mediated mechanism, after the initial nucleation of Li_2O_2 seeds, the growth of Li_2O_2 particles can proceed from solution and therefore leave behind an active cathode interface. It is the avoidance of a passivating surface film of Li_2O_2 that leads to high rates and high capacities in high DN electrolyte solutions.

High AN additives, particularly protic ones, were further found to favour large discharge capacities with big toroidal Li_2O_2 particles. The most common protic additive/contaminant is water. The reaction of water (if in large concentration) with a superoxide species can lead to the formation of insoluble LiOH deposits on the cathode interface. However, it was demonstrated that with small to medium amounts of water, the main discharge product can still be Li_2O_2 . By the addition of 0.01 to 1% water into a polyether-based electrolyte solution the discharge capacity increased by more than 25-fold. In addition, the yield and size of Li_2O_2 particles increased with added water.⁹⁴ However, Aetukuri et al. reported that discharge capacity increased 4-fold with 4000 ppm of water with increasing size of Li_2O_2 particles but the number (yield) of Li_2O_2 particles decreased.⁹⁴

The mechanism by which water induces the formation of large toroidal shaped particles is still unclear. One proposed explanation is that protic molecules increase the effective AN of the electrolyte and improve solvation of superoxide much more efficiently than aprotic solvents. This can lead to Li_2O_2 growth by a solution-based mechanism. Alternatively, protonation of O_2^- to form HO_2 can occur, which diffuses from the electrode surface and forms Li_2O_2 in solution.⁹⁵⁻⁹⁷ The presence of water can also influence the oxidation of Li_2O_2 during the OER. Li et al. suggested that the addition of water to a $\text{Li}-\text{O}_2$ cell can significantly reduce the OER overpotential.⁹⁸ They found that some Li_2O_2 was converted to LiOH during the cathode discharge (Figure 7A). However, it is not clear that LiOH can be oxidized to molecular oxygen, even catalytically, as suggested in the mechanism of Figure 7B. In order to avoid the consumption of water by Li_2O_2 , MnO_2 was introduced to regenerate water by catalysing the

reduction of H_2O_2 . However, as described below, these catalysts are known to induce decomposition reactions in the battery. A similar regenerative process was described in Li- O_2 systems that include water impurities and LiI salt.^{99–101} Therefore, the influence of water in Li- O_2 cells should be examined very carefully to avoid any misperceptions that may rise from the chemical complexity of this kind of system.

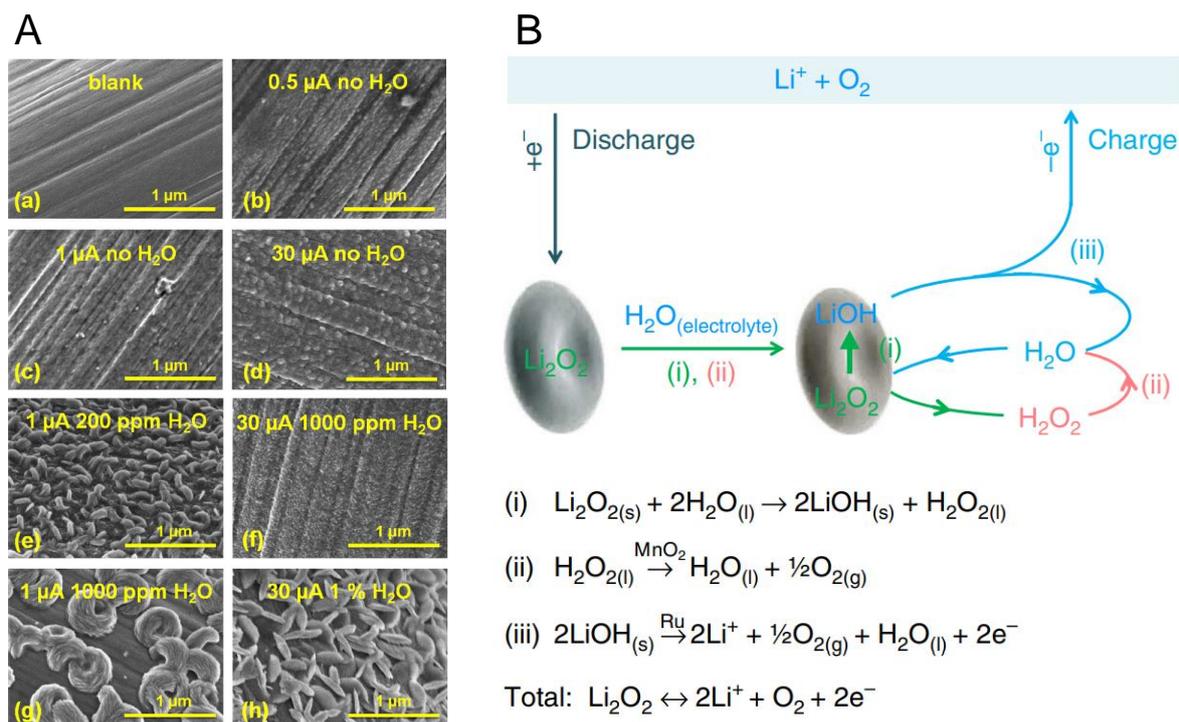


Figure 7. (A) SEM images of the carbon cathodes of Li- O_2 cell after discharged in the presence of water at different concentrations. Reproduced with permission from ref⁹⁶. Copyright 2015 The Electrochemical Society. (B) the ORR and OER mechanistic scheme in a non-aqueous Li- O_2 cell in the presence of added water. Reproduced with permission from ref⁹⁸. Copyright 2015 Springer Nature.

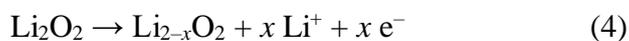
3.2(b) Li_2O_2 oxidation mechanism on charge

In contrast to the steady discharge voltage plateau during the formation of Li_2O_2 , the oxidation curve of Li_2O_2 is seemingly comprised of multiple voltage steps. A fundamental hysteresis exists because while disproportionation of LiO_2 to form Li_2O_2 dominates on discharge, the reverse process (comproportionation) cannot occur on charge. Charge usually initiates around 3.2 V and can continue up to 4.5 V in many Li- O_2 systems. In an early study,

it was suggested that the different slopes in the OER voltage curve might be associated first with the oxidation of surface Li_2O_2 , and subsequently with bulk Li_2O_2 . Using operando X-ray diffraction, Ganapathy et al. found that ORR forms both crystalline and amorphous Li_2O_2 products.¹⁰² They proposed that initially the amorphous phase starts to oxidize at low overpotentials, while the toroidal Li_2O_2 crystalline aggregates, depending on their size, begin to decompose at much higher overpotentials. The enhanced charge transport of the amorphous phase was later confirmed by charging very pure chemically synthesized amorphous Li_2O_2 .¹⁰³ It was suggested that the higher mobility of Li ions, and increased superoxide concentration within the amorphous Li_2O_2 , lead to relatively lower overpotentials. Enhancement in the local Li ion mobility was also observed in mechanically induced defects of nanocrystalline Li_2O_2 .¹⁰⁴ In practice, the OER overpotential required for complete removal of the Li_2O_2 layer is more than 1 V. During investigation of OER mechanisms it became clear that during oxidation of Li_2O_2 , different parasitic reactions take place with the electrolyte and carbon to form organic decomposition products and Li_2CO_3 . Potentials greater than 4 V vs Li/Li^+ are needed to remove those species.¹⁰⁵ OEMS studies showed that the oxidation reaction required more than $2e^-$ per O_2 , which implies that some ROS to react with cell components in an irreversible manner. A very important recent result from several groups is the recognition that the instability of the electrolyte and carbon is associated with the formation of singlet oxygen ($^1\text{O}_2$) during the OER.^{106–109} By operando EPR spectroscopy, Wandt et al. suggested that at potentials of around 3.5 V, approximately 0.5 % of highly reactive $^1\text{O}_2$ is evolved.¹⁰⁷ Subsequently, $^1\text{O}_2$ has also been shown to evolve both on discharge and from the onset of charge.¹⁰⁸ Its generation is intimately tied to the ORR and OER through superoxide disproportionation steps which generates $^1\text{O}_2$.¹⁰⁹ $^1\text{O}_2$ is the key to the degradation mechanism of Li- O_2 batteries, and thus it is discussed in detail in the following section.

Recent studies visualized the decomposition of Li_2O_2 deposits by in situ AFM and vibrational spectroscopy of sum frequency generation (SFG) measurement during the OER.^{110,111,112} They concluded that at relatively low potentials (< 3.7 V), a thin film of Li_2O_2 which is in direct contact with the cathode surface is oxidized to produce superoxide radicals. As the surface layer oxidizes, the upper Li_2O_2 layer is detached from the cathode surface.¹¹² This disconnected product layer was found to be oxidized only at potentials above 4.5 V, which is above the oxidation potentials of most aprotic solvents - thus leading to instability. In general, the mechanisms described above clearly implies that properties such as size, morphology and crystallinity of the Li_2O_2 deposited during the ORR should govern the efficiency of the OER and are controlled by the choice of electrolyte solutions.

Recently, Lu et al. introduced a comprehensive unified model that shows the solvent DN, in a similar fashion to discharge, as the governing factor for the oxidation mechanism (Figure 8).¹¹³ Using a thin film RRDE (rotating ring-disc electrode), they could show that in high DN solvents dissolved $\text{LiO}_2(\text{sol})$ is formed, while it was not detected in low DN solvents, Figure 8b. XANES showed the formation of LiO_2 at the surface upon charging in high DN solvents but its absence in low DN solvents. As for discharge, the solvent influences the Li_2O_2 decomposition pathway, via either soluble superoxide species or Li-deficient Li_2O_2 .^{90,114,115} The unified reaction mechanism is illustrated in Figure 8c and commences with delithiation:

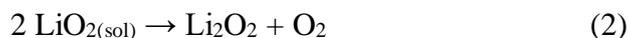


In low DN solvents $\text{Li}_{2-x}\text{O}_2$ was suggested to undergo a second oxidation:



It should be noted that disproportionation of solid superoxide species cannot be strictly excluded, however, it is hard to suggest a plausible mechanism for such a solid state reaction

¹¹⁶ In high DN solvents, they have shown $\text{Li}_{2-x}\text{O}_2$ to dissolve to $\text{LiO}_{2(\text{sol})}$ which then releases O_2 by disproportionation as depicted in [equation 2](#).



This latter step in high DN solvents was elegantly proven by SEM imaging when large particles of chemically synthesized Li_2O_2 were partly oxidized ([Figure 8a](#)). For the high DN solvents DMSO and methyl-imidazole (Me-Im), the Li_2O_2 was converted into a lamellar morphology akin to what is known from solution discharge. Whether disproportionation is the O_2 releasing step has a major impact on parasitic chemistry via its impact on $^1\text{O}_2$ formation as further discussed in the following Section.

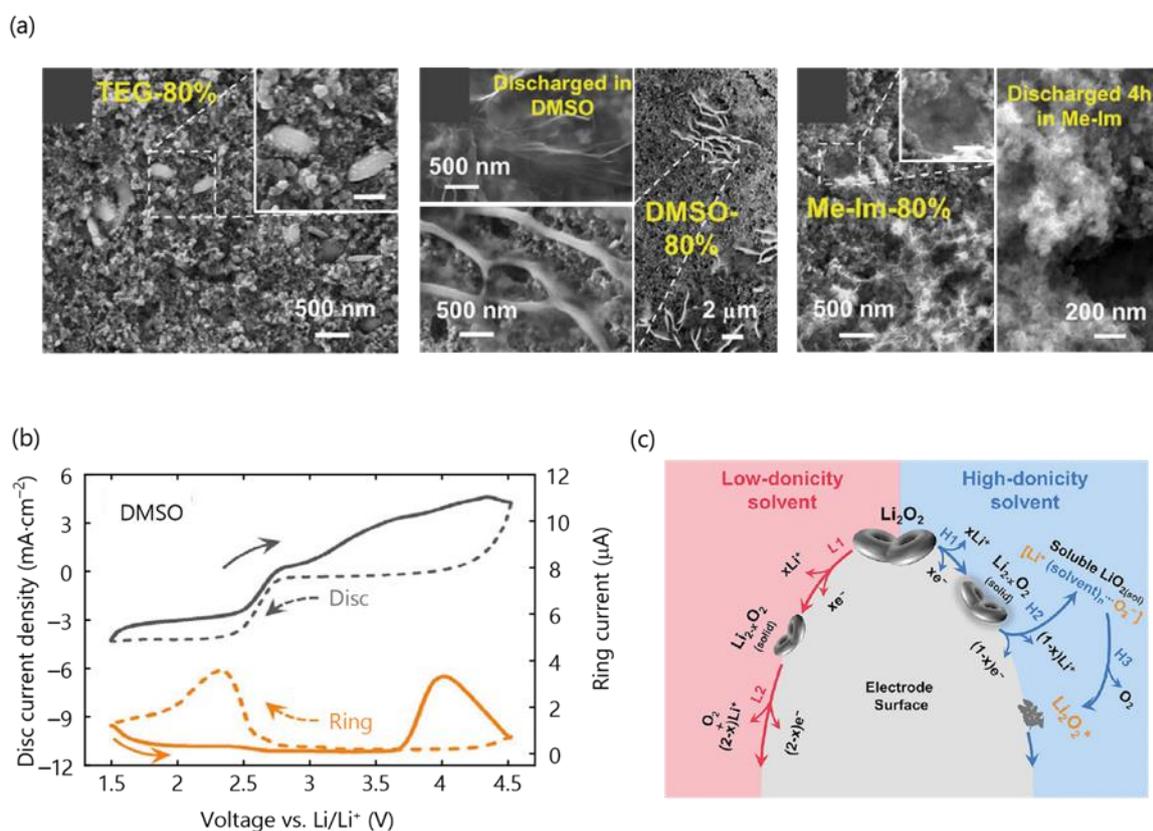
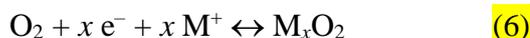


Figure 8. The solvent controlled charge mechanism of Li_2O_2 . (a) SEM images of Li_2O_2 -packed electrodes after 80% recharge in different solvents with 1 M LiTFSI. The insert shows recrystallization of Li_2O_2 due to disproportionation in DMSO and Me-Im. (b) RRDE measurement of oxygen reduction/evolution in O_2 saturated DMSO with 0.2 M LiTFSI. The ring current shows the superoxide reduction during both discharge and charge and hence presence of soluble superoxide. (c) Scheme of the charge mechanism with a surface or solution

mechanism in low or high DN solvents, respectively. Reproduced from Ref.¹¹³ with permission from Elsevier.

3.3. Parasitic chemistry

Parasitic chemistry at the cathode is the prime obstacle for reversible operation of Li-O₂ cells and hence understanding the mechanisms underpinning parasitic reactions is the most pressing research need in the field. Current understanding has resulted in a unified mechanism of parasitic chemistry across the alkali metal-O₂ cells (M = Li, Na, K) and will hence be treated together.¹⁰⁹ Metal-O₂ chemistries typically consume close to the ideal number of electrons per O₂ upon discharge according to:



The products that typically form are Li₂O₂, Na₂O₂, NaO₂, or KO₂.^{37,117–120} However, for Li-O₂ cells, significant amounts of side products such as Li₂CO₃, Li formate and Li acetate were found and the Li₂O₂ yield was significantly lower than expected from the charge passed.^{84,121–123} Similarly, on charge the amount of O₂ evolved falls below the expected amount and more side products form. Na-O₂ cells have been recognized to be more reversible if superoxide (NaO₂) is the product, but things are much worse with Na₂O₂ as the product.¹²⁴ K-O₂ cells perform best in all cases.¹²⁰ (Su)peroxide yields on discharge and O₂ yields on recharge (based on charge passed) from a wide range of studies that give quantitative values are summarized in **Table 1**. The general trend is that irreversibility becomes worse when moving from K to Na and Li, which goes along with the peroxide rather than the superoxide being the product.

Table 1 Typical M_xO₂ yields on discharge and O₂ yields on recharge of M-O₂ cells (M = Li, Na, K). Data are compiled from Refs.^{96,109,118,120,123,125–133} Definitions of reversibility and means to quantify them have been summarized in detail in Refs.^{134,135}

Product	Li-O ₂	Na-O ₂	K-O ₂
Superoxide		90 - 95%	99 - 100 %
Peroxide	50 - 93%	< 50%	

The parasitic reactions have traditionally been ascribed to the direct reactivity of electrolyte or carbon with superoxides and peroxides.^{37,118,123,127,128,130,132,133} However, the extent of side reactions would suggest the reactivity to seemingly severely grow in the order $\text{KO}_2 < \text{NaO}_2 < \text{LiO}_2$, and peroxides to be much more reactive than superoxide. In the following section we summarize the current knowledge about these reactivities and the recognition that the highly reactive singlet oxygen ($^1\text{O}_2$) forms in metal- O_2 cells, which conclusively explains the pattern of irreversibilities and guides pathways to counteract the problem.

3.3(a) Reactivity of reduced O_2 species and O_2

Reduced reactive oxygen species (RROS) have for long been known for their reactivity with many organic substrates owing to their nucleophilicity, basicity and radical nature and are recognized as a reactant or source of parasitic chemistries^{37,136,137}. The primary species are O_2^- , O_2^{2-} , $\text{HOO}\cdot$, HOO^- , and $\text{HO}\cdot$ that may react via nucleophilic substitutions, H^+ and H-atom abstraction. Superoxide is furthermore an electron transfer agent and molecular oxygen may equally react via H-atom abstraction in auto-reactions.

Salt solubility in non-aqueous media for use as electrolyte is introduced via polarity by means of heteroatoms, which however makes the solvents susceptible to the mentioned reactions, particularly at C and H atoms adjacent to the heteroatom. Along this line, nucleophilic attack has for long been considered the primary source of side reactions in metal- O_2 cells, which is certainly true for carbonates.^{84,137,138} Since also all so far investigated classes of organic solvents have shown a certain degree of decomposition, superoxide's reactivity has equally been made responsible for the reactions. Bryantsev et al.¹³⁸⁻¹⁴² pioneered DFT calculations to understand the reactivity of electrolyte solvents with reduced O_2 species and O_2 with the aim to possibly predict more stable ones. **Table 2** summarizes their data and those of several other groups for activation energies for reactions of superoxide, peroxide, and

molecular oxygen with various classes of organic solvents via nucleophilic substitution, H-atom and H⁺ abstraction.

Table 2 Reactions of organic electrolytes with reduced oxygen species and molecular oxygen and their calculated activation energy barriers. ROR' is generically used for organic moieties with polarizing heteroatoms and reactions may accordingly be translated to, e.g., N or S containing solvents. Adapted with permission from Ref. ¹³⁵ - Published by The Royal Society of Chemistry.

Reactant	Type of reaction	Reaction	E^{act} (kJ·mol ⁻¹)	References
O ₂ ⁻	nucleophilic substitution	ROR' + O ₂ ⁻ → RO ⁻ + ROO• (7)	121–144 ¹ , 105 ² , 65–95 ³	140,141,143,144
	H-atom abstraction	RH + O ₂ ⁻ → R• + HOO ⁻ (8)	129–180 ⁴ , 191 ⁵	140,143,145,146
	H ⁺ abstraction	RH + O ₂ ⁻ → R ⁻ + HOO• (9)	pKa > 30 stable ⁶	139,147
Li ₂ O ₂	nucleophilic substitution	ROR' + Li ₂ O ₂ → RO ⁻ Li ⁺ + R'OO ⁻ Li ⁺ (10)	134–192 ¹	148
	H-atom abstraction	RH + Li ₂ O ₂ → R• + [Li ₂ O ₂ - H•] (11)	96–112 ¹	148
	H ⁺ abstraction	RH + Li ₂ O ₂ → R ⁻ Li ⁺ + HOO ⁻ Li ⁺ (12)	116–311 ¹	144,148
O ₂	H-atom abstraction	RH + O ₂ → R• + HOO• (13)	163–183 ⁷ , 138–161 ⁸	140

A striking feature is that for all considered reactions the activation energies are too high to explain major reactivity. Solvent stability screening experiments with KO₂ as the superoxide source have shown that reactions with activation energies >100 kJ·mol⁻¹ can be considered not to contribute noticeably.¹⁴¹ All pathways with ethers, for example, require high activation energy and are strongly endothermic. Only one study found H-abstraction by Li₂O₂ clusters to be slightly exothermic with $E^{\text{act}} < 100$ kJ·mol⁻¹.¹⁴⁸ Generally, Li⁺ coordinated solvents were reported to be more stable against H-abstraction by O₂⁻ and O₂ than the free molecules.^{140,146}

As reviewed in **Section 3.2**, protic additives have recently been proposed to have a certain positive effect on discharge capacity despite reports of increased parasitic chemistry.^{94,96} However, in presence of proton sources HOO•, HOO⁻, and HO• may form, which are potentially more reactive than the non-protonated species¹³⁹. The HOO⁻ species may more easily abstract H⁺ as it is a stronger base than superoxide. HO• readily abstracts H-atoms

and forms C-centred radicals which initiate favourable chain reactions in the presence of O₂ akin to combustion reactions^{84,140}. Nevertheless, the described reactivity of reduced reactive oxygen species including the protonated species cannot consistently explain the observed pattern of parasitic chemistry shown in **Table 1**. For example, much stronger side reactions on charge compared to discharge directly oppose superoxide's reactivity as an explanation. Very reversible cyclability of KO₂ in DMSO for thousands of cycles further forcefully demonstrates that the direct reactivity of reduced O₂ species with cell components is not the prime source of side reactions¹²⁰.

3.3(b) Singlet oxygen in metal-O₂ cells

Evidence for singlet oxygen as the main culprit for parasitic chemistry

Chemical oxidation of peroxides is recognized as a source of the highly reactive singlet oxygen (¹O₂)¹⁴⁹. Similarly, electrochemical oxidation of Li₂O₂ was hypothesized in 2011 to possibly evolve singlet oxygen in Li-O₂ cells.¹⁵⁰ Singlet oxygen is the first excited state of triplet ground state oxygen and has an energetic difference of ~1 eV. Considering this energy difference, charging of Li₂O₂ beyond 3.5 to 3.9 V vs Li/Li⁺ was hypothesized to potentially form ¹O₂.^{106,131,150} Furthermore, disproportionation of HOO• has been considered as a source of ¹O₂ although it is now recognized to negligibly form it.^{109,151–153} The difficulty to detect ¹O₂ was the reason that this idea was only discussed in a few reports but not proven except for a seminal study by Wandt et al.¹⁰⁶ who used operando EPR to prove a small degree of ¹O₂ evolution upon charging between 3.55 and 3.75 V. The process was explained via direct 2 e⁻ oxidation according to $\text{Li}_2\text{O}_2 \rightarrow \text{O}_2 + 2 \text{Li}^+ + 2 \text{e}^-$ ^{92,106,153}. However, measurements during discharge and higher charge voltages were not possible with the used spin trap. ¹O₂ forming above ~3.5 V helped explain parasitic chemistry at high charge potentials. However, it left parasitic chemistry on discharge unexplained and particularly also why significant amounts of

side products form at the onset of charge, where always less than 1 mol O₂ evolved per 1 mole of Li₂O₂ oxidized.^{105,123,154}

To investigate involvement of ¹O₂ more comprehensively, Freunberger et al. developed a set of methods to sensitively and quantitatively detect ¹O₂ over the entire relevant voltage range during discharge and charge of metal-O₂ cells.¹⁰⁸ As direct unambiguous proof they detected the characteristic 1270 nm emission when ¹O₂ transits from the singlet to the triplet state. The process has a low quantum yield since this phosphorescent radiative decay is forbidden and is strongly influenced by the environment and is hence insensitive. Nevertheless, in-situ measurements have shown unequivocal proof for ¹O₂ from the onset of charge, **Figure 9a**¹⁰⁸. As a more sensitive and quantifiable method, 9,10-dimethylanthracene (DMA) was identified as a suitable ¹O₂ trap, which fulfils all requirements for the cell environment. DMA forms selectively and rapidly its endoperoxide form (DMA-O₂) in the presence of ¹O₂; both DMA and DMA-O₂ are electrochemically stable in the relevant voltage range between 2 and ~4 V vs Li/Li⁺; and importantly DMA does not form DMA-O₂ in contact with superoxide and peroxide alone, which otherwise could falsely indicate ¹O₂ formation. Further, DMA-O₂ forms only with ¹O₂ but not with other possibly reactive O-containing species like Li₂CO₃, O₂, or CO₂. DMA-to-DMA-O₂ can be probed in various ways such as ex-situ HPLC of electrolyte extracted from cells or in-situ fluorescence of DMA (**Figure 9b**).

These studies proved that ¹O₂ forms both during discharge and from the onset of charge and with growing rate as the charge voltage rises, **Figure 9b**. The rates of ¹O₂ formation resemble the rates at which parasitic reactions occur in Li-O₂ cells.¹⁵⁵ Further, it was proven that ¹O₂ accounts for the majority of parasitic reaction products on discharge and charge. Thus ¹O₂ arises as the biggest obstacle for cycling of the Li-O₂ cell by reversible formation/decomposition of Li₂O₂. The first report suggested up two possible means to counter ¹O₂-related side reactions. First, the ¹O₂ probe DMA, which acts by trapping ¹O₂, was shown

to reduce parasitic products during both discharge and charge if large concentrations were used. However, at the rate of $^1\text{O}_2$ formation, DMA is consumed within a few cycles. Second, they introduced the $^1\text{O}_2$ quencher 1,4-diazabicyclo[2.2.2]octane (DABCO) into Li- O_2 cells and have shown that it even more strongly reduced parasitic chemistry by physically deactivating $^1\text{O}_2$ to $^3\text{O}_2$ without consumption of the quencher itself. DABCO is, however, not electrochemically stable above 3.6 V and hence does not allow for full recharge, which requires more oxidation-stable quenchers as described below. Singlet oxygen was also shown to form on discharge and charge in Na- O_2 cells where NaO_2 forms as the discharge product.¹⁰⁷

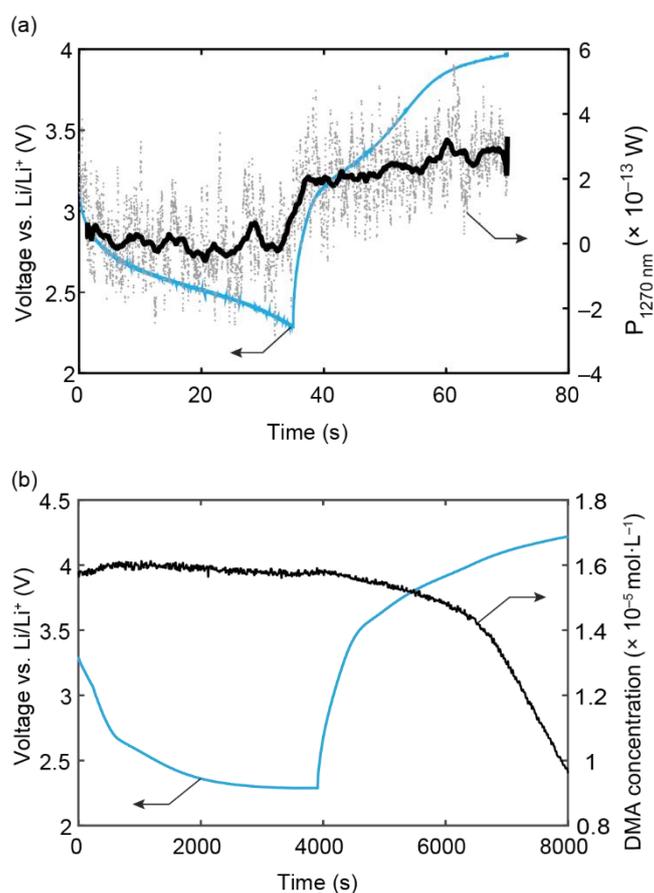


Figure 9 Singlet oxygen formation in Li- O_2 cells. (a) Operando detection of $^1\text{O}_2$ by detecting the near infra red (NIR) emission at 1270 nm during galvanostatic cycling of a Au grid electrode in O_2 saturated deuterated acetonitrile containing 0.1 M LiClO_4 and 1000 ppm D_2O . The black trace is the moving average of the raw data in grey. (b) Voltage profile and DMA concentration during galvanostatic cycling of a carbon black electrode at $25 \mu\text{A}\cdot\text{cm}^{-2}$ in O_2 saturated TEGDME containing 0.1 M LiClO_4 and 1000 ppm water. DMA consumption indicates $^1\text{O}_2$ formation. Adapted from Ref. ¹⁰⁸ with permission from Springer Nature.

$^1\text{O}_2$ is not only responsible for electrolyte and carbon decomposition but has also been shown to be the prime source of redox mediator decomposition (see section 3.5 for a further discussion on the use of redox mediators). Kwak et al. studied the deactivation of two representative redox mediators by O_2 , superoxide, peroxide, and $^1\text{O}_2$ and found $^1\text{O}_2$ to dominate decomposition by far.¹⁵⁶ Its reactivity could conclusively be explained by electrophilic addition of $^1\text{O}_2$ to electron rich moieties. In turn, DABCO, as mentioned above, has shown to preserve the mediator.¹⁵⁷

Pathways towards singlet oxygen

A direct two-electron oxidation step along $\text{Li}_2\text{O}_2 \rightarrow \text{O}_2 + 2 \text{Li}^+ + 2 \text{e}^-$ was initially suggested by Hassoun et al.¹⁵⁰ and Wandt et al.¹⁰⁶ to possibly form $^1\text{O}_2$ above 3.9 V or ~3.55 V. However, more recent insights into the Li_2O_2 oxidation mechanism suggest no direct two-electron process exists as discussed above in Figure 8. Instead, superoxide disproportionation is now accepted to be involved on discharge and charge of the peroxides.^{37,90,102,113,114,116,117,158–160} Based on this recognition, recently, a unified mechanism for $^1\text{O}_2$ generation in metal- O_2 cells was established,¹⁰⁹ which links the Lewis acidity of cations in the electrolyte with the production of $^1\text{O}_2$. Whether the initial one-electron reduction product superoxide reacts further to peroxide is governed by the relative thermodynamic stability of peroxide and superoxide with the respective cation as illustrated in Figure 10a, which also gives rise to the favoured products in Table 1. Strong Lewis acids such as Li^+ or Na^+ favour the peroxide, albeit only slightly for sodium.^{119,124,128,160–162} K^+ and even weaker Lewis acids (e.g., quaternary amines like tetrabutylammonium (TBA^+) and imidazolium favour the superoxide.^{125,159,163,164} The latter constitute often-used ionic liquid electrolytes.

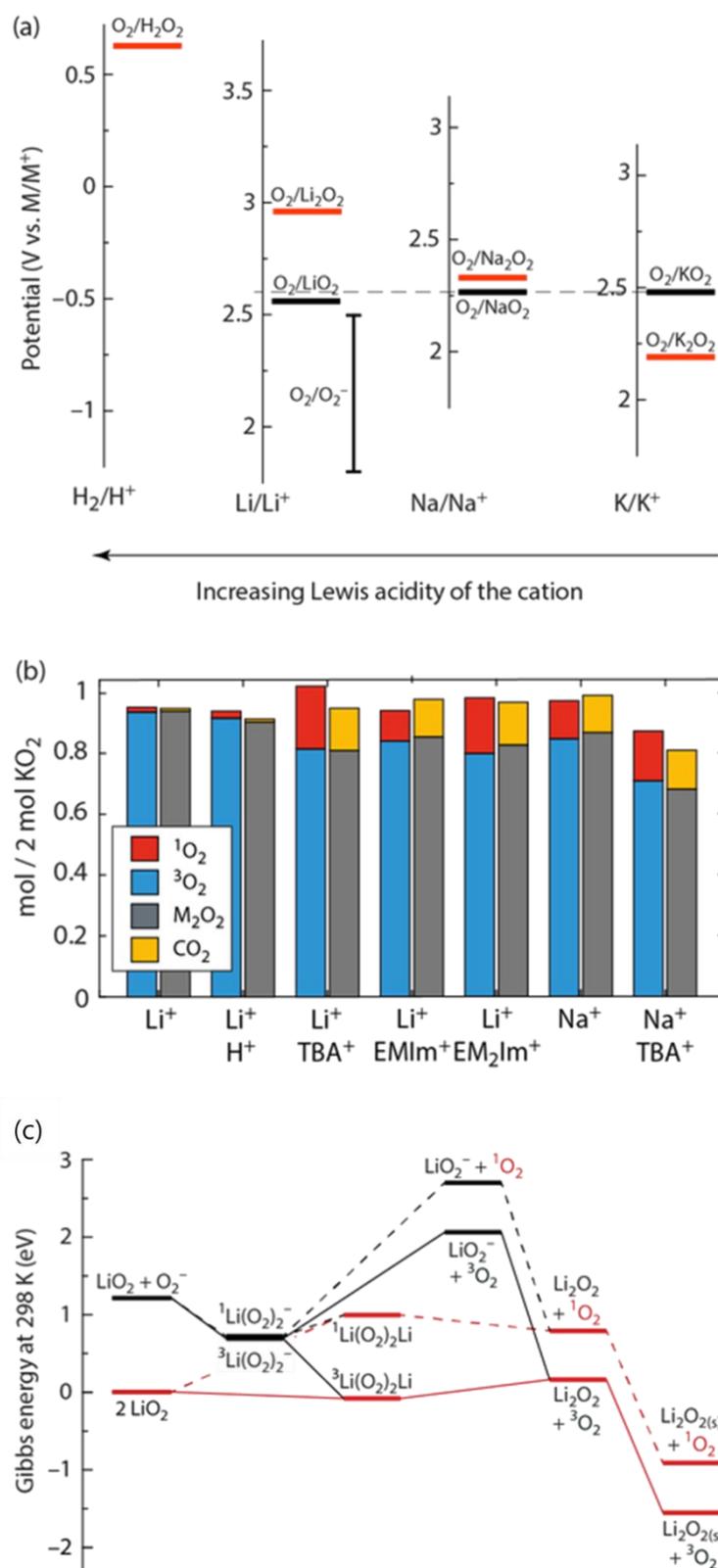
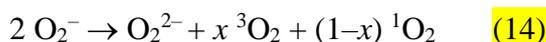


Figure 10 Singlet oxygen formation mechanism. (a) Thermodynamics of alkali peroxides and superoxides and the Lewis acidity of the involved cations. Standard potentials of the O_2/MO_2 and O_2/M_2O_2 redox couples on the M/M^+ scales with $M = Li, Na, K$ as well as for the O_2/H_2O_2 couple. The scales are brought to a common scale based on their M/M^+ standard potentials. The dashed horizontal line indicates the O_2/KO_2 couple. The O_2/LiO_2 potential is adopted from

Ref. ¹⁶⁵, but also values between 2.29 and 2.46 V vs. Li/Li⁺ have been reported.^{114,128} O₂/O₂⁻ denotes the potential range reported for O₂ reduction in TBA⁺ electrolytes.^{90,158,159,165,166} (b) ¹O₂ from superoxide disproportionation in presence of various cations. Obtained O₂, ¹O₂, and Li₂O₂ (or Na₂O₂) upon reacting KO₂ in TEGDME that contained equimolar 18-crown-6, 30 mM DMA, 0.5 M Li⁺ (or Na⁺), and either no additive, 0.1 M TBA⁺, EMIm⁺, or EM₂Im⁺, or F₃CCOOH. Ideally 1 mol O₂ and 1 mol M₂O₂ would form according to $\text{tO}_2 \text{K} \text{O}_2 + 2 \text{M}^+ \rightarrow \text{M}_2\text{O}_2 + \text{O}_2 + 2\text{K}^+$. (c) Reaction free energy profiles for LiO₂ disproportionation with itself or O₂⁻ to Li₂O₂ and molecular oxygen. Pathways to release ³O₂ and ¹O₂ are indicated by full and dashed lines, respectively. Adapted from Ref. ¹⁰⁹ - Published by The Royal Society of Chemistry.

Generally, the disproportionation reaction:



in presence of the alkali metal cations, H⁺ or weakly Lewis acidic cations was found to give rise to varying fractions of ¹O₂ as shown in [Figure 10b](#). According to [Figure 10a](#), the overall reaction is driven by the stability of peroxide and thus favoured by stronger Lewis acidic cations H⁺, Li⁺, and Na⁺. However, the ¹O₂ fraction grows with decreasing Lewis acidity of the cation, giving rise to insignificant ¹O₂ with H⁺ and strongly growing ¹O₂ fractions with Li⁺ and Na⁺. Importantly and surprisingly, weakly Lewis acidic cations such as TBA⁺ alone do not drive disproportionation, but, when combined with strong Lewis acids, cause substantially larger fractions of ¹O₂. This is generally true for prototypical motifs of cations used in ionic liquid electrolytes such as quaternary ammonium and imidazolium. How weak Lewis acids act to favour ¹O₂ can be explained by DFT calculations as shown in [Figure. 10c](#). ¹O₂ release from LiO₂ disproportionation faces a barrier of ~ 1eV (red path with dashed lines). However, when a weak Lewis acid is present next to Li⁺ in the electrolyte, asymmetric pairs of LiO₂ and, e.g, TBAO₂ will disproportionate. TBAO₂ may be seen as a free solvated O₂⁻ anion. Triplet and singlet paths pass via ³Li(O₂)₂⁻ and ¹Li(O₂)₂⁻ dimers that are stabilized versus LiO₂ + O₂⁻ by ~0.5 eV. Ongoing pathways are much more facile than for symmetric LiO₂ pairs: the Li(O₂)₂⁻ dimers exchange TBA⁺ for Li⁺ and go on to the symmetric Li(O₂)₂Li pathways. Crucially, the weak Lewis acid curbs the largest barrier towards ¹O₂, the endergonicity to the ¹Li(O₂)₂Li

dimer, from ~1 eV to only 0.27 eV. Overall, weak Lewis acids allow for pathways that bypass the most unfavourable reaction steps and hence strongly facilitate $^1\text{O}_2$ evolution. Moreover, TBA^+ arises as a probe for disproportionation steps. If the presence of TBA^+ increases the $^1\text{O}_2$ yield, then disproportionation must be involved.

This insight was used to probe $^1\text{O}_2$ formation during discharge and charge. Larger $^1\text{O}_2$ yields in mixed Li^+/TBA^+ electrolytes as compared to pure Li^+ electrolytes verified that the described pathways are active both on discharge and charge. Strong sensitivity of the $^1\text{O}_2$ yield towards the presence of TBA^+ confirms discharge to significantly occur via disproportionation, although possible second one-electron reduction of the LiO_2 intermediate is not ruled out. On charge, the initial charging step forms a LiO_2 -like surface species as discussed above^{113–115,167} and its disproportionation forms $^1\text{O}_2$ analogously. This pathway is active from the onset of charge. Another possible pathway is a second one-electron oxidation of the superoxide intermediate that could give $^1\text{O}_2$ above ~3.5 V.

Electrochemical oxidation of Li_2CO_3 has recently been revealed as another pathway towards $^1\text{O}_2$. Li_2CO_3 is a common side product on discharge and charge^{105,154} and its formation known to contribute to the ever increasing voltage on charge because of its high oxidation potential^{84,85,105,123,131,168,169}. The burden of Li_2CO_3 formation was seemingly harnessed in rechargeable metal- O_2/CO_2 batteries based on the observation that Li_2CO_3 can be electrochemically decomposed.^{84,166,170,171} For long, the fate of the third O-atom in CO_3^{2-} has led to various theories, since no O_2 evolves as expected from the formal oxidation reaction:^{84,171–173}



To explain the absence of O_2 , the formation of superoxide or “nascent oxygen” was proposed, which could react with cell components¹⁷³ (without, however, definite proof for these

mechanisms). Mahne et al. provided evidence that the electrochemical oxidation of Li_2CO_3 forms $^1\text{O}_2$, which, through the parasitic reaction of $^1\text{O}_2$ with battery components, explains the absence of O_2 evolution.¹⁷⁴ Since Li_2CO_3 is also a common residual component on transition metal oxide (TMO) cathode materials,¹⁷² $^1\text{O}_2$ from Li_2CO_3 oxidation has far-reaching implications for the surface reactivity of TMO cathodes that operate beyond 3.8 V and couples parasitic reactions upon recharging metal- O_2 and metal- O_2/CO_2 batteries.

$^1\text{O}_2$ generation upon oxidation of Li_2CO_3 was probed by charging cells with Li_2CO_3 -packed working electrodes in presence of DMA. HPLC analysis of the extracted electrolyte (Figure 11a) shows that DMA- O_2 is formed from 3.8 V onwards and hence from the onset of Li_2CO_3 oxidation. In quantitative terms, the amount of $^1\text{O}_2$ shows that the majority of the “missing O_2 ” from the electrochemical Li_2CO_3 oxidation forms $^1\text{O}_2$, explaining its absence in the gas phase. Absent O_2 evolution implies that the formed $^1\text{O}_2$ reacts with cell components rather than being, even in part, deactivated to $^3\text{O}_2$. In presence of the $^1\text{O}_2$ quencher DABCO, O_2 evolution started together with CO_2 evolution at ~ 3.8 V with a similarly growing rate as the voltage rose (Figure 11b). Hence, with a suitable quencher Li_2CO_3 could potentially be oxidized without the detrimental effects of $^1\text{O}_2$ formation. The $^1\text{O}_2$ formation mechanism was suggested to occur via a peroxodicarbonate intermediate that forms from two Li_2CO_3 moieties via two $\text{LiO}_2\text{CO}\cdot$ moieties, which combine to $\text{LiO}_2\text{COOCO}_2\text{Li}$. Its further oxidation and decarboxylation to yield LiCO_4 could in turn yield $^1\text{O}_2$.

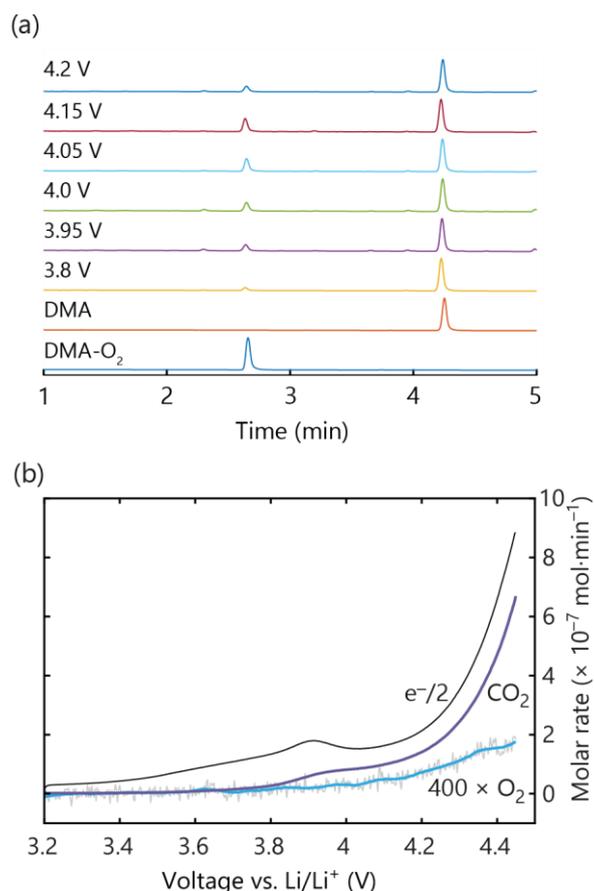


Figure. 11 Singlet oxygen from Li₂CO₃ oxidation. (a) HPLC determination of DMA-O₂ upon polarizing Li₂CO₃-packed electrodes at different potentials in DME containing 0.1 M LiTFSI and 30 mM DMA. (b) CO₂ and O₂ evolution during electrochemical oxidation of Li₂CO₃ in presence of the quencher DABCO. Adapted from adapted from Ref. ¹⁷⁴ with permission from Wiley-VCH.

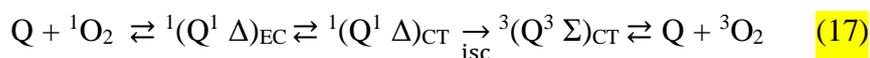
Quenching singlet oxygen

While superoxide disproportionation is now known to be the source of ¹O₂ via Eq. 15, it is also known to readily deactivate it via quenching.¹⁷⁵



The net formation of singlet oxygen will hence not simply scale with the superoxide concentration. It will rather depend on the relative kinetics of all superoxide sources and sinks with ¹O₂ being involved in both. Many factors may influence the kinetics, which may be governed by electrolyte, electrodes, current or potential, for example.

Since cell-inherent quenching by superoxide is too inefficient to deactivate $^1\text{O}_2$, quenchers may be added. Early work has shown that DABCO significantly reduces $^1\text{O}_2$ -related side reactions on discharge and charge. However, its oxidation stability of 3.6 V is too low for recharge¹⁰⁸. Physical quenchers deactivate $^1\text{O}_2$ to $^3\text{O}_2$ without quencher consumption and no new products accumulate. Radiationless physical quenching converts the excess energy of $^1\text{O}_2$ into heat via three mechanisms¹⁷⁵: electronic-to-vibrational (e-v) energy transfer ($^1\text{O}_2$ quenching by solvents, slow), charge transfer (CT) induced quenching ($\sim 10^7$ times faster than e-v), and electronic energy transfer (faster yet, unsuitable for electrochemical systems because of too low oxidation stability). Suitable quenchers for electrochemical systems use the CT mechanism which passes via a singlet encounter complex $^1(\text{Q}^1\Delta)_{\text{EC}}$ and a singlet charge transfer complex $^1(\text{Q}^1\Delta)_{\text{CT}}$, where electronic charge is partially transferred to the oxygen. Intersystem crossing (isc) to the triplet ground state complex $^3(\text{Q}^3\Sigma)_{\text{CT}}$ releases energy. The latter dissociates to Q and $^3\text{O}_2$



This mechanism was first suggested for DABCO and has generally been shown to be applicable for electron-rich quenchers such as amines. The partial charge transfer causes the quenching efficiency to correlate logarithmically with the ionization potential and hence oxidation potential.¹⁷⁵ A suitable, sufficiently oxidation-stable quencher should therefore allow for high concentrations to balance for the inevitably lower molar activity.

Monoalkylating DABCO to form the DABCONium salt has recently been shown to shift its oxidation stability to ~ 4.2 V vs. Li/Li^+ while still quenching effectively.¹⁷⁶ Previously known quenchers are strongly Lewis basic amines with too low an oxidative stability. DABCONium is a cation of a non-volatile ionic liquid, highly soluble in electrolytes, stable against superoxide and peroxide, and compatible with lithium metal. The electrochemical

stability covers the required range for metal-O₂ batteries and greatly reduces ¹O₂ related parasitic chemistry as demonstrated for the Li-O₂ cell. DABCONium was shown to drastically reduce the amount of side products during discharge and charge. Its effect was especially clear during charge, where O₂ evolved at the ideal rate of 2 e⁻/O₂ up to the oxidative stability limit of ~4.2 V, which proves that all electrons are extracted from Li₂O₂ rather than partly from side products.

Given the stellar importance of ¹O₂ for the reversibility of metal-O₂ cells, further research into ways to prevent its formation appear highly appropriate. Recognizing that ¹O₂ formation is deeply rooted in the way current metal-O₂ cells operate has serious consequences for aspects to avoid and on directions that should be taken. First, caution must be exercised with weak Lewis acids as electrolytes or additives. Given that ionic liquid cations suitable for electrolytes are most typically weak Lewis acids, the effect can likely be generalized. Favored ¹O₂ formation explains why quantitative studies of metal-O₂ chemistry with a broad variety of ionic liquids have shown worse parasitic chemistry on discharge and charge than molecular electrolytes.^{131,177,178} Second, protic additives drive ¹O₂ formation insignificantly but may drive parasitic chemistry in other ways. Third, reversible chemistry based on alkali metal carbonate formation/decomposition (metal-O₂/CO₂ cells) is questionable based on the involvement of ¹O₂ and other reactive species.^{174-176,179} The most prominent consequence is that reactions involving superoxide disproportionation must be avoided. Cells based on metastable LiO₂ or NaO₂ as target products lack the practically required tolerance to slow discharge and rest periods as the superoxides gradually convert to peroxide and side products.^{107,116,118,132,161,162,180} Peroxide products are preferred as they are much higher in energy density and are the thermodynamically stable products^{124,128,181}. Cycling them highly reversibly requires finding routes to form and decompose them without superoxide disproportionation steps. It is not yet clear whether ¹O₂ -related parasitic chemistry can be completely avoided but practical

realization of aprotic Li-O₂ and Na-O₂ batteries will, in our opinion, stand or fall with mastering ¹O₂ formation.

3.4. Cell components

3.4(a) Electrolytes.

In this section, we will divide our discussion into two categories: (1) aprotic solvents; and (2) Li- salts. As an initial note it must be stated that the vast majority of work described in this section was published before recognizing the importance of ¹O₂ in parasitic chemistry (see the previous section). Therefore, proposed side reactions and mitigation strategies were based on the assumption that the nucleophilicity, basicity, or radical nature of reduced O₂ species governed reactivity.

Aprotic Solvents

As noted previously, early research on Li-O₂ batteries utilized organic carbonates as the solvents, following their dominance in Li-ion batteries.⁸¹ In 2006, Ogasawara et al.⁸³ demonstrated a rechargeable Li-O₂ battery with ~50 cycles using alkyl carbonates. Importantly, they also demonstrated for the first time that Li₂O₂ can be electrochemically oxidized to O₂. A few years later, Mizuno et al.¹⁸² demonstrated a Li-O₂ battery that was discharged and recharged for 100 cycles - with almost 60% capacity retention. To explain the large voltage hysteresis of over 1.4 V, they suggested the formation of significant quantities of Li₂CO₃, as well as Li₂O₂, as the discharge products, the former originating from the decomposition of carbonates.¹⁸² Later, Freunberger et al. illustrated the cathodic decomposition pathways of carbonates in oxygen environment specifically in the presence of the superoxide radical.⁸⁴ Proposing the degradation mechanism during the discharge and charge shown in **Figure 12**, they showed that the prolonged cyclability of Li-O₂ cells containing alkyl carbonate-based electrolyte solutions was attributed to the irreversible decomposition and oxidation of Li₂CO₃ along with other species, namely C₃H₆(OCOOli)₂, CH₃COOLi, HCOOLi, CO₂ and H₂O with

no evidence of reversible oxygen reduction to Li_2O_2 . Furthermore, they reported the accumulation of side products on the cathode in a polymer like film, resulting in increased over-potential, capacity fading and ultimately cell failure. The dramatic instability of organic carbonates in the highly reactive environment of Li-O_2 led to studies of other solvents and catalyzed the search for better and more stable solvents.

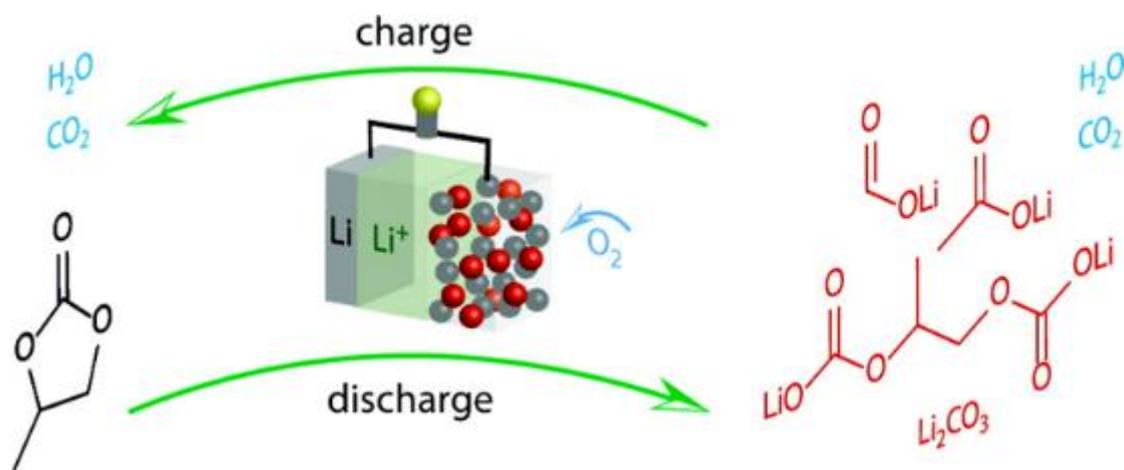
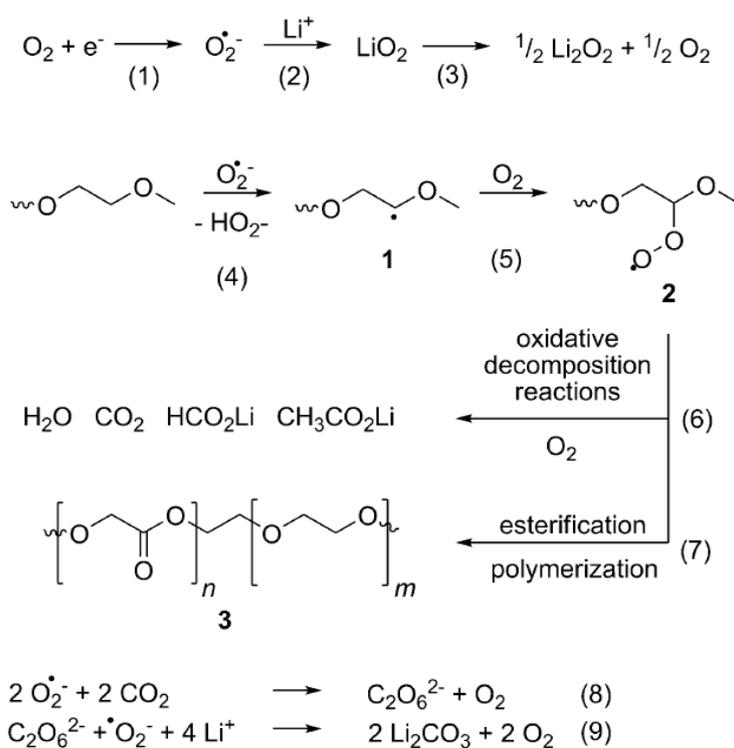


Figure 12. Graphical representation of charge –discharge of Li-O_2 cell with 1 M LiPF_6 in PC. Reproduced with permission from ref⁸⁴. Copyright 2011 American Chemical Society.

Building on their initial experience⁸² with the carbonate instability, Read et al.⁸⁰ investigated the use of ethers as electrolyte solvents. Using DME (dimethoxyethane) and DOL (1,3-dioxolane), they described a Li-O_2 cell showing both good stability and excellent rate capability. They attributed these properties to the high oxygen solubility and low viscosity which results in facilitated oxygen transport in the Li-O_2 cell. Later McCloskey et al.¹⁸³ studied the fundamental chemistry, in particular, the discharge product during the recharging of Li-O_2 cells. They confirmed that in DME-based solvents, Li_2O_2 is the major discharge product; however, DME undergoes oxidative degradation during recharge in the presence of Li_2O_2 . Sharon et al and Freunberger et al.^{121,184} further confirmed the degradation of DME and other ether-based solvents and proposed the formation of CO_2 , HCOOLi , CH_3COOLi , LiOCOOLi , $\text{LiOCH}_2\text{CH}_2\text{OLi}$ following the mechanisms shown in **Scheme 1**. Compared to the carbonates, ether-based electrolytes are relatively more stable - attributable to the ether solvents' greater

resistance to nucleophilic attack and higher cathodic stability.¹⁸³ Hence, they are considered to be the best compromise for such active systems. Indeed, progress^{185,186} in Li-O₂ batteries would not have been possible without glyme based electrolyte solutions. It is important to note that various glymes differ in their reactivity towards the oxygen species formed in Li-O₂ cells. In a study summarized in **Figure 13** it was clearly found that diglyme is the most stable glyme solvent in Li-O₂ cells. The reason for that may relate to the nature of the active sites in the polyether molecules. This depends on the chain length, number of oxygens, the steric conformations and the interactions of the oxygen atoms with the Li ions.



Scheme 1. Proposed mechanism for the decomposition of ether based solvents during the oxygen reduction reactions in Li-O₂ batteries. Reproduced with permission from ref¹²¹. Copyright 2011 WILEY-VCH Verlag GmbH & Co. KGaA, Weinheim.

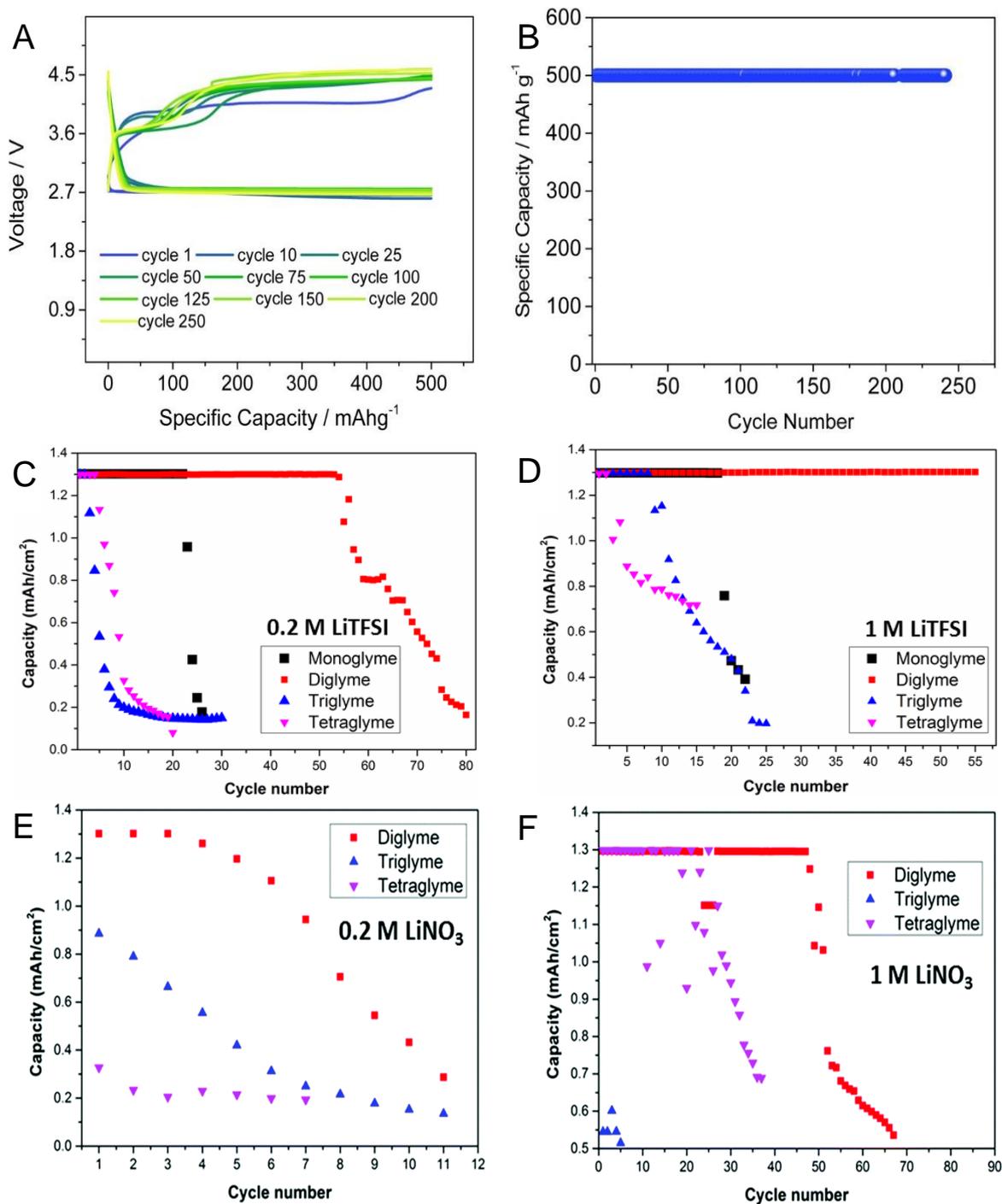


Figure 13. The performance of Li-O₂ cells with glyme based electrolyte solutions and the relative stability of glymes solvents: (A) Galvanostatic voltage profiles and (B) corresponding cycling trend of the lithium oxygen batteries using the TEGDME-LiTFSI (1 mol kg⁻¹) electrolyte and the N₂@MWCNT electrode. Reproduced with permission from ref¹⁸⁶. Copyright 2018 American Chemical Society. Cycling behavior of Li-O₂ cells in different glymes containing (C) 0.2 M LiTFSI, (D) 1 M LiTFSI (E) 0.2 M LiNO₃, and (F) 1 M LiNO₃ electrolyte solutions. Reproduced with permission from ref¹⁸⁵. Copyright 2017 Royal Society of Chemistry.

While the glyme polyether solvents were found indeed to be more stable than the carbonates, there was a lack of unanimity on the extent of this stability. Their relatively non-polar nature also keeps the concentration of Li salts rather low. Crucially, they resulted in cells that exhibited low rates and capacities, which we now understand is because they promote the formation of Li_2O_2 films on the electrode surface on discharge, which are insulating and limited to approx. 7 nm^{187} as discussed in the section describing the mechanism of $\text{O}_2/\text{Li}_2\text{O}_2$ above. This low rate and capacity promoted examination of other solvents including higher DN solvents such as dimethyl sulfoxide (DMSO). Peng et al. reported on Li- O_2 cells with DMSO as the solvent, with a porous gold electrode, exhibiting $\sim 95\%$ capacity retention for 100 cycles (Figure 14A).¹⁸⁸ Subsequently, several other groups reported excellent reversibility of Li- O_2 cells, with Li_2O_2 as the discharge product.^{189,190} However, the capacity of the Au electrode was intrinsically low due the limited pore volume. Investigation with larger pore volume carbon electrodes revealed significant degradation of DMSO, attributed to the susceptibility of DMSO to nucleophilic attack by reduced oxygen species, forming sulfone, and dimethyl sulfone - following the mechanism outlined in Figure 14B.¹⁹¹ It should, however, be noted that excellent cyclability of KO_2 in DMSO has been reported recently,¹²⁰ which casts doubt at the major influence of direct reactivity of superoxide.

Nevertheless, Liu et al. have recently reported a super-concentrated salt/DMSO based Li- O_2 cell which showed excellent stability and reversibility even with an unprotected Li anode. Since DMSO molecules are corrosive to the Li metal anode and unstable to $\text{O}_2^{\bullet-}$ attack, the better performance of the super-concentrated solutions can be attributed to the absence of free DMSO solvent molecules and existence of only $\text{TFSI}^- \text{-Li}^+ \text{-(DMSO)}_3$ complexes.

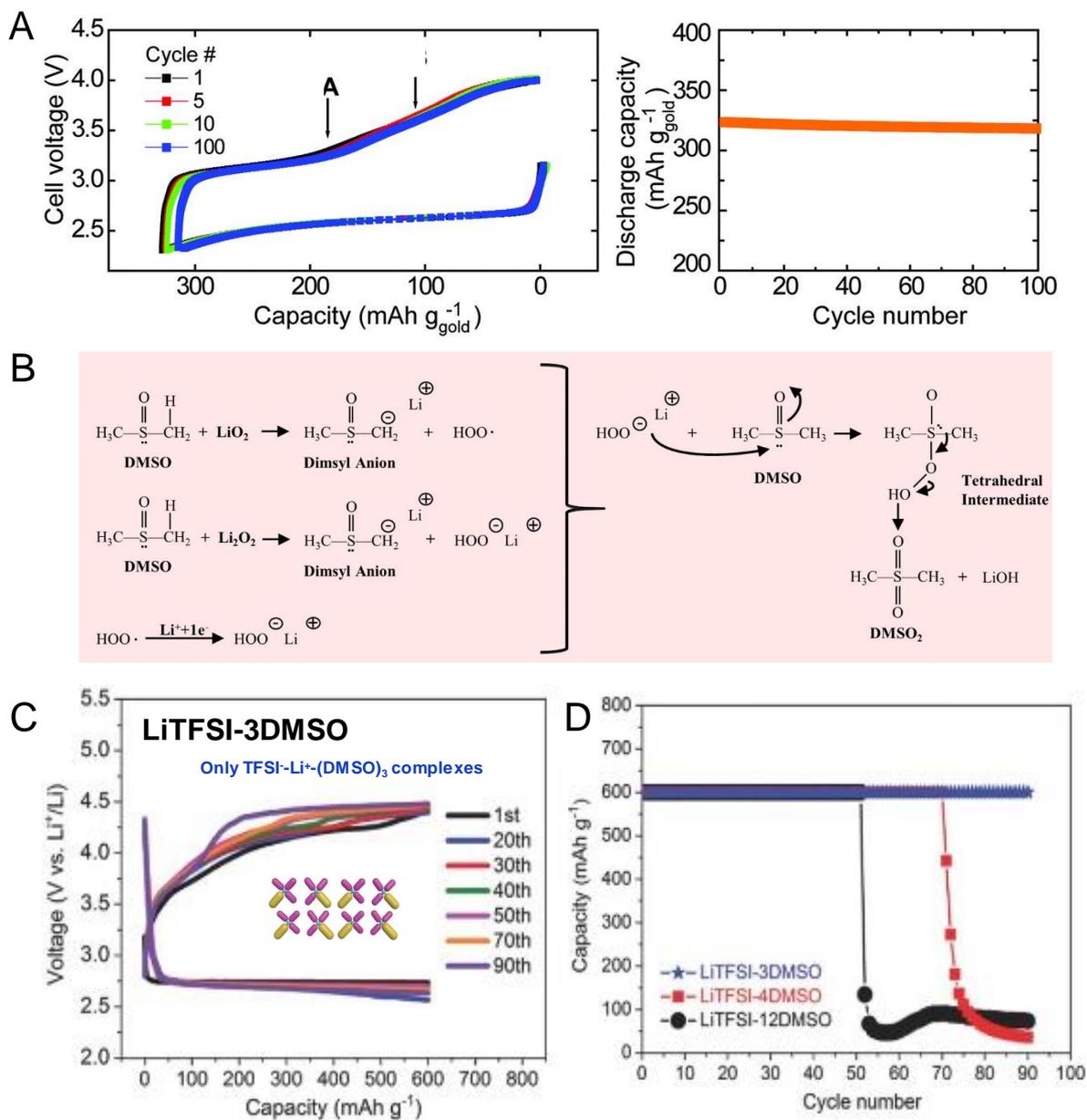
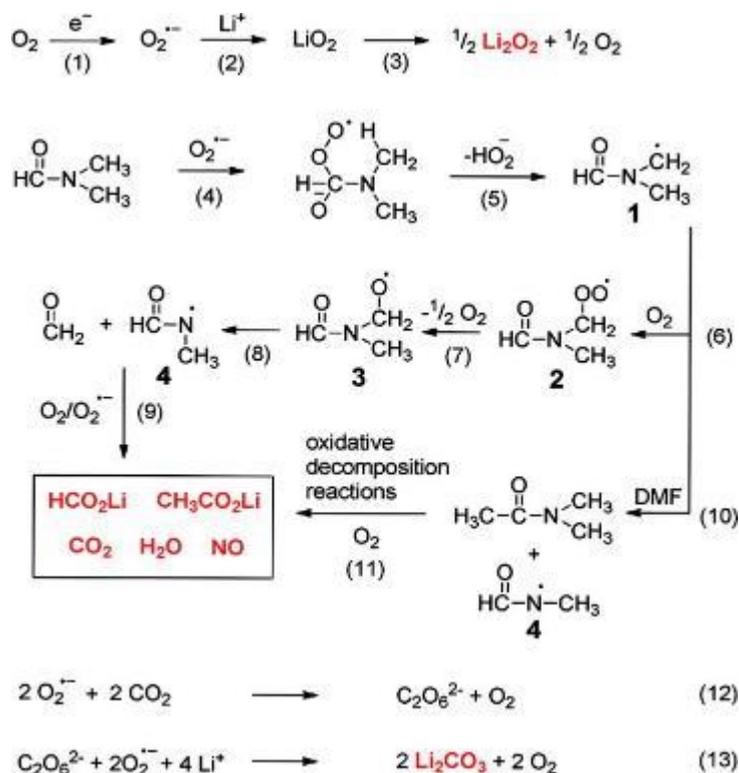


Figure 14. On the stability of DMSO: (A) Performance of Li-O₂ battery with DMSO as solvent and a porous gold cathode. Reproduced with permission from ref¹⁸⁸. Copyright 2012 The American Association for the Advancement of Science. (B) proposed mechanism of degradation of DMSO. Reproduced with permission from ref¹⁹¹. Copyright 2013 American Chemical Society. Voltage profiles for (C) LiTFSI-3DMSO electrolyte and (D) cycling stability of the three electrolytes with different concentration. Reproduced with permission from ref¹⁹². Copyright 2017 WILEY-VCH Verlag GmbH & Co. KGaA, Weinheim.

The unsatisfactory stability of DMSO spurred further search of new solvent candidates showing stability toward ROS. Amide-based solvents like DMA, DMF, and NMP were next investigated because of their presumed stability towards the superoxide radical. However, the amines do have facile reactivity toward lithium metal. To minimize the latter issue, Uddin et

al.¹⁹³ proposed the use of LiNO₃ as a passivation agent for metallic lithium. Similarly, exploiting these passivating properties of the LiNO₃, Walker et al.¹⁹⁴ showed reversible cycling of a Li-O₂ battery for more than 2000 h in N,N-dimethylacetamide (DMA), with Li₂O₂ as the main discharge product. Unfortunately, this amide-based solvent cycling stability was found to be unsuitable for oxygen reduction applications. As outlined in **Scheme 2**, studies by Chen et al.¹⁹⁵ and Sharon et al.¹⁹⁶ demonstrated the oxidative cleavage of amides and their further degradation to various species including dimethylamine, acetate, carbonate, and various N-O species. Although the stability results of amide-based solvents are debatable, it may be concluded that amide stability toward reduced oxygen species is not sufficient for rechargeable aprotic Li-O₂ cells. Another popular polar solvent, acetonitrile,¹³¹ also showed some success owing to its stability towards superoxide; however, because of its high volatility and high reactivity toward lithium metal, it was not extensively explored for Li-O₂ batteries.



Scheme 2. Proposed mechanism for degradation of dimethylacetamide during oxygen reduction reactions. Reproduced with permission from ref¹⁹⁵. Copyright 2012 American Chemical Society.

In summary, almost all the available aprotic solvents degrade to a greater or lesser extent under the highly reactive environment of Li-O₂ batteries. Four modes of degradation mechanisms can be listed: nucleophilic attack, H-atom abstraction, acid/base reactions, and reductive degradation at the metallic lithium anodes. The strongly nucleophilic superoxide anion (O₂⁻) readily attacks the electron deficient sites of polar solvents such as carbonyl or sulfoxides groups. These include carbonates, amides, as well as sulfoxides (like DMSO) which undergo degradation via nucleophilic attacks. In other modes of action, superoxide can mediate α -hydrogen atom and/or β -proton abstraction via acid/base reactions leading to solvent deprotonation and decomposition. Ethers, which are generally stable to nucleophilic attack, can undergo some autoxidative decomposition in the presence of these reactive oxygen species. Hence, the ethers solvents are proven to be the most appropriate class of solvents (among the available options) for Li-O₂ batteries.

It is important to emphasize here that the extent of degradation also depends on the possible stabilization of reduced oxygen species by the electrolyte solutions (nature of the solvent, salt anions, additives). Such a stabilization leads to prolonged exposure of electrolyte solution to the reactive species thereby enhancing the probability of nucleophilic attack, H-abstraction etc. This leads to the conundrum in which solutions that promote a top-down growth of the Li₂O₂ during discharge - leading to enhanced discharge capacity – may be also exposed to accelerated decomposition (stabilization of reduced oxygen species increase their retention time, thus their ability to attach solution species). The various degradation pathways utilized by the various families of solvents are listed in **Figure 15**.⁶⁶ New solvents continue to be synthesized and tested for use in Li-O₂ batteries – as discussed in the next section.

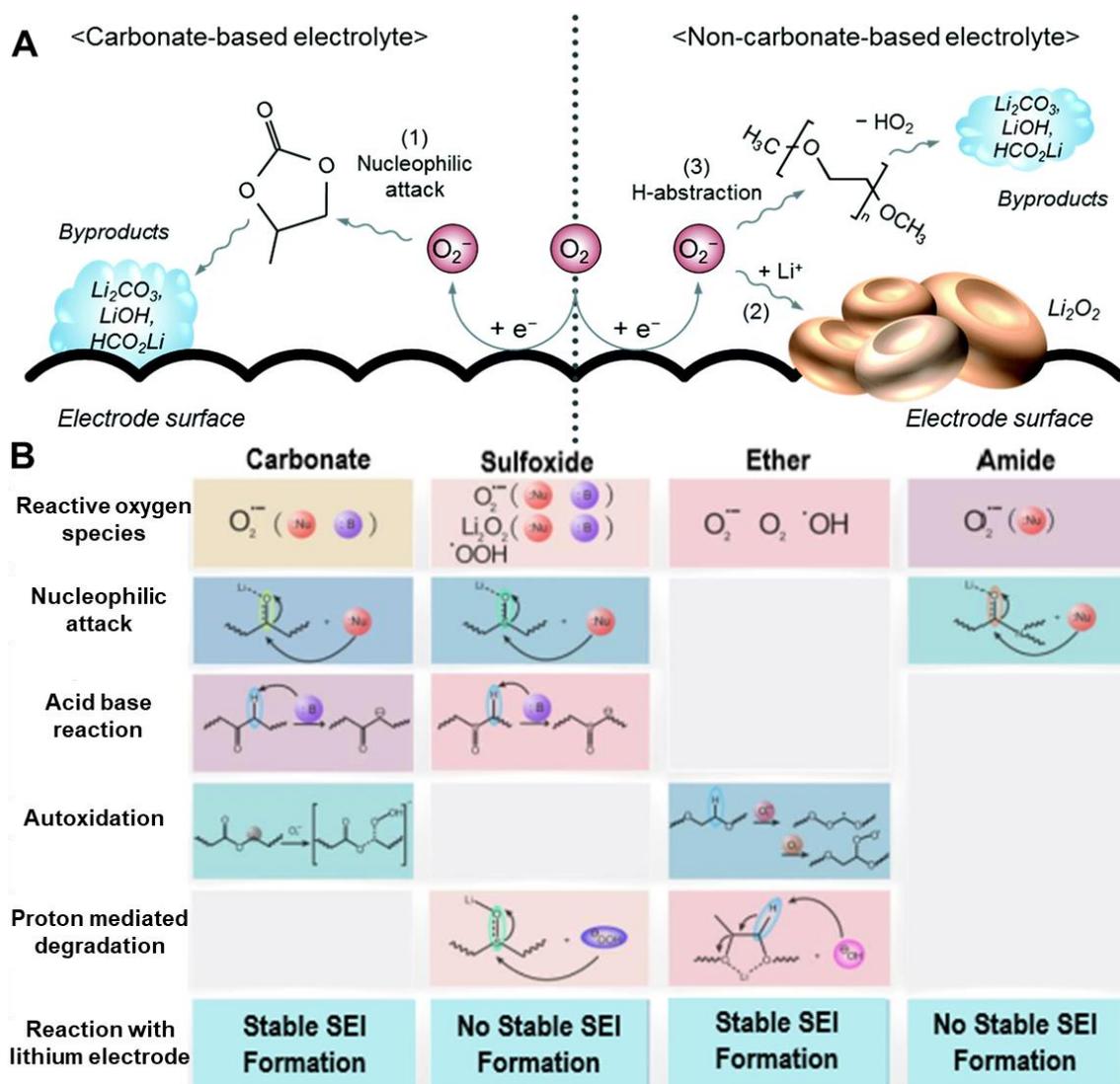


Figure 15. (A) Schematic illustration of possible discharge reactions involving oxygen radicals in different electrolyte systems. (1) Nucleophilic attack of an oxygen radical in a carbonate-based electrolyte and the subsequently produced byproducts. (2) Reaction between oxygen radicals and Li ions to produce Li_2O_2 in a non-carbonate-based electrolyte. (3) H-abstraction reaction induced by oxygen radicals and the subsequently produced byproducts in a non-carbonate-based electrolyte. Reproduced with permission from ref.¹⁹⁷ Copyright 2017 Royal Society of Chemistry. (B) Possible degradation pathways of the solvents. Reproduced with permission from ref.⁶⁶. Copyright 2019 WILEY-VCH Verlag GmbH & Co. KGaA, Weinheim.

Salts

Li salts are an indispensable component of the electrolyte solutions which strongly influence the course of ORR/OER. As already pointed out, the donor number of the solvents

as well as the counter anion (as expressed in the ionic association/dissociation level of the salt) play important roles in guiding the discharge product, especially in terms of its morphology.

LiNO_3 is one of the most explored salts in Li-O_2 batteries. In addition to promoting solution mediated growth, LiNO_3 also exhibits a bi-functional role. It catalyzes the OER^{198,199} and passivates the carbon electrode against oxidative damage.²⁰⁰ A myriad of other salts have also been employed in Li-O_2 battery electrolyte solutions and their discharge voltage profiles are presented in **Figure 16A**. Distinct discharge capacity and voltage plateaus are observed for each of the different anions. LiTFSI exhibited the maximum discharge capacity followed by LiTf , and LiPF_6 . LiTFSI is also among the most frequently salts reported for Li-O_2 batteries because of its good chemical and electrochemical stability and considerable ionic conductivity. However, it also undergoes partial decomposition to form LiF on the cathode surface, similar to LiTf and LiPF_6 .²⁰¹

Using XRD and XPS, Nasybulin et al.²⁰¹ studied various salts in Li-O_2 batteries and showed that almost every salt undergoes at least partial decomposition, reacting with the in situ generated reactive oxygen species. LiBF_4 and LiBOB suffered from complete decomposition to LiF , boron–oxygen compounds,^{66,201} and Li_2CO_3 on the cathode surface. The degradation of LiBF_4 and LiBOB is reflected in the abnormally high discharge plateaus seen in **Figure 16A**. By contrast, LiTFSI , LiNO_3 , LiClO_4 , and LiBr showed relatively stable performance. The beneficial role of LiBr and LiI as redox mediators has also been reported recently, and is discussed further below.^{202,203} This emphasizes that in describing the performance of Li-O_2 battery chemistry, the involvement of the salts holds an equal importance to the stability of the solvents. Consequently, this aspect should be comprehensively investigated in order to come up with stable electrolyte solutions.

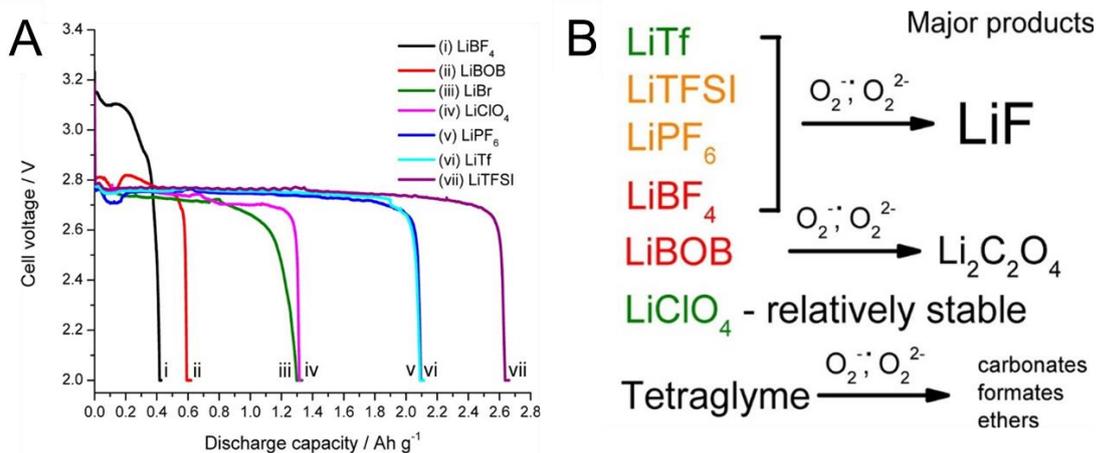


Figure 16. (A) Discharge profiles of Li-O₂ battery with various anions of Li-salt dissolved in TEGDME, (B) Decomposition products of Li-salts in the presence of reactive oxygen species. Reproduced with permission from ref²⁰¹, copyright American Chemical Society, 2013.

3.3(b) Development of new solvents for Li- oxygen batteries.

From the discussion above, the importance of finding appropriate solvents is crucial for the development of durable Li-O₂ batteries. As shown in [Figure 17](#), the various characteristics of solvents that can be suitable for Li-O₂ batteries have to be considered.

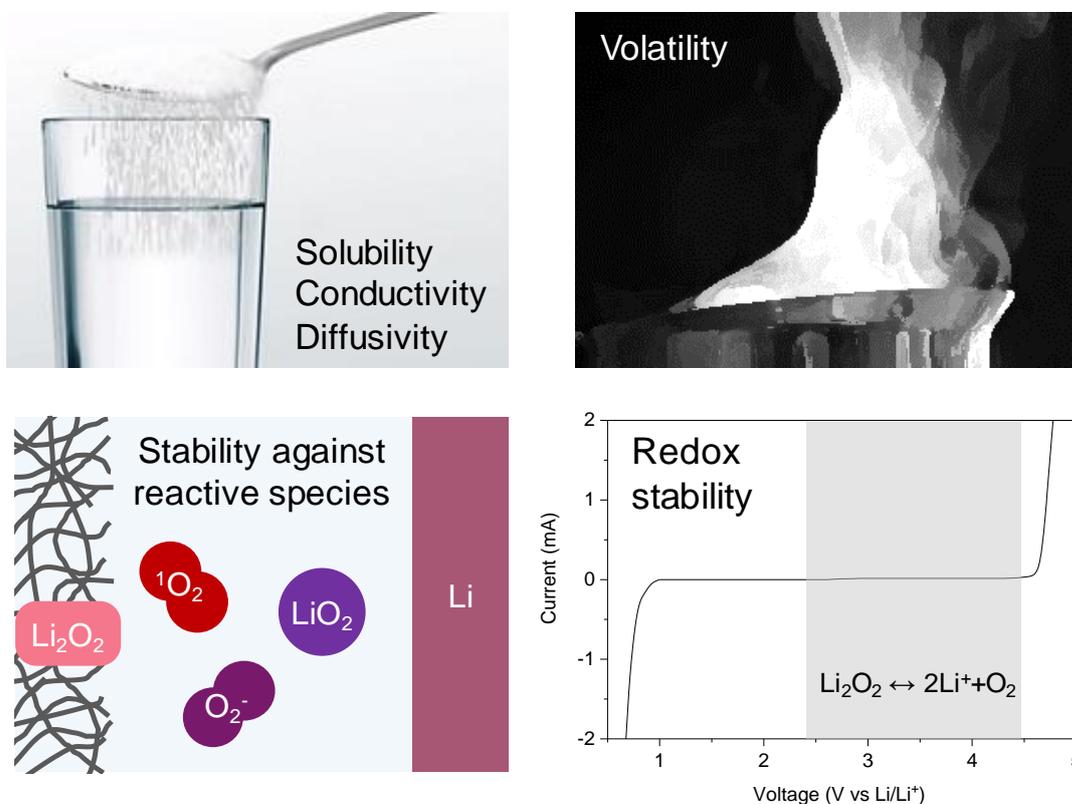


Figure 17. Important parameters for promising solvents for Li-O₂ batteries.

The solubility of the Li salt must be considered, since the solution must be capable of ion transfer. In addition, the solubility of O₂ gas should also be considered because the O₂ gas is the cathode's active mass which should be dissolved in the electrolyte solution in order to react electrochemically. The conductivity and diffusivity, which are critical properties of any electrolyte solution, can also be controlled by the nature of the solvent, and thereby should be taken into account. The volatility of the solvents is also a factor that should be considered for long-term operation. Non-volatile solvents like tetraglyme are indeed widely used in these systems.

It is also important to have enough redox stability to prevent decomposition during discharge and charge, but it should be noted that practical lithium-air batteries would aim to operate below 3.5 V, and significant anodic stability isn't requisite. However, when the stability window of the electrolyte is exceeded, decomposition of the solvent occurs and such side reactions not only interfere with the desired decomposition of Li₂O₂, but also decreases the cyclability of the batteries due to the accumulation of byproducts on the cathode.

Hence, in general, the reversibility of Li-O₂ cells is very poor if the solvent is decomposed by possible reactions with O₂⁻, LiO₂, or the recently reported ¹O₂ - produced by the formation and decomposition of Li₂O₂. Since Li₂O₂ is an indispensable product of the battery reaction, stability of solvents against Li₂O₂ must be ensured. Therefore, the stability of any solvent examined for Li-O₂ batteries to ROS should be first verified. Several candidate solvents that have been reported as more stable than glyme solvents in Li-O₂ cells are summarized in **Figure 18** below.

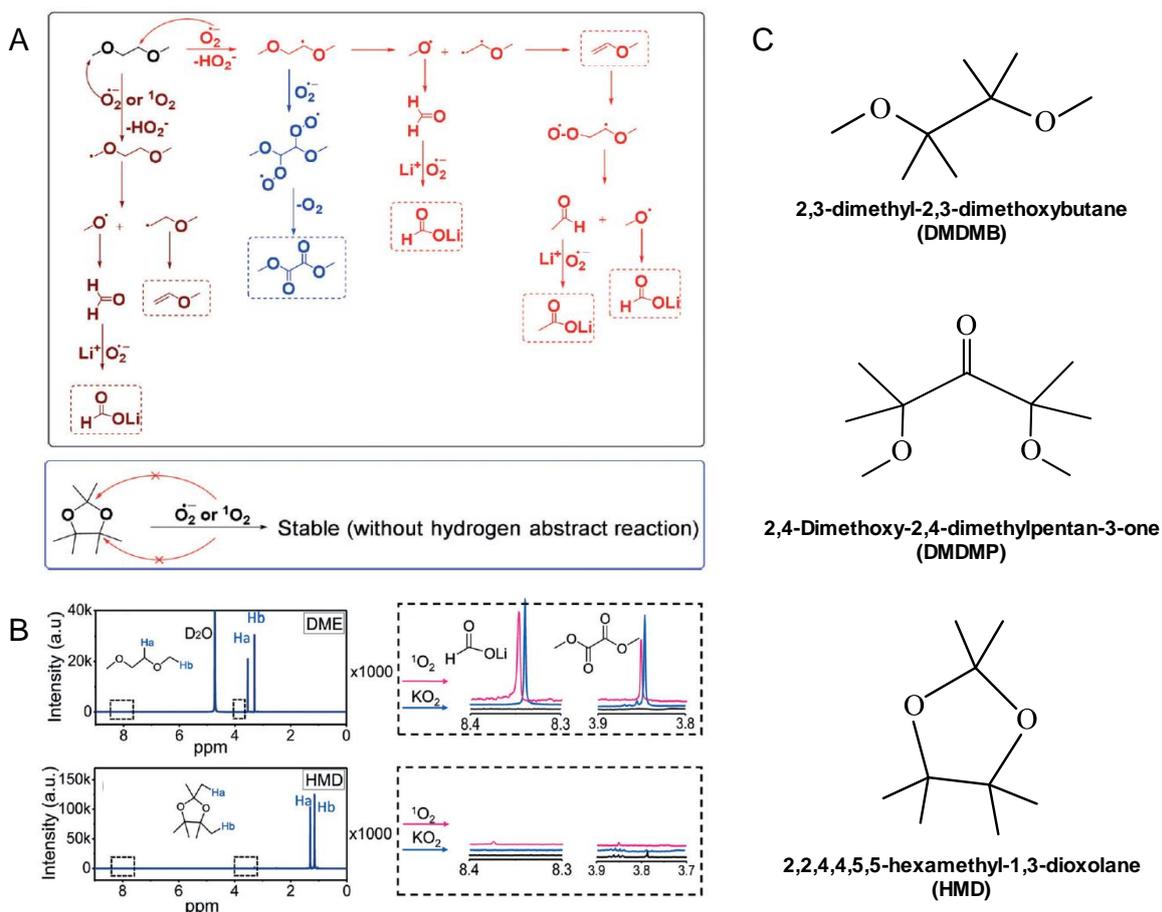


Figure 18. (A) Possible decomposition mechanism of DME (top) and HMD (down). (B) ^1H NMR spectra (left) with magnified views (right) for checking stability of DME (top) and HMD (down) against $^1\text{O}_2$ and KO_2 . (C) Proposed candidates as stable solvents in Li-O₂ batteries in previous articles.^{133,204,205} Reproduced with permission from ref.²⁰⁵. Copyright 2019 John Wiley & Sons, Inc.

Adams et al.¹³³ synthesized 2,3-dimethyl-2,3-dimethoxybutane (DMDMB). In this compound, the internal β -hydrogens of DME (1,2-dimethoxyethane) are replaced with methyl groups. As shown in **Figure 18A**, this prevents facile elimination by preventing abstraction of these hydrogens. In their paper, DMDMB reduced the amount of generated CO_2 by 10 fold and did not form byproducts such as lithium formate compared to DME.

Sharon et al.²⁰⁴ synthesized 2,4-dimethoxy-2,4-dimethylpentan-3-one (DMDMP) by replacing the acidic hydrogens alpha to the carbonyl of 1,3-dimethoxyacetone with methyl groups. The DMDMP is stable to nucleophilic attacks and H-abstraction by reduced oxygen species.

Huang et al.²⁰⁵ proposed 2,2,4,4,5,5-hexamethyl-1,3-dioxolane (HMD), which does not have any hydrogen atoms on the alpha-carbon, also preventing H-abstraction by superoxide. Compared to DMDMB and DMDMP, NMR data demonstrated that HMD is stable to both reduced oxygen species and singlet oxygen.

The common strategy in all the above studies is to substitute labile hydrogens, which are the starting point for solvent decomposition, with methyl or methoxy groups, generating structures more stable to ROS. However, since only a few examples of such compounds appear in the literature, more examples are required. Furthermore, in addition to improving the stability by using steric hindrance, various other properties such as solubility, conductivity, and volatility of these solvents must be examined.

It is expected that stable and reliable solvents, which can improve the actual performance, can be found as substitutes for polyethers (glymes), which were predominantly used until now as the best compromised family of solvents. However, the development of a suitable solvent that meets all conditions may be difficult in practice.

Strategies that can compensate for instability of solvents should be considered, such a high concentration of electrolytes (the concept of solvent in salt). Liu et al.^{146,192} proposed highly-concentrated solutions as more stable electrolytes. These include 3 M solutions of LiTFSI in DME and highly-concentrated LiTFSI-DMSO (1/3 n/n) which in either case contain no free solvent molecules and which could in comparison to less concentrated solutions demonstrate improved stability toward Li-metal anodes and air cathodes.

Dong et al.²⁰⁶ have also seemingly solved the problem of parasitic reactions in Li-O₂ cells by replacing the organic solvent with super-concentrated LiTFSI (21 m) in H₂O. Li₂O₂ was confirmed as the only discharge product, with no LiOH formation, although H₂O is used as the solvent in this system.

Zhao et al.²⁰⁷ mixed tetraglyme with 1,1,2,2-tetrafluoroethyl 2,2,3,3-tetrafluoropropyl ether (TTE) in a 1: 1 volume ratio. This was done in order to improve the electrochemical stability, as well as to increase the ionic conductivity and wettability, reduce viscosity, and flammability. In addition, the stability of the Li metal anodes in the cells were improved due to formation of passivating surface films, thanks to the presence of the fluorinated co-solvent. In summary, this section definitely shows some light at the end of the tunnel, demonstrating interesting attempts to develop new solvents in light of understanding the mechanisms of side reactions. Also, the use of super-concentrated solutions can also help to stabilize Li-O₂ batteries.

3.4(c) Cathode materials and their stability in Li-O₂ cells.

In contrast to lithium-ion batteries, in lithium-air batteries the full mass of the storage material Li₂O₂ is fully formed and decomposed in every cycle. The role of the cathode is to provide a framework within which this can occur. The cathode must possess high electronic conductivity, high open porosity to allow for Li₂O₂ storage and O₂ and Li⁺ transport. The accessible surface area governs the true surface area discharge current and impacts the ORR; larger current densities form thinner deposits, which calls for properly balancing inner surface area and pore space. This must be accomplished while introducing minimal additional mass and volume to the battery.^{208,181} Perhaps most importantly, and yet rarely acknowledged, is the need to fill large fractions of the cathode volume with Li₂O₂ during discharge (ORR).^{208,53} In our opinion the cathodes should be able to host Li₂O₂ deposits after discharge that fill at least half of their volume while not impeding battery performance, to achieve the ambitious performance targets envisaged for the lithium-air battery. Unfortunately, most work in the field fails to consider this requirement and it is common practice to measure the capacity per unit mass of the cathode material.

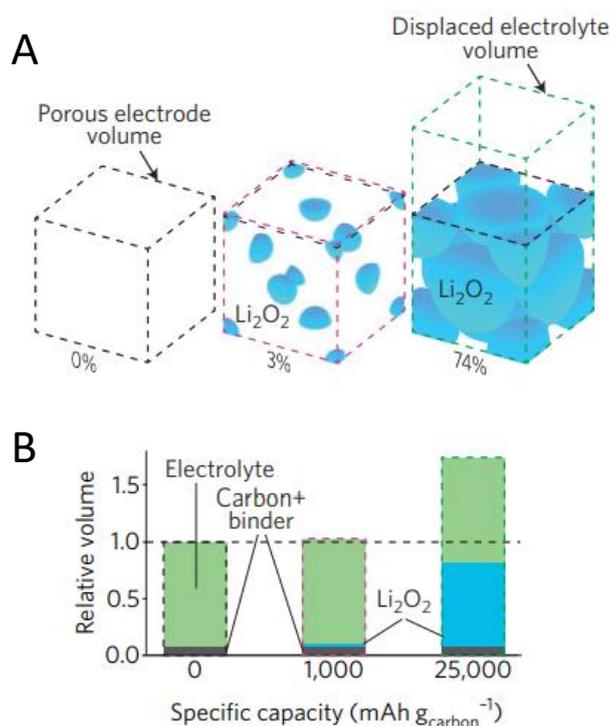


Figure 19. (A) Space filling of spherical Li₂O₂ particles inside a porous electrode and the displaced electrolyte volume at discharge capacities of 0, 1,000, and 25,000 mAh g⁻¹_{carbon}. (B) Volumes of the electrode components at these capacities normalized to the full electrode volume in the delithiated state. Reproduced with permission from ref¹⁸¹. Copyright 2017 Springer Nature.

In too many publications, the mass and volume of all other non-active components (electrolyte solution, displaced volume, binder, accumulated lithium peroxide etc.) are commonly disregarded in the calculations of specific capacity or energy. As shown in **Figure 19** and the excellent perspective article from Freunberger,¹⁸¹ such calculations mislead the true evaluation of the metrics of Li-O₂ cells. Below we attempt to review advances in cathodes in the context of their true electrochemical performance in lithium-air batteries.

Carbon cathodes and their stability

Carbon cathodes are ubiquitous in Li-O₂ batteries due to their high conductivity, low weight, low cost, and flexible structure. Typically, carbons have been selected to provide a large surface area and pore volume in which the Li₂O₂ ORR product can be hosted. Early work in the lithium-air field employed carbon blacks such as Super P, Vulcan XC 72R and Ketjen

black.^{83,209,153} These materials, while having relatively moderate surface area, possess reasonably high specific pore volume. Consequently, calculating specific discharge capacity of such carbons in cathodes of Li-O₂ cells may provide capacities >2000 mAh g⁻¹_{carbon}.²¹⁰ Ding et al. used silica templates to form a series of hierarchical carbons with varying pore sizes and showed that a general correlation exists between the pore volume and the discharge capacity.²¹¹ However, they also showed that pores of ultra-high surface area mesoporous carbons are not accessible for Li₂O₂ storage.²¹¹ Nie et al. demonstrated the importance of large macropores that remain open and able to delivery oxygen through the cathode.²¹² Graphene-based carbons have been reported to offer capacities in excess of 10,000 mAh g⁻¹_{carbon},²¹³ but there is also an unanswered question regarding the long-term stability of graphene-based cathodes. However, graphene based-materials should not be dismissed, as rigid, free standing graphene foams and papers meet many of the requirement for housing Li₂O₂ deposits (Figure 20).²¹⁴ In addition, the lack of a binder may increase stability.

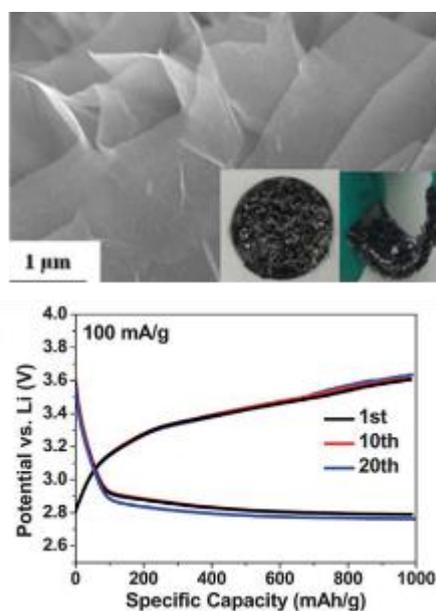


Figure 20. SEM image of a graphene foam cathode (inset optical images) and the resulting cycling profile of Li-O₂ cells when it is used as the cathode. Reproduced with permission from ref²¹⁴. Copyright 2013 Royal Society of Chemistry.

Unfortunately, most studies have failed to account for the mass of Li_2O_2 . When additional considerations of the electrolyte solution are included in the volume of highly porous cathodes few, if any, reported cathodes offer suitable performance for viable lithium-air batteries (in terms of attractive specific capacity). This is exemplified by a laudable contribution from Mitchell et al. (Figure 21)²¹⁵ where the capacity per mass of carbon only and per mass of carbon + Li_2O_2 is provided for a lithium- O_2 cell containing a CNT cathode. The CNT cathode capacities are $4720 \text{ mAh g}^{-1}_{\text{carbon}}$ versus $944 \text{ mAh g}^{-1}_{\text{carbon+Li}_2\text{O}_2}$, demonstrating the fallacy of measuring capacity as a function of carbon mass only. Zhao et al.²¹⁶ calculated the capacity of a graphene cathode to be ca. $1000 \text{ mAh g}^{-1}_{\text{carbon+Li}_2\text{O}_2}$. These variations in performance metrics have resulted in a move towards presenting real capacities and the use of fibrous carbon gas diffusion electrodes previously used in fuel cells.²¹⁷ These materials are not practical due to their large mass. But they do offer physical properties similar to those presumably required in a practical lithium-air cathode, i.e. large porous volume, large pores for O_2 delivery, reasonable surface area, and a rigid and robust 3D structure. Thus, they offer a suitable testbed. With the advent of redox mediators able to direct significant discharge via a solution route, Gao et al.²⁰⁸ have presented discharge capacity as a volume fraction, which provides a more robust measure of improvement, and paves the way for meaningful comparisons of cathode performance.

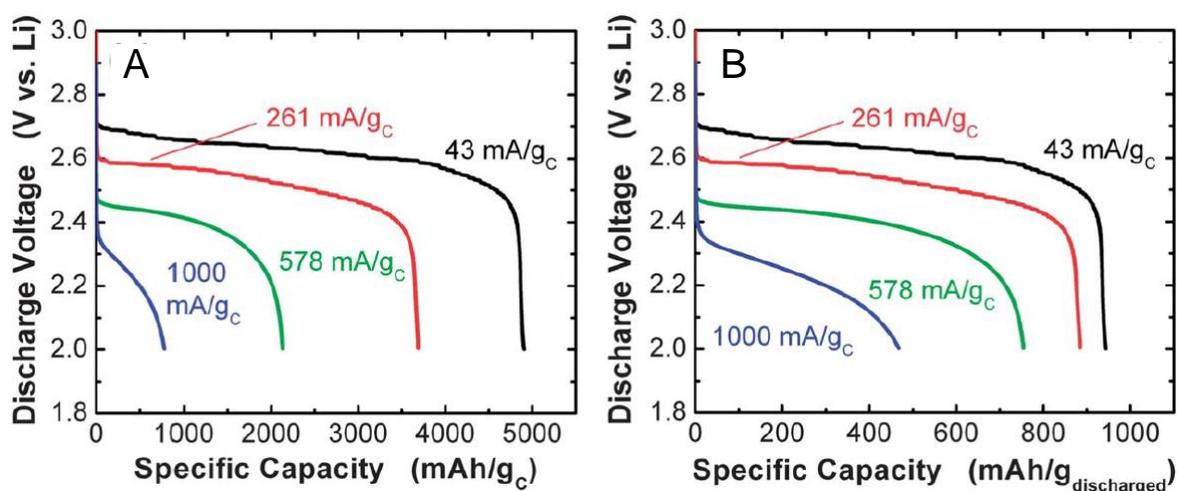


Figure 21. First-discharge profiles of Li-O₂ cells containing CNF electrodes with the capacity normalized to (A) the weight of the carbon and (B) to the carbon + Li₂O₂. Reproduced with permission from ref²¹⁵. Copyright 2011 Royal Society of Chemistry.

Stability is the Achilles' heel of carbon cathodes. Gallant et al. reported the formation of lithium carbonate at the interface between Li₂O₂ and carbon.²¹⁸ XANES spectra of discharged cathodes showed significant amounts of lithium carbonate after discharge to low capacities, which gives way to predominantly lithium peroxide as the discharge capacity increases (Figure 22). Moreover, the surface of carbon was found to be oxidized. They interpreted this as a reaction between the lithium peroxide in close proximity with the carbon followed by growth of higher purity lithium peroxide on the outer layer (Figure 22C). During subsequent cycling, they identified an accumulation of lithium carbonate, accompanied by a shift to higher charging voltages due to oxidation of lithium carbonate.

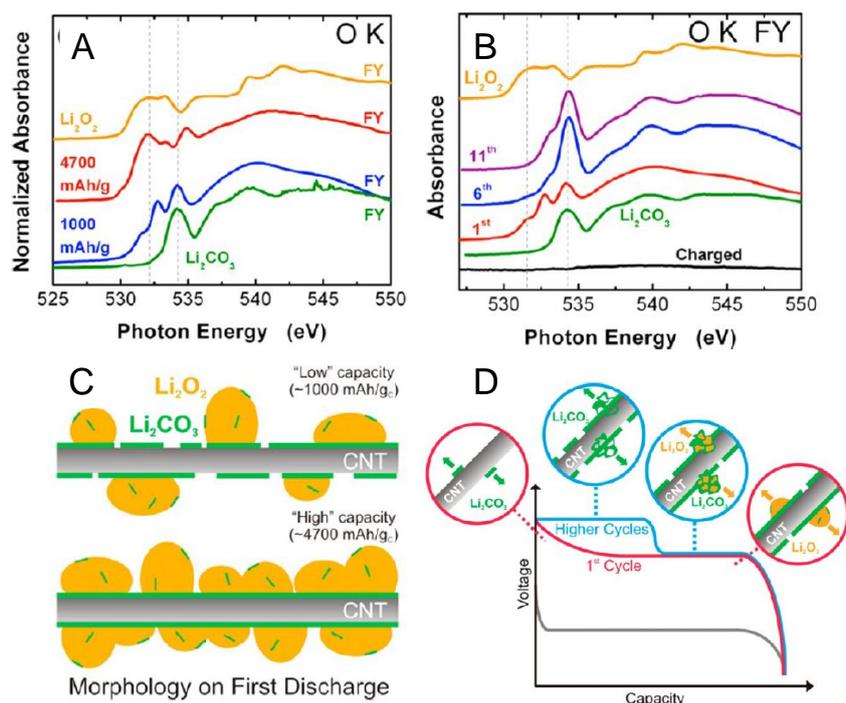


Figure 22. O K-edge FY spectra of Li-O₂ electrodes (A) discharged to 1000 (at 250) and 4700 mAh g⁻¹C (at 100 mA g⁻¹C) on the 1st discharge and (B) cycled electrodes ending on the 1st, 6th, or 11th discharge. (C) Schematic of discharge products formed at low and high capacity on carbon on the 1st discharge and (D) during 1st and higher cycle numbers and the corresponding influence on the charging voltage. Reproduced with permission from ref²¹⁸. Copyright 2012 American Chemical Society.

Thotiyl et al. demonstrated the evolution of $^{13}\text{CO}_2$ when charging a lithium- O_2 battery containing a cathode containing ^{13}C , conclusively demonstrating carbon decomposition in this battery (Figure 23).¹⁵⁴ Analysis of the cathode at various states of discharge and charge showed that carbon decomposition predominantly occurred during charging and at high voltages. Oxidation of Li_2CO_3 and CO_2 evolution was shown to occur below 4 V in accord with other studies.^{84,174} Analysis over successive cycles indicated that Li_2CO_3 from carbon decomposition accumulates in the cell. Comparison of hydrophobic and hydrophilic carbons showed that the latter offers higher stability due to the lack of reactive surface groups.¹⁵⁴ The preferential reactivity of functional groups on carbons has been observed also by others and is also relevant during the discharge processes.^{219,220}

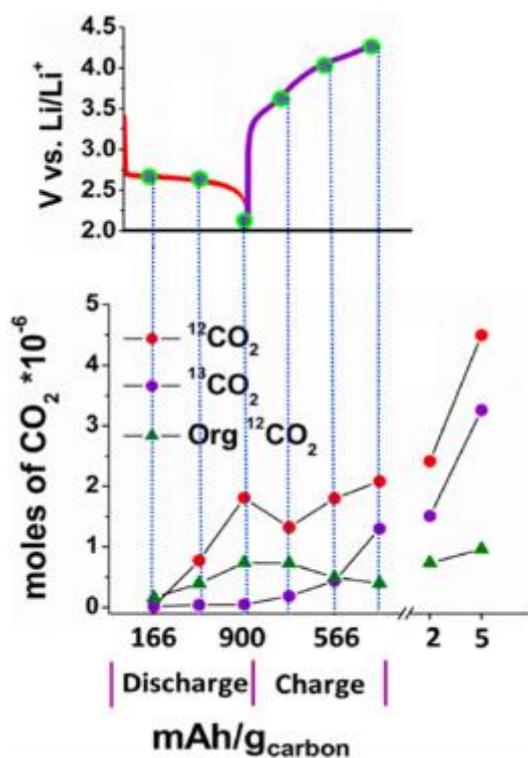


Figure 23. (upper) Discharge–charge curves on the first cycle for Li- O_2 cells containing a tetraglyme-based electrolyte and a ^{13}C cathode. (lower) Moles of CO_2 evolved during chemical treatment of the removed cathodes due to decomposition of (red) the electrolyte, (purple) the cathode, and (green) electrolyte (organic fragments) during cycling to the point indicated by the markers on the plot. Reproduced with permission from ref¹⁵⁴. Copyright 2012 American Chemical Society.

Summarizing the above results, it seemed that carbonaceous materials, despite all their promising attributes, suffered instability in lithium-air batteries. This led to a search for alternative stable materials (reviewed below). However, when using redox mediators for the charging processes of Li-O₂ batteries, a significant decrease in degradation of the battery components is observed.²²¹ Gao et al. demonstrated a dual mediator cell with both discharge and charge mediators, including cathodes containing ¹³C, and showed that the use of redox mediators increase the stability of the carbon cathodes.²⁰⁸ This is shown in **Figure 24** where the Li₂CO₃ content of discharged cathodes was analyzed ex situ. Li₂CO₃ rapidly accumulates in standard cells, but cells using dual mediators showed markedly less decomposition and accumulation.

The authors assigned this improved stability to three factors: 1) when using soluble discharge redox mediators, the products forms in solution and lacks intermediate contact with the carbon thus lowering the propensity for decomposition; 2) the redox mediators force the cells to undergo charging at 3.6 V, thus avoiding the highly positive potentials that drive carbon oxidation; 3) the redox mediators are less affected by surface area loss from passivation by Li₂CO₃ that would otherwise increase the OER overpotentials. These recent results using redox mediators rekindle the possibility of using carbon in the lithium-air cell.

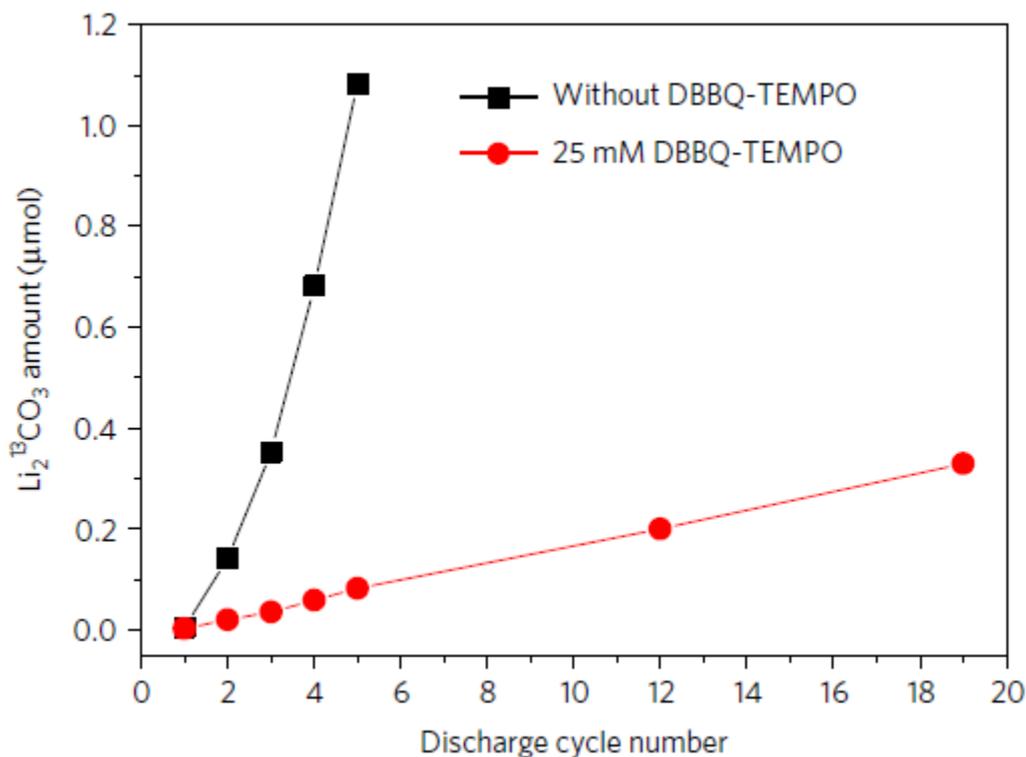


Figure 24. Amounts of $\text{Li}_2^{13}\text{CO}_3$ at the end of discharge of successive cycles of cells contain ^{13}C cathodes, with and without dual mediators. Reproduced with permission from ref ²⁰⁸. Copyright 2017 Springer Nature.

Non-carbon cathodes

Alternative stable cathode materials have been sought to replace carbon. As proof of concept, Peng et al.¹⁸⁸ showed that a cell containing an Au cathode could deliver improved cycling efficiency. In the absence of carbon, the evolution of CO_2 upon charge was much reduced, indicating less decomposition. A search for more practical alternatives identified metal carbides as promising cathode materials, in particular TiC.²²² These studies showed that the surface of TiC forms an oxide layer, which is stable during cycling. This was supported by theoretical work showing that a TiC surface in contact with Li_2O_2 would lead to termination with a stable oxide layer.²²³ Adams et al. showed that the thickness of these surface oxide layers is critically important. A complete oxide coverage can passivate the cathode and stop the charge reaction.²²⁴ However, TiC is one of the few metal carbides identified that is stable during cycling.²²⁵

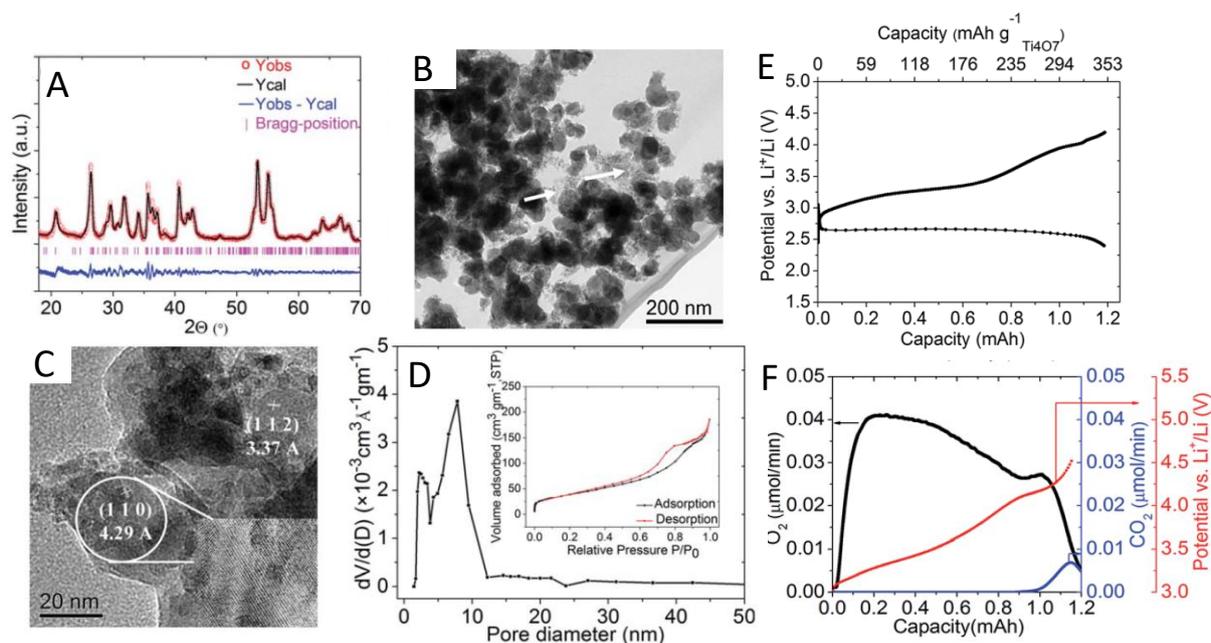


Figure 25. Characterization of a Magneli phase Ti_4O_7 cathode material including (A) XRD pattern, (B,C) TEM images and (D) pore size analysis and the (E) cycling profiles for Li- O_2 cells containing magneli phase Ti_4O_7 cathode and (F) gas analysis during cycling of the cell during charge. Reproduced with permission from ref ²²⁶. Copyright 2015 Royal Society of Chemistry.

The surface of metal carbides during cycling is predominantly oxide based, suggesting that conducting metal oxides may be considered as viable cathode materials. Kundu et al. reported the use of Magneli phase Ti_4O_7 as a cathode in the lithium- O_2 battery (Figure 25).²²⁶ The Ti_4O_7 was synthesized with a particle size of 80-100 nm consisting of smaller 10-20 nm crystallites with 2-10 nm pores. The cycling performance and accompanying on-line electrochemical mass spectrometry on charge is shown in Figures 25E and F. The overpotential during charge was below 3.5 V for the bulk of the charge process, and CO_2 release is greatly reduced. The e^-/O_2 ratio was shown to be 2.6, which is an improvement compared to what is measured with Vulcan carbon cathodes (indicating much less side reactions). A number of other metal oxides have been used in the Li- O_2 cells, such as MnO_2 ,²²⁷ Co_3O_4 ,²²⁸ Mn_3O_4 ,²²⁹ and Al_2O_3 .²³⁰ However, without careful analysis, similar to those described above, it is difficult to judge the efficacy of these materials. Other materials considered include precious metal

catalysts and metal sulfide materials, but in this review, we consider these in the context of heterogeneous catalysis.

On the stability of polymeric binders in the cathodes

The stability of polymers as binders in the cathodes was evaluated by Amanchukwu et al.²³¹ by checking their reactivity with Li_2O_2 . The polymeric binders PAN, PVC, PVdF, PVdF-HFP, and PVP were found to be unstable to Li_2O_2 , in agreement with previous studies revealing the instability of PVdF.²³² In contrast, PMMA, PTFE, and Nafion polymers were found to be chemically stable under these conditions. PEO is generally stable; however, some cross-linking can occur. The instability of polymeric binders like PVdF may complicate the analysis of composite cathodes. For instance, the side products formed by the decomposition of PVdF show Raman spectroscopic absorptions which resemble those of LiO_2 moieties; this in turn, can lead to a severe misunderstanding of the results.²³³ Hence, it is important to note that the use of unstable materials may cause errors in the interpretation of the cathodes' reaction mechanisms. This is on top of the poor stability that could obviously lead to poor electrochemical performance.

3.4(d) Negative electrodes.

As mentioned above, the Li- O_2 battery has generated much interest because of its 4 to 5-fold higher theoretical specific energy compared to conventional LIBs. This high energy density depends primarily on the use of metallic lithium as the anode, which has a very high theoretical specific capacity ($\sim 3860 \text{ mAh g}^{-1}$) and the most negative reduction potential (-3.040 V versus the standard hydrogen electrode). To date, numerous efforts have been invested in introducing lithium as a practical anode in order to meet the progressing demand of advanced, lightweight, high-energy batteries. But unfortunately, all the impressive traits of this metallic anode come at the cost of three important factors, namely, reactivity, efficiency, and safety, which impede its successful application in rechargeable energy storage devices.

The high Fermi energy level and negative reduction potential of lithium, on the one hand, offers the benefits of high operational voltage. However, these factors also make the lithium electrode extremely chemically reactive towards almost all liquid electrolyte solutions. Undesirable parasitic reactions occur as soon as the electrolyte solution is exposed to a lithium surface. These spontaneous reactions result in the formation of a fragile solid electrolyte interphase (SEI) that temporarily passivates the lithium from further reaction. The oxygen environment of the Li-O₂ system complicates matters further, as the presence of reactive oxygen species (ROS) promote parallel degradation pathways proceeding towards the accumulation of insulating by-products (Li₂O/Li₂O₂, Li₂CO₃, and various types of insoluble ionic Li compounds) on the anode surface. Additionally, during cycling, the crossover of CO₂, trace levels of H₂O (arising from electrolyte decomposition), and redox mediators (used for catalysing OER/ORR), from the cathode to the negative electrode, have a detrimental effect on the performance of Li metal electrode by direct reactions. These species also affect the composition and morphology of the SEI at the lithium/electrolyte solution interface. Recently, the effect of O₂ crossover on SEI formation has been investigated in ether-based solvents.⁶⁴ With the help of in-situ XRD and computational studies, Assary et al.⁶⁴ have revealed the formation of LiOH, aldehydes, and carbonates at the anode. These impede the migration of lithium ions, leading to increased charge transfer resistance and failure of the electrode (Figure 26).

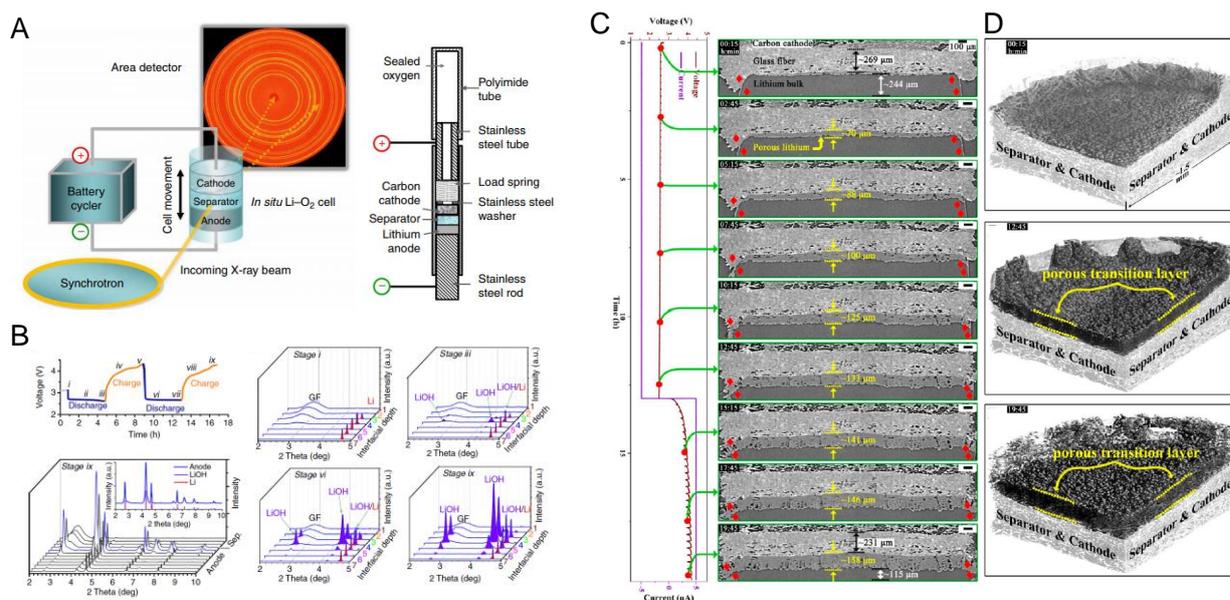


Figure 26. Observation of Li metal degradation in Li-O₂ batteries. (A) The experimental setup for operando XRD and schematic design of the cell. (B) The change of lithium anode during discharge-charge cycling in Li-O₂ batteries. XRD patterns exhibit LiOH formation during cycling tests. (C) Electrochemical and morphological results of cell during discharge and recharge process in situ measured at EDDI beamline. (D) Selected 3D demonstrations of the reconstructed data sets shown in C with time stamps (top) 00:15, (middle) 12:45, and (bottom) 19:45. Reproduced with permission (A, B) from ref. ²³⁴, copyright Springer, 2013. (C) from ref. ²³⁵. Copyright 2019 American Chemical Society.

In general, the formation of the SEI and other side reactions between lithium metal and crossover contaminants, irreversibly consumes lithium and the electrolyte solution - which results in limited coulombic efficiency (CE). If the formed SEI is self-limiting and stable throughout the cycle life, the CE of the Li-anode will be greatly improved after the first cycle. But, unfortunately, with repeated stripping/plating, the morphology of the lithium surface changes and the natural SEI undergoes repeated rupture and repair cycles. This leads to both undesirable electrolyte solution consumption, and inhomogeneous and uncontrolled deposition of Li as sharp needles or in branch like patterns commonly known as dendrites. These severely reduce the CE as illustrated in [Figure 27](#). These dendrites act as hotspots and in further cycles, become the preferred site for Li-ion deposition because of fast reaction kinetics and shorter diffusion lengths. This results in their continuous growth, which not only leads to large volumetric and morphological changes, but also creates a safety issue by short circuiting the

cell. This poses a threat of thermal runaway and explosion. To deal with the low CE of the lithium anode, and dendritic growth at high current densities, researchers typically carry out experiments using a large mass of lithium. Consequently, technical breakthroughs focusing on the use of a massive excess of lithium or its complete replacement by a stable substitute, ignore the consequences of the proposed protocols in a full practical Li-O₂ cells and are of less practical merit. It is important to emphasize that to attain practical goals, a protected Li anode is needed for Li-O₂ and many other Li metal batteries.

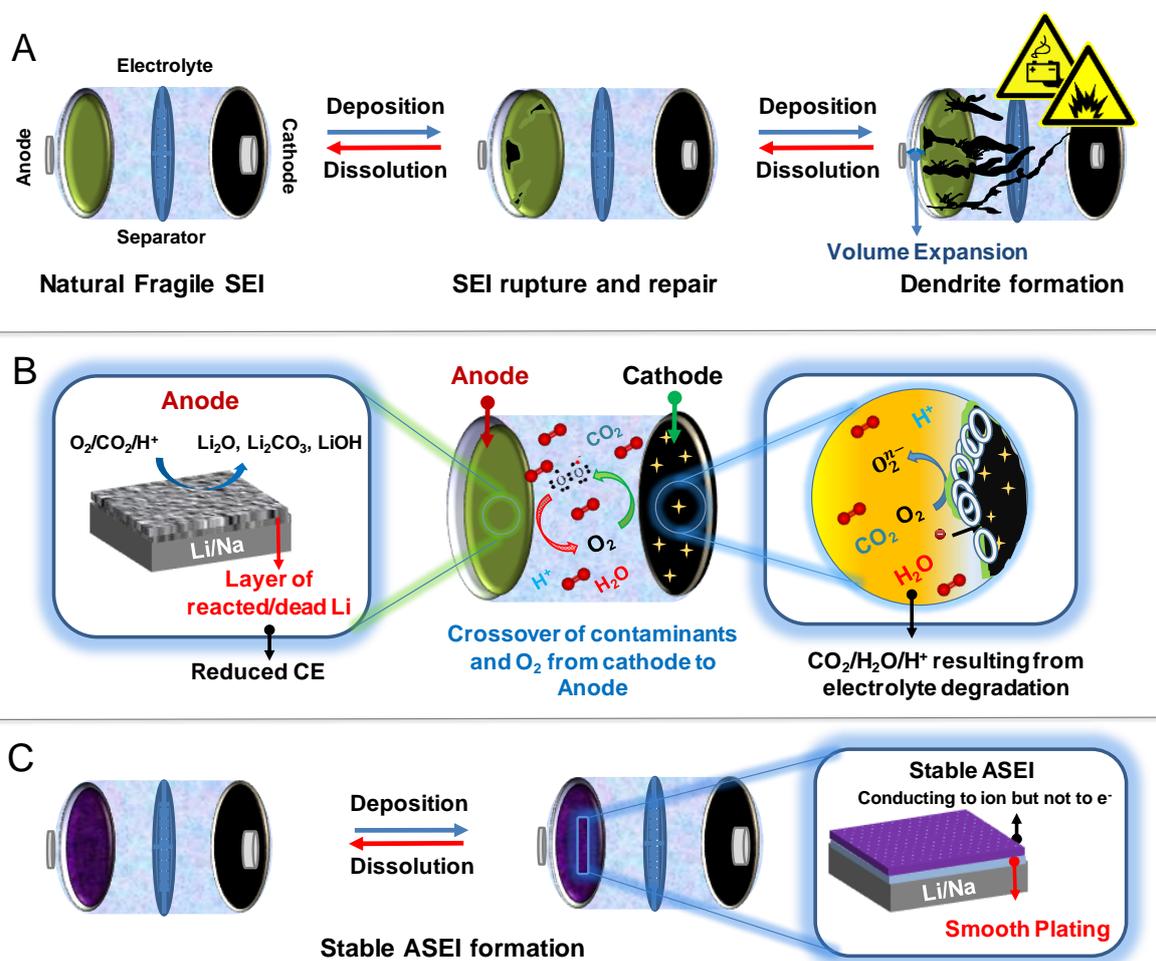


Figure 27. Graphic representation of the mechanistic issues associated with the use of metallic anodes. The figure illustrates: (A) dendrite formation and concomitant safety issues originating from the rupture and repair of the fragile solid electrolyte interface (SEI); (B) effect of O₂ crossover on the coulombic efficiency (CE); and, (C) the benefits of an artificial SEI (ASEI) in stabilizing the metal anode.

Nevertheless, despite the formidable challenges associated with metallic lithium, the recent development of the Li-O₂/air battery has undoubtedly hastened the possible discovery of practical solutions to surmount these challenges. Specifically, in the field of Li-O₂/air batteries, where metallic lithium is an indispensable component, researchers have utilized three approaches to mitigate the various issues: (1) Stabilizing the SEI - either by designing the anode/electrolyte interface (artificial SEI) and/or by optimizing the salt-solvent coordination. The latter can be accomplished by using super concentrated electrolyte solutions with fewer free solvent molecules, and an ionic liquid-like bulk nature; (2) developing selective membranes to restrict the crossover of O₂, CO₂, redox mediators, and other contaminants; and, (3) engineering the metallic Li anode structure to overcome morphology and volume changes.

The first approach relies predominantly on controlling the composition of the SEI which can, in turn, stabilize the lithium metal by suppressing the parasitic side reactions, and, by controlling the metal ion flux.^{236–241} Many kinds of approaches have been studied for direct coating of Li metal anodes by artificial protective layers to avoid contact of Li metal with reactive electrolyte solutions. Among these, liquid or gas treatments have been reported.^{242–244} In addition, the use of electrolyte solution additives or co-solvents like FEC were adapted for elaborating a stable artificial SEI on Li metal electrodes.²⁴⁵

It is now well known that a smartly designed interface can ensure uniform metal deposition and high coulombic efficiency by reducing rupture/repair of reinforced interfaces. Recently, Kozen et al.²⁴⁶ experimentally demonstrated the success of this approach by designing a hybrid organic/inorganic artificial SEI on a lithium surface, which exhibited stable cycling over 110 cycles at a current density as high as 2 mA cm⁻². Additionally, by optimizing the electrolyte composition, the solubility and diffusion of ions, gases, and other contaminants can be controlled. This is accomplished by exploiting the different solubility of reaction intermediates/discharge products in the different solvents or salts depending on their Guttmann

number and ionic strength. This can be clearly inferred from the stabilizing ability of super concentrated electrolyte solutions based on LiTFSI-DMSO (1:3 molar ratio)¹⁹², 3 M LiTFSI-DME¹⁴⁶, and 4 M LiNO₃-DMSO.²⁴⁷ With the above compositions, it has been demonstrated that Li-O₂ batteries with super concentrated solutions exhibit significantly enhanced cycle life with a CE > 90. Enhanced reversibility of Li-O₂ batteries containing an advanced Li salt, LiTNFSI, for formation of a stable SEI layer, was also reported by Tong et al.²⁴⁸ The use of gel or solid electrolytes instead of standard liquid solutions are also known to help in eliminating side reactions and suppressing dendrites formation (Figure 28). However, solid electrolytes are associated with low migration kinetics of the Li ions.

Another way to prolong the cycle life of the negative electrode is to utilize membranes with selective permeability, or functionalized separators.^{59,249–251} These can create barriers to the crossover of contaminants that accelerate the deterioration responsible for lithium instability. By restricting the gases and radical species reaching the lithium surface, such membranes can attenuate the multiple degradation pathways and accumulation of side products on the metal interface. This approach is particularly important for the Li-air system, where the cathode's active mass forms highly reactive moieties on reduction. In these systems, it is important to protect the Li anode from all other active atmospheric contaminants like N₂, CO₂ and moisture that can instantly corrode the negative electrode and shorten the battery life. The use of bi-compartment cells in which the Li anodes and the carbon cathodes can operate in the most appropriate liquid electrolyte solutions (relevant for them) separated by a membrane (ceramic or polymeric) which can transport only Li ions from side to side were proposed in order to limit the cross-talk. The bi-compartment configuration (Figure 28 B) fully prevents detrimental crossover of reactive products between the reactive electrodes. While such a configuration may not be practical due to the relatively low specific energy of these cells, nevertheless, it is excellent for systematic studies of the behavior of components in Li-O₂ cells

without interference of side reactions. For instance, with such cells it was possible to study separately the effectiveness of redox mediators at the cathode side and effective ways to stabilize the Li metal anode.²⁵² Alternatively, functional polymer membranes that fully block detrimental crossover between the electrodes can be used in similar bi-compartment Li-O₂ cells. Kim et al. reported on the use of a polyurethane based modified separator for stable and reversible Li-O₂ batteries.²⁵³ From the viewpoint of the cell's energy density, using modified separators in Li-O₂ cells is more advantageous than using bi-compartment cells separated by ceramic membranes.

In addition to the above described problems and solutions (related to the reactivity of Li metal anodes), there is a great need to accommodate the massive volume expansion observed on stripping/plating of lithium. In the past few years, material scientists have explored the problem of volume expansion and dendritic growth via 3D architecture hosts (such as reduced graphene oxide, and silicon nanostructures) to accommodate Li deposition.^{240,251,254,255} Due to the availability of free spaces and voids, such structures can control volume expansion and suppress dendritic growth - with an additional advantage of restricted contact between Li and the electrolyte, leading to fewer side reactions and high cycling efficiency. However, these 3D architectures help in accomplishing safe operation of LOB's at the expense of decreasing energy density.

The growing attention and the increasing number of papers focussing on the lithium electrode gives a good measure of the importance of the metallic negative electrode in LOBs. Despite some progress, the achieved cyclability and CE is still far away from the set milestones. A better understanding of the surface phenomena on Li metal anodes (surface reactions, the structure of the surface films) seems to be critical for enabling high-efficiency deposition of metallic Li in the corrosive environment developed in Li-O₂ cells.

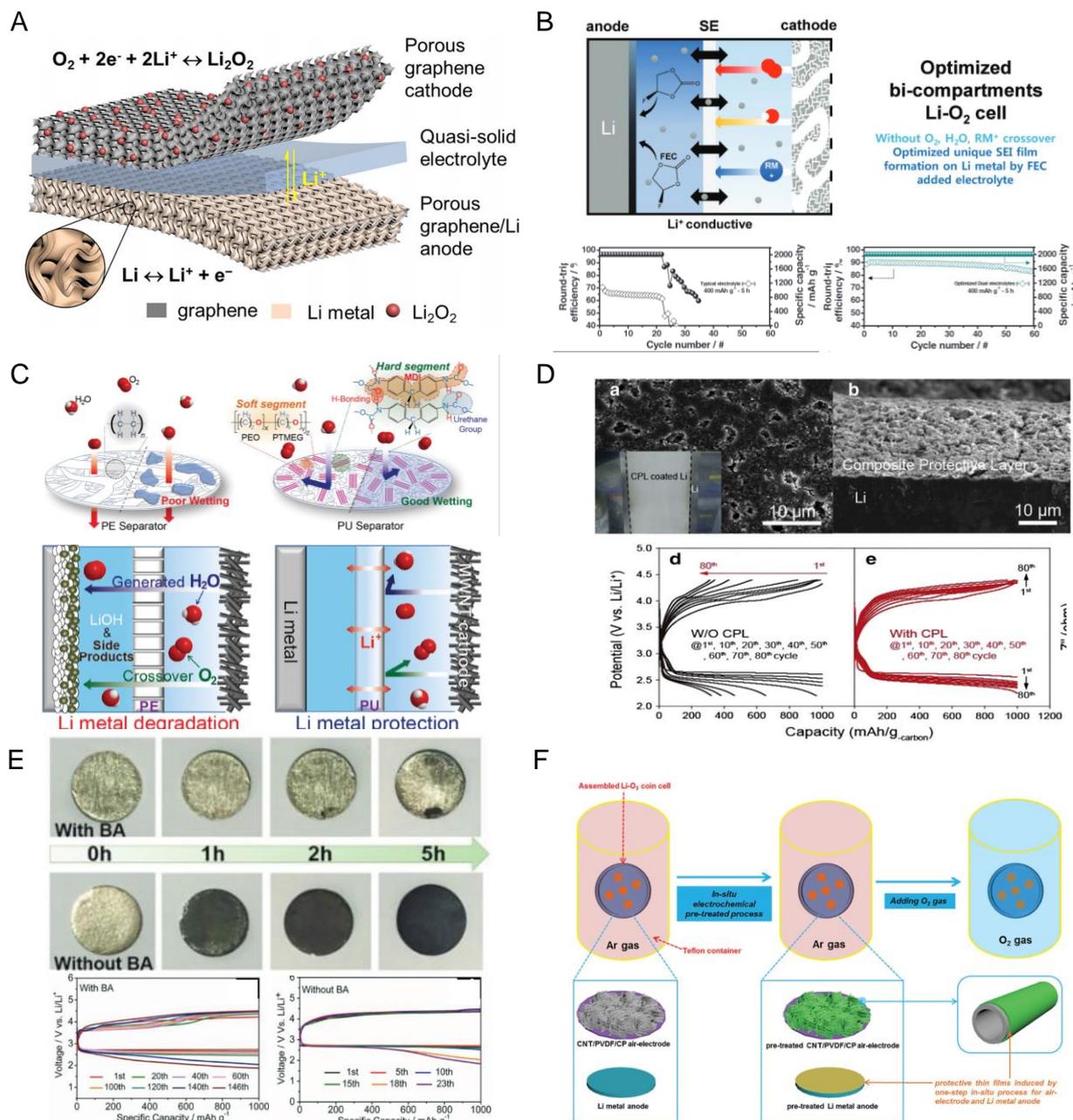


Figure 28. Representative methods for protecting Li metal in Li-O₂ batteries. (A) Gel or solid electrolyte. Reproduced with permission from ref.²⁵⁶. Copyright 2018 Springer. (B) Separation of electrolytes by lithium conductive membrane. Reproduced with permission from ref.²⁵². Copyright 2017 John Wiley & Sons, Inc. (C) Separator. Reproduced with permission from ref.²⁵³. Copyright 2016 John Wiley & Sons, Inc. (D) Coating composite materials on Li metal. Reproduced with permission from ref.²⁵⁷. Copyright 2014 Elsevier. (E) Pre-electrochemical treat. Reproduced with permission from ref.²⁵⁸. Copyright 2018 John Wiley & Sons, Inc. (F) Additive for protection of Li. Reproduced with permission from ref.²⁴⁴. Copyright 2018 John Wiley & Sons, Inc.

While Li anodes are the first choice, nevertheless, several studies have focused on alternative anodes including lithiated Si, Sn, Al, and carbons. Lithiated Si anodes were

predominantly used to satisfy the required anode capacity. These alternative anodes are electrochemically or chemically lithiated before they are used in the Li-O₂ cells (Figure 29A).

It was generally realized that Li-O₂ cells using alternative anodes behave less reversibly than typical Li-O₂ cells containing Li metal anodes (Figure 29B). This is because these lithiated alternative anodes contain a much lower amount of Li (Figure 29C), which quickly deteriorates and becomes inactive due to side reactions (Figure 29D).²⁵⁹ Hence, the use of alternative lithiated anodes (like Li-Si) may also require their surface protection.

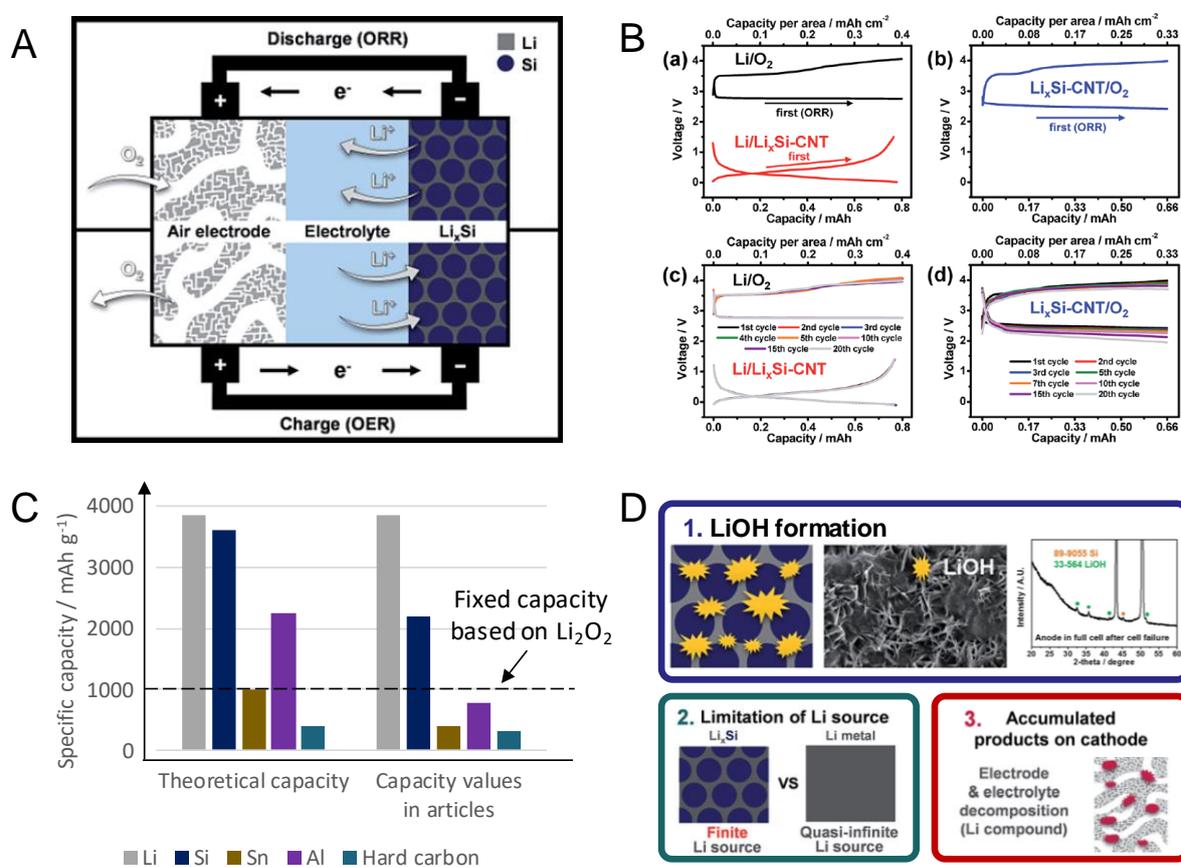


Figure 29. Alternative electrodes for Li metal in Li-O₂ batteries: (A) Schematic images for the discussion of the Li-O₂ full-cell with lithiated silicon as the negative electrode. (B) Galvanostatic cycling test results of half-cells and full-cells. (C) Theoretical and practical capacities of Li and alternative electrodes. Capacity for formation and decomposition Li₂O₂ is fixed as 1168 mAh g⁻¹ based on the reaction (2Li⁺ + O₂ + 2e⁻ ↔ Li₂O₂). The values in the chart come from ref. ^{254,260–263}. (D) Major issues for the Li-O₂ full-cell related to Li loss. Panels A, C and D were reproduced with permission from ref. ²⁵⁹. Copyright 2016 Royal Society of Chemistry.

From an experimental/didactic perspective, alternative anodes can be useful in experiments concentrating on the evaluation of the reversibility of cathodes.²⁶⁴ However, for practical purposes, the use of protected Li-metal as the anode is inevitable.

3.5. Catalysis.

3.5(a) Heterogeneous catalysis for the ORR and OER.

Heterogeneous catalysis (or electrocatalysis) is a contentious topic in the lithium-air battery field. Firstly, there is a reasonable argument that any solid catalyst should be of a limited use because it should be quickly passivated by insulating lithium peroxide, or only catalyze oxidation of lithium peroxide in direct contact. The use of redox mediators designed to avoid contact between the cathode surface and the bulk of the lithium peroxide only exacerbates this problem. Secondly, there is a strong evidence that the introduction of a heterogeneous catalyst induces decomposition. For example, during the early development of the Li-O₂ battery, MnO₂ catalysts were thought to offer improved performance,²²⁷ but have since been shown to induce significant decomposition of the electrolyte.⁸⁵ It should be noted that these earlier studies generally used carbonate-based electrolytes, which are now known to be unstable. McCloskey et al. studied the impact of a series of heterogeneous catalysts including Pt, MnO₂, Au and showed, through analysis of gas evolution, that both Pt and MnO₂ induced considerable decomposition and an increase in CO₂ evolution (Figures 30A-C).⁸⁵ Neither Au nor MnO₂ appeared to demonstrate any catalytic activity, although we highlight that the primary purpose of Au cathodes is stability. Studies by Ma et al. have similarly demonstrated that both Pd and Ru based catalysts increase the degree of irreversibility in Li-O₂ cells.²⁶⁵ Genorio et al. studied oxygen reduction in dimethoxyethane at well-defined single crystal electrodes (Figure 30).¹⁴³ A decrease in the electrochemical reversibility of the redox couple is observed in the order Au > Pt > Ir and chemical analysis of the electrode surface showed formation of a solid electrolyte interphase in the case of Pt and Ir. This is shown schematically in Figure 30E. These data

suggest that the only transition metal suitable for use in lithium-air cathodes is Au and makes any use of a metal catalyst questionable.

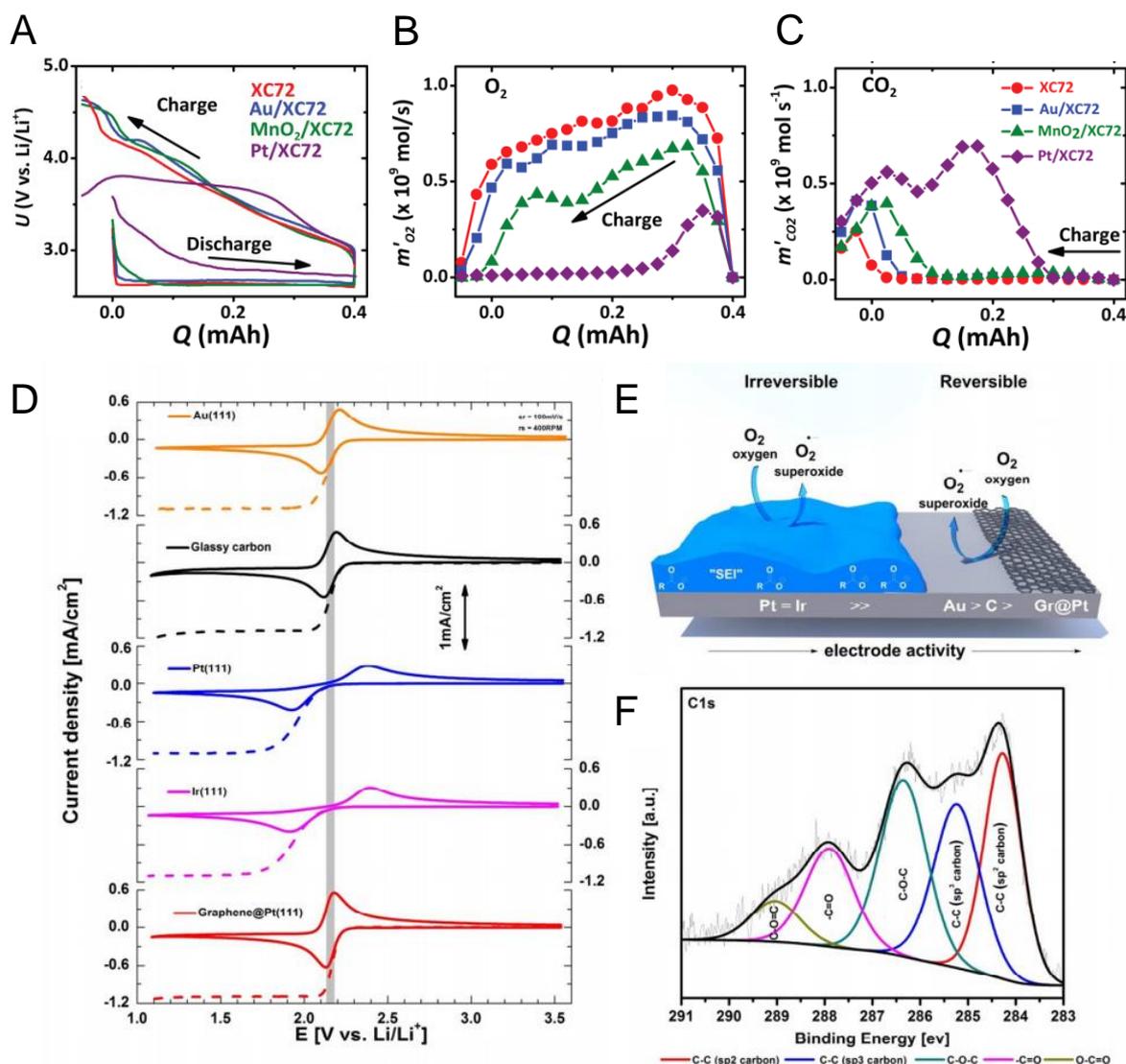


Figure 30. (A) Cycling profile of Li-O₂ cells containing DME and various cathode catalysts and the corresponding (B) O₂ and (C) CO₂ evolution during charging. Reproduced with permission from ref.⁸⁵. Copyright 2011 American Chemical Society. (D) Cyclic voltammetry (sweep rate; 100 mV s⁻¹) and polarization curves by rotating disk electrode voltammetry (rotation rate; 400 rpm) of O₂ electrochemistry in dry 0.15 M TBAPF₆/DME saturated with 10% O₂ in Ar at various single crystal metal electrodes and carbons. (E) Schematics of solid electrolyte interface (SEI) formation with corresponding electrode activity increase. (F) X-ray photoelectron C 1s spectra of the SEI formed during O₂ electrochemistry in DME on Pt(111). Reproduced with permission from ref.¹⁴³. Copyright 2016 American Chemical Society.

3.5(b) Homogenous catalysis with redox mediators

Due to the problems associated with the use of heterogeneous catalysts, mobile homogenous catalysts - dubbed “Redox Mediators” (RM) – were developed for rechargeable aprotic Li-O₂ batteries. These RMs act as mobile charge carriers in the electrolyte solution, which facilitate electron transfer between the Li₂O₂ and the electrode surface. During charging, the RM is oxidized to RM⁺ at the electrode surface and is then able to diffuse in the electrolyte where it oxidizes the solid Li₂O₂, resulting in oxygen evolution and regeneration of the RM (Figure 31A). This greatly lowers the charging potential during the cycling (Figure 31B, C). The decreased charging potential not only increases the energy efficiency, but also promotes cell durability.²⁰⁸

The first generation of mediators were reported by Giordani et al.²⁶⁶ and Chen et al. who demonstrated the concept with LiI and tetrathiafulvalene (TTF), respectively.²⁶⁷ Li-O₂ cells using TTF in the electrolyte solution exhibited suppressed charge potential, improved rate capability and cyclic stability. Since these earlier studies, several other charge mediators have been proposed.^{268,269} Bergner et al. demonstrated the use of 2,2,6,6-tetramethylpiperidinyloxy (TEMPO),²⁷⁰ and achieved enhanced cycle life compared to LiI or TTF. This work was extended to include other structurally similar mediators based on the nitroxyl group.²⁷¹ 10-methylphenothiazine (MPT)²⁷², tris[4-(diethylamino)phenyl]amine (TDPA)²⁷³, 2-phenyl-4,4,5,5-tetramethylimidazoline-1-oxyl-3-oxide (PTIO)²⁷⁴, transition metal complexes (metal porphyrins and metal phthalocyanines)^{275–277}, and various other organic materials have also been proposed as potential redox mediators for oxidation of Li₂O₂. While LiI has largely been abandoned as a mediator for Li₂O₂ oxidation, it has found use in lithium-oxygen cells where LiOH is the discharge product.¹⁰⁰

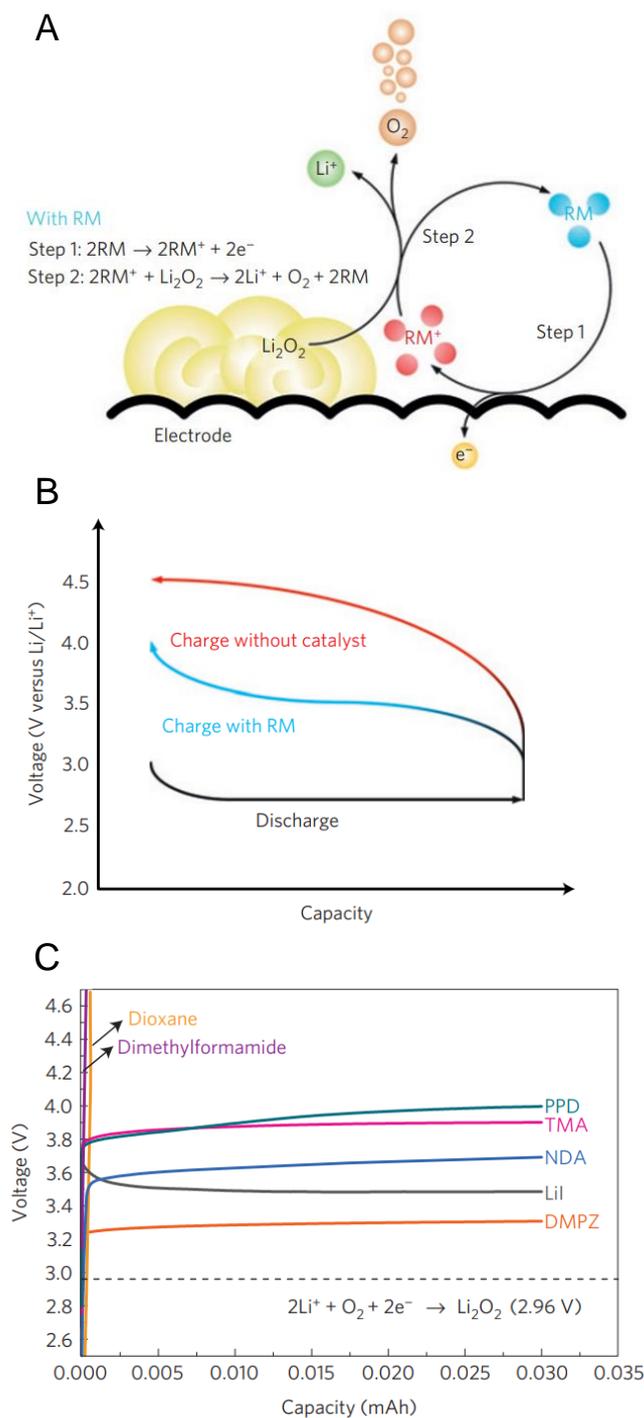


Figure 31. Redox mediators as soluble catalysts for decomposition of Li_2O_2 during recharge. (A) Schematic illustration of the reaction mechanism of redox mediator for decomposition of Li_2O_2 during charge. (B) Schematic discharge (black line) and charge profiles of Li– O_2 batteries with (blue line) and without RM (red line). (C) Charge curves with some examples of RMs. Molecular orbital energies of redox mediators and TEGDME. HOMO and LUMO energies of original RMs (black bar) and first oxidized RMs (red bar) in TEGDME electrolyte based on DFT calculations. Adapted with permission from ref. ²⁷⁸. Copyright 2016 Springer.

OER RMs were mostly evaluated for their charging potential and partly for the electron-to-O₂ ratio. The kinetics of mediated Li₂O₂ oxidation has only recently received attention.^{279–281} Chen et al. studied a variety of organic RMs for the kinetics of Li₂O₂ decomposition using scanning electrochemical microscopy (SECM) and demonstrated that the rate is highly dependent the RM used (Figure 32).²⁷⁹ Among the studied RMs, nitroxides including TEMPO were found to be the fastest. Notably, no clear ratio between heterogeneous rate constant of the RM at Au and C electrodes and the apparent rate constant k_{app} for Li₂O₂ oxidation was found. Equally, no simple proportionality between RM potential and k_{app} was found. Both were attributed to dominating steric effects on k_{app} . An important conclusion was that even the slowest mediators allow for charging rates in porous electrodes that by far exceed practically required values. Importantly however, the detailed mechanism of mediated oxidation and particularly pathways therein that cause electron-to-O₂ ratios to differ from the ideal of two are unknown.

Just as RMs for OER help the decomposition of Li₂O₂ during charging, there are compounds that can serve as mediators for ORR, assisting in the formation of Li₂O₂ deposits during discharging.^{282–287} The first generation of discharge mediators was reported by Lacey et al.²⁶⁹ They used ethyl viologen ditriflate as a redox shuttle to transfer electrons from the electrode surface to oxygen, which then forms O₂^{•-} followed by disproportionation to Li₂O₂. While these results were promising, due to the simple shuttle mechanism, this mediator operates at the same potential as the ORR reaction, did not eliminate parallel O₂ reduction at the cathode and hence the intrinsic problem of electrode passivation on discharge persisted.

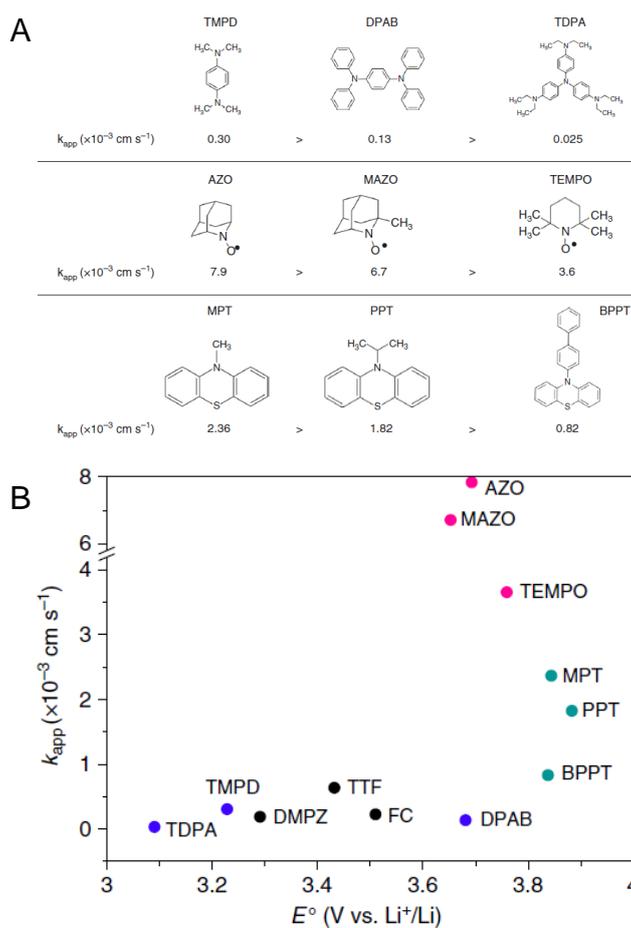


Figure 32. Li_2O_2 oxidation kinetics of redox mediators. (A) Structures of a selection of oxidation mediators and their kinetics of Li_2O_2 oxidation. (B) Comparison of the apparent rate constants (k_{app}) for the reaction between the redox mediators and Li_2O_2 grouped by structure. Reproduced with permission from ref.²⁷⁹. Copyright 2018 Springer.

In contrast, the mediator 2,5-di-tert-butyl-1,4-benzoquinone (DBBQ) as reported by Gao et al.²⁸⁶ raises the voltage of mediated above direct O_2 reduction at the cathode. Unlike previous ORR mediators, DBBQ operates as a catalyst, altering the discharge route, **Figure 33(A)**. Following reduction, DBBQ formed $\text{Li}^+\text{-DBBQ-O}_2$ in solution instead of the reactive intermediate LiO_2 . This process occurs at voltages positive of those needed for direct ORR and thus the problematic oxygen redox chemistry that limits lithium-air capacity is removed. As a result, the cell capacity increased 80- to 100-fold with a higher discharge voltage, in comparison to the case without DBBQ (**Figure 33A**). Vitamin K2²⁸³, a biological quinone, has

been reported to work in a fashion similar to DBBQ. As discussed previously, protic additives such as H₂O can improve the discharge performance. Inspired by this, Gao et al demonstrated that phenol is able to improve ORR in Li-O₂ batteries by solvating Li₂O₂ and promoting growth of large Li₂O₂ particles, while avoiding surface passivation.²⁸⁷ The result is a greatly enhanced capacity (35-fold increase) as shown in **Figure 33B**.

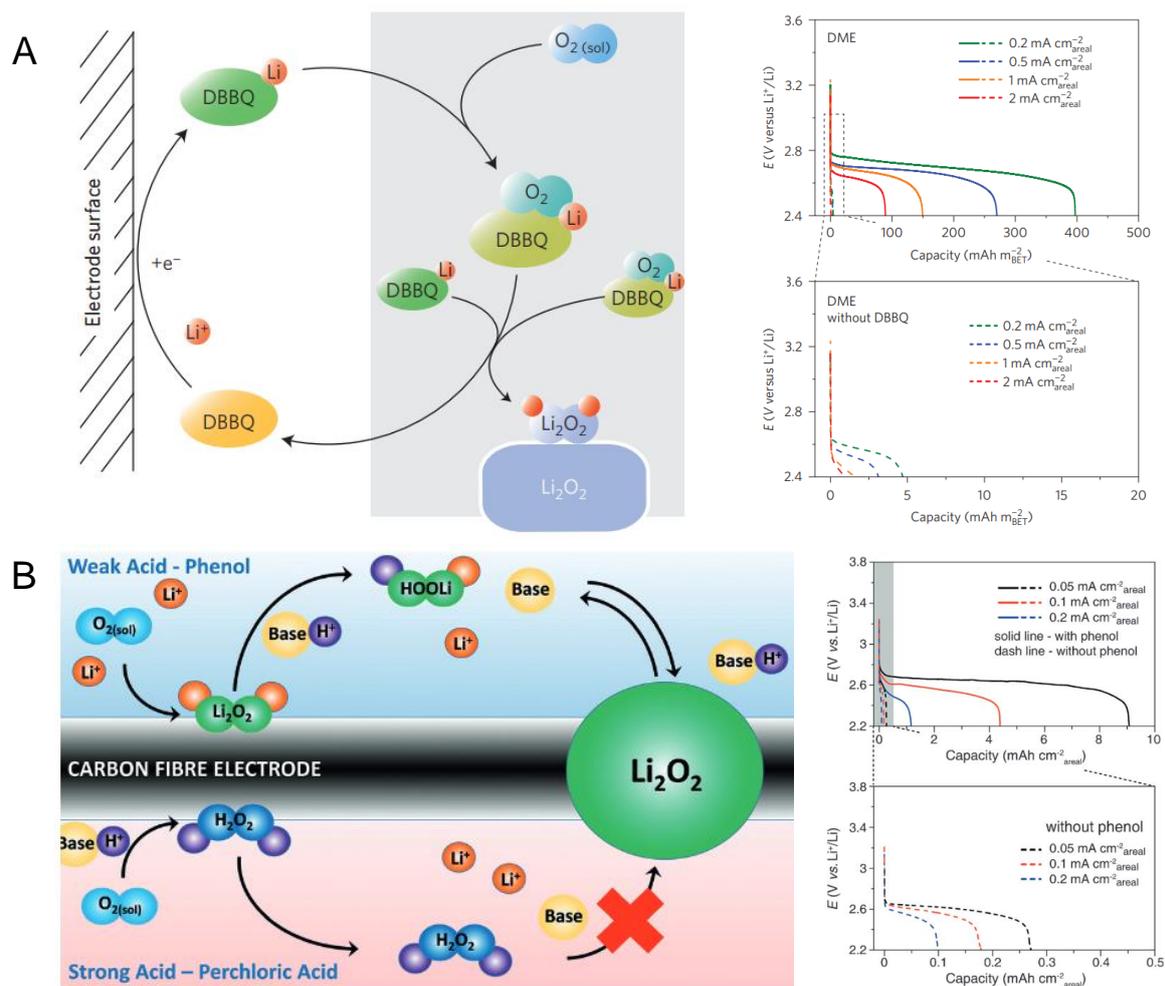


Figure 33. Li₂O₂ formation mediators in Li-O₂ batteries. (A) DBBQ. Reproduced with permission from ref. ²⁸⁶. Copyright 2016 Springer. (B) Phenol. Reproduced with permission from ref. ²⁸⁷. Copyright 2017 John Wiley & Sons, Inc.

In order to facilitate Li₂O₂ formation and decomposition during both discharge and recharge processes in Li-O₂ cells, several bifunctional red-ox mediators were suggested and tested.^{273,274,288,289}

The combined use of ORR mediator and OER mediator at the same time has shown synergistic effects.^{208,290,291} Gao et al demonstrated a cell containing DBBQ and TEMPO as discharge and charge mediators respectively, as shown in **Figure 34**.²⁰⁸ This dual redox mediator cell operates without any direct oxygen redox chemistry at the electrode surface and thus was able to significantly suppress parasitic chemistry and enhanced cycling stability.

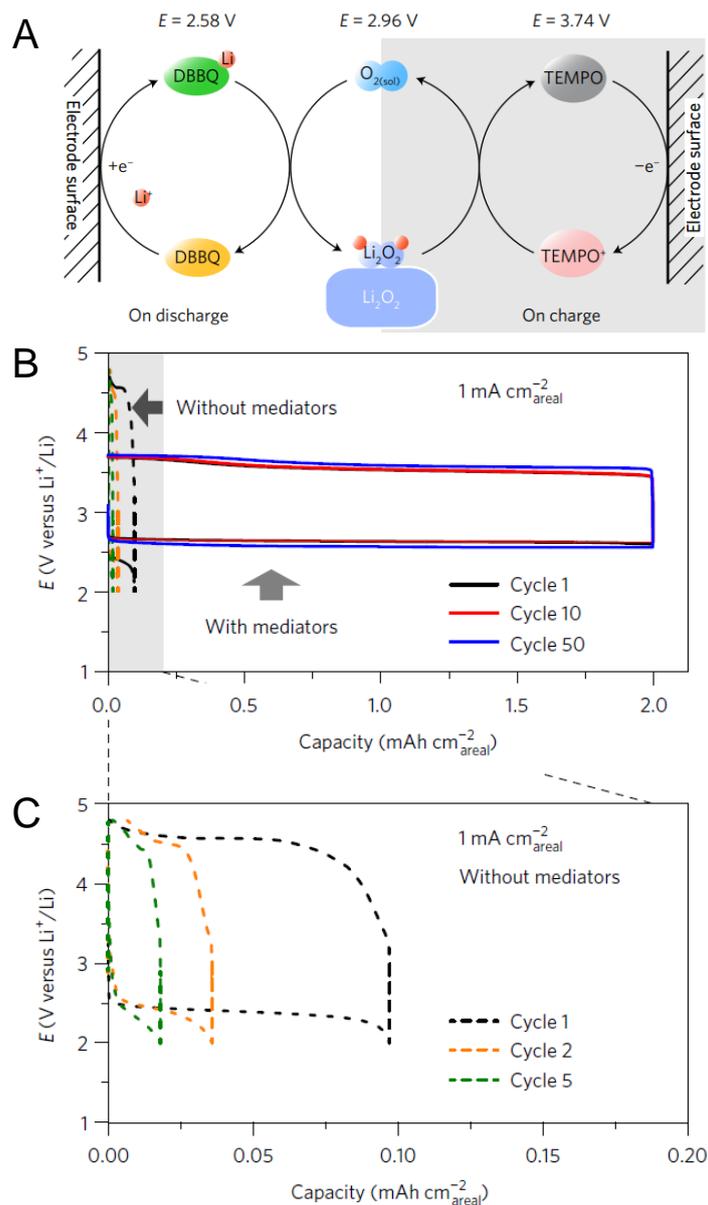


Figure 34. Dual mediators for ORR and OER in Li-O₂ batteries. (A) Schematics of positive electrode reactions on discharge and charge in the presence of DBBQ and TEMPO. (B) Discharge-charge curves of Li-O₂ batteries with (solid lines) and without mediators (dashed lines). (C) Enlarged section of the discharge-charge curves recorded without DBBQ-TEMPO in (B). Reproduced with permission from ref.²⁰⁸. Copyright 2017 Springer.

Building on this concept, Liang et al. exploited the mobile nature of the discharge redox mediator to promote growth of lithium peroxide away from the cathode, thus avoiding degradation and cell failure (Figure 35).²⁹¹ This is achieved by inserting an electrolyte soaked inert storage layer between the electrochemically active carbon cathode and the electrolyte-O₂ interface. The electrolyte contains both ORR and OER RMs, DBBQ and TEMPO in the present example. At the stationary state, on discharge, a retarded O₂ supply through the storage layer causes O₂ to be fully reduced within the storage layer by the outward-diffusing reduced form of the ORR RM. This way the carbon cathode only serves to reduce (oxidize) the ORR (OER) RM but does not perform O₂ electrochemistry.

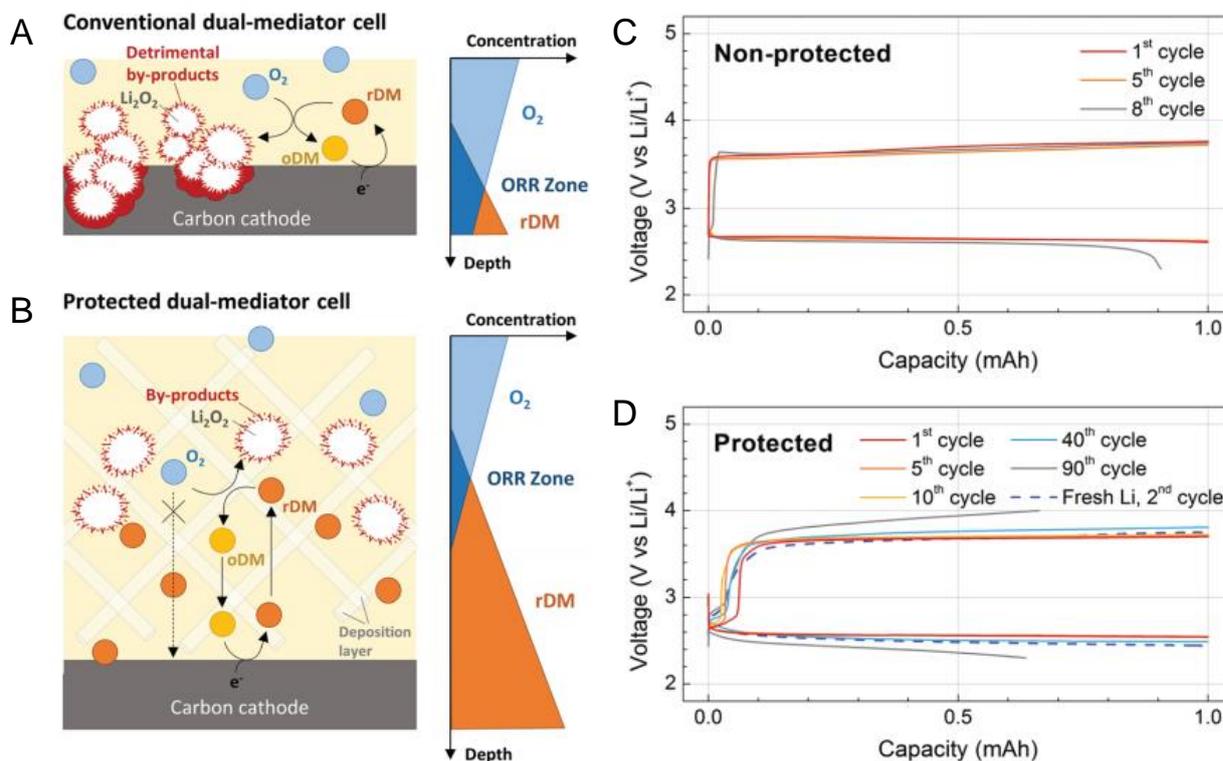


Figure 35. Strategy to protect the carbon electrode from Li₂O₂ using dual mediators in Li-O₂ batteries. During discharge in (A) a conventional dual-mediator cell, the carbon cathode is corroded by reactive oxygen species and passivated by the by-products. By contrast, the cathode in (B) a protected dual-mediator cell is isolated from both the Li-O₂ reaction intermediates and products by a dynamic O₂ shield consisting of reduced dual mediator (rDM). Comparison of cycling stability of the (C) non-protected cell and (D) protected cell. Reproduced with permission from ref.²⁹¹. Copyright 2018 Royal Society of Chemistry.

Instability issues with mediators

Despite their advantages, RMs introduce further complexity to the cell and not least potential instability. First, they may themselves not be stable with Li metal or may introduce a redox shuttle that curbs coulombic efficiency. Second, as the RMs are in many cases organic molecules, they may be subject to similar decomposition as the electrolyte or carbon.

Redox mediators can cross over to the reactive lithium metal anode due to their mobile nature. Such a phenomenon is referred to as a “redox shuttle”^{272,292,293} and could be detrimental to lithium metal, the redox mediator itself, and cycling stability. Similarly, oxidized redox mediators formed during the Li–O₂ cell charging step can diffuse to the anode site and be reduced by reaction with the lithium metal – which constitutes self-discharge (Figure 36). Ha et al.²⁹⁴ also investigated the failure mechanism of the Li–O₂ cell containing o-methylphenothiazine (MPT). They showed that the redox mediator decomposes the lithium oxide layer within the Li metal SEI through the same catalytic reaction mechanism. This exposed Li metal then directly reacts with the redox mediator resulting in irreversible decomposition, and loss of Li metal. Due to these problems, it is accepted that the lithium metal surface will need to be protected with an artificial SEI or solid electrolyte.

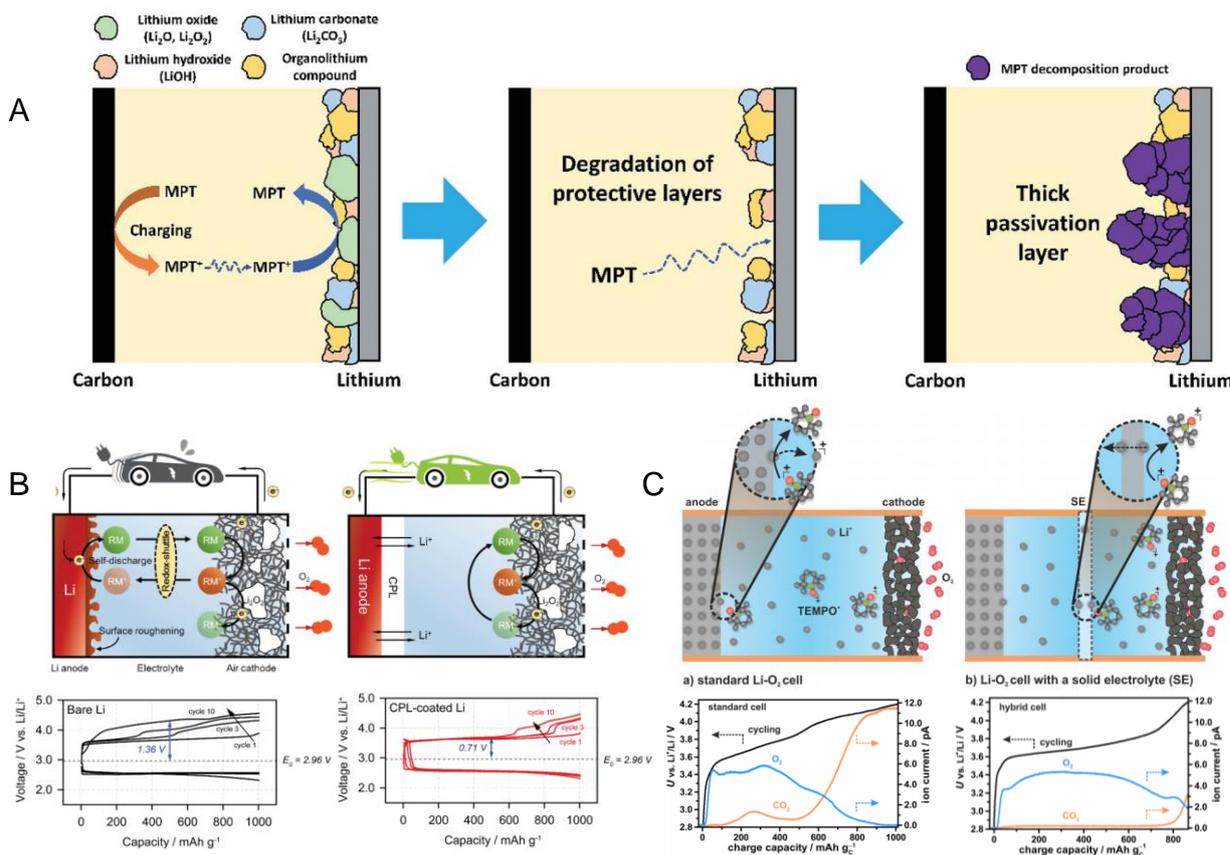


Figure 36. Issue and solution for side reaction of RM with Li metal. (A) Schematic of the failure mechanism of the Li–O₂ cell with redox mediator. Reproduced with permission from ref.²⁹⁴. Copyright 2017 Royal Society of Chemistry. (B) Coating protective layer on Li metal to prohibit parasitic reaction between RM and Li metal. Reproduced with permission from ref.²⁹⁵. Copyright 2016 John Wiley & Sons, Inc. (C) Solid electrolyte as Li⁺ conductive membrane to prohibit crossover of RM to Li metal. Reproduced with permission from ref.²⁹⁶. Copyright 2016 American Chemical Society.

Lim et al. investigated representative organic RMs to find key factors that determine the catalytic activity and electrochemical stability with solvents.²⁷⁸ They demonstrated that the ionization energy (IE) is a suitable proxy to predict the energy level (redox potential) of a candidate RM, helping to indicate to what extent it can interact with oxygen species in Li-O₂ cells, without promoting electrolyte solution degradation. They have shown that RMs with their singly occupied orbitals (SOMO) of the oxidized form close to the HOMO of the solvent are prone to extract electrons from the solvent and decompose it (Figure 37). This is seen in irreversible electrochemistry of these RMs. Based on experimental data and DFT calculations,

they suggested dimethylphenazine (DMPZ) as a low-voltage RM which is not attacking the solvent.

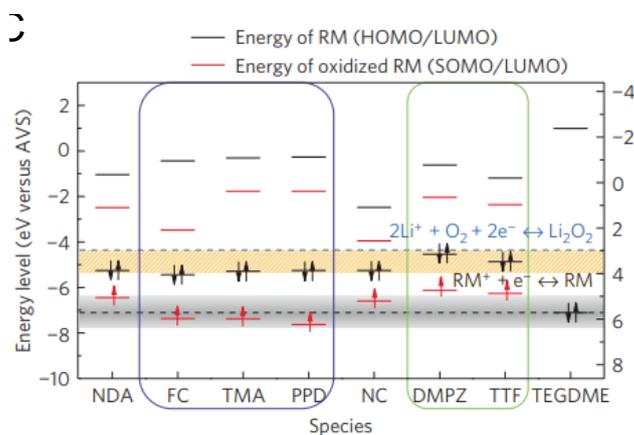


Figure 37. Molecular orbital energies of redox mediators and TEGDME. HOMO and LUMO energies of original RMs (black bar), singly occupied orbital (SOMO) energies of oxidized RMs (red bars with arrows) and HOMO energies of TEGDME electrolyte based on DFT calculations. Adapted with permission from ref. ²⁷⁸. Copyright 2016 Springer Nature.

In addition to the problem of the RM being degraded by reaction with Li metal, and considering the potential instability of RMs at the cathode, Kwak et al.²⁹⁷ have shown that several commonly used RMs already decompose in an O₂-free environment upon prolonged cycling. **Figure 38A** shows, for example, that TTF readily decomposes when cycled over both of its oxidation steps, while it proved more stable when only the first step was accessed. Such instabilities could arise from the oxidized mediator oxidizing the electrolyte as shown by Kang et al.²⁷⁸ as discussed above. Since most mediators are organic molecules, they are susceptible to the same decomposition as organic electrolytes or carbon by reactive O₂ species. This was tentatively ascribed to instability with superoxide or peroxide.²⁷⁰ Recently, Kwak et al.¹⁵⁶ investigated the reactivity of TTF and DMPZ towards various O₂ species to clarify possible degeneration routes in typical etheral electrolyte solutions (**Figure 38B**). Analogous to the predominant decomposition of solvents by ¹O₂ (see above), it was confirmed that singlet oxygen (¹O₂) predominantly causes RM decomposition. The electrophilic nature of ¹O₂ addition reactions render reduced RMs particularly susceptible to attack. While the negative,

detrimental effect of $^1\text{O}_2$ formation in Li-O $_2$ cells on the stability of electrolyte solutions and RMs is evident, understanding the exact mechanisms requires further work.

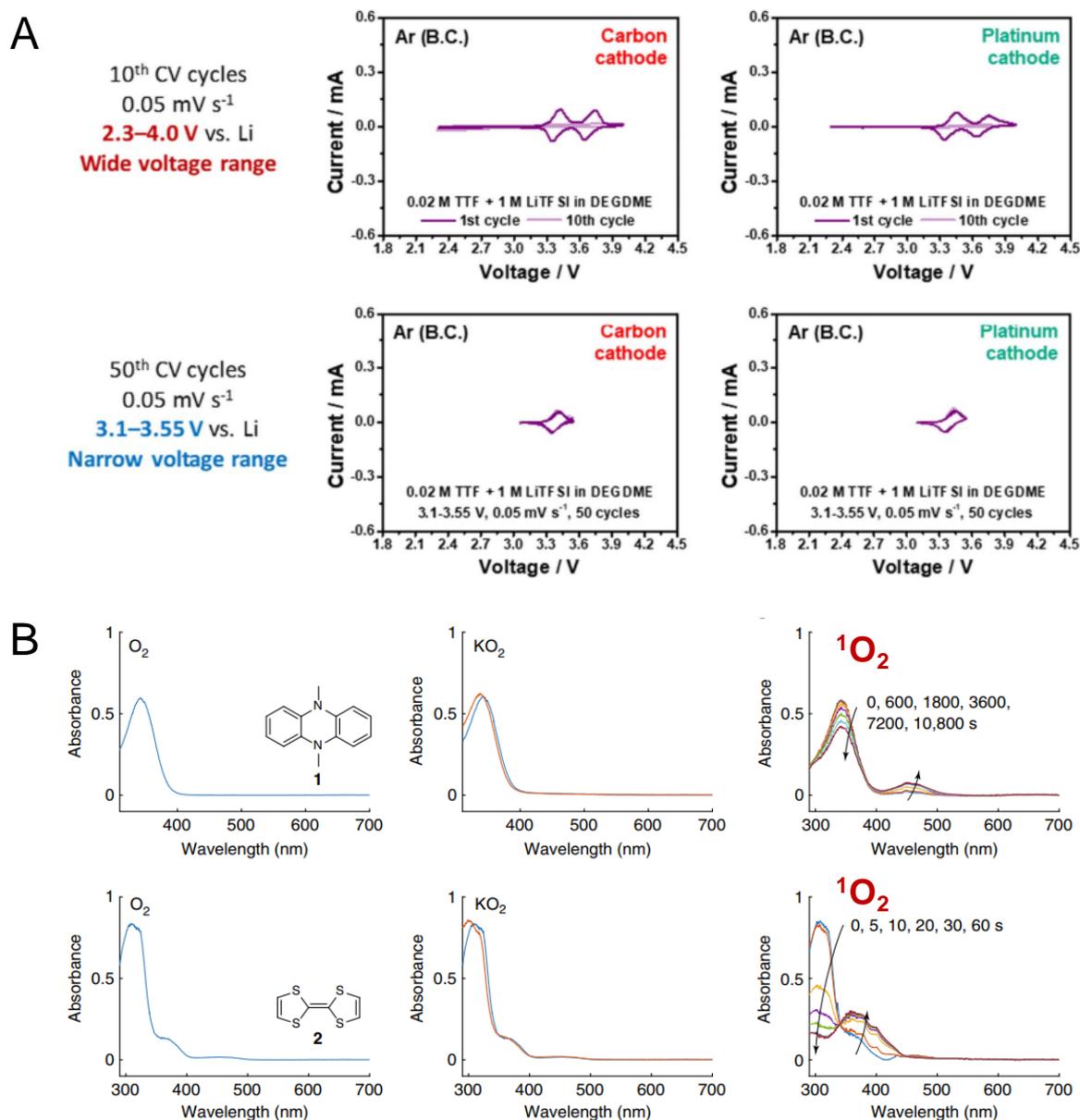


Figure 38. RM instability at the cathode. (A) Instability of doubly oxidized RMs while the singly oxidized form may cycle reversibly. Data were collected under Ar at C cathodes. Reproduced with permission from ref. ²⁹⁷. Copyright 2018 The Electrochemical Society. (B) Decomposition by singlet oxygen. Reproduced with permission from ref. ¹⁵⁶. Copyright 2019 Springer.

3.6 The important analogous systems

3.6(a) Sodium-oxygen batteries

As the higher analogue of Li-O₂, non-aqueous sodium-air batteries (also known as Na-O₂ cells) open up interesting possibilities. Hartman et al. first demonstrated that sodium superoxide (NaO₂) can be produced as the main discharge product by using a diglyme-based electrolyte.¹¹⁹ Although Na-O₂ cells offer a lower formal specific energy (1105 Wh kg⁻¹ based on NaO₂) than Li-O₂ cells, they feature a much more abundant, sustainable element, sodium, a lower charge overpotential (typically below 0.2 V), and higher reversibility (Figures 39A and B). These merits suggest that Na-O₂ batteries may have practical advantages, and they have attracted great research interest.

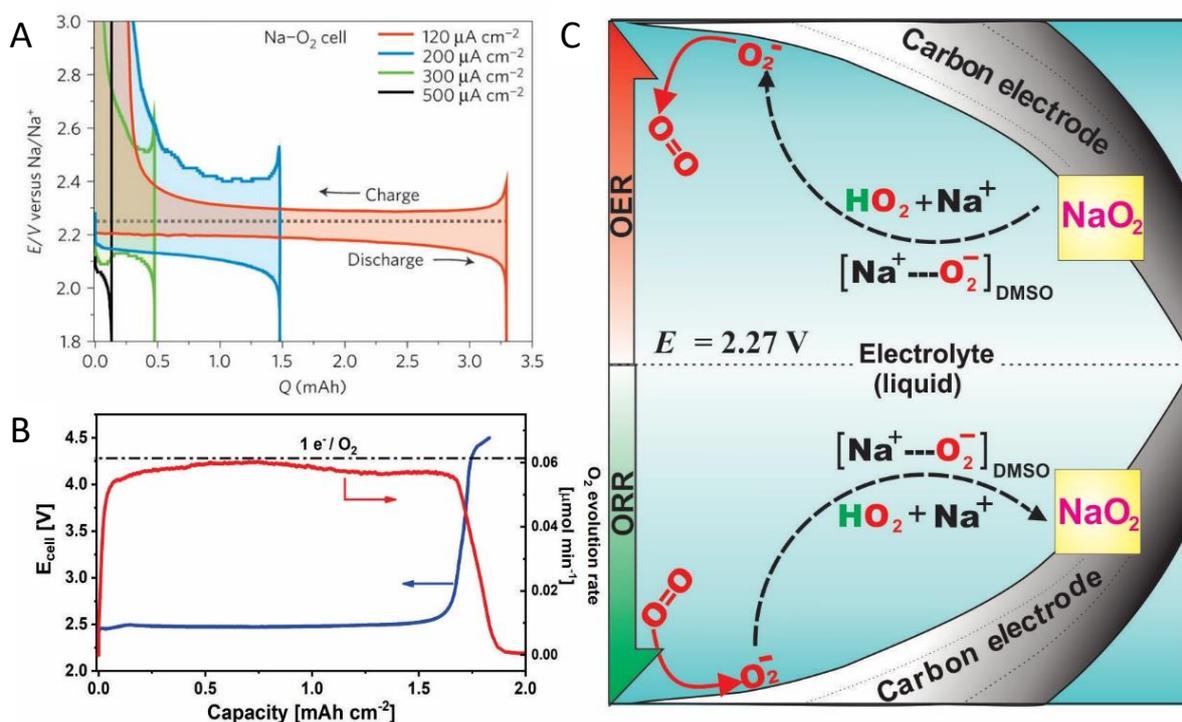


Figure 39. A rechargeable Na-O₂ battery based on NaO₂. (A) Discharge–charge cycles of Na–O₂ cells at various current densities. Cut-off potentials were set to 1.8V for discharge and 3.6V for charge. Dotted line: $E^0(\text{NaO}_2)=2.27$ V. Reproduced with permission from ref¹¹⁹. Copyright 2012 Springer Nature. (B) Online mass spectrometry analysis of oxygen evolution rate (red curve) upon charge (blue curve is the charge profile). The dashed line indicates an oxygen evolution rate corresponding to a $1 e^-/\text{O}_2$ process. (C) Schematic illustration of the solution-mediated superoxide transfer mechanism during discharge (bottom) and charge (upper). Superoxide anion is solvated as ion pairs in DMSO $[\text{Na}^+ \cdots \text{O}_2^-]$ or by generating HO₂ intermediate upon reaction with proton-based phase transfer catalysts. Once supersaturation is achieved,

NaO₂ precipitates forming cubic-shaped crystals. The charge process follows the reverse pathway of discharge process. B and C Reproduced with permission from ref ⁹⁵. Copyright 2015 Springer Nature.

Superoxide versus peroxide stability is a function of the alkali cations. On the basis of the Hard-Soft Acid-Base (HSAB) theory, the Na⁺ ion with a larger radius possesses lower charge density as compared to Li⁺ and is therefore more effective in stabilizing O₂⁻.⁷¹ It suppresses the spontaneous disproportionation of NaO₂ to sodium peroxide (Na₂O₂). On the basis of rotating ring-disk electrode (RRDE) and *in-situ* X-ray diffraction (XRD) studies, it was found that solution-mediated superoxide transfers accounts for the growth and oxidation of cube-shaped NaO₂ products.^{95,298} As shown in **Figure 39C**, superoxide species are exclusively formed at the cathode surface upon discharge and then transferred to the electrolyte. Once supersaturation is achieved, NaO₂ precipitates and forms cube-shaped crystals. During charge, superoxide species are re-solubilized in electrolyte and diffuse to the cathode for oxidation. This reversible solution-mediated pathway with a lower energy barrier leads to a large discharge capacity and a low charge overpotential. It is reported that the Gutmann DN_s of organic solvents and the addition of phase transfer catalysts significantly influence this solution process, albeit through different mechanisms. Based on surface enhanced Raman spectroscopy (SERS) results, Aldous et al. suggested that dimethyl sulfoxide (DMSO) with a large DN facilitates the solution-mediated growth of NaO₂ by forming soluble DMSO-solvated [Na⁺-O₂⁻] ion pairs.²⁸⁸ Xia et al. demonstrated that trace amount of weak acids (*e.g.*, H₂O, benzoic acid) function as proton-based phase transfer catalysts, enabling the solubilizing of superoxide via the formation of an hydroperoxyl (HO₂) intermediate.^{95,299} They further provided direct evidence of HO₂ species existing in a diglyme-based electrolyte by using *in-situ* electron spin resonance (ESR).³⁰⁰ Clearly, such reversible solution processes play a critical role in Na-O₂ chemistry with a lower overpotential and higher rechargeability. By

contrast, the spontaneous disproportionation of LiO_2 leads to irreversible discharge/charge pathways and the poor rechargeability and high overpotential of Li-O_2 cells.

In addition to NaO_2 , sodium peroxide (Na_2O_2) and its dihydrate form ($\text{Na}_2\text{O}_2 \cdot \text{H}_2\text{O}$) were also reported as discharge products of Na-O_2 cells, although which of these is preferred is not fully understood.³⁰¹ Thermodynamically, the formation of Na_2O_2 is more favourable than NaO_2 ($-449.7 \text{ kJ mol}^{-1} (\text{Na}_2\text{O}_2)^{-1}$ vs. $-437.5 \text{ kJ mol}^{-1} (2 \text{ NaO}_2)^{-1}$).¹²⁸ On the other hand, NaO_2 may be the kinetically preferred product due to the requirement of only a 1 e^- transfer (compared to two electrons for Na_2O_2). Thus, generally disproportionation in presence of Na^+ or H^+ will trigger the conversion of superoxide into peroxide (see Figure. 10a). Solvent dependent NaO_2 disproportionation has been shown by Sheng et al. to be accelerated with high DN solvents.¹⁶⁰ Although Na-O_2 cells based on NaO_2 products exhibit the desired electrochemical behaviour, their long-term stabilities are still a topic of concern, due to the associated parasitic chemistry. A number of studies have shown instability of NaO_2 upon prolonged rest.^{118,130,132,162,180} Although details vary, they generally report gradual conversion of NaO_2 into $\text{Na}_2\text{O}_2 \cdot \text{H}_2\text{O}$ accompanied by electrolyte degradation to form NaOH , Na_2CO_3 and organic compounds including Na formate and acetate. Explanations included nucleophilic attack, H^+ and H-atom abstraction by superoxide. Black et al. reported that 10 mol% of discharge products are such side products as quantified by H-NMR and acid-titration measurements.¹³⁰ By using an isotopically labelled ^{13}C model cathode, they also demonstrated that the carbon cathode is oxidized by NaO_2 , irreversibly forming carbonate.

Singlet oxygen ($^1\text{O}_2$) is now recognized an essential trigger for parasitic reactions as is formed at all stages of cycling.¹⁰⁷ Recently, poor Na_2O_2 yield and pronounced degradation upon NaO_2 disproportionation has been explained within a unified $^1\text{O}_2$ formation mechanism that shows growing $^1\text{O}_2$ yields as superoxide disproportionates in presence of increasingly weakly Lewis acidic cations (Lewis acidity: $\text{H}^+ > \text{Li}^+ > \text{Na}^+$; $^1\text{O}_2$ yield $\text{H}^+ < \text{Li}^+ < \text{Na}^+$).^[134] A major

conclusion from this mechanism in conjunction with all papers showing instability of NaO₂-based Na-O₂ cells is that NaO₂ as a thermodynamically metastable discharge product must be avoided as it lacks the required tolerance to slow discharge and rest periods; it will gradually convert to Na₂O₂ and side products. Cycling the Na-O₂ cell highly reversible requires therefore controlled formation/decomposition of Na₂O₂ without superoxide disproportionation steps.

In addition, similar to a Li metal anode, sodium deposits easily grow into a dendritic pattern, resulting in safety concerns.³⁰² Oxygen species (including oxygen and superoxide) diffuse through electrolyte solution and contaminate the surface of Na anode. Thus, a variety of protective membranes (e.g., polymer, ceramic, etc.) have been used to restrict growth of sodium dendrites and prevent the diffusion of oxygen species.^{303,304} Alternatively, Bender et al. replaced sodium metal with a sodiated carbon anode to improve the reversibility of the anode and avoid dendritic growth.³⁰⁵ However, the long-term reliability of these approaches needs to be further examined.

3.6(b) The potassium-oxygen battery

In seeking alternative metal-oxygen chemistry to overcome the challenges in Li-O₂ batteries, the potassium-oxygen battery based on potassium superoxide (KO₂) as the product was first established by Ren et al.³⁰⁶ The net battery reaction is $\text{K} + \text{O}_2 \leftrightarrow \text{KO}_2$ ($\Delta G^0 = -239.4$ kJ mol⁻¹, $E^0 = 2.48$ V), which gives a formal specific energy of 935 Wh kg⁻¹ (based on KO₂). Differing from LiO₂ and NaO₂, KO₂ is a thermodynamically favourable product (see Figure. 10a). This allows K-O₂ cells to operate exclusively via the O₂/O₂⁻ redox couple, presenting a high superoxide yield above 98 % (vs. 95 % for Na-O₂ cells), and a low overpotential of 50 mV (vs. 200 mV for Na-O₂ cells).¹²⁶ After aging in a glyme-based electrolyte for 30 days, electrochemically formed KO₂ remains stable in the cell and demonstrates 94% coulombic efficiency. By contrast, NaO₂ quickly undergoes irreversible decomposition upon aging in the

electrolyte. Thus, the higher stability and reversibility of KO_2 , as compared with LiO_2 and NaO_2 , offers K- O_2 chemistry significant advantages.

Although ether-based electrolytes are considered to be relatively stable against O_2^- attack, their decomposition in the K- O_2 cell is still problematic. By using ion chromatography, Xiao et al. reported that 0.3 % of the KO_2 product is consumed with the decomposition of dimethoxyethane (DME) solvent after the first discharge.¹²⁶ More seriously, potassium metal presents much higher reactivity with ether-based electrolytes than lithium and sodium.³⁰⁷ This owes to the formation of highly reactive solvated electrons and/or potasside (negatively charged potassium ions) by dissolution of K metal in ether solvents. These dissolved species cleave the C–O bonds in ether molecules leading to a very thick passivating layer ($\sim 500 \mu\text{m}$) precipitated on the potassium surface. Moreover, the crossover O_2 is continuously reduced to superoxide species by potassium, which also accelerates the decomposition of electrolyte. Thus, forming a stable solid electrolyte interface (SEI) is highly desired for achieving reversible potassium stripping/plating. By using appropriate electrolyte salts, applying an artificial SEI, and employing a Nafion- K^+ membrane, the cyclability of K- O_2 cells has been significantly improved from a few cycles to hundreds of cycles.^{307–309} In addition, potassium-based alloys (*e.g.*, K–Sb alloy and liquid Na–K alloy) as a new class of anode materials demonstrate advantages over K metal in terms of SEI stabilization and elimination of dendrites.^{310,311} Very recently, Lu et al. reported on an organic-oxygen cell using a potassium biphenyl complex (BpK)/DME anode and a DMSO-mediated potassium superoxide cathode (Figure 40A). Based on cyclic voltammetry studies, Lu et al. found that Bp^-/Bp redox is stable and reversible. The cell exhibits an unprecedented cycle life (3,000 cycles) with a superior average Coulombic efficiency of more than 99.84% at a high current density of 4.0 mA cm^{-2} (Figures 40B and C).¹²⁰ Clearly, developing alternative anode materials with higher stability

and reversibility is a promising approach to increase the cycle life of K-O₂ cells; more importantly, it is likely applicable to other metal-oxygen battery systems.

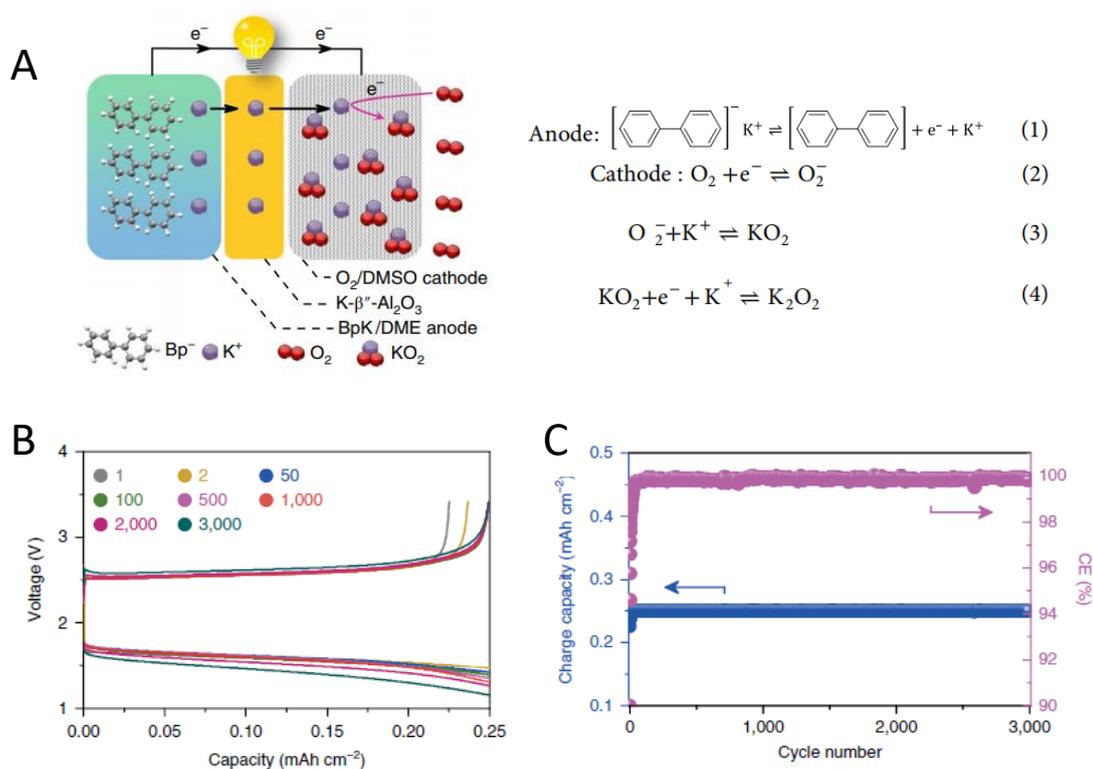


Figure. 40. A long-life potassium biphenyl complex–oxygen (BpK-O₂) cell. (A) Cell configuration and working principle of BpK-O₂ chemistry. (B) Galvanostatic voltage profiles of the BpK–O₂ cell at a current density of 4.0 mA cm⁻² during cycling (C) The corresponding charge capacities (dark blue) and CEs (purple) of the BpK–O₂ cell in B. Reproduced with permission from ref ¹²⁰. Copyright 2019 Springer Nature.

3.6(c) The magnesium-oxygen battery

Metallic magnesium is an attractive negative-electrode material for battery applications due to its high abundance in the earth's crust, low cost, and better safety in comparison to alkali metals (Li and Na).³¹² A magnesium-oxygen battery based on magnesium oxide (MgO) products theoretically delivers a volumetric and a gravimetric energy density of 14 kW h L⁻¹ and 3.9 kW h kg⁻¹, respectively. These values are even higher than those of Li-O₂ cells (8.0 kW h L⁻¹ and 3.4 kW h kg⁻¹). This makes Mg–O₂ batteries compelling candidates for large scale and sustainable energy storage devices required for grid integration.

In the past decade, aqueous Mg-O₂ cells were mainly studied as primary batteries owing to the non-rechargeability of the magnesium hydroxide (Mg(OH)₂) product.³¹³ Shiga et al. first reported a secondary Mg-O₂ cell using a DMSO-based electrolyte that operates at elevated temperature (60 °C).³¹⁴ The discharge product was assumed to be MgO that is rechargeable in the presence of redox mediators, *e.g.*, iodine (I₂) and 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO).^{314,315} Based on XRD and Raman analysis, Vardar et al. identified that the discharge products were composed of a mixture of crystalline MgO and amorphous magnesium peroxide (MgO₂), when a Grignard reagent-based electrolyte was used in a Mg-O₂ cell.³¹⁶ They speculated that O₂⁻ is formed first upon discharge followed by its disproportionation to MgO₂ and MgO. After charge, MgO₂ is fully oxidized, whereas some MgO remains. The incomplete oxidation of MgO causes low Columbic efficiency and rapid capacity fading. Thus, exploring more effective OER catalysts and redox mediators to oxidize MgO is necessary. Meanwhile, more studies are required to examine the reversibility of Mg-O₂ cells and reveal the underlying chemistry.

3.7. Novel electrolytes and electrodes.

3.7(a) The possibilities and development of active metal (Li, Na) protection.

As discussed above, various approaches and results have been reported for Li metal stabilization in Li-O₂ batteries. Metal stabilization is required in order to control overall battery performance including stability, lifetime and efficiency of the battery. In contrast to Li-metal and Li-S batteries, the Li-O₂ battery involves reactive oxygen, and more attention needs to be paid to the details. In this chapter, we will discuss the important aspects of Li-metal protection in Li-O₂ batteries and directions for the future.

Li metal deterioration, rather than dendrite formation, is a primary issue in Li-O₂ cells. It results from side reactions with oxygen species, moisture, reaction byproducts, and soluble additives which can crossover from the cathode side. However, most studies have been focused

on the advantage of protective layers to protect the Li metal without consideration of their stability to reactive oxygen species. To verify the importance of the stability of protective layers against parasitic reactions of reactive oxygen species (ROS), Kwak et al. evaluated the chemical stability of different layers and oxygen concentration during recharge in Li-O₂ batteries. It was previously reported that a protective layer containing PVDF or PEO exhibited chemical change and continuous decrease of oxygen. Nevertheless, a composite layer (NCL) based on Nafion - which is known to be stable against oxygen attack - gave effective metal protection without any parasitic reactions by reactive oxygen (O_2^- , 1O_2) as shown in **Figure 41A**. Future studies need to confirm these important results.

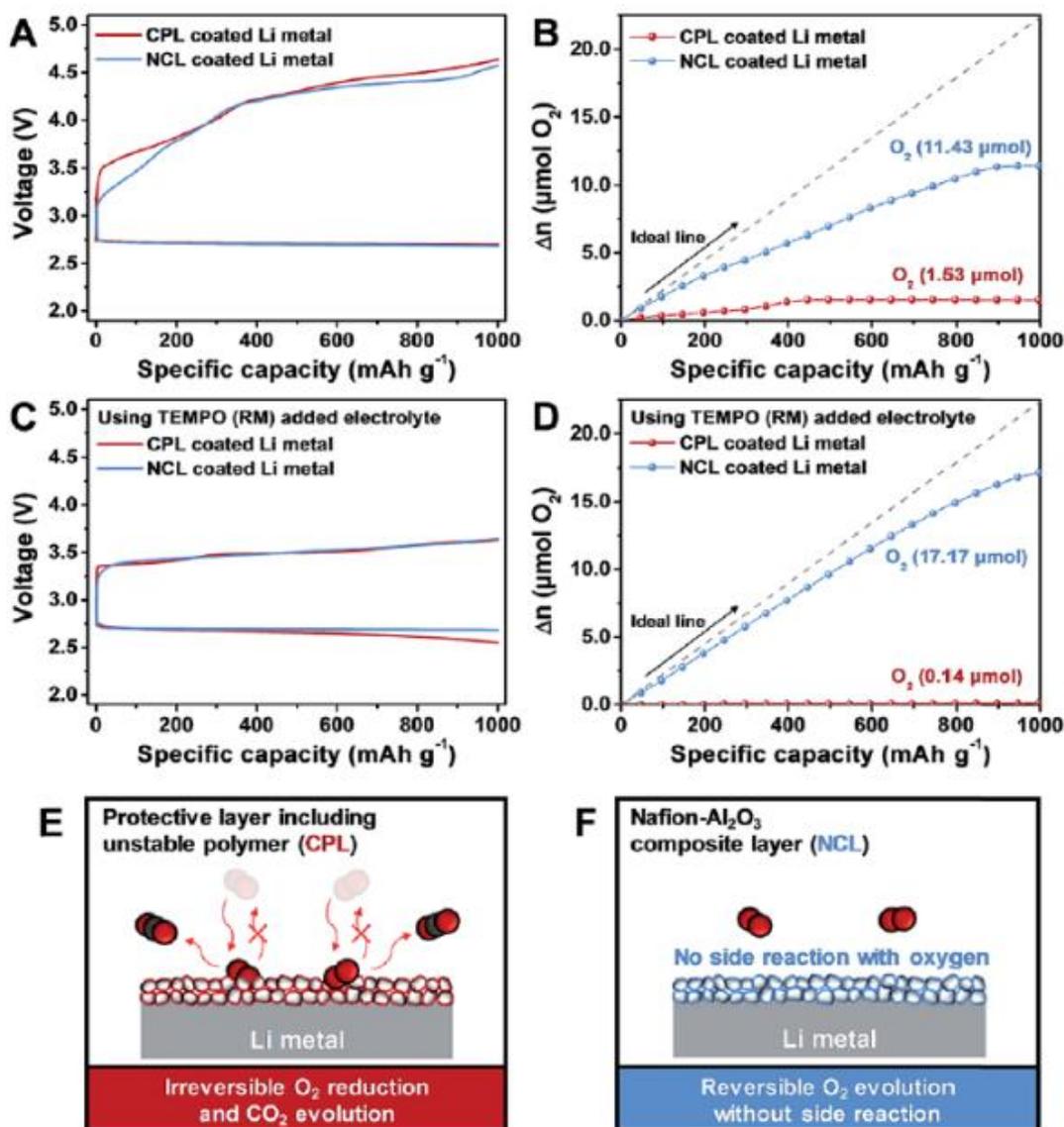


Figure 41. Chemical stability of protective layer on Li metal in Li-O₂ batteries: (A) Galvanostatic voltage profiles and (B) integration of O₂ gas amount for the corresponding charge process. (C) Galvanostatic voltage profiles with redox mediator in the same electrolyte solution and at the same current rate, and (D) integration of the amount of O₂ gas for the corresponding charge process. The red and blue graphs indicate PVdF-based CPL and Nafion-based NCL, respectively. The dashed line in (B), (D) indicates the ideal total amount of O₂ gas. (E, F) schematic illustrations of side reactions on the CPL and the negligible parasitic reaction on the inert NCL, respectively. Reproduced with permission from ref. ³¹⁷. Copyright 2019 Royal Society of Chemistry.

Another problem sometimes overlooked in the protection of lithium metal in Li-O₂ batteries is the formation of dendrites. Currently, this problem is not so serious because the actual capacity value (mAh) and current (mA) are relatively low compared to practical Li-ion batteries, although the specific capacity based on the mass of cathode is high.

In a scale up study³¹⁷⁻³¹⁹ using large area electrodes and pouch cells (Figure 42A), it was found that the dendritic growth of Li metal was much more pronounced than that of a small sized system based on either coin cells or Swagelok[®] cells (Figures 42B-E). As the area of the electrode increases, uniform distribution of current is difficult, and dendrite problems become aggravated due to uneven local reactions. This problem leads to an imbalance in Li₂O₂ formation, which greatly reduces the reversibility of the system (Figure 42F).

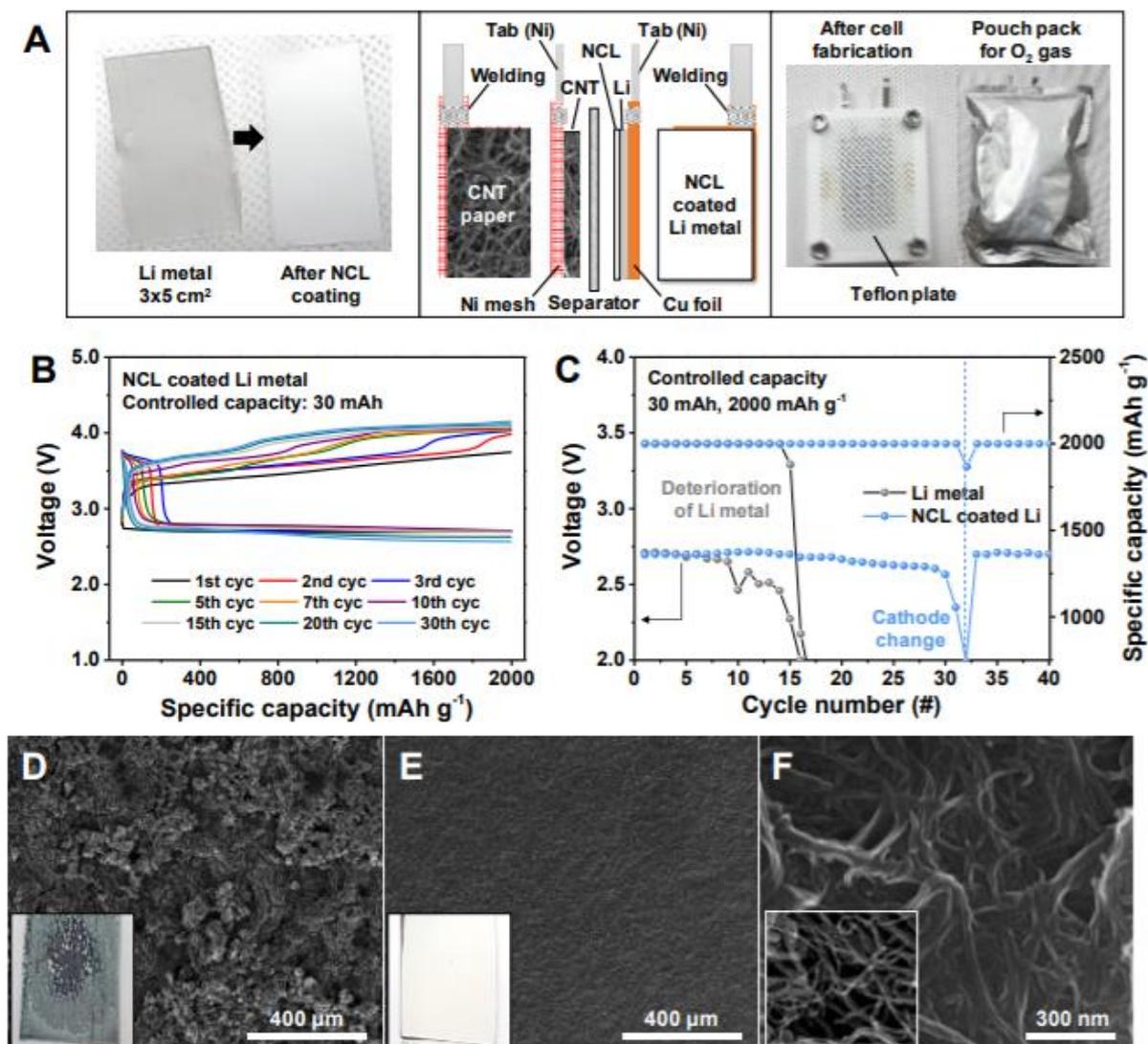


Figure 42. Protective layer for Li metal in large sized Li-O₂ cell. (A) Photograph of a Li electrode before and after NCL coating for pouch-type cell fabrication (left), schematic image of the cell configuration of the Li-O₂ pouch cell (middle), and a photograph of a Li-O₂ pouch cell before and after pouch packing with O₂ gas (right). (B) Voltage profiles during the galvanostatic cycling test of a Li-O₂ pouch cell using NCL-coated Li metal. (C) Cycling performance (i.e., end voltage of discharge and specific capacity) of Li-O₂ pouch cells using Li metal and NCL-coated Li metal as a function of the cycle number. (D-F) SEM images and photographs of (D) Li metal (E) NCL-coated Li metal after the cycling test of Li-O₂ pouch cells. (F) SEM image of a CNT paper electrode in a Li-O₂ pouch cell using NCL-coated Li metal after the cycling test until cell failure. Reproduced with permission from ref. ³¹⁷. Copyright 2019 Royal Society of Chemistry.

As a result, not only must the side reactions of the Li metal surface be considered, but also dendrite growth and volume changes. To make the battery more practical, along with the

chemical and physical stability of the protective film, a study of the effect of local current is also required. Future research directions must include scale up through controlling the local current density³²⁰ or using pouch type cells.^{317–319}

3.7(b) Solid state Li-air and Na-air batteries.

Studies on solid state Li-O₂ batteries using solid electrolytes (SE) have been conducted in order to overcome many problems (see below) in Li-O₂ batteries which contain electrolytes in organic solvents. In SEs, Li-O₂ cells, comprised of ceramic and/or polymeric materials which conduct Li ions, convey Li ions between the carbonaceous oxygen cathode and the lithium metal anode. Dozens of papers have been published in this area over the last decade and several characteristic examples are mentioned below. The use of SEs in these systems have the aim of solving chemical and electrochemical problems of instability, flammability, toxicity and volatility of organic electrolyte solutions. The SE can protect the Li or Na metal anodes and prevent crossover of reactive oxygen reduction products from the cathode to the anode. Using polymeric electrolytes also enables fabrication of flexible batteries^{321,322}

The first demonstration of solid-state Li-O₂ cells was presented by Kumar et al.³²³ Their design was comprised of a ceramic-glass electrolyte (LAGP) placed between polymeric membranes based on polyethylene oxide (PEO) blended with lithium salt. The cells demonstrated stable performance for 40 cycles and excellent thermal stability up to 105 °C.

Research on solid state Li-O₂ batteries can be divided into two categories: one direction promotes the use of pure solid state electrolytes and does not contain any liquid component; in the other, quasi-solid state systems are developed, including mixtures comprised of polymeric matrices (sometimes including ceramic particles) and liquid electrolyte solutions based on organic solvents or ionic liquids. An example of the first category is the work of Wang et al., in which pure solid electrolyte Li-O₂ cells utilizing the Li⁺ conducting materials

$\text{Li}_{2+2x}\text{Zn}_{1-x}\text{GeO}_4$ (LiSICON),³²⁴ $\text{Li}_{1.3}\text{Al}_{0.3}\text{Ti}_{1.7}(\text{PO}_4)_3$ (LATP), $\text{Li}_{1.5}\text{Al}_{0.5}\text{Ge}_{1.5}(\text{PO}_4)_3$ (LAGP) were used. This cell also emphasized the problem of the generally low ionic conductivity of the solid electrolytes, which results in low rate capability. In addition, there are severe interfacial contact failure problems between the electrodes and the solid electrolytes during the cycling of these cells.³²⁵ A possible solution to problems of contact between solid electrolytes and cathodes could be their integration into a single solid mixture, as suggested by Zhu et al.³²⁶

An important use of solid electrolytes in Li-O₂ cells is in developing protected Li anodes. Once the Li metal anodes are protected by artificial pre-designed solid electrolyte interphases, the reactivity problems related to the negative electrodes can be fully alleviated.^{327,328}

Another approach in design and elaboration of solid state Li-O₂ or Na-O₂ cells is the use of polymeric electrolyte membranes that separate the anode and cathode in the cells.^{329,330} These systems may be fully polymeric or quasi-solid, gel-type, comprising polymeric matrices and electrolyte solution included therein.³³⁰⁻³³²

It is worth mentioning the pioneering work of Peled et al. which was a first step in developing solid Na-O₂ battery systems; namely polymeric matrices containing NaClO₄, and a carbonaceous oxygen cathode.³³³ Peled's preliminary Na-O₂ cells were operated at 105 °C which is above the sodium melting point of 97.8 °C. The liquid sodium metal cells demonstrated 100 cycles with limited normalized capacity. That work was a first attempt in the direction of using sodium anodes in a solid state configuration.

Using quasi-solid polymeric matrices as the electrolytes in Li-O₂ or Na-O₂ cells reflects a compromise. These matrices are less reactive with the electrodes and the oxygen reduction products than the liquid electrolytes. Using such cells at elevated temperatures enable achievement of reasonable rate capabilities. Hence, one may be able to obtain longer cycle-life compared to Li-O₂ cells with liquid electrolytes solutions. However, their use does not fully

mitigate side reactions and there are obviously contact problems that impede fully uniform current distribution in these cells.

An interesting use of polymeric membranes as in Li-O₂ cells was demonstrated by Wu et al. who fabricated and tested Li-O₂ cells containing a Super-Hydrophobic Quasi-Solid Electrolyte (SHQSE) incorporating poly-isobutylene, silicon dioxide treated with di-methyl-dichlorosilane, and ionic liquids.³³⁴ The main advantages of these cells included a flexible configuration, good protection of the Li metal anodes and blockage of detrimental flux of oxygen reduction products from the cathode to the anode. An important advantage in the use of such combinations is the possibility of solving the contact problems, from which most types of SE based Li(Na)-O₂ cells suffer. At first glance, while solving some instability problems, the complexity of solid state Li/Na-O₂ batteries may set aside many possible advantages of such systems, regarding energy density and rate. A great advantage of this system may be the ability to use ambient air as the feed gas, instead of pure oxygen. Since a solid electrolyte at elevated temperatures is used in this case, the presence of water and CO₂ contaminations in air may be much less detrimental for basic Li-O₂ cell operation as compared to Li-O₂ cells based on liquid electrolyte solutions. Work in this direction continues.

3.7(c) On the use of ionic liquids and molten salts

Ionic liquids (ILs), also known as room temperature molten salts, possess negligible volatility, a wider electrochemical window, and lower flammability than conventional organic solvents.³³⁵ Kuboki et al. first used imidazolium-based electrolytes for primary Li-air cells.³³⁶ The hydrophobic ILs with zero vapor pressure avoid the evaporation of electrolyte and the hydrolysis of lithium anode in moist air. These make ILs excellent solvents for the lithium-air batteries operated under ambient conditions. Subsequently, a variety of IL-based electrolytes (*e.g.*, pyrrolidinium and ammonium-based electrolytes) were developed for secondary Li-O₂ cells.³³⁷ Nevertheless, the chemical stability of ILs is still under debate. Based on rotating ring-

disk electrode (RRDE) studies, Herranz et al. found that the pseudo-first-order reaction constant of 1-butyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide (Pyr₁₄TFSI) with O₂⁻ is at least three orders of magnitude lower than that of propylene carbonate (PC).³³⁸ By contrast, Piana et al., through DEMS and NMR studies, concluded that Pyr₁₄TFSI does not have sufficient long-term stability against superoxide attack.³³⁹ Katayama et al. also reported the instability of imidazolium-based IL electrolytes based on cyclic voltammetry measurements.³⁴⁰ In addition, the viscosities of ILs are usually one magnitude higher than those of conventional organic solvents. They lower the ionic conductivity of electrolyte and decrease the diffusion coefficient of oxygen/superoxide, which results in a low discharge capacity and poor rate capability.³⁴¹ Therefore, more efforts are needed to explore novel ILs with higher stability and lower viscosity.

Inorganic molten salts (*e.g.*, molten nitrate eutectic) typically melt above 100 °C and thereby allow Li-O₂ cells to operate at higher temperatures. The first molten salt Li-O₂ cell was reported by Giordani et al. using the LiNO₃-KNO₃ eutectic.³⁴³ The cell shows a very low overpotential of 50 mV and an improved rate capacity at 150 °C due to the reversible transfer of the Li₂O₂ product in the molten nitrate electrolyte. Based on *in situ* pressure and gas analyses, they found that Li-O₂ chemistry is reversible in this medium, which reflects on the high stability of the inorganic electrolyte. Recently, Xia et al. reported on a Li-O₂ cell using a molten nitrate electrolyte and solid ceramic membrane that reversibly forms and oxidizes Li₂O in a stepwise four-electron redox process (Figures 43A-C).³⁴² As shown in Figure 43D, the cell exhibits stable cycling performance for 150 cycles with a very low polarization of 0.2 V and a high Coulombic efficiency close to 100%. By tuning the operating temperature and using a bifunctional lithiated nickel oxide catalyst, this Li-O₂ battery overcomes the barriers of thermodynamics and kinetics, leading to the electrochemically reversible formation of octahedral lithium oxide (Li₂O) crystals instead of hexagonally-shaped Li₂O₂ crystals that are

formed on carbon cathodes (Figures 43E and F). The use of chemically stable inorganic electrolytes, and a non-carbonaceous cathode circumvents the degradation of organic electrolyte and carbon corrosion, which are the main failure mechanisms for nonaqueous Li-O₂ cells. It demonstrates that Li-O₂ electrochemistry is not intrinsically limited, once problems of electrolyte, superoxide, and cathode host are overcome.

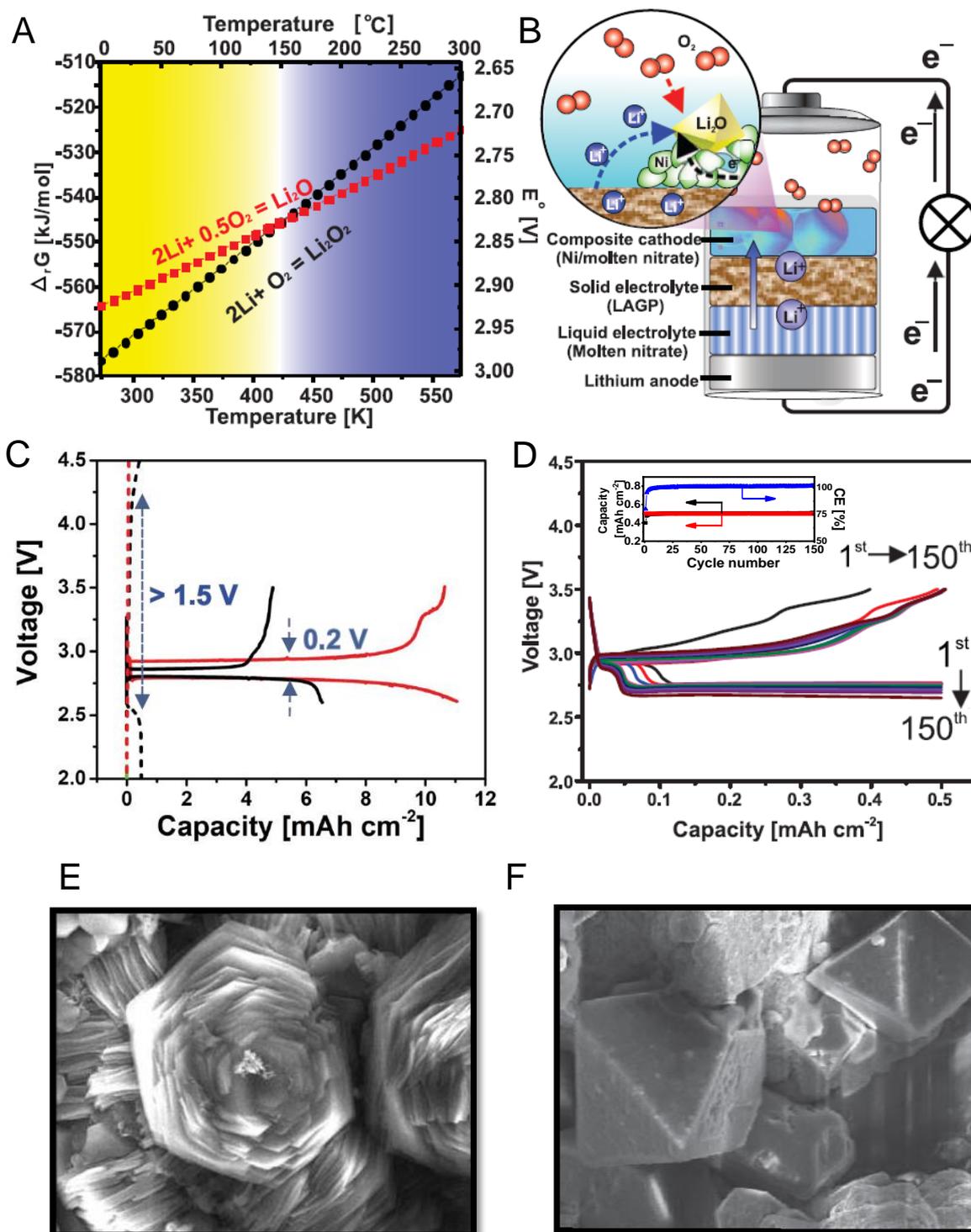


Figure 43. Li₂O based Li-O₂ cell using molten salt-based electrolytes. (A) Thermodynamics and configuration of the Li-O₂ cell. (A) Gibbs reaction energy for formation of Li₂O and Li₂O₂ as a function of temperature. (B) Configuration of the inorganic electrolyte Li-O₂ cell and schematic illustration of Li₂O formation during discharge. (C) Voltage profiles of Li-O₂ cells with a carbon cathode (black) and a Ni-based cathode (red) using aprotic electrolyte, 0.5 M LiTFSI in TEGDME, were examined at 25°C (dashed lines), whereas the cells using the molten nitrate electrolyte were measured at 150°C (solid lines). (D) Cycling performance of a molten nitrate Li-O₂ cell with a lithiated-nickel oxide catalyst at 150°C. SEM images of the discharged

cathodes using molten nitrate electrolyte and carbon cathode(E), and molten nitrate electrolyte and lithiated-nickel oxide catalyst(F), respectively. Reproduced with permission from ref. ³⁴². Copyright 2018 The American Association for the Advancement of Science.

3.7(d) On the possible use of solid Li-oxide cathodes and the connection to lithiated transition metals.

Intensive studies have been carried out on high capacity cathodes for Li ion batteries with the formulae $\text{Li}_{1+x}\text{Mn}_y\text{Co}_z\text{Ni}_{1-x-y-z}\text{O}_2$ ($x > 0$, $y > 0.5$), known as Li and Mn rich NCM cathodes. These studies revealed that their redox mechanisms, leading to a very high specific capacity (approaching 300 mAh g^{-1}), involve electrochemical activity of oxygen species. The synthesis of these materials that includes high temperature calcination stages, naturally forms a mixture of integrated phases: an electrochemically inactive monoclinic phase, Li_2MnO_3 , and an active rhombohedral phase with the general stoichiometry $\text{LiMn}_x\text{Co}_y\text{Ni}_z\text{O}_2$. When electrodes comprising, these materials are polarized to high enough potentials (e.g. 4.8 V vs. Li), a complex activation process takes place which affects the structure of the monoclinic phase, de-lithiates the entire cathode material, and makes it fully electrochemically reversible (Figure 44). Traditional thinking attributed the activation process and the subsequent further behavior of these systems to changes in the oxidation state of the transition metal cations and related structural re-arrangements. However, recent studies, starting with the pioneering work of Sathiya et al.,³⁴⁴ discovered that the activation process of these materials involves oxidation of the oxide anions to oxygen moieties with a higher oxidation state at a potential $> 4.6\text{V}$ vs. Li. Upon repeated cycling, the process is reversible and the oxidized oxygen moieties are reduced back to oxides at around 3 V. This involvement of oxygen electrochemistry, in fact enables the high specific capacity of these cathode materials, that may reach the highest limit relevant for Li insertion cathodes with the general formula $\text{Li}_{1+x}\text{M}_{1-x}\text{O}_2$ (300 mAh g^{-1}). The process of activation and reversible operation, capacity fading mechanisms and average voltage fading mechanisms - are not fully clear yet. It appears that oxidation of the oxide moieties may

lead to irreversible evolution of molecular singlet oxygen which is detrimental. Also, due to their partial oxidation, upon charging, the resulting nucleophilic oxygen moieties are highly reactive with the electrophilic alkyl carbonate solvents. Thereby, effective surface protection of these materials by coatings (to avoid oxygen release and detrimental contact with solution species) is mandatory for their stable reversible cycling. **Figure 44** summarizes several important experimental aspects related to these cathodes. Intensive work is being currently carried out on these materials (references 355 – 374 are typical examples out of hundreds of publications that have appeared in recent years).

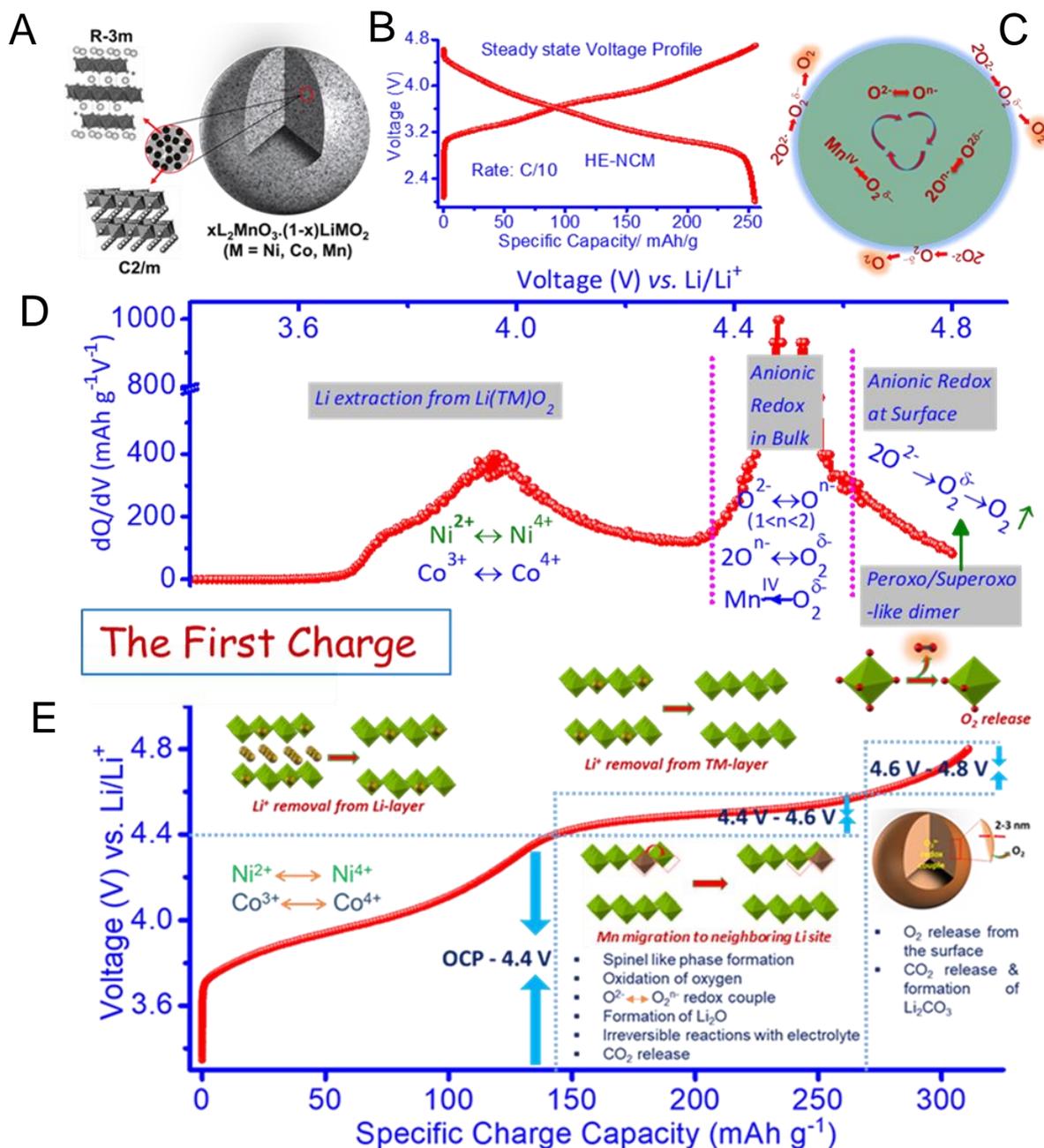


Figure 44. Li and Mn-rich cathode materials: schematically describing (A) structure, (B) steady state voltage profiles, (C) anionic oxygen redox activity in the bulk and on the surface and (D, E) the first charge dQ/dV vs. V and voltage profile demonstrating Li extraction, cationic & anionic redox activity.^{345–364}

3.8. Studies with consideration of practical metal-air batteries.

3.8(a) Li Batteries with lithium oxygen compound cathodes (and closed systems)

One of the ways to overcome the challenge of operating in air (which requires reduction in CO₂ and H₂O levels from the incoming gas stream) is to have a closed Li-O₂ system, where the oxygen is already stored in the battery. One approach is to store O₂ in a tank, while this adds weight and volume, it is offset, to a greater or lesser extent, by obviating the need for gas scrubbers. An alternative approach is to incorporate the oxygen in the solid state, e.g. to make the cell in the discharged state incorporating Li₂O₂ powder in a conductive carbon matrix to form a composite cathode. Early studies used this cathode preparation method to investigate the OER mechanism in a more controlled manner. The first composite cathode was described by Ogasawara et al. where ball-milling was used to mix commercial Li₂O₂ with high surface carbon and PTFE binder⁸³. It was demonstrated that the Li₂O₂ can be completely oxidized at high enough potentials. In a later study, Meini et al. reported that commercial Li₂O₂ contains 5% of impurities such as LiOH and Li₂CO₃.³⁶⁵ They explained that these impurities can influence the OER process and can introduce undesirable side-products. As mentioned above, several studies have described the chemical and electrochemical differences between crystalline and amorphous Li₂O₂ phases.^{102,103} We note that these studies did not investigate the cyclability of these types of preloaded Li₂O₂ cathodes.

Bhargav et al. attempted to cycle a closed Li-O₂ battery comprised of a composite cathode prepared from commercial Li₂O₂ and nano-graphite to form a composite cathode.³⁶⁶ The first discharge process showed a high specific capacity of more than 1000 mAh g⁻¹. However, after four cycles, the discharge capacity dropped by more than 50%. The capacity fading was attributed to the purity and size of the Li₂O₂ particles and their dispersion in the carbon matrix. The same group observed improvement in the battery cycle life when the cathode was made from highly pure synthesized nanocrystalline Li₂O₂ particles that were

embedded in carbon nanofibers.³⁶⁷ The influence of the Li_2O_2 loading in the carbon cathode can also be a limiting factor in maximizing the energy density of the cathode. Luo et al. showed that the impedance of the preloaded cathodes increased with more insulating Li_2O_2 . They concluded that 20 wt. % Li_2O_2 in the carbon black matrix delivered the highest energy efficiency. The crossover of oxygen evolved from Li_2O_2 particles to the lithium anode can result in different decomposition reactions which can degrade the Li-O_2 performance.^{64,368}

In a recent study, Zhu et al. proposed a new alternative closed Li-O_2 system where a cathode made of Co_3O_4 nanoporous matrix is filled with nanolithia (Li_2O).³⁶⁹ In this type of cell, the cathode cycles between the discharged Li_2O and the $\text{Li}_2\text{O}_2/\text{LiO}_2$ formed during the charge process. The authors proposed that the large OER overpotential observed in standard Li-O_2 cells is due to the phase transition between the solid products (Li_2O_2) and the evolving O_2 gas which lead to a severe kinetic bottleneck. Unlike a standard Li-O_2 cell, the proposed lithia based system is based on redox chemistry in the condensed state. As can be seen from the cyclic voltammogram in Figure 45, the difference in voltage between the reduction and oxidation peaks is only 0.24 V. As seen from Figure 45, a cell comprised of a lithia cathode and lithium metal anode exhibited 200 cycles at around 600 Ah kg^{-1} capacity with low capacity loss.

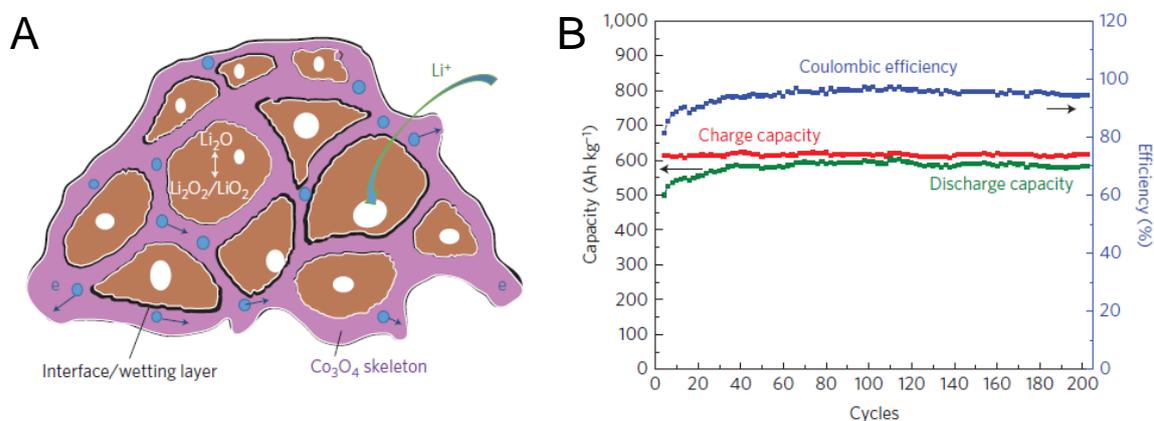


Figure 45. Cyclic voltammogram of nanolithia cathode between 2.35 and 3.0V with a scan rate of 0.05mV (b) Cycling performance and coulombic efficiency of nanolithia cathode vs. Li metal anode. Reproduced with permission from ref.³⁶⁹. Copyright 2016 Springer Nature.

3.8(b) Challenges of capacity and kinetics.

The use of redox mediators has overcome many of the challenges introduced by the intrinsic oxygen electrochemistry, especially the otherwise limited rate and capacity of Li-O₂ cells when operated in low DN solvents such as ethers, as described above in section 3.5b. However, oxygen redox limitations are only one aspect of the problem. Lithium peroxide preferentially forms at the air interface of the cathode and blocks the electrode pores (Figure 46A).²⁰⁸ The bulk of the cathode volume remains relatively free of discharge product which is detrimental to capacity. Simulated discharge curves by Liu shows that oxygen depletion and pore clogging is mainly responsible for limited capacity and cell death.³⁷⁰ Figure 46 presents this schematically, showing the O₂ concentration through the cathode at high and low rates. It indicates that the O₂ availability is limited to the air interface and much of the cathode is starved and inactive. Gittleston et al. explored the effect of O₂ and Li⁺ transport in different solvents and concluded that low oxygen solubility and pore blocking at the air interface were now the major factors limiting capacity and inducing cell death.³⁷¹

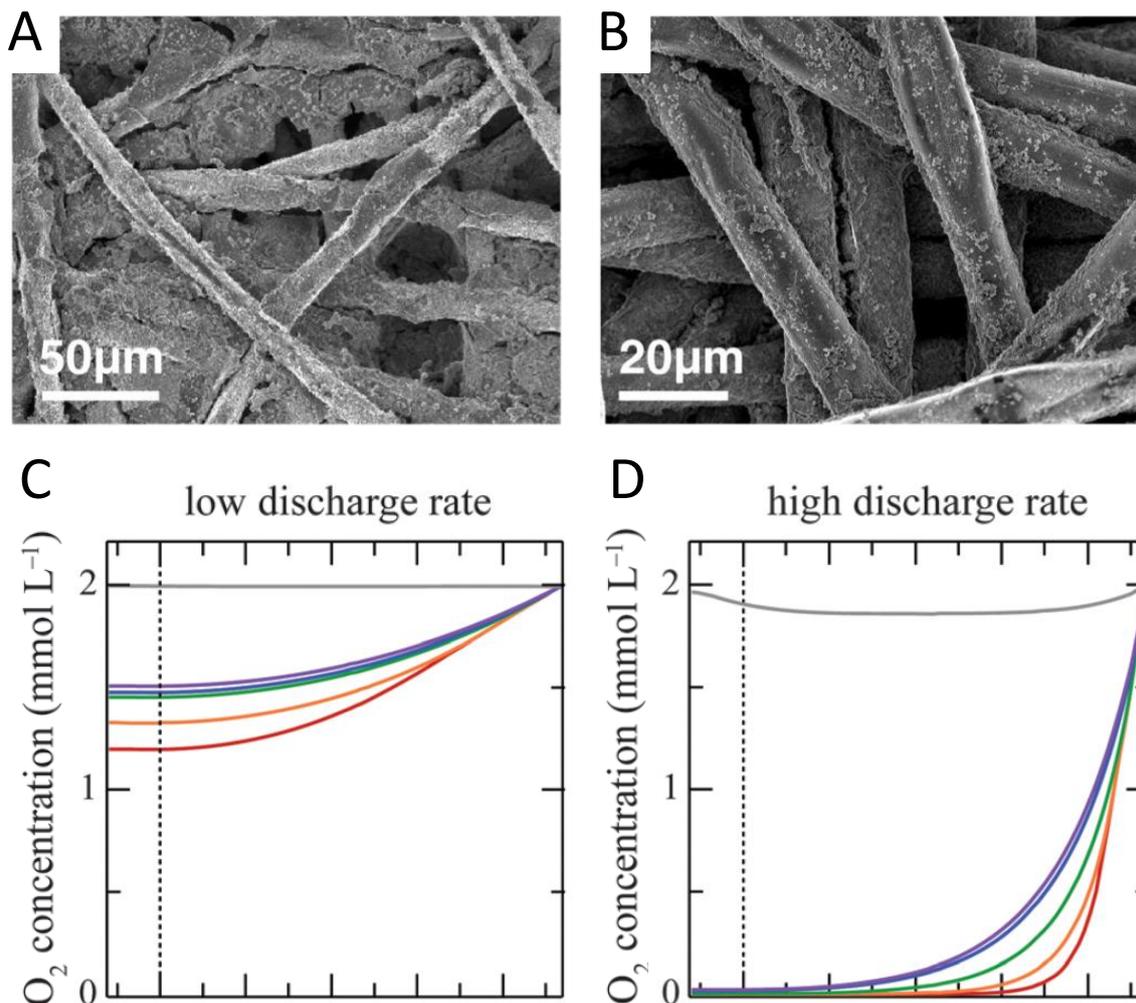


Figure 46. SEM images of a dual mediator cathode at (A) the O₂ interface and (B) the separator interface. Reproduced with permission from ref.²⁰⁸. Copyright 2017 Springer Nature. (C,D) Plots of O₂ concentration distribution in the electrolyte of the cathode at six depths of discharge. Reproduced with permission from ref.³⁷⁰. Copyright 2016 Royal Society of Chemistry.

These results highlight the need for methods to enhance O₂ mass transport in the lithium-air battery. It is unlikely that strategies can be directly translated from the aqueous fuel cell as the solvents share little common chemistry. In addition, operation of the two systems differ considerably. The product of the aqueous fuel cell is the solvent, while the Li-air battery forms a new solid phase (lithium peroxide). New engineering and chemical methods for oxygen delivery are required for the lithium-air battery. One such example of a chemical solution to this problem is the use of the perfluorocarbons (PFC), which increase oxygen solubility and mass-transport. PFC additives can be classified as miscible^{372,373,374} and immiscible.^{375,376}

Nishikami et al. showed that $\text{BrC}_6\text{F}_{13}$ forms a miscible solution with TEGDME up to 60 wt% improving the discharge capacity by 50%,³⁷⁴ although at the expense of electrolyte conductivity. Contentiously, Wijaya et al. suggested that the enhancement observed with miscible PFCs is due to parasitic currents from PFC attack by superoxide.³⁷⁷ Immiscible PFCs have been proposed to avoid this problem and introduce three phase regions into the cathode where the PFC acts as oxygen channels.³⁷⁷ This method doubled the discharge capacity at low current densities but had a negative impact on capacity at higher rates. The authors proposed that this was due to greater pore blocking. Enhancement is also seen when immiscible PFCs are introduced to a dual mediator cell.³⁷⁶

3.8(c) On the validity of energy density calculation of Li (Na) – oxygen batteries.

The exceptional theoretical specific energy of the Li-air cell led many people to rush to the conclusion that this would translate into practical specific energies and energy densities many times that of Li-ion. Such confusion is often compounded by reports that use only the carbon mass to normalize the capacity. By analysing current reported protocols for Li- O_2 cells evaluation in the literature, Noked et al. demonstrated that many of the reports fall short in specifying the exact cell composition and cycling conditions.⁶¹ When trying to clarify the practical potential of the Li-air battery system, we need to keep two major aspects in mind: (1) careful examination of reported values in the literature, (2) what can we expect as performance from a Li-air battery - even if we solve all the stability and kinetic limitation issues of these systems. Of course, Li-ion batteries are not immune from this problem, for example the often quoted capacity of Si, $3579 \text{ mAhg}_{\text{Si}}^{-1}$ ($8334 \text{ mAhcm}_{\text{Si}}^{-3}$), is misleading as it ignores the Li in $\text{Li}_3.75\text{Si}$, and the need for volume expansion is also often ignored. The true figures are 1857 mAhg^{-1} and 2290 mAhcm^{-3} with respect to the lithiated phase. The Li-air battery is very different from Li-ion, most notably the former requires a balance of plant associated with gas

handling, and the need to reduce the CO₂ and H₂O content in the gas stream. These uncertainties led some researchers to question whether even if the problems of Li-air can be solved, it would outperform Li-ion and Li-S. The first rigorous attempt to truly compare Li-ion, Li-S and Li-air batteries (as opposed to cells), was made by Gallagher et al in 2014.⁵³ They modified the Argonne Batpac modelling software to permit prediction of the specific energy and energy density of a practical Li-air battery, including the air handling/balance of plant. They concluded that while Li-air was shown to have a higher specific energy compared with Li-ion and Li-S, this would not be the case volumetrically compared with a cell containing a Li metal anode and a Li rich LMNC cathode. However, as detailed in this review, our understanding of the Li – air battery has been transformed over the last 6 years and the introduction of redox mediators and recognition that significant quantities of H₂O can be tolerate, alleviating the need to complex, heavy and bulky air handling, changes the assumptions required to model performance. Taking account of recent advances including the use of redox mediators, reduced balance of plant overhead, and a protected Li anode, which reduces the need for a large excess of Li, we predict the specific energy and the energy density of a full Li-air battery, including the balance of plant, to be 610 Whkg⁻¹ and 680 WhL⁻¹ respectively, compared with 450 for each in the case of Li-S and 300 Wh kg⁻¹ and 500 Wh L⁻¹ for a Li anode and Li rich NMC. Larger energy batteries increase this difference as the mass fraction associated with the balance of plant diminishes relative to the pack and conversely, Li-air is not competitive for small batteries such as consumer electronics. The major advantage of Li-air is in specific energy. As a result, it might prove particularly useful for challenges such as aviation where, mass and specific energies beyond what Li-ion can deliver are essential.

In summary, the gap between the very high theoretical energy density of Li-air batteries - and practical expectations is very wide, reducing the advantage of Li-air over other Li batteries, however Li-air is still predicted to offer performance beyond the other battery

technologies, especially in terms of mass, and this is based on full cell analysis and operation in air. As a result, while Li-air is still some way from a technology, the advances indicate that work on these systems should be further promoted as discussed in the conclusion section.

3.8(d) From oxygen to air.

The lithium-air battery must operate in air if it is to reach its maximum performance. Most lab scale cells operate using pure oxygen, but a scaled practical battery pack using pure O₂ would offer significantly lower specific energy compared to an open battery.⁵³ Moving to an atmospheric gas introduces a number of challenges, but surprisingly also brings some benefits. Lowering the O₂ partial pressure has been shown to impact capacity.³⁷⁸ As discussed above, as yet unimagined strategies are needed for oxygen delivery which may help to eliminate this problem. Atmospheric nitrogen will not be compatible with the lithium metal anode, but it is generally accepted that lithium will have to be protected with an artificial SEI or ceramic interphase in any case.

Contrary to established thinking, a series of papers from Meini, Schwenke and Aetukuri showed that H₂O can be an advantageous additive in lithium-air electrolytes, by increasing capacity and lowering the charge overpotential.^{94,96,97} Later this was convincingly shown in work by Qiao et al. (Figure 47) where increasing the H₂O content from 0 to 30% results in increasing discharge capacity, and lowers the charge potential to below 3.5 V.³⁷⁹ SEM analyses shows that the lithium peroxide particle size increases with increasing H₂O content and X-ray diffraction confirms formation of lithium peroxide. We note however that LiOH has been observed at higher H₂O concentrations.⁹⁶ The large particles and enhanced capacity confirm discharge by a solution route and two theories have been proposed to explain this. Aetukuri et al. suggest that the acceptor number of the H₂O additive results in enhanced solvation of superoxide.⁹⁴ Schwenke et al. proposed that protonation of lithium peroxide by water results in soluble peroxide.⁹⁶ Both mechanisms effectively enable formation of a soluble product able to

nucleate large particles and avoid electrode surface passivation. H₂O is also beneficial for performance when using Quinone based discharge mediators and more than doubles the discharge capacity.³⁸⁰ It is proposed that this is due to stabilization of the reduced Quinone and quinone-O₂ complex by H₂O, which effectively increases diffusion lengths and supports peroxide filling of the cathode structure. Finally, in the presence of an LiI catalyst, H₂O can promote formation of LiOH in a 4e⁻ reduction step.^{100,381} Decomposition due to H₂O is an area that needs further consideration.

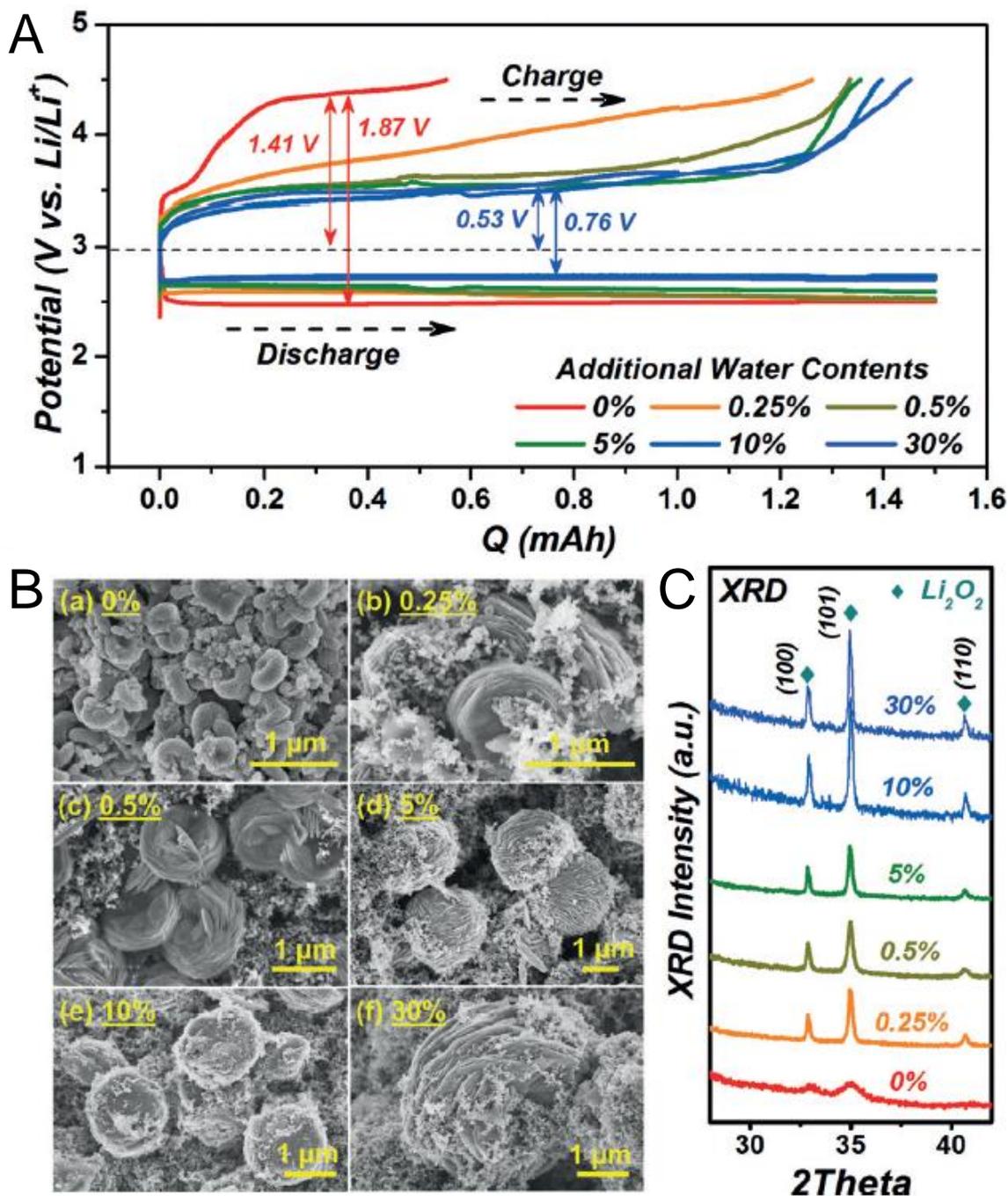


Figure 47. Discharge profiles of lithium-O₂ cells with varying H₂O content and SEM images and XRD patterns of the discharge cathodes. Reproduced with permission from ref³⁷⁹. Copyright 2017 Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim.

In contrast to H₂O, CO₂ may be detrimental to the performance of the lithium-air battery. Thermodynamically the reaction between lithium peroxide and carbon dioxide to form lithium carbonate is favourable.⁶³ Early work in this area by Takechi et al. showed that

introduction of CO_2 into the lithium-air battery results in the formation of lithium carbonate, and proposed that this occurred via formation of a peroxydicarbonate ion ($\text{C}_2\text{O}_6^{2-}$).³⁸²

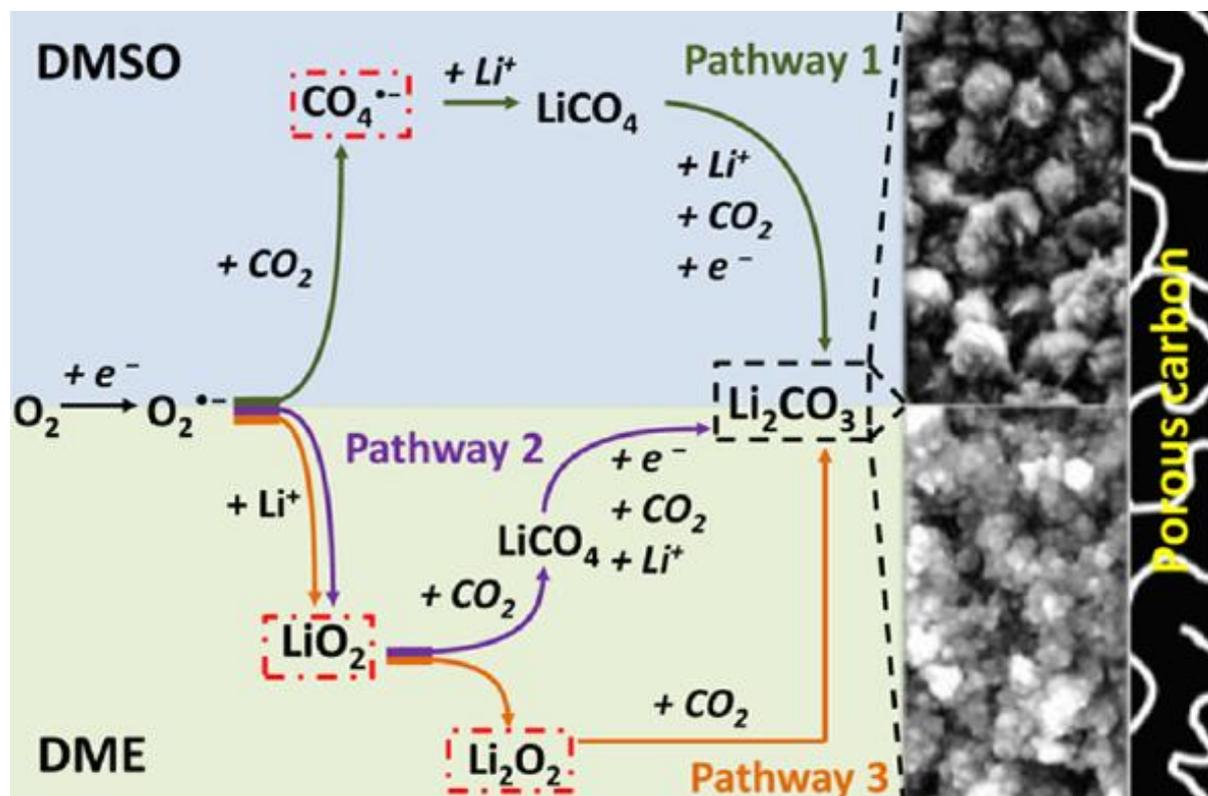


Figure 48. Reaction routes for the most likely paths in Li- O_2/CO_2 battery. Reproduced with permission from ref³⁸³. Copyright 2016 American Chemical Society.

Unfortunately, these studies used carbonate-based electrolytes which are now known to be unstable and decompose to Li_2CO_3 and CO_2 .⁶³ Gowda et al. performed similar studies but using a more stable DME-based electrolyte, which also clearly demonstrated the formation of Li_2CO_3 .³⁸⁴ Yin et al. proposed a holistic reactions mechanism for Li- O_2/CO_2 in these Li-active gas batteries, **Figure 48**.³⁸³ Three possible pathways are described that can be summarized as reaction of CO_2 with either O_2^- , LiO_2 or Li_2O_2 . Theoretical and experimental studies by Lim et al. imply that the solvent can influence the reaction of superoxide with CO_2 . The reaction of CO_2 with Li_2O_2 suggests that the formation of Li_2CO_3 in lithium-air batteries is inevitable and that CO_2 must be removed from the oxygen feed gas. However, Liu et al. has demonstrated the use of redox mediators to remove accumulated Li_2CO_3 in the lithium-air battery.³⁸⁵ Recently

there has also been interest in harnessing the Li-O₂/CO₂ reaction, either as a higher performance battery or as a means of CO₂ capture.^{386,387,388,382}

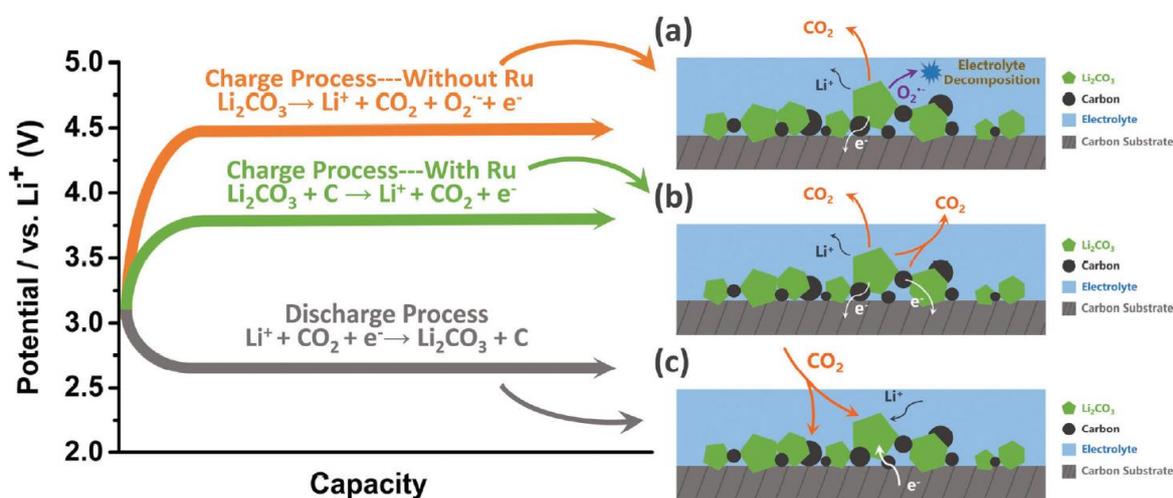


Figure 49. Proposed reaction routes for Li-CO₂ batteries without O₂. Reproduced with permission from ref³⁸⁹. Copyright 2017 Royal Society of Chemistry.

These studies have enabled the expansion of this research field to the study of new batteries using pure CO₂ gas as the cathode's active material, having Li₂CO₃ as the main discharge product. This new system can be denoted as a Li-CO₂ battery (Figure 49). However, several research groups have cast doubts on their technological viability because batteries which use pure CO₂ gas exhibit low discharge capacity as compared to batteries that use a mixture of O₂/CO₂ as the cathode active mass. Therefore, there is a serious debate about possible mechanisms that can enable effective electrochemical reduction of CO₂ in Li-CO₂ batteries without the presence/involvement of O₂.¹⁷³

It is important to note that the study of Li-O₂/CO₂ batteries and Li-CO₂ batteries is rather new. As a result, there are no established mechanisms that explain how these systems can work reversibly. There are controversial opinions whether the reduction of pure CO₂ is possible during discharge and which mechanism is reliable for this system. It seems that possible reaction mechanisms for such systems can be determined only by the use of *in-situ* spectro-electrochemical techniques.

3.8(e) Configuration of Li-air cells and the balance of plant.

The lithium-air field has rightly focused on advancing the basic science and understanding of battery chemistry. Impractical coin and Swagelok cells account for the bulk of the studies. Thus, an established cell, pack and system design has yet to emerge. Due to the pace of advancement in the field over the past 5 years, this is perhaps not unfortunate as designs would quickly become obsolete. For example, a lithium-air system designed today would benefit from the development of redox mediators and an increased tolerance to H_2O . The excellent system model by Gallagher et al. provides the most robust design of a full lithium-air battery, which considers the complete system including the pack design and the balance of plant.⁵³ Figure 50 shows the proposed layer structure for a practical lithium-air cell.⁵³ The layer consists of a lithium anode on copper foil protected from the cathode by a ceramic electrolyte. Excess lithium must be minimized to achieve the highest performance. The cathode consists of a porous carbon and would house 60 vol% of lithium peroxide. Use of a thick cathode is essential to minimizing excessive current collector mass. An aluminum foil flow field is used for oxygen delivery.

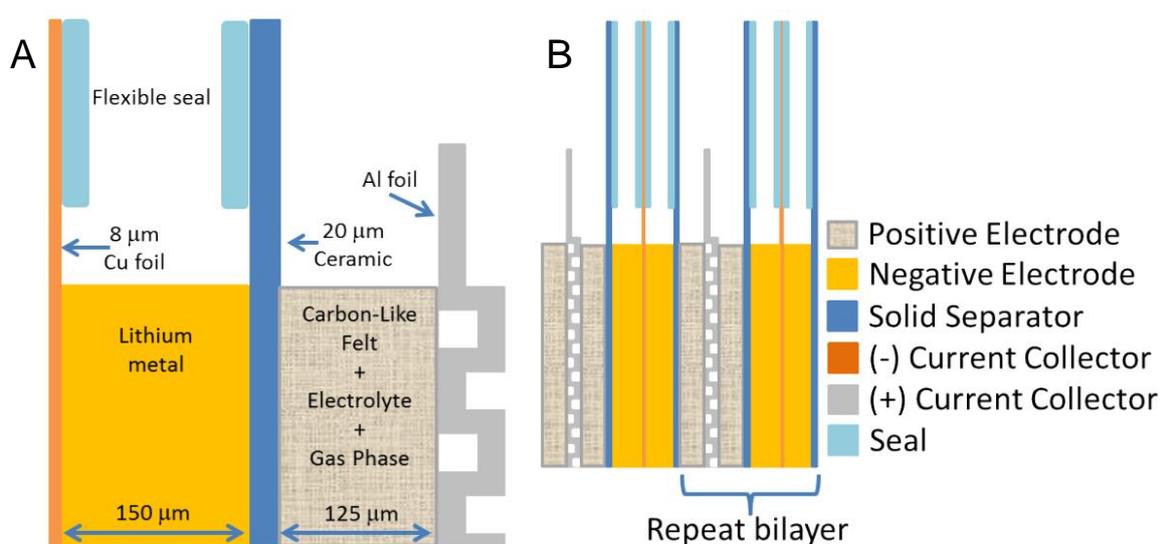


Figure 50. A schematic of a (A) single layer and (B) multiple layers of a lithium-air cell. Reproduced with permission from ref⁵³. Copyright 2014 Royal Society of Chemistry.

An alternative cell design (Figure 51) has been proposed by Samsung, which employed a folded cell architecture that minimizes inactive component volume and mass.³¹⁹ Here the carbon cathode fulfills the additional roles of oxygen flow field and current collector. Performance of $1214 \text{ Wh kg}^{-1}_{\text{cell}}$ and $896 \text{ Wh L}^{-1}_{\text{cell}}$ at the cell level is demonstrated, which exceeds that of Gallagher's cell model (Figure 51), although like-for-like comparisons between cell designs are not trivial.⁵³

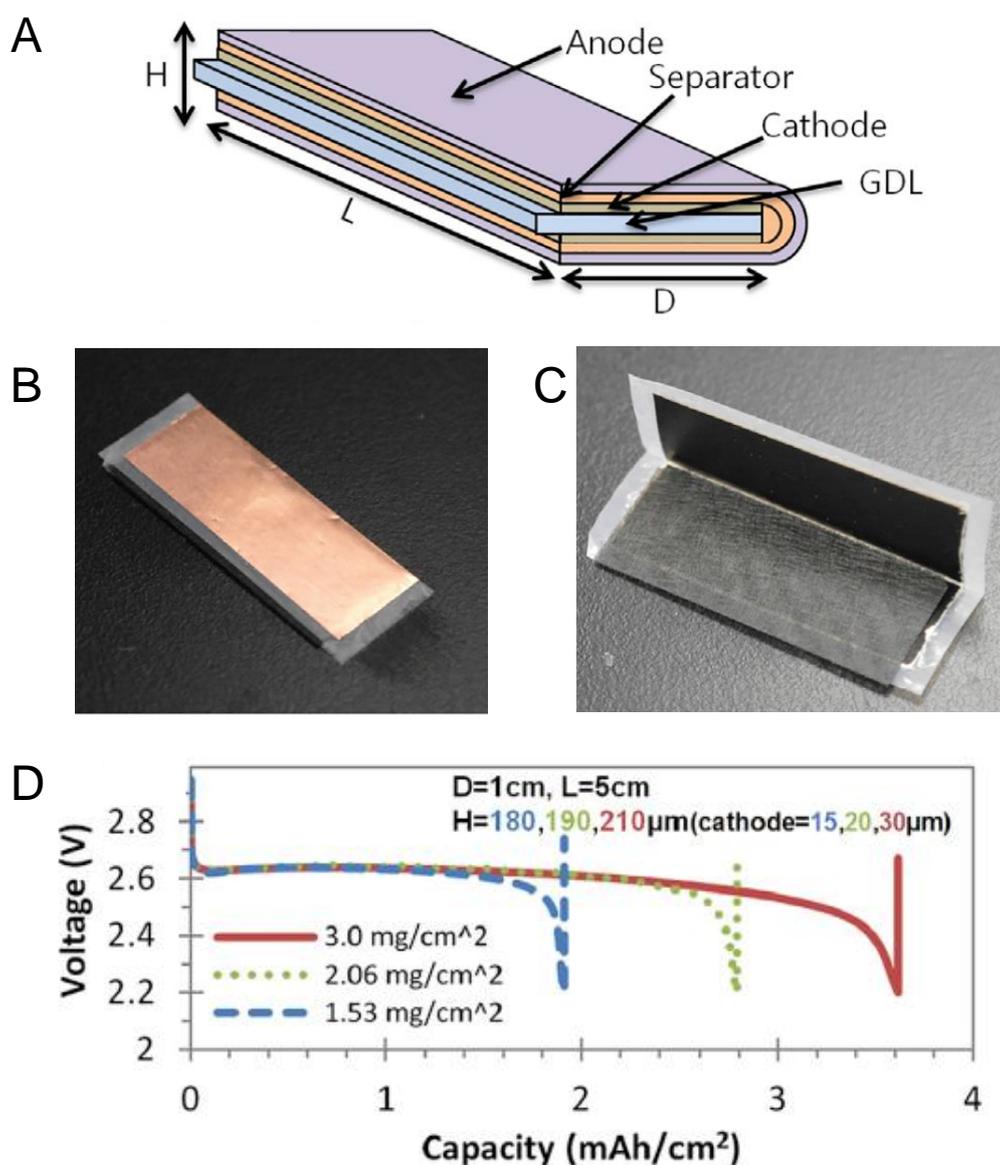


Figure 51. (A) A schematic, (B,C) images and (D) discharge curves of a single folded cell. Reproduced with permission from ref³¹⁹. Copyright 2019 Elsevier.

Repeated bilayers would form a module, typically based on a bipolar plate design similar to those used in a fuel cell,³⁹⁰ The system model by Gallagher proposes a doubling of the mass and volume from cell to pack (Figure 52). This includes consideration of thermal management and oxygen delivery. If the battery is to operate in air, then a balance of plant (BoP) will be needed to deliver a suitable oxygen and manage solvent volatility. The system proposed by Gallagher consists of a compressor, pressure swing adsorption unit and a solvent addition and capture unit.⁵³ The BoP adds an additional 70 kg to the overall system mass and has a significant impact on overall performance. Based on this overall design, and included recent advances, a favorable lithium-air system level performance of $610 \text{ Wh kg}^{-1}_{\text{system}}$ is predicted versus $300 \text{ Wh kg}^{-1}_{\text{system}}$ for a lithium metal/Li-rich NMC system. For example, the model assumes electrochemical performance based on a lithium-air cell without mediators.³⁹¹ The balance of plant is designed to achieve ppm levels of H_2O in the gas stream, but it is likely that this is not necessary and H_2O may improve performance.

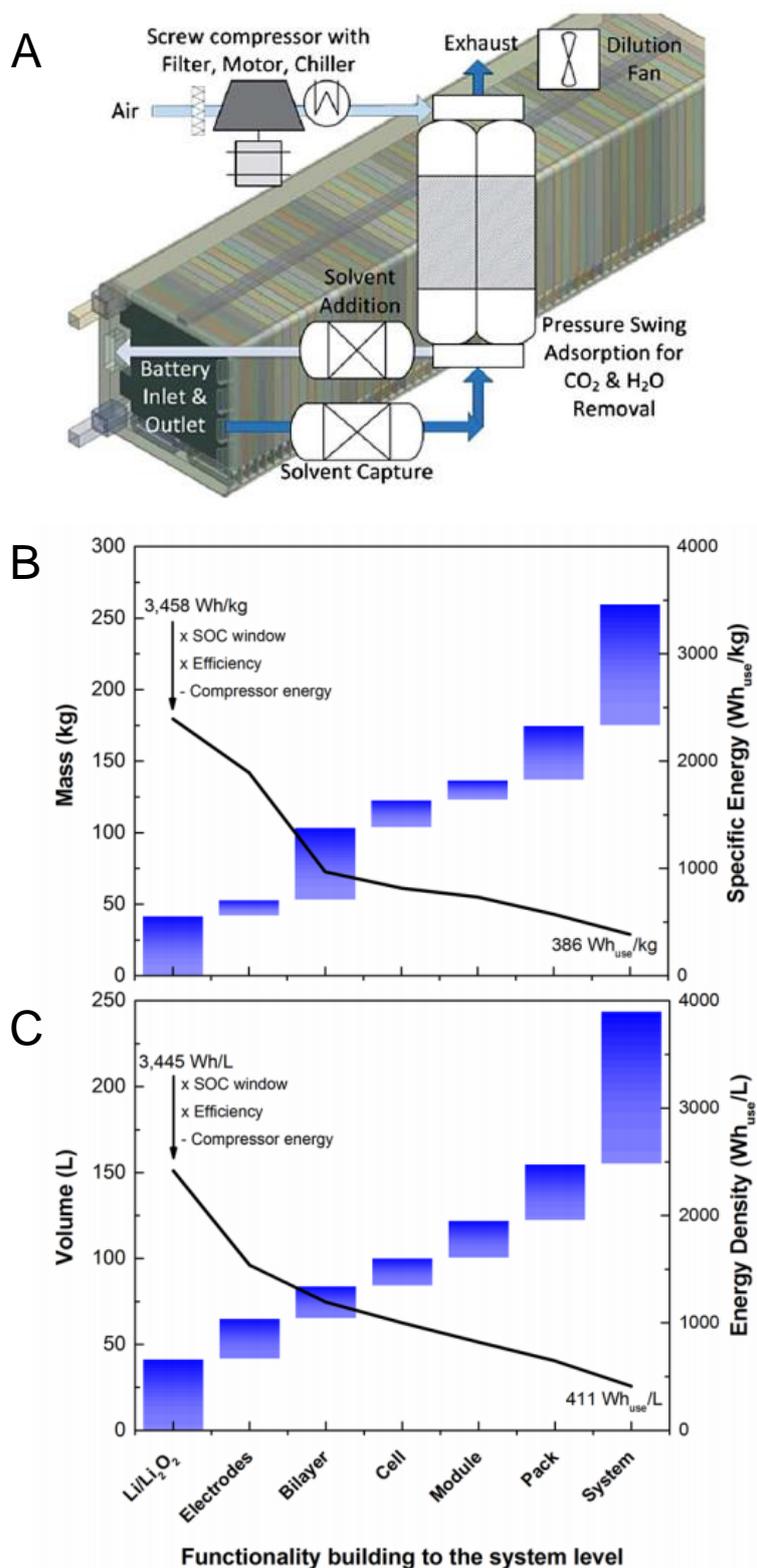


Figure 52. A proposed lithium-air system showing: (A) the components required to the balance of plant needed to operate under atmospheric gas and plots showing the (B) mass and (C) volume breakdown of the open lithium-air system (100 kWh and 80 kW) together with the specific and volumetric energies. Reproduced with permission from ref⁵³. Copyright 2014 Royal Society of Chemistry.

Further improvements may be achieved in the pressure swing adoption units by development of efficient selective gas scrubbers that are able to remove only CO₂ or fractional amounts of H₂O, e.g. metal–organic frameworks and amine based compounds.^{392,393,394} In addition, selective membranes have been proposed as methods of eliminating atmospheric impurities.³⁹⁵ Lower vapor pressure solvents such as TEGDME will help minimize the mass penalty imposed by solvent capture units, which is based on DMSO in current models.^{396,397} Perhaps most importantly, current models are designed for automotive applications, but varying the energy versus power needs can significantly lower the relative contribution of the parasitic BoP overhead and increase performance. System models must be continuously developed and updated to help direct research, identify challenges and opportunities, track progress and evaluate advancement.

Section 4: Future perspective for non-aqueous metal air batteries

A number of the significant advances made in the last five years in understanding the science of the rechargeable aprotic lithium-air battery have been reviewed here. It is important to understand that O₂ reduction to Li₂O₂ on discharge, and the reverse on charge, does not have to occur on the electrode surface, as previously thought, leading to low rates and capacities. Rather, it can occur in solution - avoiding electrode passivation and yielding significantly higher rates and capacities. This realization can begin to change the prospect for the lithium-air battery. Redox mediators can enable solution mediated O₂/Li₂O₂ to occur in more stable, low polarity solvents such as ethers. The rates and capacities in these solvents are several mA cm⁻² and mAh cm⁻², while also significantly reducing the overpotential on discharge and charge will further advance the prospects of a lithium-air technology.

Although cells may be cycled 100 times with minimal capacity loss, this does not mean that all the challenges of lithium-air have been met. Stability of the electrolyte solution is one of the main problems that must be addressed, as discussed above. However, the recent

recognition that singlet oxygen ($^1\text{O}_2$) is formed during oxidation of Li_2O_2 during charge as well as during its formation on discharge, and that this reactive species appears to be the main culprit in electrolyte degradation provides the new understanding necessary to mitigate instability. Oxidation mediators offer the opportunity to change mechanisms that form $^1\text{O}_2$, which would suppress electrolyte degradation.

In addition, it has been shown that mediators enhance the stability of carbon electrodes making them viable despite critical instability with O_2 electrochemistry directly taking place on their surface. Focus on oxidation mediators is therefore a key challenge for future research. The use of redox mediators to facilitate the $\text{O}_2/\text{Li}_2\text{O}_2$ reaction has largely addressed the limitations of the intrinsic electrochemistry at the cathode. The rate and capacity are now limited by the mass transport of O_2 and the pore space. In the flooded porous electrode this relies on the dissolution of O_2 at the gas interface, and then its diffusion throughout the electrode. As for fuel cells, a true gas diffusion electrode is required. This is because O_2 transport occurs in the gas phase entering the electrolyte solution such that it only has to diffuse short distances to the reduction mediator in order to be converted to Li_2O_2 . We cannot utilize the approach of aqueous fuel cells, which relies on hydrophobic gas channels and hydrophilic pores for the electrolyte.

New approaches and designs of electrode architecture are needed that will facilitate mass transport in aprotic Li-air cells. The design of porous gas diffusion cathodes would enable a further step change towards the realization of our Li- O_2 cell; however, alone these do not address the problem of operating in air. As discussed in the review, it is now understood that significant quantities of H_2O can be tolerated in ether-based solvents, while preserving the two-electron reduction of O_2 to Li_2O_2 ; indeed, some water can be advantageous. The ability to operate with a gas stream which does not have to be rigorously dried, potentially operating up to $20\text{ g}_{\text{H}_2\text{O}}/\text{m}^3$, makes operation in air much more viable than was thought previously. The

extensive work on CO₂ removal by, for example, porous metal organic framework solids (MOFs), driven by interest in other fields, provides potential routes to the removal of sufficient CO₂ from the gas stream.

It is important to develop further approaches to the balance of plant especially related to air handling. Much of the focus has been on the challenges associated with the reaction at the cathode. To realize the full potential of the Li-air battery it is necessary to develop a protected lithium metal anode, which avoids reactivity with O₂ and with redox mediators. While specific attention will be critical to understand the solid electrolyte/liquid electrolyte interface of a protected anode in the lithium-air battery, the anode itself is being pursued intensively because of its potential application in other lithium battery technologies. Therefore, this challenge is not specific to Li-air.

There is a well-established path to disruptive technologies. It starts with recognition that the technology would, if realized, provide a step change in performance or cost reduction or both. This drives a “gold rush” to commercialization that almost inevitably fails as the underpinning science is unknown. The attempts to commercialize do uncover the barriers facing the technology. Concentrated effort follows on the basic science, which provides the knowledge that is essential to either overcome the barriers or demonstrate that the barriers are insurmountable. Lithium-air is on this journey.

Section 5: Conclusion.

Research and development of Li-air batteries is one of the most exciting and interesting fields of modern electrochemistry, with its emphasis on power sources, energy and materials science. These systems possess the highest theoretical energy density, promising cost effectiveness and the chance to rival internal combustion engines by clean and elegant electrochemical power sources. Over the last decade, intensive work has been invested into these systems by many of the most prominent research groups worldwide. Despite these

impressive efforts, more work is required even to reach various go/no-go decisions. This is due to the highly complex nature of Li-O₂ cells (even before dealing with ambient air as the cathode feed).

While the intensive R&D efforts related to these systems have yet to reach practical goals, the electrochemistry, energy and materials science communities have gained a lot from the prolific research work carried out during recent years. The benefits and spinoffs from these studies include highly innovative contributions to electro-catalysis, active metals electrochemistry, oxygen chemistry, non-aqueous electrochemistry and solution chemistry - as we describe in this review article. The scope of this paper has included several important related analogous systems: various active metal anodes (sodium, potassium); assorted active gas cathode materials like CO₂; and recently discovered reversible oxygen redox activity in high capacity lithium transition metal oxide cathodes.

In this review, we described in detail important aspects of this field: heterogeneous electro-catalysis; homogeneous electro-catalysis by redox mediators; and possible side reactions of the relevant electrolyte solutions with the highly reactive oxygen moieties which are formed in Li-O₂ cells. Special attention was paid to the recently discovered role of the reactive singlet oxygen which forms during disproportionation of Li-superoxide to Li-peroxide.

The article covers all aspects of major components: anodes (including search for alternative, less reactive ones), cathodes, and electrolyte systems - including organic or ionic liquid-based solutions, and solid state electrolytes (ceramic and polymeric). We described also innovative efforts to develop new solvents that can be more stable towards the reactive oxygen moieties in these cells. In addition, we have covered a detailed description of basic science aspects. The article analyzes all facets of the gap between the state-of-the-art Li-O₂ cells and practical Li-air battery technology.

We have covered past, present and future perspectives of this challenging field. After great initial enthusiasm that also included some unsupported promises, the large research community working on active metals and air battery technologies, has reached an appropriate maturity and realistic approaches to the R&D efforts of these systems. Work is in progress and further work and investments of innovative efforts and means is required. It should also be noted that the field can further expand to include Li-Li₂O closed systems, high temperature solid state Li and Na air batteries and Li (Na) -CO₂ batteries.

The risk of failure is obvious, but the benefits of success are enormous. Developing practical rechargeable active metal–air batteries will revolutionize the field of electrochemical power sources and will further promote the electro-mobility revolution.

References

- (1) Zubi, G.; Dufó-López, R.; Carvalho, M.; Pasaoglu, G. The Lithium-Ion Battery: State of the Art and Future Perspectives. *Renew. Sustain. Energy Rev.* **2018**, *89*, 292–308. <https://doi.org/10.1016/J.RSER.2018.03.002>.
- (2) Thackeray, M. M.; Wolverton, C.; Isaacs, E. D. Electrical Energy Storage for Transportation—approaching the Limits Of, and Going Beyond, Lithium-Ion Batteries. *Energy Environ. Sci.* **2012**, *5*, 7854–7863. <https://doi.org/10.1039/c2ee21892e>.
- (3) Li, M.; Lu, J.; Chen, Z.; Amine, K. 30 Years of Lithium-Ion Batteries. *Adv. Mater.* **2018**, *30*, 1800561. <https://doi.org/10.1002/adma.201800561>.
- (4) Choi, S.; Wang, G. Advanced Lithium-Ion Batteries for Practical Applications: Technology, Development, and Future Perspectives. *Adv. Mater. Technol.* **2018**, *3*, 1700376. <https://doi.org/10.1002/admt.201700376>.
- (5) Sun, Y.-K.; Kim, D.-H.; Yoon, C. S.; Myung, S.-T.; Prakash, J.; Amine, K. A Novel Cathode Material with a Concentration-Gradient for High-Energy and Safe Lithium-Ion Batteries. *Adv. Funct. Mater.* **2010**, *20*, 485–491. <https://doi.org/10.1002/adfm.200901730>.
- (6) Kim, U.; Ryu, H.; Kim, J.; Mücke, R.; Kaghazchi, P.; Yoon, C. S.; Sun, Y. Microstructure-Controlled Ni-Rich Cathode Material by Microscale Compositional Partition for Next-Generation Electric Vehicles. *Adv. Energy Mater.* **2019**, *9*, 1803902. <https://doi.org/10.1002/aenm.201803902>.
- (7) Manthiram, A.; Knight, J. C.; Myung, S.-T.; Oh, S.-M.; Sun, Y.-K. Nickel-Rich and Lithium-Rich Layered Oxide Cathodes: Progress and Perspectives. *Adv. Energy Mater.* **2016**, *6*, 1501010. <https://doi.org/10.1002/aenm.201501010>.
- (8) Markevich, E.; Salitra, G.; Hartmann, P.; Kulisch, J.; Aurbach, D.; Park, K. J.; Yoon, C. S.; Sun, Y. K. New Insights Related to Rechargeable Lithium Batteries: Li Metal Anodes, Ni Rich LiNi_xCo_yMn_zO₂ Cathodes and beyond Them. *J. Electrochem. Soc.* **2019**, *166*, A5265–A5274. <https://doi.org/10.1149/2.0261903jes>.
- (9) Rosy; Sclar, H.; Evenstein, E.; Haber, S.; Maiti, S.; Sharabani, T.; Leskes, M.; Noked,

- M. Mitigating Structural Instability of High-Energy Lithium- and Manganese-Rich $\text{LiNi}_x\text{Mn}_y\text{Co}_z$ Oxide by Interfacial Atomic Surface Reduction. *Chem. Mater.* **2019**, *31*, 3840–3847. <https://doi.org/10.1021/acs.chemmater.9b00875>.
- (10) Weigel, T.; Schipper, F.; Erickson, E. M.; Susai, F. A.; Markovsky, B.; Aurbach, D. Structural and Electrochemical Aspects of $\text{LiNi}_{0.8}\text{Co}_{0.1}\text{Mn}_{0.1}\text{O}_2$ Cathode Materials Doped by Various Cations. *ACS Energy Lett.* **2019**, *4*, 508–516. <https://doi.org/10.1021/acsenergylett.8b02302>.
- (11) Kim, J.; Lee, H.; Cha, H.; Yoon, M.; Park, M.; Cho, J. Prospect and Reality of Ni-Rich Cathode for Commercialization. *Adv. Energy Mater.* **2018**, *8*, 1702028. <https://doi.org/10.1002/aenm.201702028>.
- (12) Schipper, F.; Erickson, E. M.; Erk, C.; Shin, J.-Y.; Chesneau, F. F.; Aurbach, D. Review—Recent Advances and Remaining Challenges for Lithium Ion Battery Cathodes. *J. Electrochem. Soc.* **2017**, *164*, A6220–A6228. <https://doi.org/10.1149/2.0351701jes>.
- (13) Liu, W.; Oh, P.; Liu, X.; Lee, M. J.; Cho, W.; Chae, S.; Kim, Y.; Cho, J. Nickel-Rich Layered Lithium Transition-Metal Oxide for High-Energy Lithium-Ion Batteries. *Angew. Chemie - Int. Ed.* **2015**, *54*, 4440–4457. <https://doi.org/10.1002/anie.201409262>.
- (14) Ryu, H. H.; Park, K. J.; Yoon, C. S.; Sun, Y. K. Capacity Fading of Ni-Rich $\text{Li}[\text{Ni}_x\text{Co}_y\text{Mn}_{1-x-y}]\text{O}_2$ ($0.6 \leq X \leq 0.95$) Cathodes for High-Energy-Density Lithium-Ion Batteries: Bulk or Surface Degradation? *Chem. Mater.* **2018**, *30*, 1155–1163. <https://doi.org/10.1021/acs.chemmater.7b05269>.
- (15) Han, B.; Key, B.; Lapidus, S. H.; Garcia, J. C.; Iddir, H.; Vaughey, J. T.; Dogan, F. From Coating to Dopant: How the Transition Metal Composition Affects Alumina Coatings on Ni-Rich Cathodes. *ACS Appl. Mater. Interfaces* **2017**, *9*, 41291–41302. <https://doi.org/10.1021/acsami.7b13597>.
- (16) Dixit, M.; Markovsky, B.; Schipper, F.; Aurbach, D.; Major, D. T. Origin of Structural Degradation During Cycling and Low Thermal Stability of Ni-Rich Layered Transition Metal-Based Electrode Materials. *J. Phys. Chem. C* **2017**, *121*, 22628–22636. <https://doi.org/10.1021/acs.jpcc.7b06122>.
- (17) Kim, U.-H.; Jun, D.-W.; Park, K.-J.; Zhang, Q.; Kaghazchi, P.; Aurbach, D.; Major, D. T.; Goobes, G.; Dixit, M.; Leifer, N.; et al. Pushing the Limit of Layered Transition Metal Oxide Cathodes for High-Energy Density Rechargeable Li Ion Batteries. *Energy Environ. Sci.* **2018**, *11*, 1271–1279. <https://doi.org/10.1039/C8EE00227D>.
- (18) Markevich, E.; Salitra, G.; Chesneau, F.; Schmidt, M.; Aurbach, D. Very Stable Lithium Metal Stripping–Plating at a High Rate and High Areal Capacity in Fluoroethylene Carbonate-Based Organic Electrolyte Solution. *ACS Energy Lett.* **2017**, *2*, 1321–1326. <https://doi.org/10.1021/acsenergylett.7b00300>.
- (19) Liu, D.; Zhu, W.; Kim, C.; Cho, M.; Guerfi, A.; Delp, S. A.; Allen, J. L.; Jow, T. R.; Zaghbi, K. High-Energy Lithium-Ion Battery Using Substituted LiCoPO_4 : From Coin Type to 1 Ah Cell. *J. Power Sources* **2018**, *388*, 52–56. <https://doi.org/10.1016/J.JPOWSOUR.2018.03.077>.
- (20) Li, W.; Song, B.; Manthiram, A. High-Voltage Positive Electrode Materials for Lithium-Ion Batteries. *Chem. Soc. Rev.* **2017**, *46*, 3006–3059. <https://doi.org/10.1039/C6CS00875E>.
- (21) Zhang, M.; Garcia-Araez, N.; Hector, A. L. Understanding and Development of Olivine LiCoPO_4 Cathode Materials for Lithium-Ion Batteries. *J. Mater. Chem. A* **2018**, *5*, 14483–14517. <https://doi.org/10.1039/c8ta04063j>.
- (22) Choi, J. W.; Aurbach, D. Promise and Reality of Post-Lithium-Ion Batteries with High Energy Densities. *Nat. Rev. Mater.* **2016**, *1*, 16013.

- <https://doi.org/10.1038/natrevmats.2016.13>.
- (23) Zuo, X.; Zhu, J.; Müller-Buschbaum, P.; Cheng, Y.-J. Silicon Based Lithium-Ion Battery Anodes: A Chronicle Perspective Review. *Nano Energy* **2017**, *31*, 113–143. <https://doi.org/10.1016/J.NANOEN.2016.11.013>.
- (24) Feng, K.; Li, M.; Liu, W.; Kashkooli, A. G.; Xiao, X.; Cai, M.; Chen, Z. Silicon-Based Anodes for Lithium-Ion Batteries: From Fundamentals to Practical Applications. *Small* **2018**, *14*, 1702737. <https://doi.org/10.1002/sml.201702737>.
- (25) Wu, H.; Cui, Y. Designing Nanostructured Si Anodes for High Energy Lithium Ion Batteries. *Nano Today* **2012**, *7*, 414–429. <https://doi.org/10.1016/J.NANTOD.2012.08.004>.
- (26) Croguennec, L.; Palacin, M. R. Recent Achievements on Inorganic Electrode Materials for Lithium-Ion Batteries. *J. Am. Chem. Soc.* **2015**, *137*, 3140–3156. <https://doi.org/10.1021/ja507828x>.
- (27) Wang, K.-X.; Li, X.-H.; Chen, J.-S. Surface and Interface Engineering of Electrode Materials for Lithium-Ion Batteries. *Adv. Mater.* **2015**, *27*, 527–545. <https://doi.org/10.1002/adma.201402962>.
- (28) Markevich, E.; Salitra, G.; Aurbach, D. Fluoroethylene Carbonate as an Important Component for the Formation of an Effective Solid Electrolyte Interphase on Anodes and Cathodes for Advanced Li-Ion Batteries. *ACS Energy Lett.* **2017**, *2*, 1337–1345. <https://doi.org/10.1021/acsenergylett.7b00163>.
- (29) Erickson, E. M.; Markevich, E.; Salitra, G.; Sharon, D.; Hirshberg, D.; de la Llave, E.; Shterenberg, I.; Rosenman, A.; Frimer, A.; Aurbach, D. Review—Development of Advanced Rechargeable Batteries: A Continuous Challenge in the Choice of Suitable Electrolyte Solutions. *J. Electrochem. Soc.* **2015**, *162*, A2424–A2438. <https://doi.org/10.1149/2.0051514jes>.
- (30) Delmas, C. Sodium and Sodium-Ion Batteries: 50 Years of Research. *Adv. Energy Mater.* **2018**, *8*, 1703137. <https://doi.org/10.1002/aenm.201703137>.
- (31) Larcher, D.; Tarascon, J.-M. Towards Greener and More Sustainable Batteries for Electrical Energy Storage. *Nat. Chem.* **2015**, *7*, 19–29. <https://doi.org/10.1038/nchem.2085>.
- (32) Kubota, K.; Dahbi, M.; Hosaka, T.; Kumakura, S.; Komaba, S. Towards K-Ion and Na-Ion Batteries as “Beyond Li-Ion.” *Chem. Rec.* **2018**, *18*, 459–479. <https://doi.org/10.1002/tcr.201700057>.
- (33) Vaalma, C.; Buchholz, D.; Weil, M.; Passerini, S. A Cost and Resource Analysis of Sodium-Ion Batteries. *Nature Reviews Materials.* **2018**, *3*, 18013. <https://doi.org/10.1038/natrevmats.2018.13>.
- (34) Luntz, A. Beyond Lithium Ion Batteries. *J. Phys. Chem. Lett.* **2015**, *6*, 300–301. <https://doi.org/10.1021/jz502665r>.
- (35) Shen, X.; Liu, H.; Cheng, X.-B.; Yan, C.; Huang, J.-Q. Beyond Lithium Ion Batteries: Higher Energy Density Battery Systems Based on Lithium Metal Anodes. *Energy Storage Mater.* **2018**, *12*, 161–175. <https://doi.org/10.1016/J.ENSM.2017.12.002>.
- (36) Neburchilov, V.; Zhang, J.; Zhang, J. *Metal-Air and Metal-Sulfur Batteries*; Neburchilov, V., Zhang, J., Eds.; CRC Press: Boca Raton : Taylor & Francis, CRC Press, 2016. | Series:, 2016. <https://doi.org/10.1201/9781315372280>.
- (37) Aurbach, D.; McCloskey, B. D.; Nazar, L. F.; Bruce, P. G. Advances in Understanding Mechanisms Underpinning Lithium–air Batteries. *Nat. Energy* **2016**, *1*, 16128. <https://doi.org/10.1038/nenergy.2016.128>.
- (38) Bruce, P. G.; Freunberger, S. A.; Hardwick, L. J.; Tarascon, J.-M. Li–O₂ and Li–S Batteries with High Energy Storage. *Nat. Mater.* **2012**, *11*, 19–29. <https://doi.org/10.1038/nmat3191>.

- (39) Seh, Z. W.; Sun, Y.; Zhang, Q.; Cui, Y. Designing High-Energy Lithium–sulfur Batteries. *Chem. Soc. Rev.* **2016**, *45*, 5605–5634. <https://doi.org/10.1039/C5CS00410A>.
- (40) Manthiram, A.; Chung, S.-H.; Zu, C. Lithium-Sulfur Batteries: Progress and Prospects. *Adv. Mater.* **2015**, *27*, 1980–2006. <https://doi.org/10.1002/adma.201405115>.
- (41) Zhang, X.; Wang, X.-G.; Xie, Z.; Zhou, Z. Recent Progress in Rechargeable Alkali Metal–air Batteries. *Green Energy Environ.* **2016**, *1*, 4–17. <https://doi.org/10.1016/J.GEE.2016.04.004>.
- (42) Li, Y.; Lu, J. Metal–Air Batteries: Will They Be the Future Electrochemical Energy Storage Device of Choice? *ACS Energy Lett.* **2017**, *2*, 1370–1377. <https://doi.org/10.1021/acsenergylett.7b00119>.
- (43) Blurton, K. F.; Sammells, A. F. Metal/air Batteries: Their Status and Potential — a Review. *J. Power Sources* **1979**, *4*, 263–279. [https://doi.org/10.1016/0378-7753\(79\)80001-4](https://doi.org/10.1016/0378-7753(79)80001-4).
- (44) George W Heise. *AIR DEPOLARIZED PRIMARY BATTERY*; 1933.
- (45) Liu, Y.; Sun, Q.; Li, W.; Adair, K. R.; Li, J.; Sun, X. A Comprehensive Review on Recent Progress in Aluminum–air Batteries. *Green Energy Environ.* **2017**, *2*, 246–277. <https://doi.org/10.1016/J.GEE.2017.06.006>.
- (46) Zaromb, S. The Use and Behavior of Aluminum Anodes in Alkaline Primary Batteries. *J. Electrochem. Soc.* **1962**, *109*, 1125–1130. <https://doi.org/10.1149/1.2425257>.
- (47) Hoey, G. R.; Cohen, M. Corrosion of Anodically and Cathodically Polarized Magnesium in Aqueous Media. *J. Electrochem. Soc.* **1958**, *105*, 245. <https://doi.org/10.1149/1.2428817>.
- (48) Sharon, D.; Hirshberg, D.; Afri, M.; Frimer, A. A.; Noked, M.; Aurbach, D. *Aprotic Metal-Oxygen Batteries: Recent Findings and Insights*; Springer Berlin Heidelberg, 2017; Vol. 21, pp 1861–1878. <https://doi.org/10.1007/s10008-017-3590-7>.
- (49) Abraham, K. M. A Polymer Electrolyte-Based Rechargeable Lithium/Oxygen Battery. *J. Electrochem. Soc.* **1996**, *143*, 1–5. <https://doi.org/10.1149/1.1836378>.
- (50) Visco, S. J.; Nimon, V. Y.; Petrov, A.; Pridatko, K.; Goncharenko, N.; Nimon, E.; De Jonghe, L.; Volkovich, Y. M.; Bograchev, D. A. Aqueous and Nonaqueous Lithium–Air Batteries Enabled by Water-Stable Lithium Metal Electrodes. *J. Solid State Electrochem.* **2014**, *18*, 1443–1456. <https://doi.org/10.1007/s10008-014-2427-x>.
- (51) Luntz, A. C.; McCloskey, B. D. Nonaqueous Li–Air Batteries: A Status Report. *Chem. Rev.* **2014**, *114*, 11721–11750. <https://doi.org/10.1021/cr500054y>.
- (52) Imanishi, N.; Yamamoto, O. Rechargeable Lithium–air Batteries: Characteristics and Prospects. *Mater. Today* **2014**, *17*, 24–30. <https://doi.org/10.1016/J.MATTOD.2013.12.004>.
- (53) Gallagher, K. G.; Goebel, S.; Greszler, T.; Mathias, M.; Oelerich, W.; Eroglu, D.; Srinivasan, V. Quantifying the Promise of Lithium–air Batteries for Electric Vehicles. *Energy Environ. Sci.* **2014**, *7*, 1555. <https://doi.org/10.1039/c3ee43870h>.
- (54) Ko, Y.; Park, H.; Kim, B.; Kim, J. S.; Kang, K. Redox Mediators: A Solution for Advanced Lithium–Oxygen Batteries. *Trends Chem.* **2019**, *1*, 349–360. <https://doi.org/10.1016/J.TRECHM.2019.03.016>.
- (55) Jung, K.-N.; Kim, J.; Yamauchi, Y.; Park, M.-S.; Lee, J.-W.; Kim, J. H. Rechargeable Lithium–air Batteries: A Perspective on the Development of Oxygen Electrodes. *J. Mater. Chem. A* **2016**, *4*, 14050–14068. <https://doi.org/10.1039/C6TA04510C>.
- (56) Tan, P.; Jiang, H. R.; Zhu, X. B.; An, L.; Jung, C. Y.; Wu, M. C.; Shi, L.; Shyy, W.; Zhao, T. S. Advances and Challenges in Lithium–Air Batteries. *Appl. Energy* **2017**, *204*, 780–806. <https://doi.org/10.1016/J.APENERGY.2017.07.054>.
- (57) Wang, L.; Zhou, Z.; Yan, X.; Hou, F.; Wen, L.; Luo, W.; Liang, J.; Dou, S. X.

- Engineering of Lithium-Metal Anodes towards a Safe and Stable Battery. *Energy Storage Mater.* **2018**, *14*, 22–48. <https://doi.org/10.1016/J.ENSM.2018.02.014>.
- (58) Guo, Y.; Li, H.; Zhai, T. Reviving Lithium-Metal Anodes for Next-Generation High-Energy Batteries. *Adv. Mater.* **2017**, *29*, 1700007. <https://doi.org/10.1002/adma.201700007>.
- (59) Lin, D.; Liu, Y.; Cui, Y. Reviving the Lithium Metal Anode for High-Energy Batteries. *Nat. Nanotechnol.* **2017**, *12*, 194–206. <https://doi.org/10.1038/nnano.2017.16>.
- (60) Hong, Y.; Zhao, C.; Xiao, Y.; Xu, R.; Xu, J.; Huang, J.; Zhang, Q.; Yu, X.; Li, H. Safe Lithium-Metal Anodes for Li–O₂ Batteries: From Fundamental Chemistry to Advanced Characterization and Effective Protection. *Batter. Supercaps* **2019**, *2*, 638–658. <https://doi.org/10.1002/batt.201900031>.
- (61) Noked, M.; Schroeder, M. A.; Pearse, A. J.; Rubloff, G. W.; Lee, S. B. Protocols for Evaluating and Reporting Li–O₂ Cell Performance. *J. Phys. Chem. Lett.* **2016**, *7*, 211–215. <https://doi.org/10.1021/acs.jpcclett.5b02613>.
- (62) Augustin, M.; Vullum, P. E.; Vullum-Bruer, F.; Svensson, A. M. Inside the Electrode: Looking at Cycling Products in Li/O₂ Batteries. *J. Power Sources* **2019**, *414*, 130–140. <https://doi.org/10.1016/J.JPOWSOUR.2018.12.088>.
- (63) Lim, H.-K.; Lim, H.-D.; Park, K.-Y.; Seo, D.-H.; Gwon, H.; Hong, J.; Goddard, W. A.; Kim, H.; Kang, K. Toward a Lithium–“Air” Battery: The Effect of CO₂ on the Chemistry of a Lithium–Oxygen Cell. *J. Am. Chem. Soc.* **2013**, *135*, 9733–9742. <https://doi.org/10.1021/ja4016765>.
- (64) Assary, R. S.; Lu, J.; Du, P.; Luo, X.; Zhang, X.; Ren, Y.; Curtiss, L. A.; Amine, K. The Effect of Oxygen Crossover on the Anode of a Li–O₂ Battery Using an Ether-Based Solvent: Insights from Experimental and Computational Studies. *ChemSusChem* **2013**, *6*, 51–55. <https://doi.org/10.1002/cssc.201200810>.
- (65) Zhang, X.; Mu, X.; Yang, S.; Wang, P.; Guo, S.; Han, M.; He, P.; Zhou, H. Research Progress for the Development of Li–Air Batteries: Addressing Parasitic Reactions Arising from Air Composition. *Energy Environ. Mater.* **2018**, *1*, 61–74. <https://doi.org/10.1002/eem2.12008>.
- (66) Lai, J.; Xing, Y.; Chen, N.; Li, L.; Wu, F.; Chen, R. A Comprehensive Insight into the Electrolytes for Rechargeable Lithium–Air Batteries. *Angew. Chemie Int. Ed.* **2019**, *122*, 1522–1534. <https://doi.org/10.1002/anie.201903459>.
- (67) Saito, M.; Fujinami, T.; Yamada, S.; Ishikawa, T.; Otsuka, H.; Ito, K.; Kubo, Y. Effects of Li Salt Anions and O₂ Gas on Li Dissolution/deposition Behavior at Li Metal Negative Electrode for Non-Aqueous Li–Air Batteries. *J. Electrochem. Soc.* **2017**, *164*, A2872–A2880. <https://doi.org/10.1149/2.0151713jes>.
- (68) Grande, L.; Paillard, E.; Hassoun, J.; Park, J.-B.; Lee, Y.-J.; Sun, Y.-K.; Passerini, S.; Scrosati, B. The Lithium/Air Battery: Still an Emerging System or a Practical Reality? *Adv. Mater.* **2015**, *27*, 784–800. <https://doi.org/10.1002/adma.201403064>.
- (69) Vegge, T.; Garcia-Lastra, J. M.; Siegel, D. J. Lithium–oxygen Batteries: At a Crossroads? *Curr. Opin. Electrochem.* **2017**, *6*, 100–107. <https://doi.org/10.1016/J.COEELEC.2017.10.014>.
- (70) Abraham, K. M. A Brief History of Non-Aqueous Metal–Air Batteries. In *ECS Transactions*; ECS, 2008; Vol. 3, pp 67–71. <https://doi.org/10.1149/1.2838193>.
- (71) Laoire, C. O.; Mukerjee, S.; Abraham, K. M.; Plichta, E. J.; Hendrickson, M. A. Elucidating the Mechanism of Oxygen Reduction for Lithium–Air Battery Applications. *J. Phys. Chem. C* **2009**, *113*, 20127–20134. <https://doi.org/10.1021/jp908090s>.
- (72) Laoire, C. O.; Mukerjee, S.; Abraham, K. M.; Plichta, E. J.; Hendrickson, M. A.

- Influence of Nonaqueous Solvents on the Electrochemistry of Oxygen in the Rechargeable Lithium–Air Battery. *J. Phys. Chem. C* **2010**, *114*, 9178–9186. <https://doi.org/10.1021/jp102019y>.
- (73) McCloskey, B. D.; Scheffler, R.; Speidel, A.; Girishkumar, G.; Luntz, A. C. On the Mechanism of Nonaqueous Li–O₂ Electrochemistry on C and Its Kinetic Overpotentials: Some Implications for Li–Air Batteries. *J. Phys. Chem. C* **2012**, *116*, 23897–23905. <https://doi.org/10.1021/jp306680f>.
- (74) Visco, S.J., Katz, B.D., Nimon, Y.S., Jonghe, L. C. D. Protected Active Metal Electrode and Battery Cell Structures with Non-Aqueous Interlayer Architecture, 2007.
- (75) Visco, Steven J., Nimon, Yevgeniy S., Katz, B. D. Ionically Conductive Composites for Protection of Active Metal Anodes, United States Patent Application: 0040142244, A1. 40142244, A1, 2004.
- (76) Visco; Steven J.; Nimon; Yevgeniy S.; Katz Bruce D.; De Jonghe Lutgard C. ACTIVE METAL FUEL CELLS, United States Patent Application: 0090286114, A1, 2009.
- (77) Zhang, T.; Imanishi, N.; Takeda, Y.; Yamamoto, O. Aqueous Lithium/Air Rechargeable Batteries. *Chem. Lett.* **2011**, *40*, 668–673. <https://doi.org/10.1246/cl.2011.668>.
- (78) Sunahiro, S.; Matsui, M.; Takeda, Y.; Yamamoto, O.; Imanishi, N. Rechargeable Aqueous Lithium–air Batteries with an Auxiliary Electrode for the Oxygen Evolution. *J. Power Sources* **2014**, *262*, 338–343. <https://doi.org/10.1016/J.JPOWSOUR.2014.03.016>.
- (79) Hasegawa, S.; Imanishi, N.; Zhang, T.; Xie, J.; Hirano, A.; Takeda, Y.; Yamamoto, O. Study on Lithium/air Secondary batteries—Stability of NASICON-Type Lithium Ion Conducting Glass–ceramics with Water. *J. Power Sources* **2009**, *189*, 371–377. <https://doi.org/10.1016/J.JPOWSOUR.2008.08.009>.
- (80) Read, J. Ether-Based Electrolytes for the Lithium/Oxygen Organic Electrolyte Battery. *J. Electrochem. Soc.* **2006**, *153*, A96–A100. <https://doi.org/10.1149/1.2131827>.
- (81) Read, J.; Mutolo, K.; Ervin, M.; Behl, W.; Wolfenstine, J.; Driedger, A.; Foster, D. Oxygen Transport Properties of Organic Electrolytes and Performance of Lithium/Oxygen Battery. *J. Electrochem. Soc.* **2003**, *150*, A1351–1356. <https://doi.org/10.1149/1.1606454>.
- (82) Read, J. Characterization of the Lithium/Oxygen Organic Electrolyte Battery. *J. Electrochem. Soc.* **2002**, *149*, A1190–A1195. <https://doi.org/10.1149/1.1498256>.
- (83) Ogasawara, T.; Débart, A.; Holzapfel, M.; Novák, P.; Bruce, P. G. Rechargeable Li₂O₂ Electrode for Lithium Batteries. *J. Am. Chem. Soc.* **2006**, *128*, 1390–1393. <https://doi.org/10.1021/ja056811q>.
- (84) Freunberger, S. A.; Chen, Y.; Peng, Z.; Griffin, J. M.; Hardwick, L. J.; Bardé, F.; Novák, P.; Bruce, P. G. Reactions in the Rechargeable Lithium–O₂ Battery with Alkyl Carbonate Electrolytes. *J. Am. Chem. Soc.* **2011**, *133*, 8040–8047. <https://doi.org/10.1021/ja2021747>.
- (85) McCloskey, B. D.; Scheffler, R.; Speidel, A.; Bethune, D. S.; Shelby, R. M.; Luntz, A. C. On the Efficacy of Electrocatalysis in Nonaqueous Li–O₂ Batteries. *J. Am. Chem. Soc.* **2011**, *133*, 18038–18041. <https://doi.org/10.1021/ja207229n>.
- (86) Goolsby, A. D.; Sawyer, D. T. The Electrochemical Reduction of Superoxide ION and Oxidation Of Hydroxide Ion in Dimethyl Sulfoxide. *J. Anal. Chem.* **1968**, *40*, 83–86.
- (87) Merritt, M. V.; Sawyer, D. T. Electrochemical Studies of the Reactivity of Superoxide Ion with Several Alkyl Halides in Dimethyl Sulfoxide. *J. Org. Chem.* **1970**, *35*, 2157–2159. <https://doi.org/10.1021/jo00832a011>.
- (88) Sharon, D.; Hirshberg, D.; Afri, M.; Garsuch, A.; Frimer, A. A.; Aurbach, D. Lithium-

- Oxygen Electrochemistry in Non-Aqueous Solutions. *Israel Journal of Chemistry*. 2015, pp 508–520. <https://doi.org/10.1002/ijch.201400135>.
- (89) Hirshberg, D.; Sharon, D.; Afri, M.; Lavi, R.; Frimer, A. A.; Metoki, N.; Eliaz, N.; Kwak, W.; Sun, Y.-K.; Aurbach, D. Shedding Light on the Oxygen Reduction Reaction Mechanism in Ether-Based Electrolyte Solutions: A Study Using Operando UV–Vis Spectroscopy. *ACS Appl. Mater. Interfaces* **2018**, *10*, 10860–10869. <https://doi.org/10.1021/acsami.7b18376>.
- (90) Johnson, L.; Li, C.; Liu, Z.; Chen, Y.; Freunberger, S. A.; Tarascon, J.-M.; Ashok, P. C.; Praveen, B. B.; Dholakia, K.; Bruce, P. G. The Role of LiO₂ Solubility in O₂ Reduction in Aprotic Solvents and Its Consequences for Li–O₂ Batteries. *Nat. Chem.* **2014**, *6*, 1091–1099. <https://doi.org/10.1038/nchem.2101>.
- (91) Gunasekara, I.; Mukerjee, S.; Plichta, E. J.; Hendrickson, M. A.; Abraham, K. M. A Study of the Influence of Lithium Salt Anions on Oxygen Reduction Reactions in Li–Air Batteries. *J. Electrochem. Soc.* **2015**, *162*, A1055–A1066. <https://doi.org/10.1149/2.0841506jes>.
- (92) Sharon, D.; Hirsberg, D.; Afri, M.; Chesneau, F.; Lavi, R.; Frimer, A. a.; Sun, Y.-K.; Aurbach, D. Catalytic Behavior of Lithium Nitrate in Li–O₂ Cells. *ACS Appl. Mater. Interfaces* **2015**, *7*, 16590–16600. <https://doi.org/10.1021/acsami.5b04145>.
- (93) Sharon, D.; Hirsberg, D.; Salama, M.; Afri, M.; Frimer, A. A.; Noked, M.; Kwak, W.-J.; Sun, Y.-K.; Aurbach, D. Mechanistic Role of Li⁺ Dissociation Level in Aprotic Li–O₂ Battery. *ACS Appl. Mater. Interfaces* **2016**, *8*, 5300–5307. <https://doi.org/10.1021/acsami.5b11483>.
- (94) Aetukuri, N. B.; McCloskey, B. D.; Garcia, J. M.; Krupp, L. E.; Viswanathan, V.; Luntz, A. C. Solvating Additives Drive Solution-Mediated Electrochemistry and Enhance Toroid Growth in Non-Aqueous Li–O₂ Batteries. *Nat. Chem.* **2015**, *7*, 50–56. <https://doi.org/10.1038/nchem.2132>.
- (95) Xia, C.; Black, R.; Fernandes, R.; Adams, B.; Nazar, L. F. The Critical Role of Phase-Transfer Catalysis in Aprotic Sodium Oxygen Batteries. *Nat. Chem.* **2015**, *7*, 496–501. <https://doi.org/10.1038/nchem.2260>.
- (96) Schwenke, K. U.; Metzger, M.; Restle, T.; Piana, M.; Gasteiger, H. A. The Influence of Water and Protons on Li₂O₂ Crystal Growth in Aprotic Li–O₂ Cells. *J. Electrochem. Soc.* **2015**, *162*, A573–A584. <https://doi.org/10.1149/2.0201504jes>.
- (97) Meini, S.; Solchenbach, S.; Piana, M.; Gasteiger, H. A. The Role of Electrolyte Solvent Stability and Electrolyte Impurities in the Electrooxidation of Li₂O₂ in Li–O₂ Batteries. *J. Electrochem. Soc.* **2014**, *161*, A1306–A1314. <https://doi.org/10.1149/2.0621409jes>.
- (98) Li, F.; Wu, S.; Li, D.; Zhang, T.; He, P.; Yamada, A.; Zhou, H. The Water Catalysis at Oxygen Cathodes of Lithium–oxygen Cells. *Nat. Commun.* **2015**, *6*, 7843. <https://doi.org/10.1038/ncomms8843>.
- (99) Kwak, W.-J.; Hirshberg, D.; Sharon, D.; Shin, H.-J.; Afri, M.; Park, J.-B.; Garsuch, A.; Chesneau, F. F.; Frimer, A. A.; Aurbach, D.; et al. Understanding the Behavior of Li–oxygen Cells Containing LiI. *J. Mater. Chem. A* **2015**, *3*, 8855–8864. <https://doi.org/10.1039/C5TA01399B>.
- (100) Liu, T.; Leskes, M.; Yu, W.; Moore, A. J.; Zhou, L.; Bayley, P. M.; Kim, G.; Grey, C. P. Cycling Li–O₂ Batteries via LiOH Formation and Decomposition. *Science (80-.)*. **2015**, *350*, 530–533. <https://doi.org/10.1126/science.aac7730>.
- (101) Burke, C. M.; Black, R.; Kochetkov, I. R.; Giordani, V.; Addison, D.; Nazar, L. F.; McCloskey, B. D. Implications of 4 E- Oxygen Reduction via Iodide Redox Mediation in Li–O₂ Batteries. *ACS Energy Lett.* **2016**, *1*, 747–756. <https://doi.org/10.1021/acsenergylett.6b00328>.

- (102) Ganapathy, S.; Adams, B. D.; Stenou, G.; Anastasaki, M. S.; Goubitz, K.; Miao, X. F.; Nazar, L. F.; Wagemaker, M. Nature of Li_2O_2 Oxidation in a Li- O_2 Battery Revealed by Operando X-Ray Diffraction. *J. Am. Chem. Soc.* **2014**, *136*, 16335–16344. <https://doi.org/10.1021/ja508794r>.
- (103) Zhang, Y.; Cui, Q.; Zhang, X.; McKee, W. C.; Xu, Y.; Ling, S.; Li, H.; Zhong, G.; Yang, Y.; Peng, Z. Amorphous Li_2O_2 : Chemical Synthesis and Electrochemical Properties. *Angew. Chemie - Int. Ed.* **2016**, *55*, 10717–10721. <https://doi.org/10.1002/anie.201605228>.
- (104) Dunst, A.; Epp, V.; Hanzu, I.; Freunberger, S. A.; Wilkening, M. Short-Range Li Diffusion vs. Long-Range Ionic Conduction in Nanocrystalline Lithium Peroxide Li_2O_2 —the Discharge Product in Lithium-Air Batteries. *Energy Environ. Sci.* **2014**, *7*, 2739–2752. <https://doi.org/10.1039/C4EE00496E>.
- (105) McCloskey, B. D.; Speidel, A.; Scheffler, R.; Miller, D. C.; Viswanathan, V.; Hummelshøj, J. S.; Nørskov, J. K.; Luntz, A. C. Twin Problems of Interfacial Carbonate Formation in Nonaqueous Li- O_2 Batteries. *J. Phys. Chem. Lett.* **2012**, *3*, 997–1001. <https://doi.org/10.1021/jz300243r>.
- (106) Wandt, J.; Jakes, P.; Granwehr, J.; Gasteiger, H. A.; Eichel, R. A. Singlet Oxygen Formation during the Charging Process of an Aprotic Lithium-Oxygen Battery. *Angew. Chemie - Int. Ed.* **2016**, *55*, 6892–6895. <https://doi.org/10.1002/anie.201602142>.
- (107) Schafzahl, L.; Mahne, N.; Schafzahl, B.; Wilkening, M.; Slugovc, C.; Borisov, S. M.; Freunberger, S. A. Singlet Oxygen during Cycling of the Aprotic Sodium- O_2 Battery. *Angew. Chemie - Int. Ed.* **2017**, *56*, 15728–15732. <https://doi.org/10.1002/anie.201709351>.
- (108) Mahne, N.; Schafzahl, B.; Leypold, C.; Leypold, M.; Grumm, S.; Leitgeb, A.; Strohmeier, G. A.; Wilkening, M.; Fontaine, O.; Kramer, D.; et al. Singlet Oxygen Generation as a Major Cause for Parasitic Reactions during Cycling of Aprotic Lithium-Oxygen Batteries. *Nat. Energy* **2017**, *2*, 17036. <https://doi.org/10.1038/nenergy.2017.36>.
- (109) Mourad, E.; Petit, Y. K.; Spezia, R.; Samojlov, A.; Summa, F. F.; Prehal, C.; Leypold, C.; Mahne, N.; Slugovc, C.; Fontaine, O.; et al. Singlet Oxygen from Cation Driven Superoxide Disproportionation and Consequences for Aprotic metal- O_2 Batteries. *Energy Environ. Sci.* **2019**. <https://doi.org/10.1039/C9EE01453E>.
- (110) Peng, Q.; Chen, J.; Ji, H.; Morita, A.; Ye, S. Origin of the Overpotential for the Oxygen Evolution Reaction on a Well-Defined Graphene Electrode Probed by in Situ Sum Frequency Generation Vibrational Spectroscopy. *J. Am. Chem. Soc.* **2018**, *140*, 15568–15571. <https://doi.org/10.1021/jacs.8b08285>.
- (111) Hong, M.; Yang, C.; Wong, R. A.; Nakao, A.; Choi, H. C.; Byon, H. R. Determining the Facile Routes for Oxygen Evolution Reaction by in Situ Probing of Li- O_2 Cells with Conformal Li_2O_2 Films. *J. Am. Chem. Soc.* **2018**, *140*, 6190–6193. <https://doi.org/10.1021/jacs.8b02003>.
- (112) Liu, C.; Ye, S. In Situ Atomic Force Microscopy (AFM) Study of Oxygen Reduction Reaction on a Gold Electrode Surface in a Dimethyl Sulfoxide (DMSO)-Based Electrolyte Solution. *J. Phys. Chem. C* **2016**, *120*, 25246–25255. <https://doi.org/10.1021/acs.jpcc.6b08718>.
- (113) Wang, Y.; Lai, N.-C.; Lu, Y.-R.; Zhou, Y.; Dong, C.-L.; Lu, Y.-C. A Solvent-Controlled Oxidation Mechanism of Li_2O_2 in Lithium-Oxygen Batteries. *Joule* **2018**, *2*, 2364–2380. <https://doi.org/10.1016/j.joule.2018.07.021>.
- (114) Kang, S.; Mo, Y.; Ong, S. P.; Ceder, G. A Facile Mechanism for Recharging Li_2O_2 in Li- O_2 Batteries. *Chem. Mater.* **2013**, *25*, 3328–3336.

- <https://doi.org/10.1021/cm401720n>.
- (115) Lu, Y.; Shao-Horn, Y. Probing the Reaction Kinetics of the Charge Reactions of Nonaqueous Li–O₂ Batteries. *J. Phys. Chem. Lett.* **2013**, *4*, 93–99. <https://doi.org/10.1021/jz3018368>.
- (116) Zhai, D.; Wang, H.-H.; Yang, J.; Lau, K. C.; Li, K.; Amine, K.; Curtiss, L. A. Disproportionation in Li–O₂ Batteries Based on a Large Surface Area Carbon Cathode. *J. Am. Chem. Soc.* **2013**, *135*, 15364–15372. <https://doi.org/10.1021/ja403199d>.
- (117) Lu, Y.-C.; Gallant, B. M.; Kwabi, D. G.; Harding, J. R.; Mitchell, R. R.; Whittingham, M. S.; Shao-Horn, Y. Lithium–oxygen Batteries: Bridging Mechanistic Understanding and Battery Performance. *Energy Environ. Sci.* **2013**, *6*, 750. <https://doi.org/10.1039/c3ee23966g>.
- (118) Liu, T.; Kim, G.; Casford, M. T. L.; Grey, C. P. Mechanistic Insights into the Challenges of Cycling a Nonaqueous Na–O₂ Battery. *J. Phys. Chem. Lett.* **2016**, *7*, 4841–4846. <https://doi.org/10.1021/acs.jpcclett.6b02267>.
- (119) Hartmann, P.; Bender, C. L.; Vračar, M.; Dürr, A. K.; Garsuch, A.; Janek, J.; Adelhelm, P. A Rechargeable Room-Temperature Sodium Superoxide (NaO₂) Battery. *Nat. Mater.* **2013**, *12*, 228–232. <https://doi.org/10.1038/nmat3486>.
- (120) Cong, G.; Wang, W.; Lai, N.-C.; Liang, Z.; Lu, Y.-C. A High-Rate and Long-Life Organic–oxygen Battery. *Nat. Mater.* **2019**, *18*, 390–396. <https://doi.org/10.1038/s41563-019-0286-7>.
- (121) Freunberger, S. A.; Chen, Y.; Drewett, N. E.; Hardwick, L. J.; Bardé, F.; Bruce, P. G. The Lithium-Oxygen Battery with Ether-Based Electrolytes. *Angew. Chemie Int. Ed.* **2011**, *50*, 8609–8613. <https://doi.org/10.1002/anie.201102357>.
- (122) McCloskey, B. D.; Bethune, D. S.; Shelby, R. M.; Girishkumar, G.; Luntz, A. C. Solvents' Critical Role in Nonaqueous Lithium–Oxygen Battery. *J. Phys. Chem. Lett.* **2011**, *2*, 1161–1166. <https://doi.org/10.1021/jz200352v>.
- (123) McCloskey, B. D.; Valery, A.; Luntz, A. C.; Gowda, S. R.; Wallraff, G. M.; Garcia, J. M.; Mori, T.; Krupp, L. E. Combining Accurate O₂ and Li₂O₂ Assays to Separate Discharge and Charge Stability Limitations in Nonaqueous Li–O₂ Batteries. *J. Phys. Chem. Lett.* **2013**, *4*, 2989–2993. <https://doi.org/10.1021/jz401659f>.
- (124) Bender, C. L.; Schröder, D.; Pinedo, R.; Adelhelm, P.; Janek, J. One- or Two-Electron Transfer? The Ambiguous Nature of the Discharge Products in Sodium-Oxygen Batteries. *Angew. Chemie Int. Ed.* **2016**, *55*, 4640–4649. <https://doi.org/10.1002/anie.201510856>.
- (125) Wang, W.; Lai, N.-C.; Liang, Z.; Wang, Y.; Lu, Y.-C. Superoxide Stabilization and a Universal KO₂ Growth Mechanism in Potassium-Oxygen Batteries. *Angew. Chemie Int. Ed.* **2018**, *57*, 5042–5046. <https://doi.org/10.1002/anie.201801344>.
- (126) Xiao, N.; Rooney, R. T.; Gewirth, A. A.; Wu, Y. The Long-Term Stability of KO₂ in K–O₂ Batteries. *Angew. Chemie - Int. Ed.* **2018**, *57*, 1227–1231. <https://doi.org/10.1002/anie.201710454>.
- (127) McCloskey, B. D.; Garcia, J. M.; Luntz, A. C. Chemical and Electrochemical Differences in Nonaqueous Li–O₂ and Na–O₂ Batteries. *J. Phys. Chem. Lett.* **2014**, *5*, 1230–1235. <https://doi.org/10.1021/jz500494s>.
- (128) Bender, C. L.; Hartmann, P.; Vračar, M.; Adelhelm, P.; Janek, J. On the Thermodynamics, the Role of the Carbon Cathode, and the Cycle Life of the Sodium Superoxide (NaO₂) Battery. *Adv. Energy Mater.* **2014**, *4*, 1301863. <https://doi.org/10.1002/aenm.201301863>.
- (129) Lutz, L.; Yin, W.; Grimaud, A.; Alves Dalla Corte, D.; Tang, M.; Johnson, L.; Azaceta, E.; Sarou-Kanian, V.; Naylor, A. J.; Hamad, S.; et al. High Capacity Na–O₂ Batteries: Key Parameters for Solution-Mediated Discharge. *J. Phys. Chem. C* **2016**,

- 120, 20068–20076. <https://doi.org/10.1021/acs.jpcc.6b07659>.
- (130) Black, R.; Shyamsunder, A.; Adeli, P.; Kundu, D.; Murphy, G. K.; Nazar, L. F. The Nature and Impact of Side Reactions in Glyme-Based Sodium-Oxygen Batteries. *ChemSusChem* **2016**, *9*, 1795–1803. <https://doi.org/10.1002/cssc.201600034>.
- (131) McCloskey, B. D.; Bethune, D. S.; Shelby, R. M.; Mori, T.; Scheffler, R.; Speidel, A.; Sherwood, M.; Luntz, A. C. Limitations in Rechargeability of Li-O₂ Batteries and Possible Origins. *J. Phys. Chem. Lett.* **2012**, *3*, 3043–3047. <https://doi.org/10.1021/jz301359t>.
- (132) Sayed, S. Y.; Yao, K. P. C.; Kwabi, D. G.; Batcho, T. P.; Amanchukwu, C. V.; Feng, S.; Thompson, C. V.; Shao-Horn, Y. Revealing Instability and Irreversibility in Nonaqueous sodium–O₂ Battery Chemistry. *Chem. Commun.* **2016**, *52*, 9691–9694. <https://doi.org/10.1039/C6CC02291J>.
- (133) Adams, B. D.; Black, R.; Williams, Z.; Fernandes, R.; Cuisinier, M.; Berg, E. J.; Novak, P.; Murphy, G. K.; Nazar, L. F. Towards a Stable Organic Electrolyte for the Lithium Oxygen Battery. *Adv. Energy Mater.* **2015**, *5*, 1400867. <https://doi.org/10.1002/aenm.201400867>.
- (134) McCloskey, B. D.; Burke, C. M.; Nichols, J. E.; Renfrew, S. E. Mechanistic Insights for the Development of Li–O₂ Battery Materials: Addressing Li₂O₂ Conductivity Limitations and Electrolyte and Cathode Instabilities. *Chem. Commun.* **2015**, *51*, 12701–12715. <https://doi.org/10.1039/C5CC04620C>.
- (135) Mahne, N.; Fontaine, O.; Thotiyl, M. O.; Wilkening, M.; Freunberger, S. A. Mechanism and Performance of Lithium–oxygen Batteries – a Perspective. *Chem. Sci.* **2017**, *8*, 6716–6729. <https://doi.org/10.1039/C7SC02519J>.
- (136) Sawyer, D. T.; Valentine, J. S. How Super Is Superoxide? *Acc. Chem. Res.* **1981**, *14*, 393–400. <https://doi.org/10.1021/ar00072a005>.
- (137) Aurbach, D.; Daroux, M.; Faguy, P.; Yeager, E. The Electrochemistry of Noble Metal Electrodes in Aprotic Organic Solvents Containing Lithium Salts. *J. Electroanal. Chem. Interfacial Electrochem.* **1991**, *297*, 225–244. [https://doi.org/10.1016/0022-0728\(91\)85370-5](https://doi.org/10.1016/0022-0728(91)85370-5).
- (138) Bryantsev, V. S.; Blanco, M. Computational Study of the Mechanisms of Superoxide-Induced Decomposition of Organic Carbonate-Based Electrolytes. *J. Phys. Chem. Lett.* **2011**, *2*, 379–383. <https://doi.org/10.1021/jz1016526>.
- (139) Bryantsev, V. S. Predicting the Stability of Aprotic Solvents in Li-Air Batteries: pK_a Calculations of Aliphatic C–H Acids in Dimethyl Sulfoxide. *Chem. Phys. Lett.* **2013**, *558*, 42–47. <https://doi.org/10.1016/j.cplett.2012.12.066>.
- (140) Bryantsev, V. S.; Faglioni, F. Predicting Autoxidation Stability of Ether- and Amide-Based Electrolyte Solvents for Li–Air Batteries. *J. Phys. Chem. A* **2012**, *116*, 7128–7138. <https://doi.org/10.1021/jp301537w>.
- (141) Bryantsev, V. S.; Giordani, V.; Walker, W.; Blanco, M.; Zecevic, S.; Sasaki, K.; Uddin, J.; Addison, D.; Chase, G. V. Predicting Solvent Stability in Aprotic Electrolyte Li–Air Batteries: Nucleophilic Substitution by the Superoxide Anion Radical (O₂^{•-}). *J. Phys. Chem. A* **2011**, *115*, 12399–12409. <https://doi.org/10.1021/jp2073914>.
- (142) Bryantsev, V. S.; Uddin, J.; Giordani, V.; Walker, W.; Addison, D.; Chase, G. V. The Identification of Stable Solvents for Nonaqueous Rechargeable Li-Air Batteries. *J. Electrochem. Soc.* **2013**, *160*, A160–A171. <https://doi.org/10.1149/2.027302jes>.
- (143) Genorio, B.; Staszak-Jirkovský, J.; Assary, R. S.; Connell, J. G.; Strmcnik, D.; Diesendruck, C. E.; Lopes, P. P.; Stamenkovic, V. R.; Moore, J. S.; Curtiss, L. A.; et al. Superoxide (Electro)Chemistry on Well-Defined Surfaces in Organic Environments. *J. Phys. Chem. C* **2016**, *120*, 15909–15914.

- <https://doi.org/10.1021/acs.jpcc.5b12230>.
- (144) Zhang, Z.; Lu, J.; Assary, R. S.; Du, P.; Wang, H.-H.; Sun, Y.-K.; Qin, Y.; Lau, K. C.; Greeley, J.; Redfern, P. C.; et al. Increased Stability Toward Oxygen Reduction Products for Lithium-Air Batteries with Oligoether-Functionalized Silane Electrolytes. *J. Phys. Chem. C* **2011**, *115*, 25535–25542. <https://doi.org/10.1021/jp2087412>.
- (145) Carboni, M.; Marrani, A. G.; Spezia, R.; Brutti, S. 1,2-Dimethoxyethane Degradation Thermodynamics in Li–O₂ Redox Environments. *Chem. - A Eur. J.* **2016**, *22*, 17188–17203. <https://doi.org/10.1002/chem.201602375>.
- (146) Liu, B.; Xu, W.; Yan, P.; Sun, X.; Bowden, M. E.; Read, J.; Qian, J.; Mei, D.; Wang, C. M.; Zhang, J. G. Enhanced Cycling Stability of Rechargeable Li–O₂ Batteries Using High-Concentration Electrolytes. *Adv. Funct. Mater.* **2016**, *26*, 605–613. <https://doi.org/10.1002/adfm.201503697>.
- (147) Khetan, A.; Pitsch, H.; Viswanathan, V. Solvent Degradation in Nonaqueous Li–O₂ Batteries: Oxidative Stability versus H-Abstraction. *J. Phys. Chem. Lett.* **2014**, *5*, 2419–2424. <https://doi.org/10.1021/jz501154v>.
- (148) Assary, R. S.; Lau, K. C.; Amine, K.; Sun, Y.-K.; Curtiss, L. A. Interactions of Dimethoxy Ethane with Li₂O₂ Clusters and Likely Decomposition Mechanisms for Li–O₂ Batteries. *J. Phys. Chem. C* **2013**, *117*, 8041–8049. <https://doi.org/10.1021/jp400229n>.
- (149) Li, Q.; Chen, F.; Zhao, W.; Xu, M.; Fang, B.; Zhang, Y.; Duo, L.; Jin, Y.; Sang, F. A Spectroscopic Study on Singlet Oxygen Production from Different Reaction Paths Using Solid Inorganic Peroxides as Starting Materials. *Bull. Korean Chem. Soc.* **2007**, *28*, 1656–1660. <https://doi.org/10.5012/bkcs.2007.28.10.1656>.
- (150) Hassoun, J.; Croce, F.; Armand, M.; Scrosati, B. Investigation of the O₂ Electrochemistry in a Polymer Electrolyte Solid-State Cell. *Angew. Chemie Int. Ed.* **2011**, *50*, 2999–3002. <https://doi.org/10.1002/anie.201006264>.
- (151) Aubry, J. M.; Rigaudy, J.; Ferradini, C.; Pucheault, J. Search for Singlet Oxygen in the Disproportionation of Superoxide Anion. *J. Am. Chem. Soc.* **1981**, *103*, 4965–4966. <https://doi.org/10.1021/ja00406a060>.
- (152) Nanni, E. J.; Birge, R. R.; Hubbard, L. M.; Morrison, M. M.; Sawyer, D. T. Oxidation and Dismutation of Superoxide Ion Solutions to Molecular Oxygen. Singlet vs. Triplet State. *Inorg. Chem.* **1981**, *20*, 737–741. <https://doi.org/10.1021/ic50217a019>.
- (153) Black, R.; Lee, J.-H.; Adams, B.; Mims, C. A.; Nazar, L. F. The Role of Catalysts and Peroxide Oxidation in Lithium-Oxygen Batteries. *Angew. Chemie Int. Ed.* **2013**, *52*, 392–396. <https://doi.org/10.1002/anie.201205354>.
- (154) Ottakam Thotiyl, M. M.; Freunberger, S. A.; Peng, Z.; Bruce, P. G. The Carbon Electrode in Nonaqueous Li–O₂ Cells. *J. Am. Chem. Soc.* **2013**, *135*, 494–500. <https://doi.org/10.1021/ja310258x>.
- (155) Luntz, A. C.; McCloskey, B. D. Li–air Batteries: Importance of Singlet Oxygen. *Nat. Energy* **2017**, *2*, 17056. <https://doi.org/10.1038/nenergy.2017.56>.
- (156) Kwak, W. J.; Kim, H.; Petit, Y. K.; Leybold, C.; Nguyen, T. T.; Mahne, N.; Redfern, P.; Curtiss, L. A.; Jung, H. G.; Borisov, S. M.; et al. Deactivation of Redox Mediators in Lithium-Oxygen Batteries by Singlet Oxygen. *Nat. Commun.* **2019**, *10*, 4–11. <https://doi.org/10.1038/s41467-019-09399-0>.
- (157) Kwak, W.-J.; Freunberger, S. A.; Kim, H.; Park, J.; Nguyen, T. T.; Jung, H.-G.; Byon, H. R.; Sun, Y.-K. Mutual Conservation of Redox Mediator and Singlet Oxygen Quencher in Lithium-Oxygen Batteries. *ACS Catal.* **2019**, acscatal.9b01337. <https://doi.org/10.1021/acscatal.9b01337>.
- (158) Zhang, Y.; Zhang, X.; Wang, J.; McKee, W. C.; Xu, Y.; Peng, Z. Potential-Dependent Generation of O₂ – and LiO₂ and Their Critical Roles in O₂ Reduction to Li₂O₂ in

- Aprotic Li–O₂ Batteries. *J. Phys. Chem. C* **2016**, *120*, 3690–3698. <https://doi.org/10.1021/acs.jpcc.5b12338>.
- (159) Allen, C. J.; Hwang, J.; Kautz, R.; Mukerjee, S.; Plichta, E. J.; Hendrickson, M. A.; Abraham, K. M. Oxygen Reduction Reactions in Ionic Liquids and the Formulation of a General ORR Mechanism for Li–Air Batteries. *J. Phys. Chem. C* **2012**, *116*, 20755–20764. <https://doi.org/10.1021/jp306718v>.
- (160) Sheng, C.; Yu, F.; Wu, Y.; Peng, Z.; Chen, Y. Disproportionation of Sodium Superoxide in Metal–Air Batteries. *Angew. Chemie Int. Ed.* **2018**, *57*, 9906–9910. <https://doi.org/10.1002/anie.201804726>.
- (161) Das, U.; Lau, K. C.; Redfern, P. C.; Curtiss, L. A. Structure and Stability of Lithium Superoxide Clusters and Relevance to Li–O₂ Batteries. *J. Phys. Chem. Lett.* **2014**, *5*, 813–819. <https://doi.org/10.1021/jz500084e>.
- (162) Landa-Medrano, I.; Pinedo, R.; Bi, X.; Ruiz de Larramendi, I.; Lezama, L.; Janek, J.; Amine, K.; Lu, J.; Rojo, T. New Insights into the Instability of Discharge Products in Na–O₂ Batteries. *ACS Appl. Mater. Interfaces* **2016**, *8*, 20120–20127. <https://doi.org/10.1021/acsami.6b06577>.
- (163) Chen, Y.; Jovanov, Z. P.; Gao, X.; Liu, J.; Holc, C.; Johnson, L. R.; Bruce, P. G. High Capacity Surface Route Discharge at the Potassium–O₂ Electrode. *J. Electroanal. Chem.* **2018**, *819*, 542–546. <https://doi.org/10.1016/j.jelechem.2018.03.041>.
- (164) Amanchukwu, C. V.; Chang, H.-H.; Gauthier, M.; Feng, S.; Batcho, T. P.; Hammond, P. T. One-Electron Mechanism in a Gel–Polymer Electrolyte Li–O₂ Battery. *Chem. Mater.* **2016**, *28*, 7167–7177. <https://doi.org/10.1021/acs.chemmater.6b03718>.
- (165) Kwabi, D. G.; Bryantsev, V. S.; Batcho, T. P.; Itkis, D. M.; Thompson, C. V.; Shao-Horn, Y. Experimental and Computational Analysis of the Solvent-Dependent O₂/Li⁺–O₂ – Redox Couple: Standard Potentials, Coupling Strength, and Implications for Lithium–Oxygen Batteries. *Angew. Chemie Int. Ed.* **2016**, *55*, 3129–3134. <https://doi.org/10.1002/anie.201509143>.
- (166) Peng, Z.; Freunberger, S. A.; Hardwick, L. J.; Chen, Y.; Giordani, V.; Bardé, F.; Novák, P.; Graham, D.; Tarascon, J.-M.; Bruce, P. G. Oxygen Reactions in a Non-Aqueous Li⁺ Electrolyte. *Angew. Chemie Int. Ed.* **2011**, *50*, 6351–6355. <https://doi.org/10.1002/anie.201100879>.
- (167) Mo, Y.; Ong, S. P.; Ceder, G. First-Principles Study of the Oxygen Evolution Reaction of Lithium Peroxide in the Lithium–Air Battery. *Phys. Rev. B* **2011**, *84*, 205446. <https://doi.org/10.1103/PhysRevB.84.205446>.
- (168) Huang, J.; Tong, B.; Li, Z.; Zhou, T.; Zhang, J.; Peng, Z. Probing the Reaction Interface in Li–Oxygen Batteries Using Dynamic Electrochemical Impedance Spectroscopy: Discharge–Charge Asymmetry in Reaction Sites and Electronic Conductivity. *J. Phys. Chem. Lett.* **2018**, *9*, 3403–3408. <https://doi.org/10.1021/acs.jpcclett.8b01351>.
- (169) Wang, J.; Zhang, Y.; Guo, L.; Wang, E.; Peng, Z. Identifying Reactive Sites and Transport Limitations of Oxygen Reactions in Aprotic Lithium–O₂ Batteries at the Stage of Sudden Death. *Angew. Chemie Int. Ed.* **2016**, *55*, 5201–5205. <https://doi.org/10.1002/anie.201600793>.
- (170) Yin, W.; Grimaud, A.; Lepoivre, F.; Yang, C.; Tarascon, J. M. Chemical vs Electrochemical Formation of Li₂CO₃ as a Discharge Product in Li–O₂/CO₂ Batteries by Controlling the Superoxide Intermediate. *J. Phys. Chem. Lett.* **2017**, *8*, 214–222. <https://doi.org/10.1021/acs.jpcclett.6b02610>.
- (171) Gowda, S. R.; Brunet, A.; Wallraff, G. M.; McCloskey, B. D. Implications of CO₂ Contamination in Rechargeable Nonaqueous Li–O₂ Batteries. *J. Phys. Chem. Lett.* **2013**, *4*, 276–279. <https://doi.org/10.1021/jz301902h>.

- (172) Renfrew, S. E.; McCloskey, B. D. Residual Lithium Carbonate Predominantly Accounts for First Cycle CO₂ and CO Outgassing of Li-Stoichiometric and Li-Rich Layered Transition-Metal Oxides. *J. Am. Chem. Soc.* **2017**, *139*, 17853–17860. <https://doi.org/10.1021/jacs.7b08461>.
- (173) Zhao, Z.; Huang, J.; Peng, Z. Achilles' Heel of Lithium-Air Batteries: Lithium Carbonate. *Angew. Chemie Int. Ed.* **2018**, *57*, 3874–3886. <https://doi.org/10.1002/anie.201710156>.
- (174) Mahne, N.; Renfrew, S. E.; McCloskey, B. D.; Freunberger, S. A. Electrochemical Oxidation of Lithium Carbonate Generates Singlet Oxygen. *Angew. Chemie Int. Ed.* **2018**, *57*, 5529–5533. <https://doi.org/10.1002/anie.201802277>.
- (175) Schweitzer, C.; Schmidt, R. Physical Mechanisms of Generation and Deactivation of Singlet Oxygen. *Chem. Rev.* **2003**, *103*, 1685–1758. <https://doi.org/10.1021/cr010371d>.
- (176) Petit, Y. K.; Leypold, C.; Mahne, N.; Mourad, E.; Schafzahl, L.; Slugovc, C.; Borisov, S. M.; Freunberger, S. A. DABCONium: An Efficient and High-Voltage Stable Singlet Oxygen Quencher for Metal–O₂ Cells. *Angew. Chemie - Int. Ed.* **2019**, *58*, 6535–6539. <https://doi.org/10.1002/anie.201901869>.
- (177) Das, S.; Højberg, J.; Knudsen, K. B.; Younesi, R.; Johansson, P.; Norby, P.; Vegge, T. Instability of Ionic Liquid-Based Electrolytes in Li–O₂ Batteries. *J. Phys. Chem. C* **2015**, *119*, 18084–18090. <https://doi.org/10.1021/acs.jpcc.5b04950>.
- (178) Piana, M.; Wandt, J.; Meini, S.; Buchberger, I.; Tsiouvaras, N.; Gasteiger, H. A. Stability of a Pyrrolidinium-Based Ionic Liquid in Li–O₂ Cells. *J. Electrochem. Soc.* **2014**, *161*, A1992–A2001. <https://doi.org/10.1149/2.1131412jes>.
- (179) Zhang, S.; Nava, M. J.; Chow, G. K.; Lopez, N.; Wu, G.; Britt, D. R.; Nocera, D. G.; Cummins, C. C. On the Incompatibility of lithium–O₂ Battery Technology with CO₂. *Chem. Sci.* **2017**, *8*, 6117–6122. <https://doi.org/10.1039/C7SC01230F>.
- (180) Kim, J.; Park, H.; Lee, B.; Seong, W. M.; Lim, H.-D.; Bae, Y.; Kim, H.; Kim, W. K.; Ryu, K. H.; Kang, K. Dissolution and Ionization of Sodium Superoxide in Sodium–oxygen Batteries. *Nat. Commun.* **2016**, *7*, 10670. <https://doi.org/10.1038/ncomms10670>.
- (181) Freunberger, S. A. True Performance Metrics in beyond-Intercalation Batteries. *Nat. Energy* **2017**, *2*, 1–4. <https://doi.org/10.1038/nenergy.2017.91>.
- (182) Mizuno, F.; Nakanishi, S.; Kotani, Y.; Yokoishi, S.; Iba, H. Rechargeable Li-Air Batteries with Carbonate-Based Liquid Electrolytes. *Electrochemistry* **2010**, *78*, 403–405. <https://doi.org/10.5796/electrochemistry.78.403>.
- (183) McCloskey, B. D.; Bethune, D. S.; Shelby, R. M.; Girishkumar, G.; Luntz, A. C. Solvents' Critical Role in Nonaqueous Lithium–Oxygen Battery Electrochemistry. *J. Phys. Chem. Lett.* **2011**, *2*, 1161–1166. <https://doi.org/10.1021/jz200352v>.
- (184) Sharon, D.; Etacheri, V.; Garsuch, A.; Afri, M.; Frimer, A. A.; Aurbach, D. On the Challenge of Electrolyte Solutions for Li-Air Batteries: Monitoring Oxygen Reduction and Related Reactions in Polyether Solutions by Spectroscopy and EQCM. *J. Phys. Chem. Lett.* **2013**, *4*, 127–131. <https://doi.org/10.1021/jz3017842>.
- (185) Sharon, D.; Hirshberg, D.; Afri, M.; Frimer, A. A.; Aurbach, D. The Importance of Solvent Selection in Li–O₂ Cells. *Chem. Commun.* **2017**, *53*, 3269–3272. <https://doi.org/10.1039/C6CC09086A>.
- (186) Carbone, L.; Moro, P. T.; Gobet, M.; Munoz, S.; Devany, M.; Greenbaum, S. G.; Hassoun, J. Enhanced Lithium Oxygen Battery Using a Glyme Electrolyte and Carbon Nanotubes. *ACS Appl. Mater. Interfaces* **2018**, *10*, 16367–16375. <https://doi.org/10.1021/acsami.7b19544>.
- (187) Viswanathan, V.; Thygesen, K. S.; Hummelshøj, J. S.; Nørskov, J. K.; Girishkumar, G.;

- McCloskey, B. D.; Luntz, A. C. Electrical Conductivity in Li_2O_2 and Its Role in Determining Capacity Limitations in Non-Aqueous Li-O₂ Batteries. *J. Chem. Phys.* **2011**, *135*, 214704. <https://doi.org/10.1063/1.3663385>.
- (188) Peng, Z.; Freunberger, S. A.; Chen, Y.; Bruce, P. G. A Reversible and Higher-Rate Li-O₂ Battery. *Science*. **2012**, *337*, 563–566. <https://doi.org/10.1126/science.1223985>.
- (189) Xu, D.; Wang, Z.; Xu, J.; Zhang, L.; Zhang, X. Novel DMSO-Based Electrolyte for High Performance Rechargeable Li-O₂ Batteries. *Chem. Commun.* **2012**, *48*, 6948. <https://doi.org/10.1039/c2cc32844e>.
- (190) Trahan, M. J.; Mukerjee, S.; Plichta, E. J.; Hendrickson, M. A.; Abraham, K. M. Studies of Li-Air Cells Utilizing Dimethyl Sulfoxide-Based Electrolyte. *J. Electrochem. Soc.* **2013**, *160*, 259–267. <https://doi.org/10.1149/2.048302jes>.
- (191) Sharon, D.; Afri, M.; Noked, M.; Garsuch, A.; Frimer, A. A.; Aurbach, D. Oxidation of Dimethyl Sulfoxide Solutions by Electrochemical Reduction of Oxygen. *J. Phys. Chem. Lett.* **2013**, *4*, 3115–3119. <https://doi.org/10.1021/jz4017188>.
- (192) Liu, B.; Xu, W.; Yan, P.; Kim, S. T.; Engelhard, M. H.; Sun, X.; Mei, D.; Cho, J.; Wang, C.-M.; Zhang, J.-G. Stabilization of Li Metal Anode in DMSO-Based Electrolytes via Optimization of Salt-Solvent Coordination for Li-O₂ Batteries. *Adv. Energy Mater.* **2017**, *7*, 1602605. <https://doi.org/10.1002/aenm.201602605>.
- (193) Uddin, J.; Bryantsev, V. S.; Giordani, V.; Walker, W.; Chase, G. V.; Addison, D. Lithium Nitrate as Regenerable SEI Stabilizing Agent for Rechargeable Li/O₂ Batteries. *J. Phys. Chem. Lett.* **2013**, *4*, 3760–3765. <https://doi.org/10.1021/jz402025n>.
- (194) Walker, W.; Giordani, V.; Uddin, J.; Bryantsev, V. S.; Chase, G. V.; Addison, D. A Rechargeable Li-O₂ Battery Using a Lithium Nitrate/ *N, N* -Dimethylacetamide Electrolyte. *J. Am. Chem. Soc.* **2013**, *135*, 2076–2079. <https://doi.org/10.1021/ja311518s>.
- (195) Chen, Y.; Freunberger, S. A.; Peng, Z.; Bardé, F.; Bruce, P. G. Li-O₂ Battery with a Dimethylformamide Electrolyte. *J. Am. Chem. Soc.* **2012**, *134*, 7952–7957. <https://doi.org/10.1021/ja302178w>.
- (196) Sharon, D.; Hirsberg, D.; Afri, M.; Garsuch, A.; Frimer, A. A.; Aurbach, D. Reactivity of Amide Based Solutions in Lithium-Oxygen Cells. *J. Phys. Chem. C* **2014**, *118*, 15207–15213. <https://doi.org/10.1021/jp506230v>.
- (197) Lim, H. D.; Lee, B.; Bae, Y.; Park, H.; Ko, Y.; Kim, H.; Kim, J.; Kang, K. Reaction Chemistry in Rechargeable Li-O₂ Batteries. *Chemical Society Reviews*. The Royal Society of Chemistry May 22, 2017, pp 2873–2888. <https://doi.org/10.1039/c6cs00929h>.
- (198) Walker, W.; Giordani, V.; Uddin, J.; Bryantsev, V. S.; Chase, G. V.; Addison, D. A Rechargeable Li-O₂ Battery Using a Lithium Nitrate/ *N,N*-Dimethylacetamide Electrolyte. *J. Am. Chem. Soc.* **2013**, *135*, 2076–2079. <https://doi.org/10.1021/ja311518s>.
- (199) Sharon, D.; Hirsberg, D.; Afri, M.; Chesneau, F.; Lavi, R.; Frimer, A. a.; Sun, Y.-K.; Aurbach, D. “The Catalytic Behavior of Lithium Nitrate in Li-O₂ Cells.” *ACS Appl. Mater. Interfaces* **2015**, *7*, 150709102638004. <https://doi.org/10.1021/acsami.5b04145>.
- (200) Rosy; Akabayov, S.; Leskes, M.; Noked, M. Bifunctional Role of LiNO_3 in Li-O₂ Batteries: Deconvoluting Surface and Catalytic Effects. *ACS Appl. Mater. Interfaces* **2018**, *10*, 29622–29629. <https://doi.org/10.1021/acsami.8b10054>.
- (201) Nasybulin, E.; Xu, W.; Engelhard, M. H.; Nie, Z.; Burton, S. D.; Cosimbescu, L.; Gross, M. E.; Zhang, J. G. Effects of Electrolyte Salts on the Performance of Li-O₂ Batteries. *J. Phys. Chem. C* **2013**, *117*, 2635–2645. <https://doi.org/10.1021/jp311114u>.
- (202) Nakanishi, A.; Thomas, M. L.; Kwon, H. M.; Kobayashi, Y.; Tatara, R.; Ueno, K.; Dokko, K.; Watanabe, M. Electrolyte Composition in Li/O₂ Batteries with LiI Redox

- Mediators: Solvation Effects on Redox Potentials and Implications for Redox Shuttling. *J. Phys. Chem. C* **2018**, *122*, 1522–1534. <https://doi.org/10.1021/acs.jpcc.7b11859>.
- (203) Kwak, W.-J.; Hirshberg, D.; Sharon, D.; Afri, M.; Frimer, A. A.; Jung, H.-G.; Aurbach, D.; Sun, Y.-K. Li–O₂ Cells with LiBr as an Electrolyte and a Redox Mediator. *Energy Environ. Sci.* **2016**, *9*, 2334–2345. <https://doi.org/10.1039/C6EE00700G>.
- (204) Sharon, D.; Sharon, P.; Hirshberg, D.; Salama, M.; Afri, M.; Shimon, L. J. W.; Kwak, W. J.; Sun, Y. K.; Frimer, A. A.; Aurbach, D. 2,4-Dimethoxy-2,4-Dimethylpentan-3-One: An Aprotic Solvent Designed for Stability in Li–O₂ Cells. *J. Am. Chem. Soc.* **2017**, *139*, 11690–11693. <https://doi.org/10.1021/jacs.7b06414>.
- (205) Huang, Z.; Zeng, H.; Xie, M.; Lin, X.; Huang, Z.; Shen, Y.; Huang, Y. A Stable Lithium–Oxygen Battery Electrolyte Based on Fully Methylated Cyclic Ether. *Angew. Chemie - Int. Ed.* **2019**, *58*, 2345–2349. <https://doi.org/10.1002/anie.201812983>.
- (206) Dong, Q.; Yao, X.; Zhao, Y.; Qi, M.; Zhang, X.; Sun, H.; He, Y.; Wang, D. Cathodically Stable Li–O₂ Battery Operations Using Water-in-Salt Electrolyte. *Chem* **2018**, *4*, 1345–1358. <https://doi.org/10.1016/j.chempr.2018.02.015>.
- (207) Zhao, Q.; Zhang, Y.; Sun, G.; Cong, L.; Sun, L.; Xie, H.; Liu, J. Binary Mixtures of Highly Concentrated Tetraglyme and Hydrofluoroether as a Stable and Nonflammable Electrolyte for Li–O₂ Batteries. *ACS Appl. Mater. Interfaces* **2018**, *10*, 26312–26319. <https://doi.org/10.1021/acsami.8b08346>.
- (208) Gao, X.; Chen, Y.; Johnson, L. R.; Jovanov, Z. P.; Bruce, P. G. A Rechargeable Lithium–oxygen Battery with Dual Mediators Stabilizing the Carbon Cathode. *Nat. Energy* **2017**, *2*, 17118. <https://doi.org/10.1038/nenergy.2017.118>.
- (209) Lu, Y.-C.; Gasteiger, H. A.; Parent, M. C.; Chiloyan, V.; Shao-Horn, Y. The Influence of Catalysts on Discharge and Charge Voltages of Rechargeable Li–Oxygen Batteries. *Electrochem. Solid-State Lett.* **2010**, *13*, A69–A72. <https://doi.org/10.1149/1.3363047>.
- (210) Woo, H.; Kang, J.; Kim, J.; Kim, C.; Nam, S.; Park, B. Development of Carbon-Based Cathodes for Li–Air Batteries: Present and Future. *Electron. Mater. Lett.* **2016**, *12*, 551–567. <https://doi.org/10.1007/s13391-016-6183-1>.
- (211) Ding, N.; Chien, S. W.; Hor, T. S. A.; Lum, R.; Zong, Y.; Liu, Z. Influence of Carbon Pore Size on the Discharge Capacity of Li–O₂ Batteries. *J. Mater. Chem. A* **2014**, *2*, 12433–12441. <https://doi.org/10.1039/c4ta01745e>.
- (212) Nie, H.; Zhang, Y.; Li, J.; Zhou, W.; Lai, Q.; Liu, T.; Zhang, H. Synthesis of a Meso–macro Hierarchical Porous Carbon Material for Improvement of O₂ Diffusivity in Li–O₂ Batteries. *RSC Adv.* **2014**, *4*, 17141. <https://doi.org/10.1039/c4ra01940g>.
- (213) Xiao, J.; Mei, D.; Li, X.; Xu, W.; Wang, D.; Graff, G. L.; Bennett, W. D.; Nie, Z.; Saraf, L. V.; Aksay, I. A.; et al. Hierarchically Porous Graphene as a Lithium–Air Battery Electrode. *Nano Lett.* **2011**, *11*, 5071–5078. <https://doi.org/10.1021/nl203332e>.
- (214) Zhang, W.; Zhu, J.; Ang, H.; Zeng, Y.; Xiao, N.; Gao, Y.; Liu, W.; Hng, H. H.; Yan, Q. Binder-Free Graphene Foams for O₂ Electrodes of Li–O₂ Batteries. *Nanoscale* **2013**, *5*, 9651–9658. <https://doi.org/10.1039/c3nr03321j>.
- (215) Mitchell, R. R.; Gallant, B. M.; Thompson, C. V.; Shao-Horn, Y. All-Carbon-Nanofiber Electrodes for High-Energy Rechargeable Li–O₂ Batteries. *Energy Environ. Sci.* **2011**, *4*, 2952–2958.
- (216) Zhao, C.; Yu, C.; Liu, S.; Yang, J.; Fan, X.; Huang, H.; Qiu, J. 3D Porous N-Doped Graphene Frameworks Made of Interconnected Nanocages for Ultrahigh-Rate and Long-Life Li–O₂ Batteries. *Adv. Funct. Mater.* **2015**, n/a–n/a. <https://doi.org/10.1002/adfm.201503077>.
- (217) Gao, X.; Chen, Y.; Johnson, L.; Bruce, P. G. Promoting Solution Phase Discharge in

- Li-O₂ Batteries Containing Weakly Solvating Electrolyte Solutions. *Nat. Mater.* **2016**, *15*, 882–888. <https://doi.org/10.1038/nmat4629>.
- (218) Gallant, B. M.; Mitchell, R. R.; Kwabi, D. G.; Zhou, J.; Zuin, L.; Thompson, C. V.; Shao-Horn, Y. Chemical and Morphological Changes of Li–O₂ Battery Electrodes upon Cycling. *J. Phys. Chem. C* **2012**, *116*, 20800–20805. <https://doi.org/10.1021/jp308093b>.
- (219) Carboni, M.; Brutti, S.; Marrani, A. G. Surface Reactivity of a Carbonaceous Cathode in a Lithium Triflate/Ether Electrolyte-Based Li-O₂ Cell. *ACS Appl Mater Interfaces* **2015**, *7*, 21751–21762. <https://doi.org/10.1021/acsami.5b05202>.
- (220) Itkis, D. M.; Semenenko, D. A.; Kataev, E. Y.; Belova, A. I.; Neudachina, V. S.; Sirotina, A. P.; Hävecker, M.; Teschner, D.; Knop-Gericke, A.; Dudin, P.; et al. Reactivity of Carbon in Lithium-Oxygen Battery Positive Electrodes. *Nano Lett.* **2013**, *13*, 4697–4701. <https://doi.org/10.1021/nl4021649>.
- (221) Liang, Z.; Lu, Y. C. Critical Role of Redox Mediator in Suppressing Charging Instabilities of Lithium-Oxygen Batteries. *J. Am. Chem. Soc.* **2016**, *138*, 7574–7583. <https://doi.org/10.1021/jacs.6b01821>.
- (222) Ottakam Thotiyl, M. M.; Freunberger, S. A.; Peng, Z.; Chen, Y.; Liu, Z.; Bruce, P. G. A Stable Cathode for the Aprotic Li-O₂ Battery. *Nat. Mater.* **2013**, *12*, 1050–1056. <https://doi.org/10.1038/nmat3737>.
- (223) Wang, Z.; Sun, J.; Cheng, Y.; Niu, C. Adsorption and Deposition of Li₂O₂ on TiC {111} Surface. *J. Phys. Chem. Lett.* **2014**, *5*, 3919–3923.
- (224) Adams, B.; Black, R.; Radtke, C.; Williams, Z.; Mehdi, B. L.; Browning, N. D.; Nazar, L. The Importance of Nanometric Passivating Films on Cathodes for Li Air Batteries. *ACS Nano* **2014**, *8*, 12483–12493.
- (225) Kundu, D.; Black, R.; Adams, B.; Harrison, K.; Zavadil, K.; Nazar, L. F. Nanostructured Metal Carbides for Aprotic Li–O₂ Batteries: New Insights into Interfacial Reactions and Cathode Stability. *J. Phys. Chem. Lett.* **2015**, *6*, 2252–2258. <https://doi.org/10.1021/acs.jpcclett.5b00721>.
- (226) Kundu, D.; Black, R.; Berg, E. J.; Nazar, L. F. A Highly Active Nanostructured Metallic Oxide Cathode for Aprotic Li–O₂ batteries. *Energy Environ. Sci.* **2015**, *8*, 1292–1298. <https://doi.org/10.1039/c4ee02587c>.
- (227) Débart, A.; Paterson, A. J.; Bao, J.; Bruce, P. G. α -MnO₂ Nanowires: A Catalyst for the O₂ Electrode in Rechargeable Lithium Batteries. *Angew. Chemie Int. Ed.* **2008**, *47*, 4521–4524. <https://doi.org/10.1002/anie.200705648>.
- (228) Black, R.; Lee, J.-H.; Adams, B.; Mims, C. A.; Nazar, L. F. The Role of Catalysts and Peroxide Oxidation in Lithium–Oxygen Batteries. *Angew. Chemie Int. Ed.* **2013**, *52*, 392–396. <https://doi.org/10.1002/anie.201205354>.
- (229) Wang, Y.; Zhou, H. A Lithium-Air Battery with a Potential to Continuously Reduce O₂ from Air for Delivering Energy. *J. Power Sources* **2010**, *195*, 358–361. <https://doi.org/10.1016/j.jpowsour.2009.06.109>.
- (230) Lu, J.; Lei, Y.; Lau, K. C.; Luo, X.; Du, P.; Wen, J.; Assary, R. S.; Das, U.; Miller, D. J.; Elam, J. W.; et al. A Nanostructured Cathode Architecture for Low Charge Overpotential in Lithium-Oxygen Batteries. *Nat. Commun.* **2013**, *4*, 2383. <https://doi.org/10.1038/ncomms3383>.
- (231) Amanchukwu, C. V.; Harding, J. R.; Shao-Horn, Y.; Hammond, P. T. Understanding the Chemical Stability of Polymers for Lithium-Air Batteries. *Chem. Mater.* **2015**, *27*, 550–561. <https://doi.org/10.1021/cm5040003>.
- (232) Black, R.; Oh, S. H.; Lee, J.-H.; Yim, T.; Adams, B.; Nazar, L. F. Screening for Superoxide Reactivity in Li-O₂ Batteries: Effect on Li₂O₂ /LiOH Crystallization. *J. Am. Chem. Soc.* **2012**, *134*, 2902–2905. <https://doi.org/10.1021/ja2111543>.

- (233) Papp, J. K.; Forster, J. D.; Burke, C. M.; Kim, H. W.; Luntz, A. C.; Shelby, R. M.; Urban, J. J.; McCloskey, B. D. Poly(vinylidene Fluoride) (PVDF) Binder Degradation in Li-O₂ Batteries: A Consideration for the Characterization of Lithium Superoxide. *J. Phys. Chem. Lett.* **2017**, *8*, 1169–1174. <https://doi.org/10.1021/acs.jpcclett.7b00040>.
- (234) Shui, J.-L.; Okasinski, J. S.; Kenesei, P.; Dobbs, H. A.; Zhao, D.; Almer, J. D.; Liu, D.-J. Reversibility of Anodic Lithium in Rechargeable Lithium–oxygen Batteries. *Nat. Commun.* **2013**, *4*. <https://doi.org/10.1038/ncomms3255>.
- (235) Sun, F.; Gao, R.; Zhou, D.; Osenberg, M.; Dong, K.; Kardjilov, N.; Hilger, A.; Markötter, H.; Bieker, P. M.; Liu, X.; et al. Revealing Hidden Facts of Li Anode in Cycled Lithium–Oxygen Batteries through X-Ray and Neutron Tomography. *ACS Energy Lett.* **2019**, *4*, 306–316. <https://doi.org/10.1021/acsenergylett.8b02242>.
- (236) Tikekar, M. D.; Choudhury, S.; Tu, Z.; Archer, L. A. Design Principles for Electrolytes and Interfaces for Stable Lithium-Metal Batteries. *Nat. Energy* **2016**, *1*, 16114. <https://doi.org/10.1038/nenergy.2016.114>.
- (237) Kozen, A. C.; Lin, C. F.; Pearse, A. J.; Schroeder, M. A.; Han, X.; Hu, L.; Lee, S. B.; Rubloff, G. W.; Noked, M. Next-Generation Lithium Metal Anode Engineering via Atomic Layer Deposition. *ACS Nano* **2015**, *9*, 5884–5892. <https://doi.org/10.1021/acsnano.5b02166>.
- (238) Fan, L.; Zhuang, H. L.; Gao, L.; Lu, Y.; Archer, L. A. Regulating Li Deposition at Artificial Solid Electrolyte Interphases. *J. Mater. Chem. A* **2017**, *5*, 3483–3492. <https://doi.org/10.1039/c6ta10204b>.
- (239) Ma, L.; Kim, M. S.; Archer, L. A. Stable Artificial Solid Electrolyte Interphases for Lithium Batteries. *Chem. Mater.* **2017**, *29*, 4181–4189. <https://doi.org/10.1021/acs.chemmater.6b03687>.
- (240) Liu, W.; Lin, D.; Pei, A.; Cui, Y. Stabilizing Lithium Metal Anodes by Uniform Li-Ion Flux Distribution in Nanochannel Confinement. *J. Am. Chem. Soc.* **2016**, *138*, 15443–15450. <https://doi.org/10.1021/jacs.6b08730>.
- (241) Rosy; Noked, M. Multifunctional Interphase. *Nat. Energy* **2018**, *3*, 253–254. <https://doi.org/10.1038/s41560-018-0114-3>.
- (242) Ma, G.; Wen, Z.; Wu, M.; Shen, C.; Wang, Q.; Jin, J.; Wu, X. A Lithium Anode Protection Guided Highly-Stable Lithium-Sulfur Battery. *Chem. Commun.* **2014**, *50*, 14209–14212. <https://doi.org/10.1039/c4cc05535g>.
- (243) Wu, M.; Wen, Z.; Liu, Y.; Wang, X.; Huang, L. Electrochemical Behaviors of a Li₃N Modified Li Metal Electrode in Secondary Lithium Batteries. *J. Power Sources* **2011**, *196*, 8091–8097. <https://doi.org/10.1016/j.jpowsour.2011.05.035>.
- (244) Huang, Z.; Ren, J.; Zhang, W.; Xie, M.; Li, Y.; Sun, D.; Shen, Y.; Huang, Y. Protecting the Li-Metal Anode in a Li-O₂ Battery by Using Boric Acid as an SEI-Forming Additive. *Adv. Mater.* **2018**, *30*, 1803270. <https://doi.org/10.1002/adma.201803270>.
- (245) Liu, Q.-C.; Xu, J.-J.; Yuan, S.; Chang, Z.-W.; Xu, D.; Yin, Y.-B.; Li, L.; Zhong, H.-X.; Jiang, Y.-S.; Yan, J.-M.; et al. Artificial Protection Film on Lithium Metal Anode toward Long-Cycle-Life Lithium-Oxygen Batteries. *Adv. Mater.* **2015**, *27*, 5241–5247. <https://doi.org/10.1002/adma.201501490>.
- (246) Kozen, A. C.; Lin, C. F.; Zhao, O.; Lee, S. B.; Rubloff, G. W.; Noked, M. Stabilization of Lithium Metal Anodes by Hybrid Artificial Solid Electrolyte Interphase. *Chem. Mater.* **2017**, *29*, 6298–6307. <https://doi.org/10.1021/acs.chemmater.7b01496>.
- (247) Togasaki, N.; Momma, T.; Osaka, T. Enhanced Cycling Performance of a Li Metal Anode in a Dimethylsulfoxide-Based Electrolyte Using Highly Concentrated Lithium Salt for a Lithium–oxygen Battery. *J. Power Sources* **2016**, *307*, 98–104. <https://doi.org/10.1016/J.JPOWSOUR.2015.12.123>.

- (248) Tong, B.; Huang, J.; Zhou, Z.; Peng, Z. The Salt Matters: Enhanced Reversibility of Li–O₂ Batteries with a Li[(CF₃SO₂)(N–C₄F₉SO₂)N]-Based Electrolyte. *Adv. Mater.* **2018**, *30*, 1–6. <https://doi.org/10.1002/adma.201704841>.
- (249) Tu, Z.; Zachman, M. J.; Choudhury, S.; Wei, S.; Ma, L.; Yang, Y.; Kourkoutis, L. F.; Archer, L. A. Nanoporous Hybrid Electrolytes for High-Energy Batteries Based on Reactive Metal Anodes. *Adv. Energy Mater.* **2017**, *7*, 1602367. <https://doi.org/10.1002/aenm.201602367>.
- (250) Liu, Y.; Liu, Q.; Xin, L.; Liu, Y.; Yang, F.; Stach, E. A.; Xie, J. Making Li-Metal Electrodes Rechargeable by Controlling the Dendrite Growth Direction. *Nat. Energy* **2017**, *2*, 17083. <https://doi.org/10.1038/nenergy.2017.83>.
- (251) Liu, B.; Zhang, J.-G.; Xu, W. Advancing Lithium Metal Batteries. *Joule* **2018**, *2*, 833–845. <https://doi.org/10.1016/J.JOULE.2018.03.008>.
- (252) Kwak, W.-J.; Jung, H.-G.; Aurbach, D.; Sun, Y.-K. Optimized Bicompartiment Two Solution Cells for Effective and Stable Operation of Li–O₂ Batteries. *Adv. Energy Mater.* **2017**, *7*, 1701232. <https://doi.org/10.1002/aenm.201701232>.
- (253) Kim, B. G.; Kim, J.-S.; Min, J.; Lee, Y.-H.; Choi, J. H.; Jang, M. C.; Freunberger, S. A.; Choi, J. W. A Moisture- and Oxygen-Impermeable Separator for Aprotic Li–O₂ Batteries. *Adv. Funct. Mater.* **2016**, *26*, 1747–1756. <https://doi.org/10.1002/adfm.201504437>.
- (254) Lin, D.; Liu, Y.; Liang, Z.; Lee, H.-W.; Sun, J.; Wang, H.; Yan, K.; Xie, J.; Cui, Y. Layered Reduced Graphene Oxide with Nanoscale Interlayer Gaps as a Stable Host for Lithium Metal Anodes. *Nat. Nanotechnol.* **2016**, *11*, 626–632. <https://doi.org/10.1038/nnano.2016.32>.
- (255) Yang, C.-P.; Yin, Y.-X.; Zhang, S.-F.; Li, N.-W.; Guo, Y.-G. Accommodating Lithium into 3D Current Collectors with a Submicron Skeleton towards Long-Life Lithium Metal Anodes. *Nat. Commun.* **2015**, *6*, 8058. <https://doi.org/10.1038/ncomms9058>.
- (256) Huang, G.; Han, J.; Yang, C.; Wang, Z.; Fujita, T.; Hirata, A.; Chen, M. Graphene-Based Quasi-Solid-State Lithium–oxygen Batteries with High Energy Efficiency and a Long Cycling Lifetime. *NPG Asia Mater.* **2018**, *10*, 1037–1045. <https://doi.org/10.1038/s41427-018-0095-5>.
- (257) Lee, D. J.; Lee, H.; Song, J.; Ryou, M.-H.; Lee, Y. M.; Kim, H.-T.; Park, J.-K. Composite Protective Layer for Li Metal Anode in High-Performance Lithium–oxygen Batteries. *Electrochem. commun.* **2014**, *40*, 45–48. <https://doi.org/10.1016/j.elecom.2013.12.022>.
- (258) Liu, B.; Xu, W.; Tao, J.; Yan, P.; Zheng, J.; Engelhard, M. H.; Lu, D.; Wang, C.; Zhang, J. G. Enhanced Cyclability of Lithium–Oxygen Batteries with Electrodes Protected by Surface Films Induced via In Situ Electrochemical Process. *Adv. Energy Mater.* **2018**, *8*, 1–12. <https://doi.org/10.1002/aenm.201702340>.
- (259) Kwak, W. J.; Shin, H. J.; Reiter, J.; Tsiouvaras, N.; Hassoun, J.; Passerini, S.; Scrosati, B.; Sun, Y. K. Understanding Problems of Lithiated Anodes in Lithium Oxygen Full-Cells. *J. Mater. Chem. A* **2016**, *4*, 10467–10471. <https://doi.org/10.1039/c6ta03013k>.
- (260) Hirshberg, D.; Sharon, D.; De La Llave, E.; Afri, M.; Frimer, A. A.; Kwak, W. J.; Sun, Y. K.; Aurbach, D. Feasibility of Full (Li-Ion)-O₂ Cells Comprised of Hard Carbon Anodes. *ACS Appl. Mater. Interfaces* **2017**, *9*, 4352–4361. <https://doi.org/10.1021/acsami.6b10974>.
- (261) Wu, S.; Zhu, K.; Tang, J.; Liao, K.; Bai, S.; Yi, J.; Yamauchi, Y.; Ishida, M.; Zhou, H. A Long-Life Lithium Ion Oxygen Battery Based on Commercial Silicon Particles as the Anode. *Energy Environ. Sci.* **2016**, *9*, 3262–3271. <https://doi.org/10.1039/c6ee01512c>.
- (262) Qin, L.; Zhai, D.; Lv, W.; Yang, W.; Huang, J.; Yao, S.; Cui, J.; Chong, W.-G.;

- Huang, J.-Q.; Kang, F.; et al. A High-Performance Lithium Ion Oxygen Battery Consisting of Li_2O_2 Cathode and Lithiated Aluminum Anode with Nafion Membrane for Reduced O_2 Crossover. *Nano Energy* **2017**, *40*, 258–263. <https://doi.org/10.1016/j.nanoen.2017.08.029>.
- (263) Elia, G. A.; Bernhard, R.; Hassoun, J. A Lithium-Ion Oxygen Battery Using a Polyethylene Glyme Electrolyte Mixed with an Ionic Liquid. *RSC Adv.* **2015**, *5*, 21360–21365. <https://doi.org/10.1039/c4ra17277a>.
- (264) Zhang, W.; Shen, Y.; Sun, D.; Huang, Z.; Zhou, J.; Yan, H.; Huang, Y. Promoting Li_2O_2 Oxidation via Solvent-Assisted Redox Shuttle Process for Low Overpotential Li- O_2 Battery. *Nano Energy* **2016**, *30*, 43–51. <https://doi.org/10.1016/j.nanoen.2016.09.031>.
- (265) Ma, S.; Wu, Y.; Wang, J.; Zhang, Y.; Zhang, Y.; Yan, X.; Wei, Y.; Liu, P.; Wang, J.; Jiang, K.; et al. Reversibility of Noble Metal-Catalyzed Aprotic Li- O_2 Batteries. *Nano Lett* **2015**, *15*, 8084–8090. <https://doi.org/10.1021/acs.nanolett.5b03510>.
- (266) Gregory V. Chase, Strahinja Zecevic, T. Walker Wesley, Jasim Uddin, Kenji A. Sasaki, P. Giordani Vincent, Vyacheslav Bryantsev, Mario Blanco, D. D. A. Patentes Soluble Oxygen Evolving Catalysts for Rechargeable Metal-Air Batteries. US20120028137A1, April 25, 2011.
- (267) Chen, Y.; Freunberger, S. A.; Peng, Z.; Fontaine, O.; Bruce, P. G. Charging a Li- O_2 Battery Using a Redox Mediator. *Nat. Chem.* **2013**, *5*, 489–494. <https://doi.org/10.1038/nchem.1646>.
- (268) Park, J.-B.; Lee, S. H.; Jung, H.-G.; Aurbach, D.; Sun, Y.-K. Redox Mediators for Li- O_2 Batteries: Status and Perspectives. *Adv. Mater.* **2018**, *30*, 1704162. <https://doi.org/10.1002/adma.201704162>.
- (269) Landa-Medrano, I.; Lozano, I.; Ortiz-Vitoriano, N.; Ruiz de Larramendi, I.; Rojo, T. Redox Mediators: A Shuttle to Efficacy in metal- O_2 Batteries. *J. Mater. Chem. A* **2019**, *7*, 8746–8764. <https://doi.org/10.1039/C8TA12487F>.
- (270) Bergner, B. J.; Schürmann, A.; Peppler, K.; Garsuch, A.; Janek, J. TEMPO: A Mobile Catalyst for Rechargeable Li- O_2 Batteries. *J. Am. Chem. Soc.* **2014**, *136*, 15054–15064. <https://doi.org/10.1021/ja508400m>.
- (271) Bergner, B. J.; Hofmann, C.; Schürmann, A.; Schröder, D.; Peppler, K.; Schreiner, P. R.; Janek, J. Understanding the Fundamentals of Redox Mediators in Li- O_2 Batteries: A Case Study on Nitroxides. *Phys. Chem. Chem. Phys.* **2015**, *17*, 31769–31779. <https://doi.org/10.1039/c5cp04505c>.
- (272) Feng, N.; He, P.; Zhou, H. Enabling Catalytic Oxidation of Li_2O_2 at the Liquid-Solid Interface: The Evolution of an Aprotic Li- O_2 Battery. *ChemSusChem* **2015**, *8*, 600–602. <https://doi.org/10.1002/cssc.201403338>.
- (273) Kundu, D.; Black, R.; Adams, B.; Nazar, L. F. A Highly Active Low Voltage Redox Mediator for Enhanced Rechargeability of Lithium-Oxygen Batteries. *ACS Cent. Sci.* **2015**, *1*, 510–515. <https://doi.org/10.1021/acscentsci.5b00267>.
- (274) Xu, C.; Xu, G.; Zhang, Y.; Fang, S.; Nie, P.; Wu, L.; Zhang, X. Bifunctional Redox Mediator Supported by an Anionic Surfactant for Long-Cycle Li- O_2 Batteries. *ACS Energy Lett.* **2017**, *2*, 2659–2666. <https://doi.org/10.1021/acsenenergylett.7b00884>.
- (275) Matsuda, S.; Mori, S.; Kubo, Y.; Uosaki, K.; Hashimoto, K.; Nakanishi, S. Cobalt Phthalocyanine Analogs as Soluble Catalysts That Improve the Charging Performance of Li- O_2 Batteries. *Chem. Phys. Lett.* **2015**, *620*, 78–81. <https://doi.org/10.1016/j.cplett.2014.12.036>.
- (276) Sun, D.; Shen, Y.; Zhang, W.; Yu, L.; Yi, Z.; Yin, W.; Wang, D.; Huang, Y.; Wang, J.; Wang, D.; et al. A Solution-Phase Bifunctional Catalyst for Lithium-Oxygen Batteries. *J. Am. Chem. Soc.* **2014**, *136*, 8941–8946. <https://doi.org/10.1021/ja501877e>.

- (277) Ryu, W.-H.; Gittleston, F. S.; Thomsen, J. M.; Li, J.; Schwab, M. J.; Brudvig, G. W.; Taylor, A. D. Heme Biomolecule as Redox Mediator and Oxygen Shuttle for Efficient Charging of Lithium-Oxygen Batteries. *Nat. Commun.* **2016**, *7*, 12925. <https://doi.org/10.1038/ncomms12925>.
- (278) Lim, H.-D.; Lee, B.; Zheng, Y.; Hong, J.; Kim, J.; Gwon, H.; Ko, Y.; Lee, M.; Cho, K.; Kang, K. Rational Design of Redox Mediators for Advanced Li–O₂ Batteries. *Nat. Energy* **2016**, *1*, 16066. <https://doi.org/10.1038/nenergy.2016.66>.
- (279) Chen, Y.; Gao, X.; Johnson, L. R.; Bruce, P. G. Kinetics of Lithium Peroxide Oxidation by Redox Mediators and Consequences for the Lithium–oxygen Cell. *Nat. Commun.* **2018**, *9*, 767. <https://doi.org/10.1038/s41467-018-03204-0>.
- (280) Ko, Y.; Park, H.; Lee, B.; Bae, Y.; Park, S. K.; Kang, K. A Comparative Kinetic Study of Redox Mediators for High-Power Lithium–oxygen Batteries. *J. Mater. Chem. A* **2019**, *7*, 6491–6498. <https://doi.org/10.1039/C9TA00096H>.
- (281) Bawol, P. P.; Reinsberg, P.; Bondue, C. J.; Abd-El-Latif, A. A.; Königshoven, P.; Baltruschat, H. A New Thin Layer Cell for Battery Related DEMS-Experiments: The Activity of Redox Mediators in the Li-O₂ Cell. *Phys. Chem. Chem. Phys.* **2018**, *20*, 21447–21456. <https://doi.org/10.1039/c8cp03592j>.
- (282) Zhang, Y.; Wang, L.; Zhang, X.; Guo, L.; Wang, Y.; Peng, Z. High-Capacity and High-Rate Discharging of a Coenzyme Q10-Catalyzed Li–O₂ Battery. *Adv. Mater.* **2018**, *30*, 2–7. <https://doi.org/10.1002/adma.201705571>.
- (283) Ko, Y.; Park, H.; Kim, J.; Lim, H. D.; Lee, B.; Kwon, G.; Lee, S.; Bae, Y.; Park, S. K.; Kang, K. Biological Redox Mediation in Electron Transport Chain of Bacteria for Oxygen Reduction Reaction Catalysts in Lithium–Oxygen Batteries. *Adv. Funct. Mater.* **2019**, *29*, 1–9. <https://doi.org/10.1002/adfm.201805623>.
- (284) Yang, L.; Frith, J. T.; Garcia-Araez, N.; Owen, J. R. A New Method to Prevent Degradation of Lithium-Oxygen Batteries: Reduction of Superoxide by Viologen. *Chem. Commun.* **2015**, *51*, 1705–1708. <https://doi.org/10.1039/c4cc09208b>.
- (285) Tesio, A. Y.; Blasi, D.; Olivares-Marín, M.; Ratera, I.; Tonti, D.; Veciana, J. Organic Radicals for the Enhancement of Oxygen Reduction Reaction in Li–O₂ Batteries. *Chem. Commun.* **2015**, *51*, 17623–17626. <https://doi.org/10.1039/c5cc07242e>.
- (286) Gao, X.; Chen, Y.; Johnson, L.; Bruce, P. G. Promoting Solution Phase Discharge in Li–O₂ Batteries Containing Weakly Solvating Electrolyte Solutions. *Nat. Mater.* **2016**, *15*, 882–888. <https://doi.org/10.1038/nmat4629>.
- (287) Gao, X.; Jovanov, Z. P.; Chen, Y.; Johnson, L. R.; Bruce, P. G. Phenol-Catalyzed Discharge in the Aprotic Lithium-Oxygen Battery. *Angew. Chemie - Int. Ed.* **2017**, *56*, 6539–6543. <https://doi.org/10.1002/anie.201702432>.
- (288) Lee, S. H.; Kwak, W. J.; Sun, Y. K. A New Perspective of the Ruthenium Ion: A Bifunctional Soluble Catalyst for High Efficiency Li–O₂ Batteries. *J. Mater. Chem. A* **2017**, *5*, 15512–15516. <https://doi.org/10.1039/c7ta04070a>.
- (289) Liu, H.; Liu, M.; Yang, L.; Song, Y.; Wang, X.; Yang, K.; Pan, F. A Bi-Functional Redox Mediator Promoting the ORR and OER in Non-Aqueous Li–O₂ Batteries. *Chem. Commun.* **2019**, *55*, 6567–6570. <https://doi.org/10.1039/C9CC02596K>.
- (290) Kwak, W.-J.; Ha, S. H.; Kim, D. H.; Shin, K. H.; Sun, Y.-K.; Lee, Y. J. Synergistic Integration of Soluble Catalysts with Carbon-Free Electrodes for Li–O₂ Batteries. *ACS Catal.* **2017**, *7*, 8192–8199. <https://doi.org/10.1021/acscatal.7b02359>.
- (291) Liang, Z.; Zhou, Y.; Lu, Y. C. Dynamic Oxygen Shield Eliminates Cathode Degradation in Lithium-Oxygen Batteries. *Energy Environ. Sci.* **2018**, *11*, 3500–3510. <https://doi.org/10.1039/c8ee02067a>.
- (292) Shen, X.; Zhang, S.; Wu, Y.; Chen, Y. Promoting Li–O₂ Batteries With Redox Mediators. *ChemSusChem* **2019**, *12*, 104–114.

- <https://doi.org/10.1002/cssc.201802007>.
- (293) Wang, Y.; Xia, Y. Li-O₂ Batteries: An Agent for Change. *Nature Chemistry*. Nature Publishing Group June 2013, pp 445–447. <https://doi.org/10.1038/nchem.1658>.
- (294) Ha, S.; Kim, Y.; Koo, D.; Ha, K. H.; Park, Y.; Kim, D. M.; Son, S.; Yim, T.; Lee, K. T. Investigation into the Stability of Li Metal Anodes in Li-O₂ batteries with a Redox Mediator. *J. Mater. Chem. A* **2017**, *5*, 10609–10621. <https://doi.org/10.1039/c7ta02210g>.
- (295) Lee, D. J.; Lee, H.; Kim, Y.-J.; Park, J.-K.; Kim, H.-T. Sustainable Redox Mediation for Lithium-Oxygen Batteries by a Composite Protective Layer on the Lithium-Metal Anode. *Adv. Mater.* **2016**, *28*, 857–863. <https://doi.org/10.1002/adma.201503169>.
- (296) Bergner, B. J.; Busche, M. R.; Pinedo, R.; Berkes, B. B.; Schröder, D.; Janek, J. How To Improve Capacity and Cycling Stability for Next Generation Li–O₂ Batteries: Approach with a Solid Electrolyte and Elevated Redox Mediator Concentrations. *ACS Appl. Mater. Interfaces* **2016**, *8*, 7756–7765. <https://doi.org/10.1021/acsami.5b10979>.
- (297) Kwak, W.-J.; Kim, H.; Jung, H.-G.; Aurbach, D.; Sun, Y.-K. Review—A Comparative Evaluation of Redox Mediators for Li-O₂ Batteries: A Critical Review. *J. Electrochem. Soc.* **2018**, *165*, A2274–A2293. <https://doi.org/10.1149/2.0901810jes>.
- (298) Pinedo, R.; Weber, D. A.; Bergner, B.; Schröder, D.; Adelhelm, P.; Janek, J. Insights into the Chemical Nature and Formation Mechanisms of Discharge Products in Na–O₂ Batteries by Means of *Operando* X-Ray Diffraction. *J. Phys. Chem. C* **2016**, *120*, 8472–8481. <https://doi.org/10.1021/acs.jpcc.6b00903>.
- (299) Aldous, I. M.; Hardwick, L. J. Solvent-Mediated Control of the Electrochemical Discharge Products of Non-Aqueous Sodium-Oxygen Electrochemistry. *Angew. Chemie Int. Ed.* **2016**, *55*, 8254–8257. <https://doi.org/10.1002/anie.201601615>.
- (300) Xia, C.; Fernandes, R.; Cho, F. H.; Sudhakar, N.; Buonacorsi, B.; Walker, S.; Xu, M.; Baugh, J.; Nazar, L. F. Direct Evidence of Solution-Mediated Superoxide Transport and Organic Radical Formation in Sodium-Oxygen Batteries. *J. Am. Chem. Soc.* **2016**, *138*, 11219–11226. <https://doi.org/10.1021/jacs.6b05382>.
- (301) Yadegari, H.; Li, Y.; Banis, M. N.; Li, X.; Wang, B.; Sun, Q.; Li, R.; Sham, T.-K.; Cui, X.; Sun, X. On Rechargeability and Reaction Kinetics of Sodium–air Batteries. *Energy Environ. Sci.* **2014**, *7*, 3747–3757. <https://doi.org/10.1039/C4EE01654H>.
- (302) Medenbach, L.; Bender, C. L.; Haas, R.; Mogwitz, B.; Pompe, C.; Adelhelm, P.; Schröder, D.; Janek, J. Origins of Dendrite Formation in Sodium-Oxygen Batteries and Possible Countermeasures. *Energy Technol.* **2017**, *5*, 2265–2274. <https://doi.org/10.1002/ente.201700326>.
- (303) Bi, X.; Ren, X.; Huang, Z.; Yu, M.; Kreidler, E.; Wu, Y. Investigating Dendrites and Side Reactions in Sodium–oxygen Batteries for Improved Cycle Lives. *Chem. Commun.* **2015**, *51*, 7665–7668. <https://doi.org/10.1039/C5CC00825E>.
- (304) Hartmann, P.; Bender, C. L.; Sann, J.; Dürr, A. K.; Jansen, M.; Janek, J.; Adelhelm, P. A Comprehensive Study on the Cell Chemistry of the Sodium Superoxide (NaO₂) Battery. *Phys. Chem. Chem. Phys.* **2013**, *15*, 11661–11672. <https://doi.org/10.1039/c3cp50930c>.
- (305) Bender, C. L.; Jache, B.; Adelhelm, P.; Janek, J. Sodiated Carbon: A Reversible Anode for Sodium–oxygen Batteries and Route for the Chemical Synthesis of Sodium Superoxide (NaO₂). *J. Mater. Chem. A* **2015**, *3*, 20633–20641. <https://doi.org/10.1039/C5TA06640A>.
- (306) Ren, X.; Wu, Y. A Low-Overpotential Potassium–Oxygen Battery Based on Potassium Superoxide. *J. Am. Chem. Soc.* **2013**, *135*, 2923–2926. <https://doi.org/10.1021/ja312059q>.
- (307) Ren, X.; Lau, K. C.; Yu, M.; Bi, X.; Kreidler, E.; Curtiss, L. A.; Wu, Y. Understanding

- Side Reactions in K-O₂ Batteries for Improved Cycle Life. *ACS Appl. Mater. Interfaces* **2014**, *6*, 19299–19307. <https://doi.org/10.1021/am505351s>.
- (308) Xiao, N.; McCulloch, W. D.; Wu, Y. Reversible Dendrite-Free Potassium Plating and Stripping Electrochemistry for Potassium Secondary Batteries. *J. Am. Chem. Soc.* **2017**, *139*, 9475–9478. <https://doi.org/10.1021/jacs.7b04945>.
- (309) Xiao, N.; Gourdin, G.; Wu, Y. Simultaneous Stabilization of Potassium Metal and Superoxide in K-O₂ Batteries on the Basis of Electrolyte Reactivity. *Angew. Chemie Int. Ed.* **2018**, *57*, 10864–10867. <https://doi.org/10.1002/anie.201804115>.
- (310) Yu, W.; Lau, K. C.; Lei, Y.; Liu, R.; Qin, L.; Yang, W.; Li, B.; Curtiss, L. A.; Zhai, D.; Kang, F. Dendrite-Free Potassium–Oxygen Battery Based on a Liquid Alloy Anode. *ACS Appl. Mater. Interfaces* **2017**, *9*, 31871–31878. <https://doi.org/10.1021/acsami.7b08962>.
- (311) McCulloch, W. D.; Ren, X.; Yu, M.; Huang, Z.; Wu, Y. Potassium-Ion Oxygen Battery Based on a High Capacity Antimony Anode. *ACS Appl. Mater. Interfaces* **2015**, *7*, 26158–26166. <https://doi.org/10.1021/acsami.5b08037>.
- (312) Sun, X.; Bonnick, P.; Nazar, L. F. Layered TiS₂ Positive Electrode for Mg Batteries. *ACS Energy Lett.* **2016**, *1*, 297–301. <https://doi.org/10.1021/acsenergylett.6b00145>.
- (313) Li, C.-S.; Sun, Y.; Gebert, F.; Chou, S.-L. Current Progress on Rechargeable Magnesium-Air Battery. *Adv. Energy Mater.* **2017**, *7*, 1700869. <https://doi.org/10.1002/aenm.201700869>.
- (314) Shiga, T.; Hase, Y.; Kato, Y.; Inoue, M.; Takechi, K. A Rechargeable Non-Aqueous Mg–O₂ Battery. *Chem. Commun.* **2013**, *49*, 9152. <https://doi.org/10.1039/c3cc43477j>.
- (315) Shiga, T.; Hase, Y.; Yagi, Y.; Takahashi, N.; Takechi, K. Catalytic Cycle Employing a TEMPO–Anion Complex to Obtain a Secondary Mg–O₂ Battery. *J. Phys. Chem. Lett.* **2014**, *5*, 1648–1652. <https://doi.org/10.1021/jz500602r>.
- (316) Vardar, G.; Nelson, E. G.; Smith, J. G.; Naruse, J.; Hiramatsu, H.; Bartlett, B. M.; Sleightholme, A. E. S.; Siegel, D. J.; Monroe, C. W. Identifying the Discharge Product and Reaction Pathway for a Secondary Mg/O₂ Battery. *Chem. Mater.* **2015**, *27*, 7564–7568. <https://doi.org/10.1021/acs.chemmater.5b03608>.
- (317) Kwak, W. J.; Park, J.; Nguyen, T. T.; Kim, H.; Byon, H. R.; Jang, M.; Sun, Y. K. A Dendrite-and Oxygen-Proof Protective Layer for Lithium Metal in Lithium-Oxygen Batteries. *J. Mater. Chem. A* **2019**, *7*, 3857–3862. <https://doi.org/10.1039/c8ta11941d>.
- (318) Shin, H.-J.; Kwak, W.-J.; Aurbach, D.; Sun, Y.-K. Large-Scale Li-O₂ Pouch Type Cells for Practical Evaluation and Applications. *Adv. Funct. Mater.* **2017**, *27*, 1605500. <https://doi.org/10.1002/adfm.201605500>.
- (319) Lee, H. C.; Park, J. O.; Kim, M.; Kwon, H. J.; Kim, J. H.; Choi, K. H.; Kim, K.; Im, D. High-Energy-Density Li-O₂ Battery at Cell Scale with Folded Cell Structure. *Joule* **2019**, *3*, 542–556. <https://doi.org/10.1016/j.joule.2018.11.016>.
- (320) Ye, L.; Liao, M.; Sun, H.; Yang, Y.; Tang, C.; Zhao, Y.; Wang, L.; Xu, Y.; Zhang, L.; Wang, B.; et al. Stabilizing Lithium into Cross-Stacked Nanotube Sheets with an Ultra-High Specific Capacity for Lithium Oxygen Batteries. *Angew. Chemie - Int. Ed.* **2019**, *58*, 2437–2442. <https://doi.org/10.1002/anie.201814324>.
- (321) Zhang, Y.; Wang, L.; Guo, Z.; Xu, Y.; Wang, Y.; Peng, H. High-Performance Lithium-Air Battery with a Coaxial-Fiber Architecture. *Angew. Chemie Int. Ed.* **2016**, *55*, 4487–4491. <https://doi.org/10.1002/anie.201511832>.
- (322) Liu, Q.-C.; Liu, T.; Liu, D.-P.; Li, Z.-J.; Zhang, X.-B.; Zhang, Y. A Flexible and Wearable Lithium-Oxygen Battery with Record Energy Density Achieved by the Interlaced Architecture Inspired by Bamboo Slips. *Adv. Mater.* **2016**, *28*, 8413–8418. <https://doi.org/10.1002/adma.201602800>.
- (323) Kumar, B.; Kumar, J.; Leese, R.; Fellner, J. P.; Rodrigues, S. J.; Abraham, K. M. A

- Solid-State, Rechargeable, Long Cycle Life Lithium–Air Battery. *J. Electrochem. Soc.* **2010**, *157*, A50–A54. <https://doi.org/10.1149/1.3256129>.
- (324) Wang, Y.; Zhou, H. To Draw an Air Electrode of a Li–air Battery by Pencil. *Energy Environ. Sci.* **2011**, *4*, 1704. <https://doi.org/10.1039/c0ee00759e>.
- (325) Liu, Y.; He, P.; Zhou, H. Rechargeable Solid-State Li-Air and Li-S Batteries: Materials, Construction, and Challenges. *Adv. Energy Mater.* **2018**, *8*, 1701602. <https://doi.org/10.1002/aenm.201701602>.
- (326) Zhu, X.; Zhao, T.; Tan, P.; Wei, Z.; Wu, M. A High-Performance Solid-State Lithium-Oxygen Battery with a Ceramic-Carbon Nanostructured Electrode. *Nano Energy* **2016**, *26*, 565–576. <https://doi.org/10.1016/j.nanoen.2016.06.010>.
- (327) Koerver, R.; Zhang, W.; de Biasi, L.; Schweidler, S.; Kondrakov, A. O.; Kolling, S.; Brezesinski, T.; Hartmann, P.; Zeier, W. G.; Janek, J. Chemo-Mechanical Expansion of Lithium Electrode Materials – on the Route to Mechanically Optimized All-Solid-State Batteries. *Energy Environ. Sci.* **2018**, *11*, 2142–2158. <https://doi.org/10.1039/C8EE00907D>.
- (328) Jin, Y.; Liu, K.; Lang, J.; Zhuo, D.; Huang, Z.; Wang, C.; Wu, H.; Cui, Y. An Intermediate Temperature Garnet-Type Solid Electrolyte-Based Molten Lithium Battery for Grid Energy Storage. *Nat. Energy* **2018**, *3*, 732–738. <https://doi.org/10.1038/s41560-018-0198-9>.
- (329) Balaish, M.; Peled, E.; Golodnitsky, D.; Ein-Eli, Y. Liquid-Free Lithium-Oxygen Batteries. *Angew. Chemie - Int. Ed.* **2015**, *54*, 436–440. <https://doi.org/10.1002/anie.201408008>.
- (330) Shu, C.; Long, J.; Dou, S.-X.; Wang, J. Component-Interaction Reinforced Quasi-Solid Electrolyte with Multifunctionality for Flexible Li-O₂ Battery with Superior Safety under Extreme Conditions. *Small* **2019**, *15*, 1804701. <https://doi.org/10.1002/sml.201804701>.
- (331) Choi, Y.; Jung, K.; Kim, H. J.; Moon, J. W.; Lee, J. W. Lithium-Oxygen Batteries with Triplex Li⁺-Selective Solid Membranes. *Chem. Commun.* **2019**, *55*, 7643–7646. <https://doi.org/10.1039/c9cc03652k>.
- (332) Liu, T.; Liu, Q. C.; Xu, J. J.; Zhang, X. B. Cable-Type Water-Survivable Flexible Li-O₂ Battery. *Small* **2016**, *12*, 3101–3105. <https://doi.org/10.1002/sml.201600540>.
- (333) Peled, E.; Golodnitsky, D.; Mazor, H.; Goor, M.; Avshalomov, S. Parameter Analysis of a Practical Lithium- and Sodium-Air Electric Vehicle Battery. *J. Power Sources* **2011**, *196*, 6835–6840. <https://doi.org/10.1016/J.JPOWSOUR.2010.09.104>.
- (334) Wu, S.; Yi, J.; Zhu, K.; Bai, S.; Liu, Y.; Qiao, Y.; Ishida, M.; Zhou, H. A Super-Hydrophobic Quasi-Solid Electrolyte for Li-O₂ Battery with Improved Safety and Cycle Life in Humid Atmosphere. *Adv. Energy Mater.* **2017**, *7*, 1601759. <https://doi.org/10.1002/aenm.201601759>.
- (335) Yang, Q.; Zhang, Z.; Sun, X.-G.; Hu, Y.-S.; Xing, H.; Dai, S. Ionic Liquids and Derived Materials for Lithium and Sodium Batteries. *Chem. Soc. Rev.* **2018**, *47*, 2020–2064. <https://doi.org/10.1039/C7CS00464H>.
- (336) Kuboki, T.; Okuyama, T.; Ohsaki, T.; Takami, N. Lithium-Air Batteries Using Hydrophobic Room Temperature Ionic Liquid Electrolyte. *J. Power Sources* **2005**, *146*, 766–769. <https://doi.org/10.1016/j.jpowsour.2005.03.082>.
- (337) Bresser, D.; Paillard, E.; Passerini, S. Ionic Liquid-Based Electrolytes for Li Metal/Air Batteries: A Review of Materials and the New ‘LABOHR’ Flow Cell Concept. *J. Electrochem. Sci. Technol.* **2014**, *5*, 37–44. <https://doi.org/10.5229/JECST.2014.5.2.37>.
- (338) Herranz, J.; Garsuch, A.; Gasteiger, H. A. Using Rotating Ring Disc Electrode Voltammetry to Quantify the Superoxide Radical Stability of Aprotic Li–Air Battery

- Electrolytes. *J. Phys. Chem. C* **2012**, *116*, 19084–19094.
<https://doi.org/10.1021/jp304277z>.
- (339) Piana, M.; Wandt, J.; Meini, S.; Buchberger, I.; Tsiouvaras, N.; Gasteiger, H. A. Stability of a Pyrrolidinium-Based Ionic Liquid in Li-O₂ Cells. *J. Electrochem. Soc.* **2014**, *161*, A1992–A2001. <https://doi.org/10.1149/2.1131412jes>.
- (340) Katayama, Y.; Onodera, H.; Yamagata, M.; Miura, T. Electrochemical Reduction of Oxygen in Some Hydrophobic Room-Temperature Molten Salt Systems. *J. Electrochem. Soc.* **2004**, *151*, A59. <https://doi.org/10.1149/1.1626669>.
- (341) Knipping, E.; Aucher, C.; Guirado, G.; Aubouy, L. Room Temperature Ionic Liquids: Versus Organic Solvents as Lithium-Oxygen Battery Electrolytes. *New J. Chem.* **2018**, *42*, 4693–4699. <https://doi.org/10.1039/c8nj00449h>.
- (342) Xia, C.; Kwok, C. Y.; Nazar, L. F. A High-Energy-Density Lithium-Oxygen Battery Based on a Reversible Four-Electron Conversion to Lithium Oxide. *Science (80-.)*. **2018**, *361*, 777–781. <https://doi.org/10.1126/science.aas9343>.
- (343) Giordani, V.; Tozier, D.; Tan, H.; Burke, C. M.; Gallant, B. M.; Uddin, J.; Greer, J. R.; McCloskey, B. D.; Chase, G. V.; Addison, D. A Molten Salt Lithium–Oxygen Battery. *J. Am. Chem. Soc.* **2016**, *138*, 2656–2663. <https://doi.org/10.1021/jacs.5b11744>.
- (344) Sathiya, M.; Rousse, G.; Ramesha, K.; Laisa, C. P.; Vezin, H.; Sougrati, M. T.; Doublet, M. L.; Foix, D.; Gonbeau, D.; Walker, W.; et al. Reversible Anionic Redox Chemistry in High-Capacity Layered-Oxide Electrodes. *Nat. Mater.* **2013**, *12*, 827–835. <https://doi.org/10.1038/nmat3699>.
- (345) Erickson, E. M.; Sclar, H.; Schipper, F.; Liu, J.; Tian, R.; Ghanty, C.; Burstein, L.; Leifer, N.; Grinblat, J.; Talianker, M.; et al. High-Temperature Treatment of Li-Rich Cathode Materials with Ammonia: Improved Capacity and Mean Voltage Stability during Cycling. *Adv. Energy Mater.* **2017**, *7*, 1700708.
<https://doi.org/10.1002/aenm.201700708>.
- (346) Muhammad, S.; Kim, H.; Kim, Y.; Kim, D.; Song, J. H.; Yoon, J.; Park, J.-H.; Ahn, S.-J.; Kang, S.-H.; Thackeray, M. M.; et al. Evidence of Reversible Oxygen Participation in Anomalously High Capacity Li- and Mn-Rich Cathodes for Li-Ion Batteries. *Nano Energy* **2016**, *21*, 172–184.
<https://doi.org/10.1016/J.NANOEN.2015.12.027>.
- (347) Myeong, S.; Cho, W.; Jin, W.; Hwang, J.; Yoon, M.; Yoo, Y.; Nam, G.; Jang, H.; Han, J. G.; Choi, N. S.; et al. Understanding Voltage Decay in Lithium-Excess Layered Cathode Materials through Oxygen-Centred Structural Arrangement. *Nat. Commun.* **2018**, *9*. <https://doi.org/10.1038/s41467-018-05802-4>.
- (348) Lee, J.; Kitchaev, D. A.; Kwon, D.; Lee, C.; Papp, J. K.; Liu, Y.; Lun, Z. Lithium-Excess Cathode Materials. *Nature* **2018**, *556*, 185–190.
<https://doi.org/10.1038/s41586-018-0015-4>.
- (349) Hu, E.; Yu, X.; Lin, R.; Bi, X.; Lu, J.; Bak, S.; Nam, K.-W.; Xin, H. L.; Jaye, C.; Fischer, D. A.; et al. Evolution of Redox Couples in Li- and Mn-Rich Cathode Materials and Mitigation of Voltage Fade by Reducing Oxygen Release. *Nat. Energy* **2018**, *3*, 690–698. <https://doi.org/10.1038/s41560-018-0207-z>.
- (350) Hong, J.; Gent, W. E.; Xiao, P.; Lim, K.; Seo, D. H.; Wu, J.; Csernica, P. M.; Takacs, C. J.; Nordlund, D.; Sun, C. J.; et al. Metal–oxygen Decoordination Stabilizes Anion Redox in Li-Rich Oxides. *Nat. Mater.* **2019**, *18*, 256–265.
<https://doi.org/10.1038/s41563-018-0276-1>.
- (351) Ben Yahia, M.; Vergnet, J.; Saubanère, M.; Doublet, M. L. Unified Picture of Anionic Redox in Li/Na-Ion Batteries. *Nat. Mater.* **2019**, *18*, 496–502.
<https://doi.org/10.1038/s41563-019-0318-3>.
- (352) Assat, G.; Tarascon, J. M. Fundamental Understanding and Practical Challenges of

- Anionic Redox Activity in Li-Ion Batteries. *Nat. Energy* **2018**, *3*, 373–386. <https://doi.org/10.1038/s41560-018-0097-0>.
- (353) Li, Q.; Zhou, D.; Zhang, L.; Ning, D.; Chen, Z.; Xu, Z.; Gao, R.; Liu, X.; Xie, D.; Schumacher, G.; et al. Tuning Anionic Redox Activity and Reversibility for a High-Capacity Li-Rich Mn-Based Oxide Cathode via an Integrated Strategy. *Adv. Funct. Mater.* **2019**, *29*, 1–13. <https://doi.org/10.1002/adfm.201806706>.
- (354) Jacquet, Q.; Iadecola, A.; Saubanère, M.; Lemarquis, L.; Berg, E. J.; Alves Dalla Corte, D.; Rouse, G.; Doublet, M. L.; Tarascon, J. M. Competition between Metal Dissolution and Gas Release in Li-Rich Li₃RuyIr_{1-y}O₄ Model Compounds Showing Anionic Redox. *Chem. Mater.* **2018**. <https://doi.org/10.1021/acs.chemmater.8b02955>.
- (355) Yin, W.; Mariyappan, S.; Grimaud, A.; Tarascon, J. M. Rotating Ring Disk Electrode for Monitoring the Oxygen Release at High Potentials in Li-Rich Layered Oxides. *J. Electrochem. Soc.* **2018**, *165*, A3326–A3333. <https://doi.org/10.1149/2.0481814jes>.
- (356) Assat, G.; Foix, D.; Delacourt, C.; Iadecola, A.; Dedryvère, R.; Tarascon, J.-M. M. J. M. Fundamental Interplay between Anionic/cationic Redox Governing the Kinetics and Thermodynamics of Lithium-Rich Cathodes. *Nat. Commun.* **2017**, *8*, 2219. <https://doi.org/10.1038/s41467-017-02291-9>.
- (357) Lin, R.; Hu, E.; Liu, M.; Wang, Y.; Cheng, H.; Wu, J.; Zheng, J. C.; Wu, Q.; Bak, S.; Tong, X.; et al. Anomalous Metal Segregation in Lithium-Rich Material Provides Design Rules for Stable Cathode in Lithium-Ion Battery. *Nat. Commun.* **2019**, *10*, 1–11. <https://doi.org/10.1038/s41467-019-09248-0>.
- (358) Yang, W.; Devereaux, T. P. Anionic and Cationic Redox and Interfaces in Batteries: Advances from Soft X-Ray Absorption Spectroscopy to Resonant Inelastic Scattering. *J. Power Sources* **2018**, *389*, 188–197. <https://doi.org/10.1016/j.jpowsour.2018.04.018>.
- (359) Assat, G.; Iadecola, A.; Foix, D.; Dedryvère, R.; Tarascon, J. M. Direct Quantification of Anionic Redox over Long Cycling of Li-Rich Nmc via Hard X-Ray Photoemission Spectroscopy. *ACS Energy Lett.* **2018**, *3*, 2721–2728. <https://doi.org/10.1021/acsenergylett.8b01798>.
- (360) Xu, J.; Sun, M.; Qiao, R.; Renfrew, S. E.; Ma, L.; Wu, T.; Hwang, S.; Nordlund, D.; Su, D.; Amine, K.; et al. Elucidating Anionic Oxygen Activity in Lithium-Rich Layered Oxides. *Nat. Commun.* **2018**, *9*, 1–10. <https://doi.org/10.1038/s41467-018-03403-9>.
- (361) Yan, P.; Zheng, J.; Tang, Z. K.; Devaraj, A.; Chen, G.; Amine, K.; Zhang, J. G.; Liu, L. M.; Wang, C. Injection of Oxygen Vacancies in the Bulk Lattice of Layered Cathodes. *Nat. Nanotechnol.* **2019**, *14*, 602–608. <https://doi.org/10.1038/s41565-019-0428-8>.
- (362) Li, Q.; Yao, Z.; Lee, E.; Xu, Y.; Thackeray, M. M.; Wolverton, C.; Dravid, V. P.; Wu, J. Dynamic Imaging of Crystalline Defects in Lithium-Manganese Oxide Electrodes during Electrochemical Activation to High Voltage. *Nat. Commun.* **2019**, *10*, 1–7. <https://doi.org/10.1038/s41467-019-09408-2>.
- (363) Chen, Z.; Li, J.; Zeng, X. C. Unraveling Oxygen Evolution in Li-Rich Oxides: A Unified Modeling of the Intermediate Peroxo/Superoxo-like Dimers. *J. Am. Chem. Soc.* **2019**, *141*, 10751–10759. <https://doi.org/10.1021/jacs.9b03710>.
- (364) Susai, F. A.; Sclar, H.; Shilina, Y.; Penki, T. R.; Raman, R.; Maddukuri, S.; Maiti, S.; Halalay, I. C.; Luski, S.; Markovsky, B.; et al. Horizons for Li-Ion Batteries Relevant to Electro-Mobility: High-Specific-Energy Cathodes and Chemically Active Separators. *Adv. Mater.* **2018**, *1801348*, 1–14. <https://doi.org/10.1002/adma.201801348>.
- (365) Meini, S.; Tsiouvaras, N.; Schwenke, K. U.; Piana, M.; Beyer, H.; Lange, L.; Gasteiger, H. a. Rechargeability of Li-Air Cathodes Pre-Filled with Discharge

- Products Using an Ether-Based Electrolyte Solution: Implications for Cycle-Life of Li-Air Cells. *Phys. Chem. Chem. Phys.* **2013**, *15*, 11478–11493. <https://doi.org/10.1039/c3cp51112j>.
- (366) Bhargav, A.; Fu, Y. Lithium Peroxide-Carbon Composite Cathode for Closed System Li-O₂ Batteries. *J. Electrochem. Soc.* **2015**, *162*, A1327–A1333. <https://doi.org/10.1149/2.0851507jes>.
- (367) Bhargav, A.; Guo, W.; Fu, Y. Chemically Synthesized Lithium Peroxide Composite Cathodes for Closed System Li-O₂ Batteries. *Chem. Commun.* **2016**, *52*, 5678–5681. <https://doi.org/10.1039/c6cc01547f>.
- (368) Marinaro, M.; Balasubramanian, P.; Gucciardi, E.; Theil, S.; Jörissen, L.; Wohlfahrt-Mehrens, M. Importance of Reaction Kinetics and Oxygen Crossover in Aprotic Li-O₂ Batteries Based on a Dimethyl Sulfoxide Electrolyte. *ChemSusChem* **2015**, *8*, 3139–3145. <https://doi.org/10.1002/cssc.201500600>.
- (369) Zhu, Z.; Kushima, A.; Yin, Z.; Qi, L.; Amine, K.; Lu, J.; Li, J. Anion-Redox Nanolithia Cathodes for Li-Ion Batteries. *Nat. Energy* **2016**, *1*, 1–7. <https://doi.org/10.1038/nenergy.2016.111>.
- (370) Liu, J.; Khaleghi Rahimian, S.; Monroe, C. W. Capacity-Limiting Mechanisms in Li/O₂ Batteries. *Phys. Chem. Chem. Phys.* **2016**, *18*, 22840–22851. <https://doi.org/10.1039/c6cp04055a>.
- (371) Gittleston, F. S.; Jones, R. E.; Ward, D. K.; Foster, M. E. Oxygen Solubility and Transport in Li-Air Battery Electrolytes: Establishing Criteria and Strategies for Electrolyte Design. *Energy Environ. Sci.* **2017**, *10*, 1167–1179. <https://doi.org/10.1039/c6ee02915a>.
- (372) Wang, Y.; Zheng, D.; Yang, X. Q.; Qu, D. High Rate Oxygen Reduction in Non-Aqueous Electrolytes with the Addition of Perfluorinated Additives. *Energy Environ. Sci.* **2011**, *4*, 3697–3702. <https://doi.org/10.1039/c1ee01556g>.
- (373) Zhang, S. S.; Read, J. Partially Fluorinated Solvent as a Co-Solvent for the Non-Aqueous Electrolyte of Li/air Battery. *J. Power Sources* **2011**, *196*, 2867–2870. [https://doi.org/DOI 10.1016/j.jpowsour.2010.11.021](https://doi.org/DOI%2010.1016/j.jpowsour.2010.11.021).
- (374) Nishikami, Y.; Konishi, T.; Omoda, R.; Aihara, Y.; Oyaizu, K.; Nishide, H. Oxygen-Enriched Electrolytes Based on Perfluorochemicals for High-Capacity Lithium-Oxygen Batteries. *J. Mater. Chem. A* **2015**, *3*, 10845–10850. <https://doi.org/10.1039/c5ta02219c>.
- (375) Balaish, M.; Ein-Eli, Y. The Role of Air-Electrode Structure on the Incorporation of Immiscible PFCs in Nonaqueous Li-O₂ Battery. *ACS Appl. Mater. Interfaces* **2017**, *9*, 9726–9737. <https://doi.org/10.1021/acsami.7b00076>.
- (376) Balaish, M.; Gao, X.; Bruce, P. G.; Ein-Eli, Y. Enhanced Li-O₂ Battery Performance in a Binary “Liquid Teflon” and Dual Redox Mediators. *Adv. Mater. Technol.* **2019**, *4*, 2–7. <https://doi.org/10.1002/admt.201800645>.
- (377) Wijaya, O.; Rinaldi, A.; Younesi, R.; Yazami, R. The Origin of Li-O₂ Battery Performance Enhancement Using Fluorocarbon Additive. *J. Electrochem. Soc.* **2016**, *163*, A2660–A2664. <https://doi.org/10.1149/2.0671613jes>.
- (378) Kwon, H. J.; Lee, H. C. H. C.; Ko, J.; Jung, I. S.; Lee, H. C. H. C.; Lee, H. C. H. C.; Kim, M.; Lee, D. J.; Kim, H.; Kim, T. Y.; et al. Effects of Oxygen Partial Pressure on Li-Air Battery Performance. *J. Power Sources* **2017**, *364*, 280–287. <https://doi.org/10.1016/j.jpowsour.2017.08.052>.
- (379) Qiao, Y.; Wu, S.; Yi, J.; Sun, Y.; Guo, S.; Yang, S.; He, P.; Zhou, H. From O₂⁻ to HO₂⁻: Reducing By-Products and Overpotential in Li-O₂ Batteries by Water Addition. *Angew. Chemie - Int. Ed.* **2017**, *56*, 4960–4964. <https://doi.org/10.1002/anie.201611122>.

- (380) Liu, T.; Frith, J. T.; Kim, G.; Kerber, R. N.; Dubouis, N.; Shao, Y.; Liu, Z.; Magusin, P. C. M. M.; Casford, M. T. L.; Garcia-Araez, N.; et al. The Effect of Water on Quinone Redox Mediators in Nonaqueous Li-O₂ Batteries. *J. Am. Chem. Soc.* **2018**, *140*, 1428–1437. <https://doi.org/10.1021/jacs.7b11007>.
- (381) Liu, T.; Kim, G.; Jónsson, E.; Castillo-Martinez, E.; Temprano, I.; Shao, Y.; Carretero-González, J.; Kerber, R. N.; Grey, C. P. Understanding LiOH Formation in a Li-O₂ Battery with LiI and H₂O Additives. *ACS Catal.* **2019**, *9*, 66–77. <https://doi.org/10.1021/acscatal.8b02783>.
- (382) Takechi, K.; Shiga, T.; Asaoka, T. A Li-O₂/CO₂ Battery. *Chem. Commun.* **2011**, *47*, 3463–3465. <https://doi.org/10.1039/c0cc05176d>.
- (383) Yin, W.; Grimaud, A.; Lepoivre, F.; Yang, C.; Tarascon, J. M. Chemical vs Electrochemical Formation of Li₂CO₃ as a Discharge Product in Li-O₂/CO₂ Batteries by Controlling the Superoxide Intermediate. *J. Phys. Chem. Lett.* **2017**, *8*, 214–222. <https://doi.org/10.1021/acs.jpcclett.6b02610>.
- (384) Gowda, S. R.; Brunet, A.; Wallraff, G. M.; McCloskey, B. D. Implications of CO₂ Contamination in Rechargeable Nonaqueous Li-O₂ Batteries. *J. Phys. Chem. Lett.* **2013**, *4*, 276–279. <https://doi.org/10.1021/jz301902h>.
- (385) Liu, Z.; Zhang, Y.; Jia, C.; Wan, H.; Peng, Z. Z.; Bi, Y.; Liu, Y.; Peng, Z. Z.; Wang, Q.; Li, H.; et al. Decomposing Lithium Carbonate with a Mobile Catalyst. *Nano Energy* **2017**, *36*, 390–397. <https://doi.org/10.1016/j.nanoen.2017.04.049>.
- (386) Xu, S.; Das, S. K.; Archer, L. A. The Li-CO₂ Battery: A Novel Method for CO₂ Capture and Utilization. *RSC Adv.* **2013**, *3*, 6656–6660.
- (387) Qiao, Y.; Yi, J.; Wu, S.; Liu, Y.; Yang, S.; He, P.; Zhou, H. Li-CO₂ Electrochemistry: A New Strategy for CO₂ Fixation and Energy Storage. *Joule* **2017**, *1*, 359–370. <https://doi.org/10.1016/j.joule.2017.07.001>.
- (388) Zhao, Z.; Huang, J.; Peng, Z. Achilles' Heel of Lithium–Air Batteries: Lithium Carbonate. *Angew. Chemie - Int. Ed.* **2018**, *57*, 3874–3886. <https://doi.org/10.1002/anie.201710156>.
- (389) Yang, S.; He, P.; Liu, Y.; Cheng, Z.; Zhou, H.; Qiao, Y.; Zhu, J.-J. A Reversible Lithium-CO₂ Battery with Ru Nanoparticles as a Cathode Catalyst. *Energy Environ. Sci.* **2017**, *10*, 972–978. <https://doi.org/10.1039/C6EE03770D>.
- (390) Adams, J.; Karulkar, M. Bipolar Plate Cell Design for a Lithium Air Battery. *J. Power Sources* **2012**, *199*, 247–255. <https://doi.org/10.1016/j.jpowsour.2011.10.041>.
- (391) Viswanathan, V.; Nørskov, J. K.; Speidel, A.; Scheffler, R.; Gowda, S.; Luntz, A. C. Li–O₂ Kinetic Overpotentials: Tafel Plots from Experiment and First-Principles Theory. *J. Phys. Chem. Lett.* **2013**, *4*, 556–560. <https://doi.org/10.1021/jz400019y>.
- (392) Kumar, A.; Madden, D. G.; Lusi, M.; Chen, K. J.; Daniels, E. A.; Curtin, T.; Perry, J. J.; Zaworotko, M. J. Direct Air Capture of CO₂ by Physisorbent Materials. *Angew. Chemie - Int. Ed.* **2015**, *54*, 14372–14377. <https://doi.org/10.1002/anie.201506952>.
- (393) Ben-Mansour, R.; Habib, M. A.; Bamidele, O. E.; Basha, M.; Qasem, N. A. A.; Peedikakkal, A.; Laoui, T.; Ali, M. Carbon Capture by Physical Adsorption: Materials, Experimental Investigations and Numerical Modeling and Simulations – A Review. *Appl. Energy* **2016**, *161*, 225–255. <https://doi.org/10.1016/J.APENERGY.2015.10.011>.
- (394) Ünveren, E. E.; Monkul, B. Ö.; Sariođlan, Ş.; Karademir, N.; Alper, E. Solid Amine Sorbents for CO₂ Capture by Chemical Adsorption: A Review. **2017**, *3*, 37–50. <https://doi.org/10.1016/j.petlm.2016.11.001>.
- (395) Hamed, M.; Wigenius, J.; Tai, F.; Björk, P.; Aili, D. A High Performance O₂ Selective Membrane Based on CAU-1- NH₂@polydopamine and PMMA Polymer for Li-Air Batteries. **2010**, *51*, 4364–4367. <https://doi.org/10.1039/b000000x>.

- (396) Qian, X.; Han, B.; Liu, Y.; Yan, H.; Liu, R. Vapor Pressure of Dimethyl Sulfoxide and Water Binary System. *J. Solution Chem.* **1995**, *24*, 1183–1189.
<https://doi.org/10.1007/BF00972964>.
- (397) Tang, S.; Zhao, H. Glymes as Versatile Solvents for Chemical Reactions and Processes: From the Laboratory to Industry. *RSC Adv.* **2014**, *4*, 11251–11287.
<https://doi.org/10.1039/c3ra47191h>.

Biodata

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